3/14 P.S.

(25) 50ml gas 293 → 303 20°C → 30°C new V = ?

$\frac{V_1}{T_1} = \frac{V_2}{T_2}$

$\frac{303}{1} * \frac{50}{293} = \frac{V_2}{303} * \frac{303}{1}$

(26) $\frac{0.5 mole}{11.2 l} = \frac{1 mole}{22.4 l}$ STP

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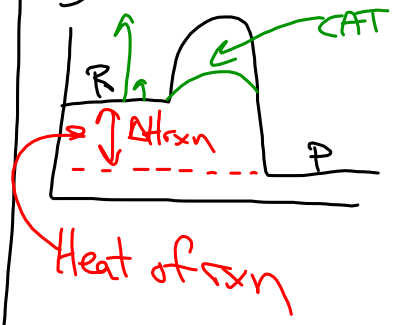
(27) $\frac{PV}{T} = \frac{PV}{T}$

$\frac{(101.3)(1,400)}{20} = \frac{(50.65)V_2}{40}$

(40)



Heat of rxn

(28) 7 Mole A + 3 mole B $P_T = 1$ atm

$P_A = X_A(P_T) = \frac{7}{10}(1) = 0.7$ atm

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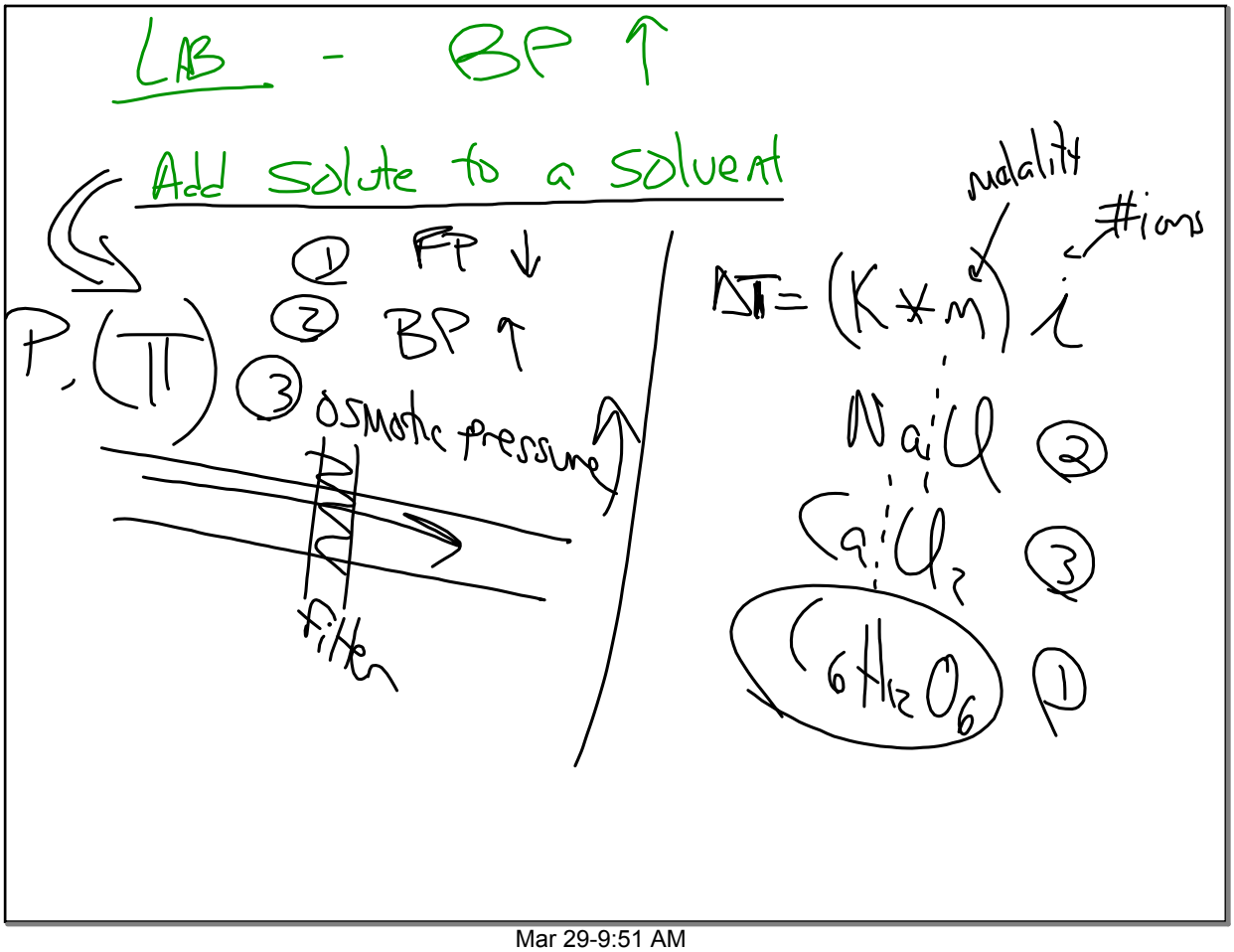

