

MOLARITY (M) = $\frac{\text{moles solute}}{\text{L solution}}$

Concentration

Low ← Solvent > Solute
"Dilute"
"Watery"

→ High
More concentrated

(Solution)

Dec 16-8:43 AM

KOOL - AID (Jim Jones
Guyana Tragedy)

Regular 0.395M (MOLAR)

Dec 16-9:35 AM

Regular
 $\frac{0.395 \text{ moles } C_6H_{12}O_6}{1 \text{ SOLUTION}} = 0.395 M$

↑
molar

Find #grams needed = 71.1g Kool-Aid ?

$0.395 \text{ mole } C_6H_{12}O_6$	$180g \text{ } C_6H_{12}O_6$	= 71.1g $C_6H_{12}O_6$
$1 \text{ mole } C_6H_{12}O_6$		

$\frac{71.1g}{1000ml} = \frac{7.11g \text{ } C_6H_{12}O_6}{100ml}$ Regular strength

Dec 16-9:44 AM

$\frac{538g \text{ } C_6H_{12}O_6}{2 \text{ gallons}}$

 $\frac{\text{moles}}{l}$
MOLARITY
"Molar solution"

$538g \text{ } C_6H_{12}O_6$	$1g/l$	$1 \text{ mole } C_6H_{12}O_6$	= $\frac{0.395 \text{ mole}}{l}$
2 gallons	$3.78l$	$180g \text{ } C_6H_{12}O_6$	

$0.395 M$

Dec 16-9:41 AM

XXX Kool Aid

0.395M 1.195M

$\frac{7.11g}{100ml}$ \Rightarrow $\frac{21.33g}{100ml}$

Reg XXX

Dec 16-9:56 AM

Color
Transparency
Smell
Taste
Consistency (Swirl)

Dec 16-9:59 AM

P460 1,3,5,7
TEXT

Dec 16-10:02 AM