

⑦

170 g NH₃	1 Mole NH₃	22.4 L	=
17 g NH ₃	1 Mole NH ₃	2.24 L	

Dec 21-9:21 AM

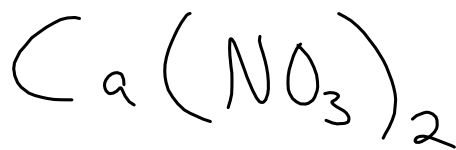
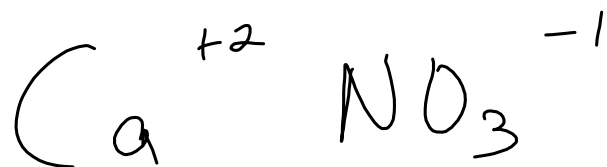
⑧

3 mole O ₂	32 g O ₂	=	96 g O ₂
1 mole O ₂	32 g O ₂		

⑩

2 g He	1 mole He	22.4 L	=	11.2 L
4 g He	1 mole He	22.4 L		

Dec 21-9:36 AM



$$40 + 2(14) + 6(16) = 164$$

Dec 21-9:38 AM

② 40 pounds $\text{C}_3\text{H}_8(\text{g})$, $\frac{453.6\text{g}}{1\text{pound}}$, — 2

40 pounds C_3H_8	453.6g C_3H_8	1 mole C_3H_8	22.4L
	1 pound C_3H_8	44g C_3H_8	1 mole :

9236.95L
 C_3H_8

Dec 21-9:40 AM

(EC) $4.88 \times 10^{-3} \text{ g}$ $\text{Al}^{+3} \text{NO}_3^{-1}$ } — atoms O
 $\text{Al}(\text{NO}_3)_3$

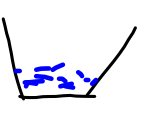


$4.88 \times 10^{-3} \text{ g Al}(\text{NO}_3)_3$	1 mole $\text{Al}(\text{NO}_3)_3$	9 mole O	6×10^{23} atoms O
$213 \text{ g Al}(\text{NO}_3)_3$	1 mole $\text{Al}(\text{NO}_3)_3$	1 mole O	1 mole O


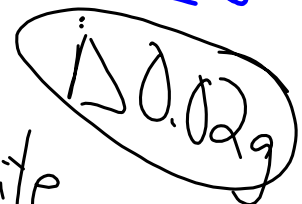
* 1.24×10^{20} atoms O

Dec 21-9:45 AM

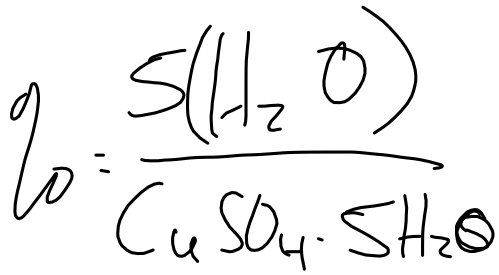
% H₂O in a hydrate $\frac{\text{H}_2\text{O}}{\text{CuSO}_4 \cdot 5\text{H}_2\text{O}} \times 100$

$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

$\approx 3 \text{ g}$   

CuSO_4  Anhydrous ← white 

Dec 21-9:54 AM



Dec 21-10:01 AM