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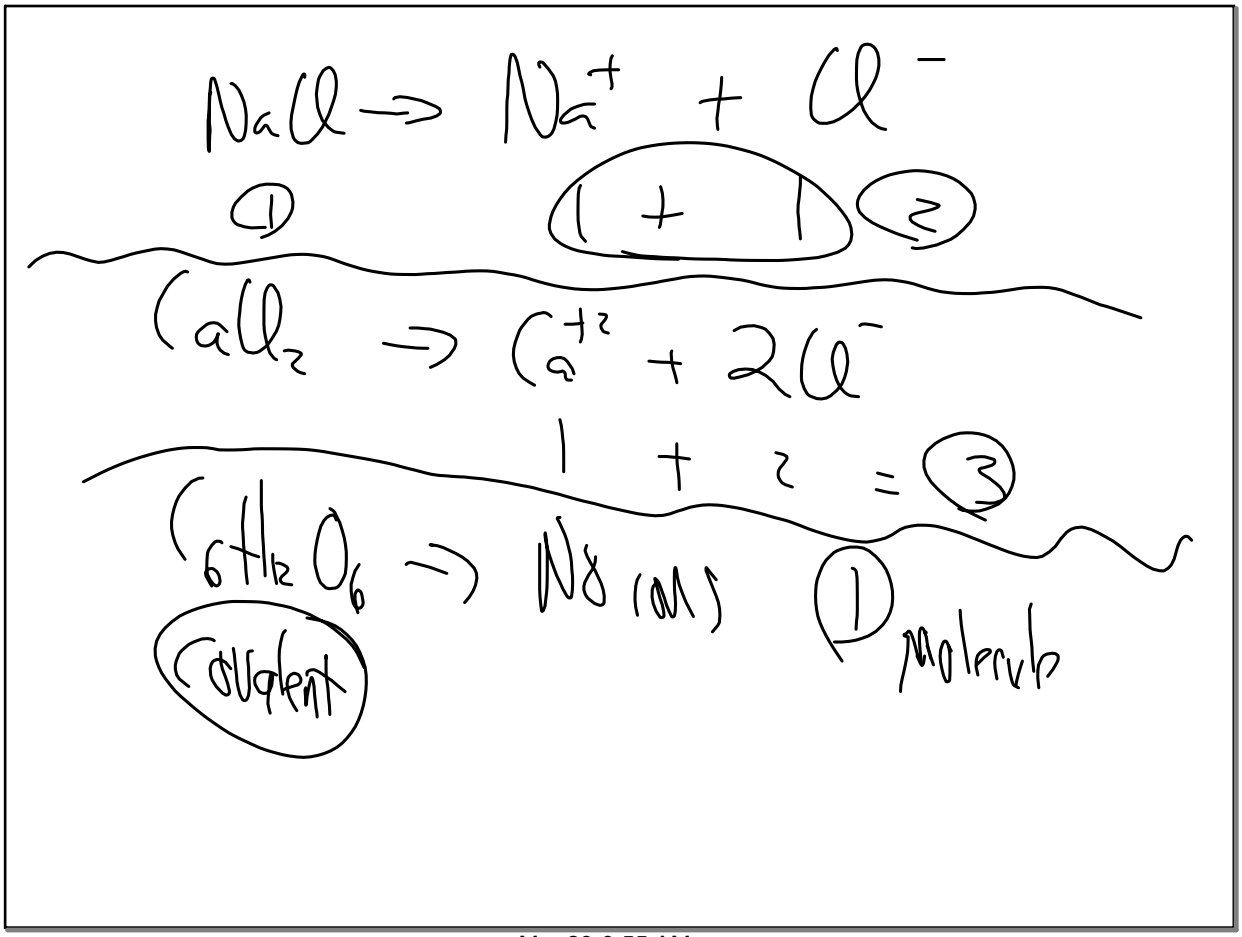

