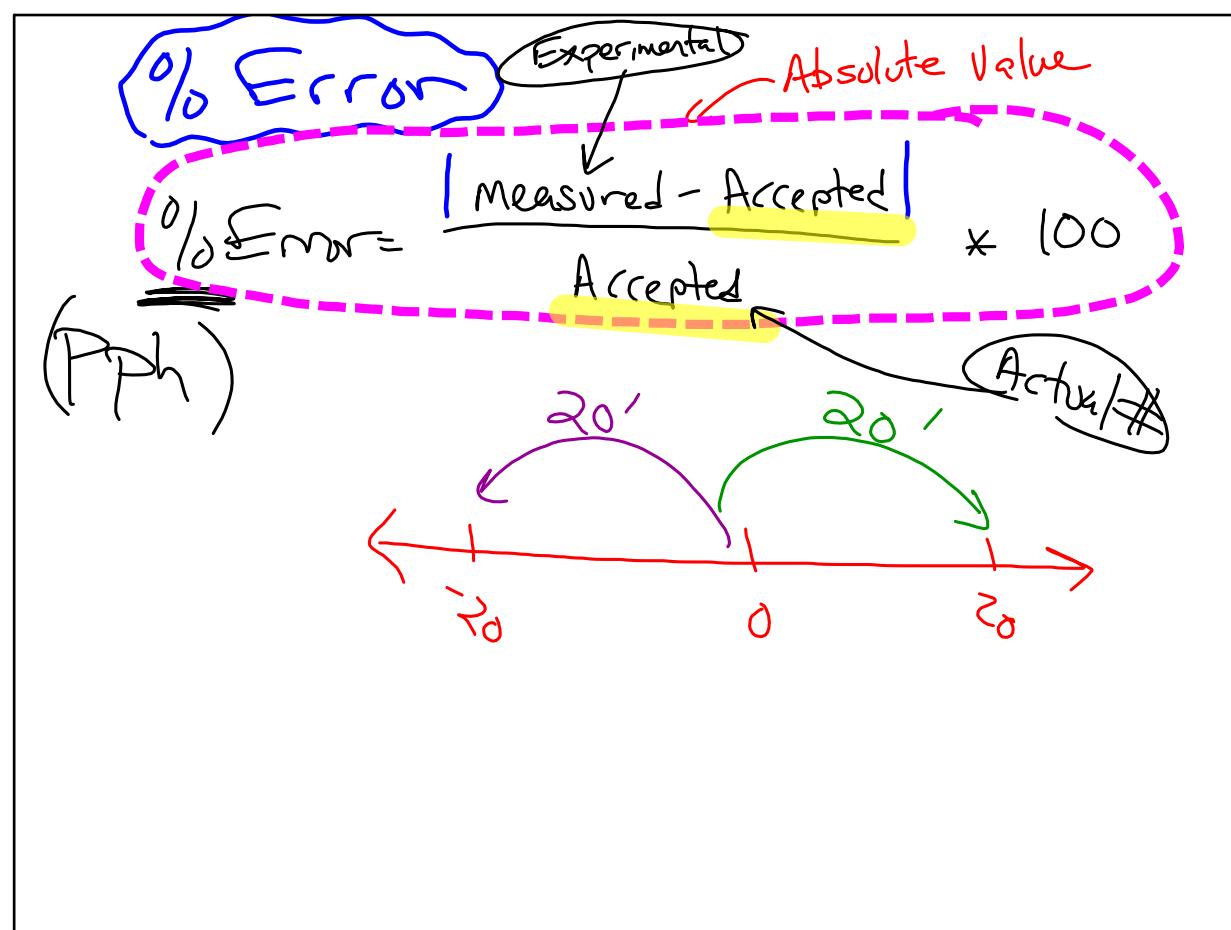


Dec 14-8:34 AM



Dec 14-9:00 AM

$$\text{Percent Yield} = \frac{\text{measured}}{\text{accepted}} * 100$$

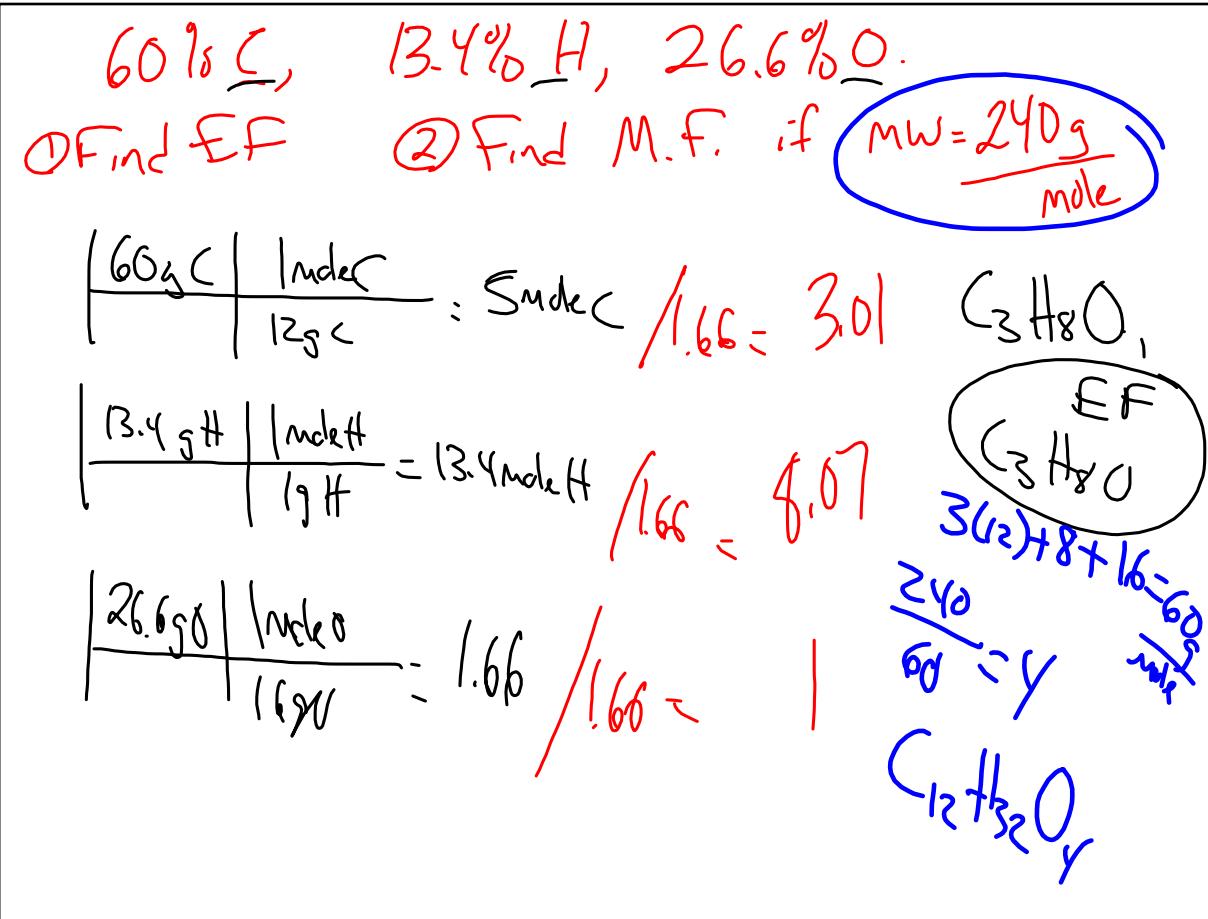
$$= \frac{\text{Experimental}}{\text{actual}} * 100$$

9.5
12
79.2% Yield

Nick
9.5g C
Should have
12g C

% error = $\frac{9.5 - 12}{12} = -20.8\% \text{ error}$

Dec 14-9:10 AM



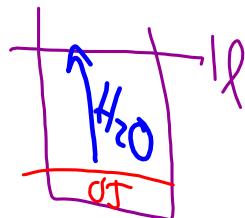
Dec 14-9:15 AM

(M) MOLARITY → measure of concentration

$$M = \frac{\text{Moles solute}}{\ell \text{ of solution}}$$

[OJ concentrate] [water]

Water + OJ concentrate
Homogeneous Mixture



Dec 14-9:30 AM

Kool-Aid

Sugar ($C_6H_{12}O_6$) ↑

Solution

1 can ⇒ 8 gts.

