$$\overset{0.09}{Mg} + \overset{0.02}{O_2} \rightarrow Mg\,\overset{0.11}{O} \qquad \frac{0.02}{0.11} \times 100$$

0.07g   0.09   0.16g   A.S.   18%
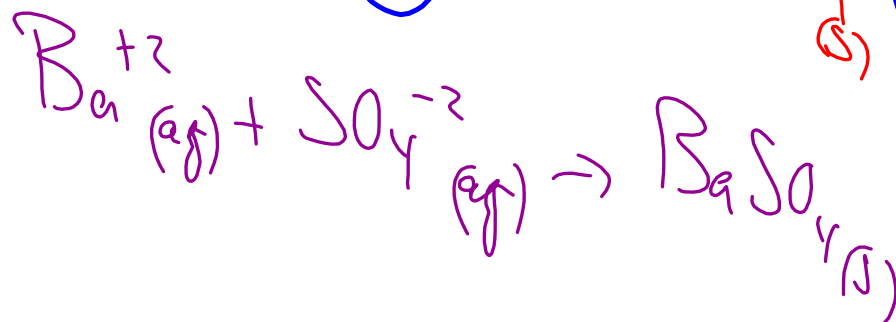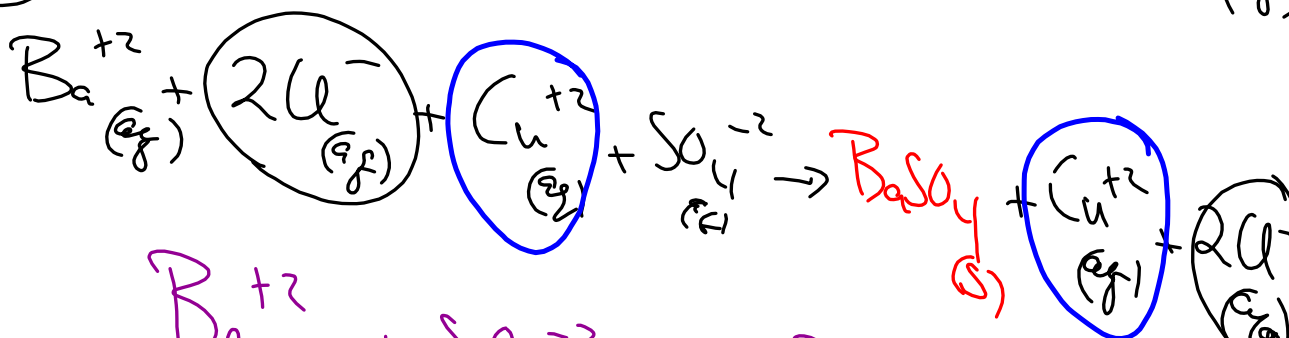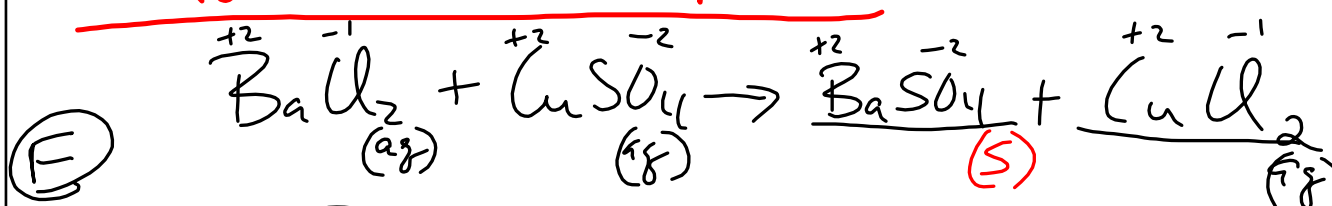
% O in MgO $\frac{O}{MgO}$

Actual on PT $\rightarrow$ $\frac{16}{24+16} \times 100 = 40\%$

EXP $\frac{O}{MgO} \rightarrow$ $\frac{0.09}{0.16} = 56.25\%$ Cart heating.

$$\%E = \frac{|A - E|}{A} \times 100$$

Actual   Experment

Actual

Mole test chap 8+9

$$\overset{+2}{Ba}\overset{-1}{Cl_2}_{(aq)} + \overset{+2}{Cu}\overset{-2}{SO_4}_{(aq)} \rightarrow \overset{+2}{Ba}\overset{-2}{SO_4}_{(S)} + \overset{+2}{Cu}\overset{-1}{Cl_2}_{(aq)}$$

(F)

$$Ba^{+2}_{(aq)} + 2Cl^-_{(aq)} + Cu^{+2}_{(aq)} + SO_4^{-2}_{(aq)} \rightarrow BaSO_4_{(S)} + Cu^{+2}_{(aq)} + 2Cl^-_{(aq)}$$

$$Ba^{+2}_{(aq)} + SO_4^{-2}_{(aq)} \rightarrow BaSO_4_{(S)}$$

Review mole test, midterm, heating curve   **2**

㉓ $2 Mg_{(s)} + O_{2(g)} \rightarrow 2MgO$

—— g $2.39 g$

$$\dfrac{2.39 g\ MgO}{} \cdot \dfrac{1\ mole\ MgO}{40\ g\ MgO} \cdot \dfrac{2\ Mole\ Mg}{2\ Mole\ MgO} \cdot \dfrac{24_5\ Mg}{1\ Mole\ Mg} =$$

Review mole test, midterm, heating curve   **3**

② $N_2 + 3H_2 \rightarrow 2NH_3$

$\underline{\quad} g \qquad 1 mole$

$$\frac{1 mole\ NH_3}{} \cdot \frac{3\ mole\ H_2}{2\ mole\ NH_3} \cdot \frac{2g\ H_2}{1\ mole\ H_2} =$$

Feb 2-9:08 AM
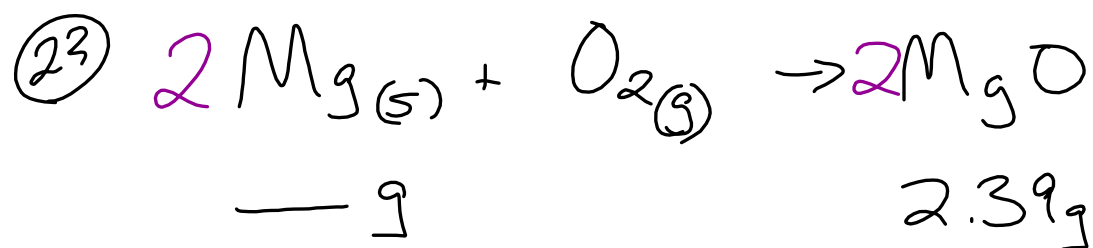
$$K = °C + 273$$

°C		K
100		373
0		273
-273		0

$$D = \frac{M}{V}$$

$$\frac{10.04}{8.21 \, cm^3} \leftarrow (4)$$

(3)

(35)

$$.75 \left(35\right) + .25 \left(37\right)$$

$$+ \begin{array}{c} 1.230 \\ 1.1 \\ \hline 1.758 \\ \hline 4.088 \end{array}$$

(4.1)

Thermochemistry

H_2O

NO △KE
only △PE
Phase △
→

100℃ BP

TEMP ℃
Avg KE

NO △KE
△PE
Phase △
S → ℓ

$H_f = 334\ J/g$
$Q = mH_f$

MP/FP
0℃

ℓ　　　　　g

$H_v = 2260\ J/g$

$Q = mH_v$

$Q = mc\Delta T$

2

Time

6 min

12