
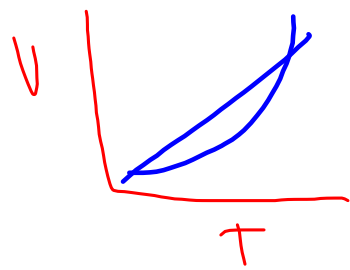


## Combined Gas Law

$$\frac{PV}{T}$$


Mult ( $P \times V = \text{const}$ ) Indirect

Division ( $\frac{P}{T} = \text{const}$   $\frac{V}{T} = \text{const}$ )  
→ Direct



Feb 23-8:28 AM

## Boyles Law Lab

$P_1 =$  \_\_\_\_\_

$P_2 =$  \_\_\_\_\_

$P_3 =$  \_\_\_\_\_

$V_1$

Volume  
in  
Syringe

$V_2 =$  \_\_\_\_\_

$V_3 =$  \_\_\_\_\_

10 → 4.5

4.5

Feb 23-8:48 AM

Boyles LAW

DUE THURS

① Marshmallow

Procedure + obs.



② Soda Bottle + Syringe

P vs Volume left in Syringe

Procedure + graph (Best fit)

Discussion (error analysis)

P and V rel. ↑

Feb 23-8:51 AM

GAS Packet

PIS

②

$$V_1 = 17\ell, P_1 = 2.3 \text{ atm}, T_1 = 299 \text{ K}$$

$$\frac{PV}{T} = \text{constant}$$

$$V_2 = ?, P_2 = 1.5 \text{ atm}, T_2 = 350 \text{ K}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{(2.3)(17)}{299} = \frac{(1.5)V_2}{350}$$

$$V_2 = 30.5 \ell$$

Feb 23-8:58 AM

Charles Law Lab

Feb 23-9:08 AM