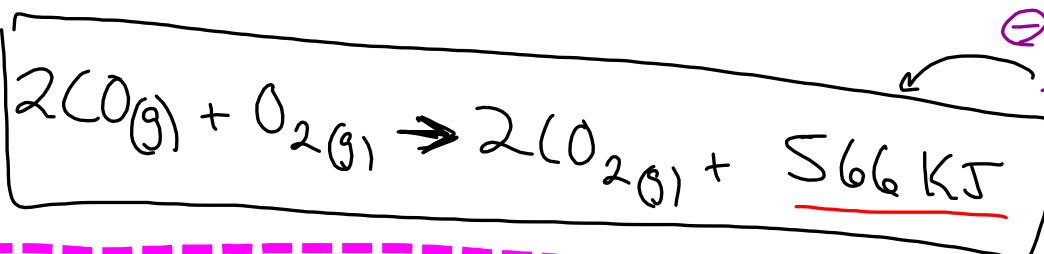




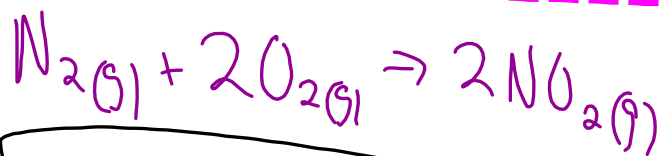
$\Delta H = -566 \text{ kJ}$

⊖ means exothermic

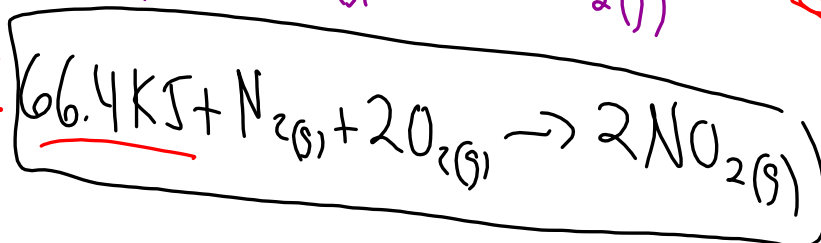


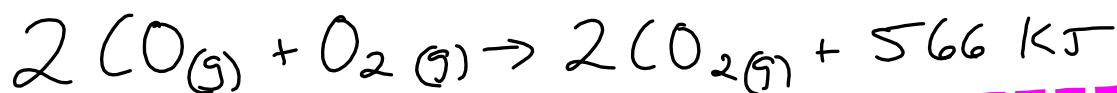
heat released

end heat gained.



$\Delta H = +66.4 \text{ kJ}$





① Find ΔH for $\text{CO}_2 \text{(g)}$ in kJ/mole .

566 kJ	1 mole O_2
2 mole CO_2	

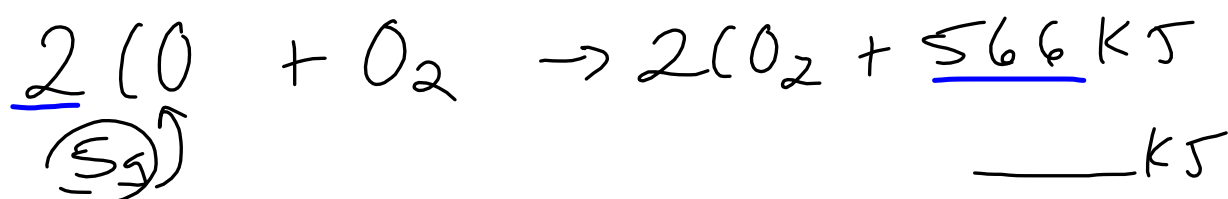
$$\frac{283 \text{ kJ}}{\text{mole}}$$

ΔH is part of the
MOLE RATIO

$$\frac{566 \text{ kJ}}{2 \text{ mole } \text{O}_2} = \frac{283 \text{ kJ}}{1 \text{ mole } \text{O}_2}$$

2 mole O_2	566 kJ
1 mole O_2	

2 mole O_2
$2(566) = 1132 \text{ kJ}$



5g CO	1 mole CO	566 kJ	=	50.5 kJ
	28g CO	2 mole CO		