538g → 2 gal Given

1 gallon = 3.78 l

more dilute

less Kool-Aid Powder

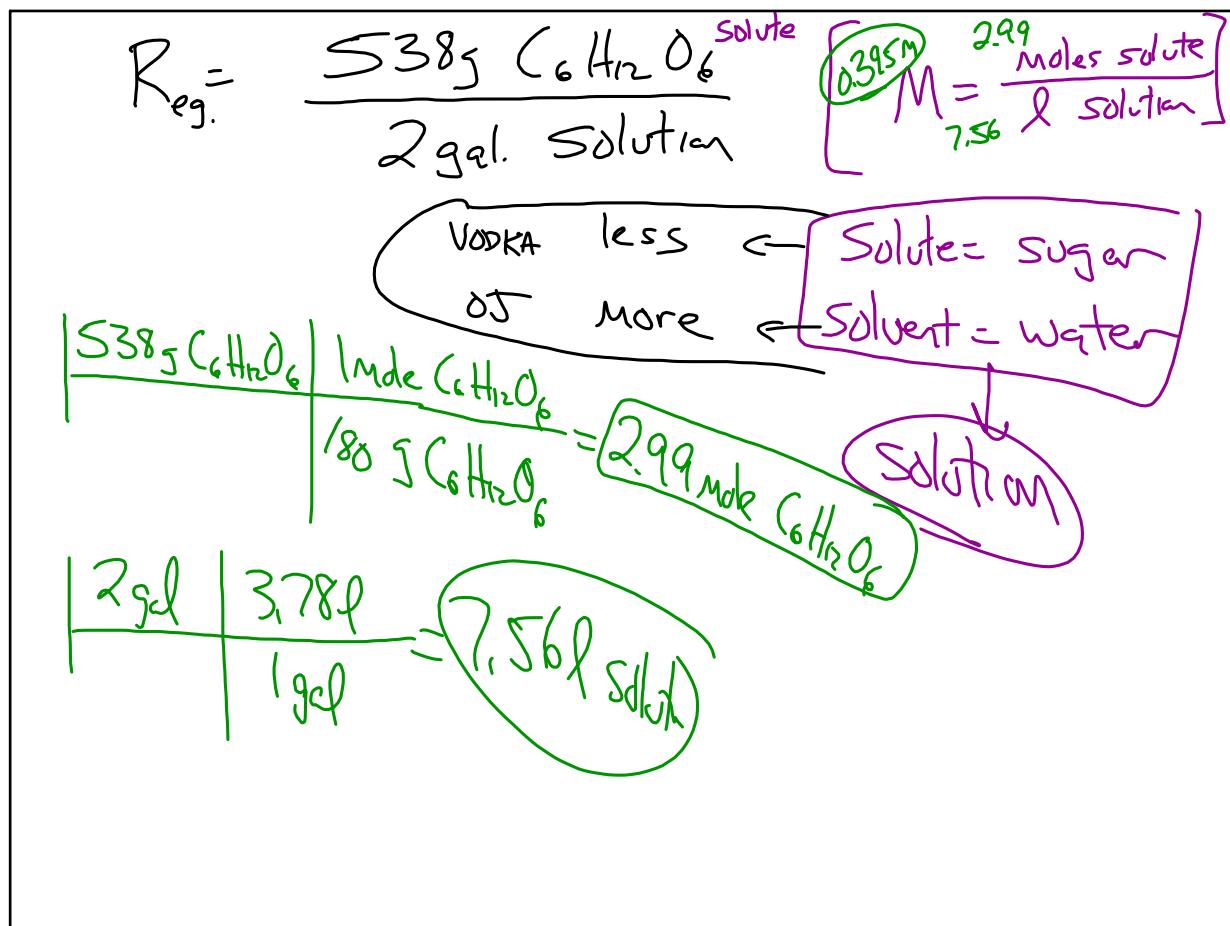
↑ Concentrate

