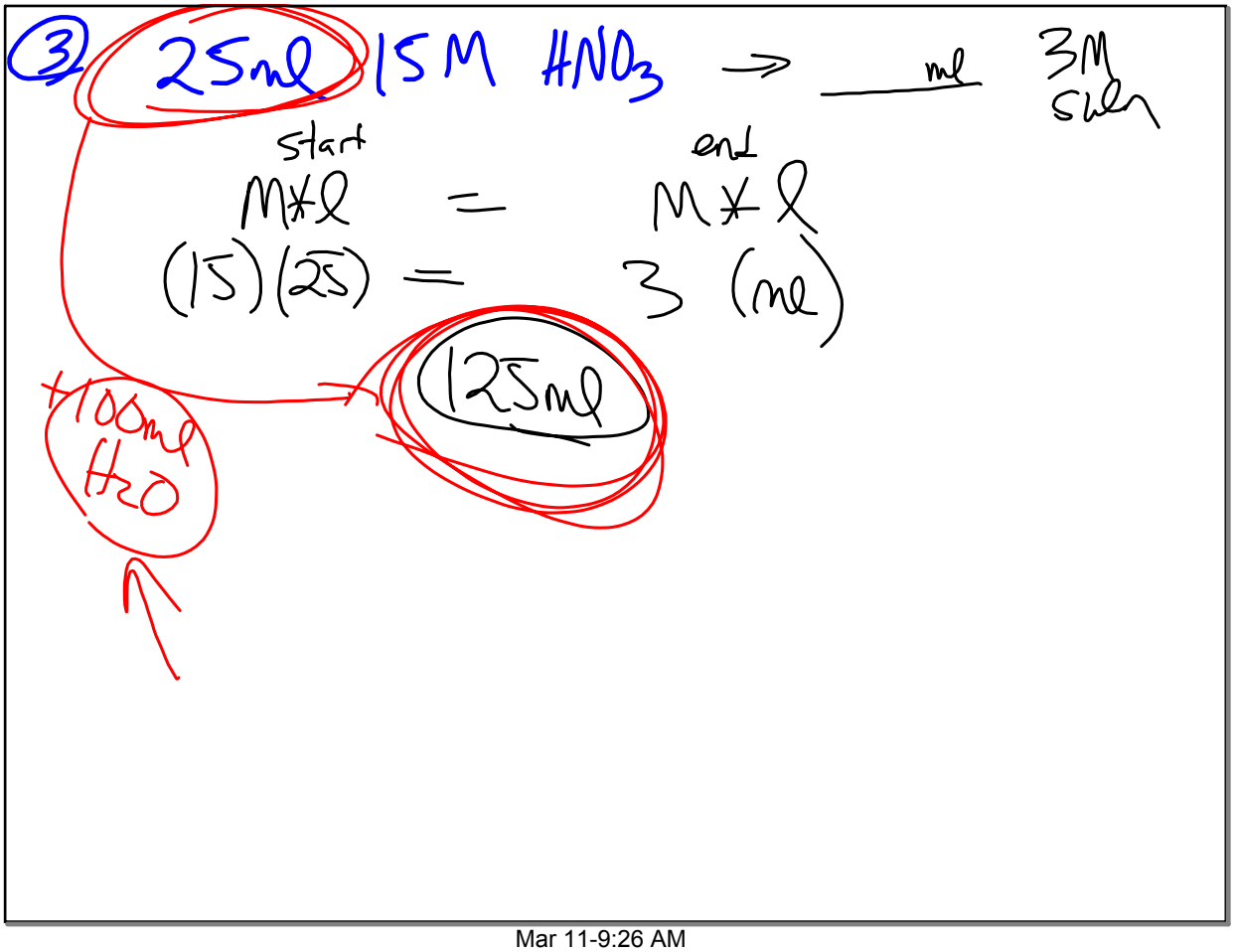


Mar 11-9:26 AM

Molality = moles solute / Kg Solvent = m

25g A + 200g B Find m

Solvent is ALWAYS present in greater amount.