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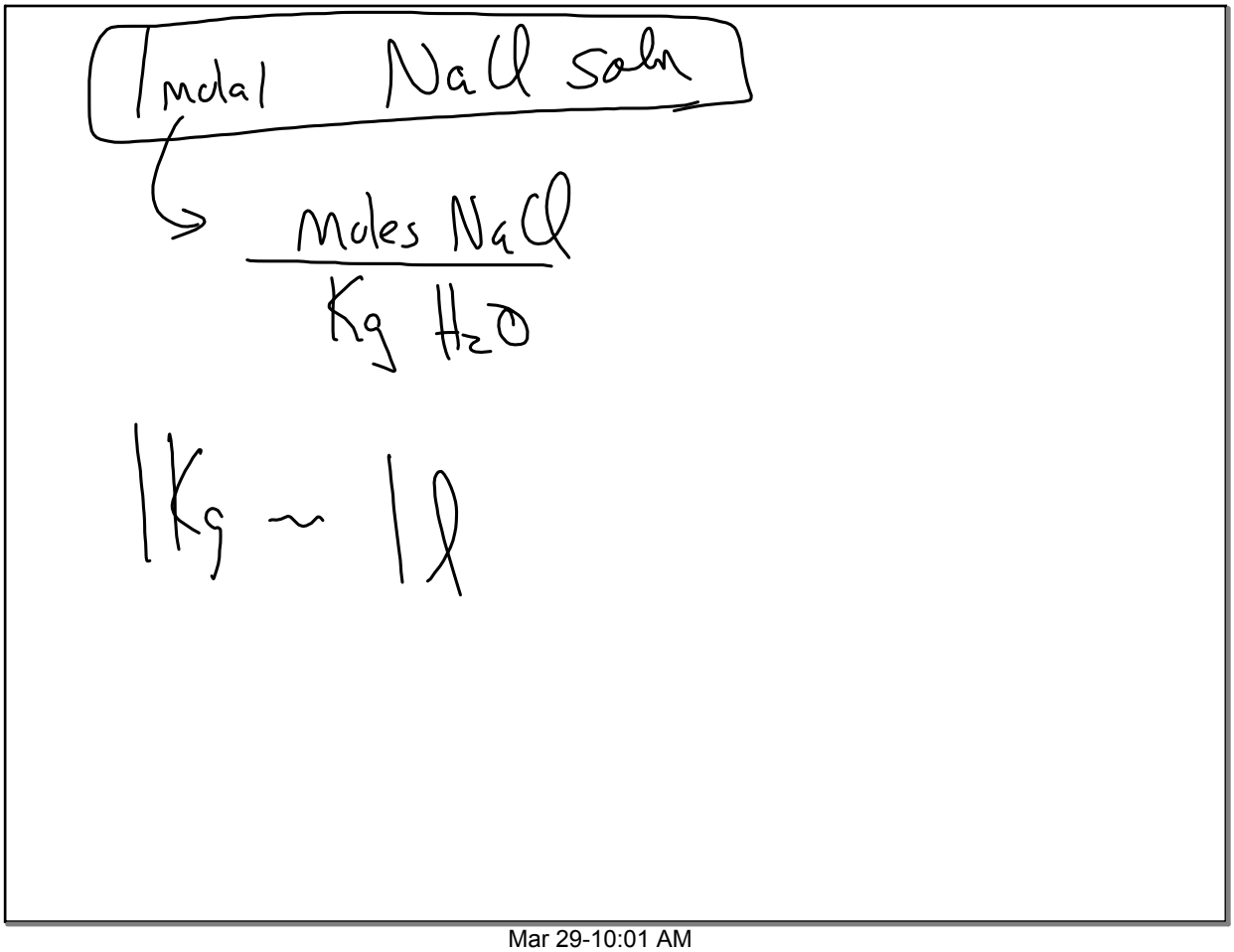

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