Add more Kool-Aid Powder

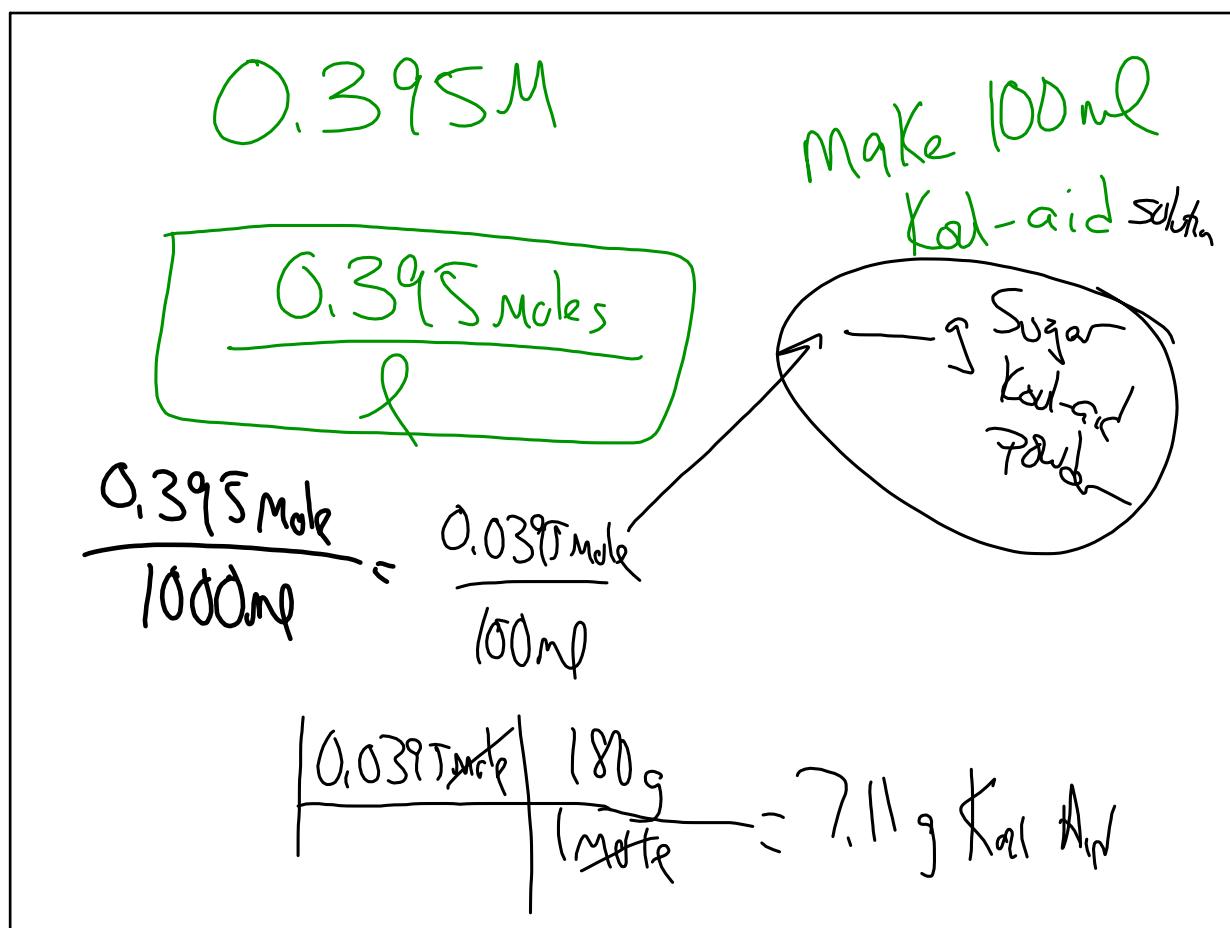
Regular Kool-Aid

M

Dec 14-9:41 AM



Dec 14-9:49 AM



Dec 14-9:59 AM