Mar 11-9:36 AM

300g NaCl in 0.5L H_2O @ 25°C

0.5Kg (500g)

$d_{H_2O} = 1\ g/ml$

$m = \frac{\text{Moles solute}}{\text{Kg solvent}}$

$\frac{1g}{1ml} \frac{*1000}{*1000} = \frac{1Kg}{1L}$

$\frac{300\ g\ NaCl}{0.5\ Kg} \left| \frac{1\ Mole\ NaCl}{58\ g\ NaCl} \right. = \frac{10.34\ moles}{Kg} = 10.34\ m$

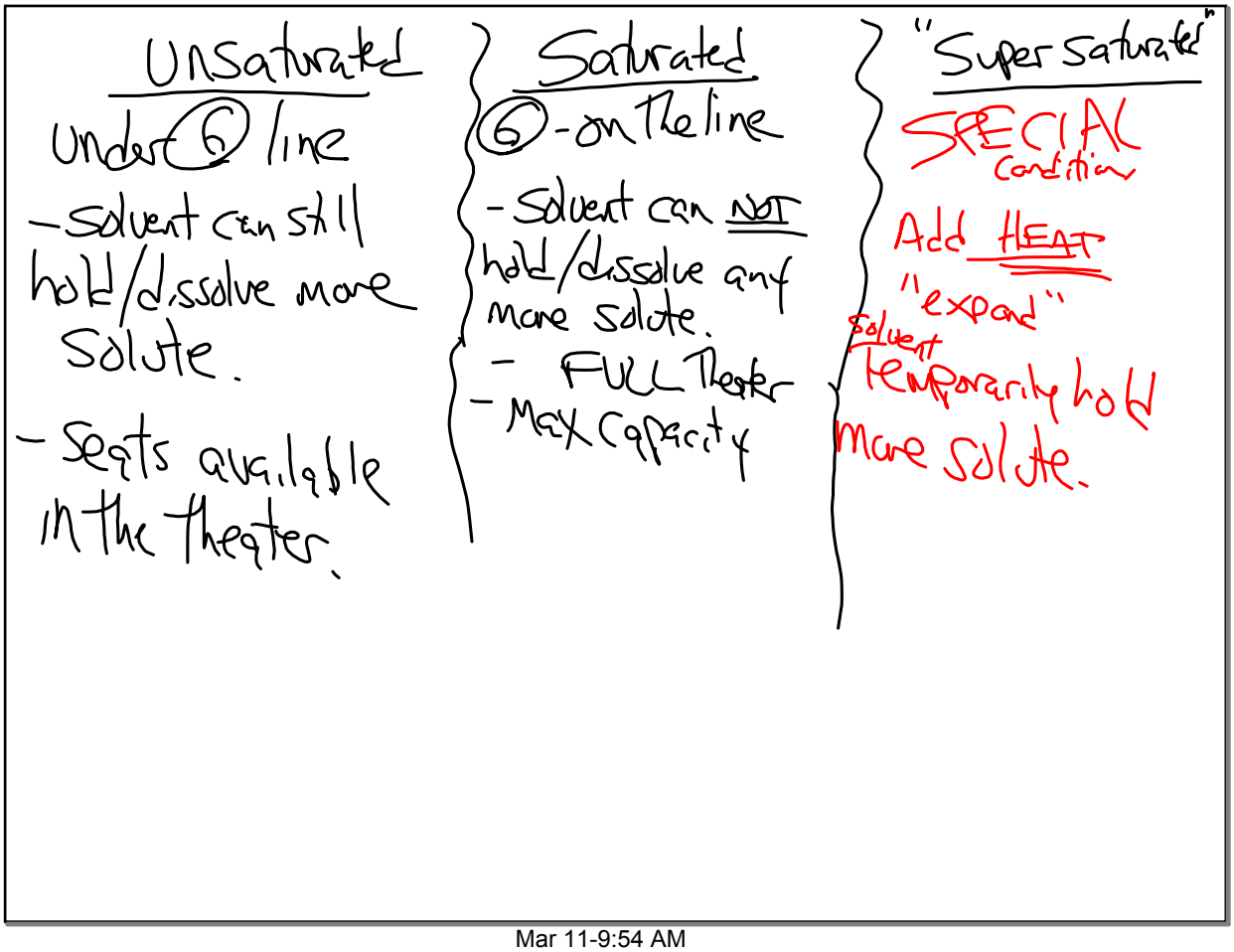
Molal soln

Mar 11-9:39 AM

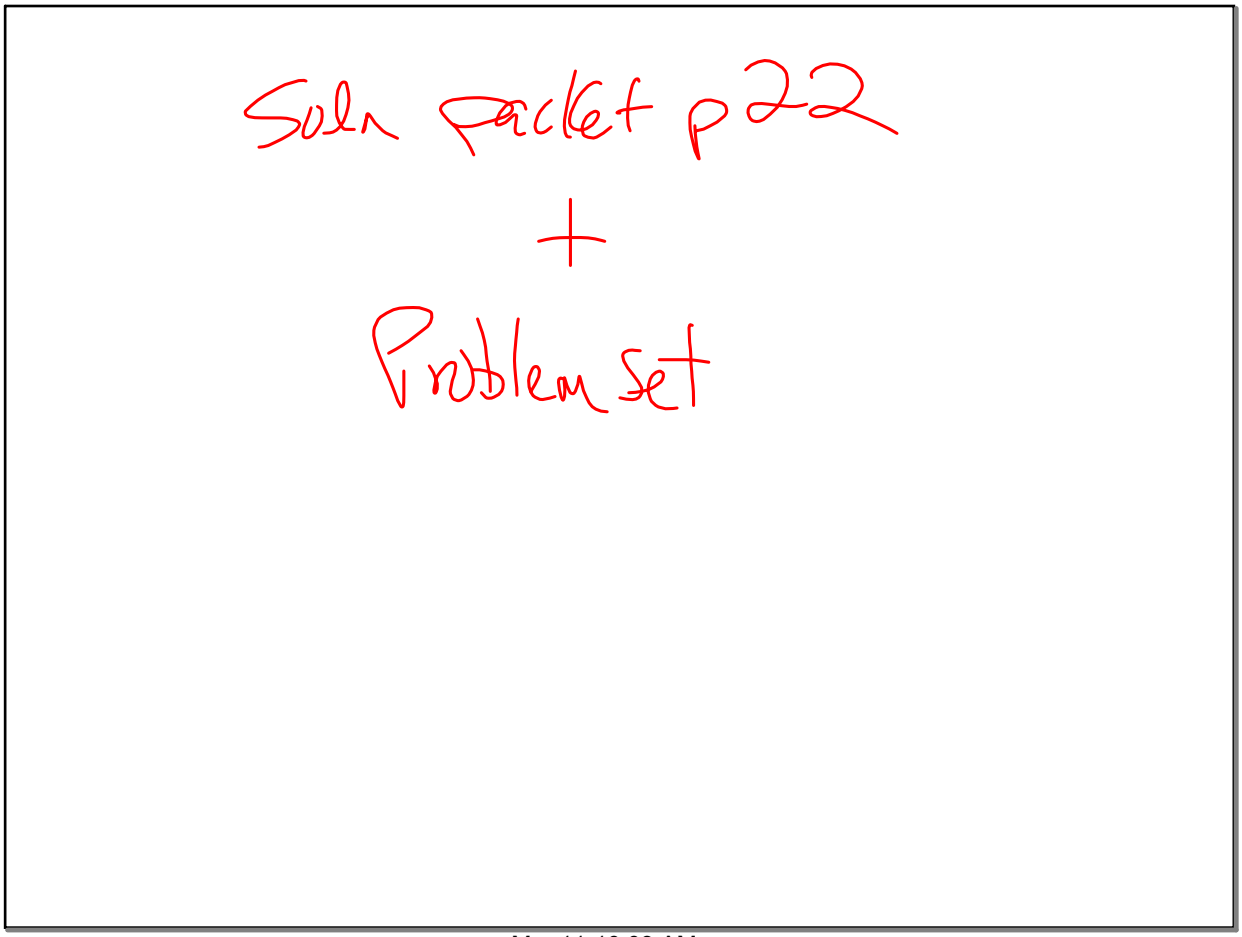
Steps to make a solute dissolve Quicker in the solvent

1. ↑T (Heat da wawa) → ↑KE (More collisions)
2. STIR
3. ↑P of a gas. (on a gas)
4. ↑ Surface area (Cube → granulated → powder SUGAR)
5. ↓ Concentration already in soln.

Mar 11-9:48 AM



Mar 11-9:54 AM



Mar 11-10:03 AM