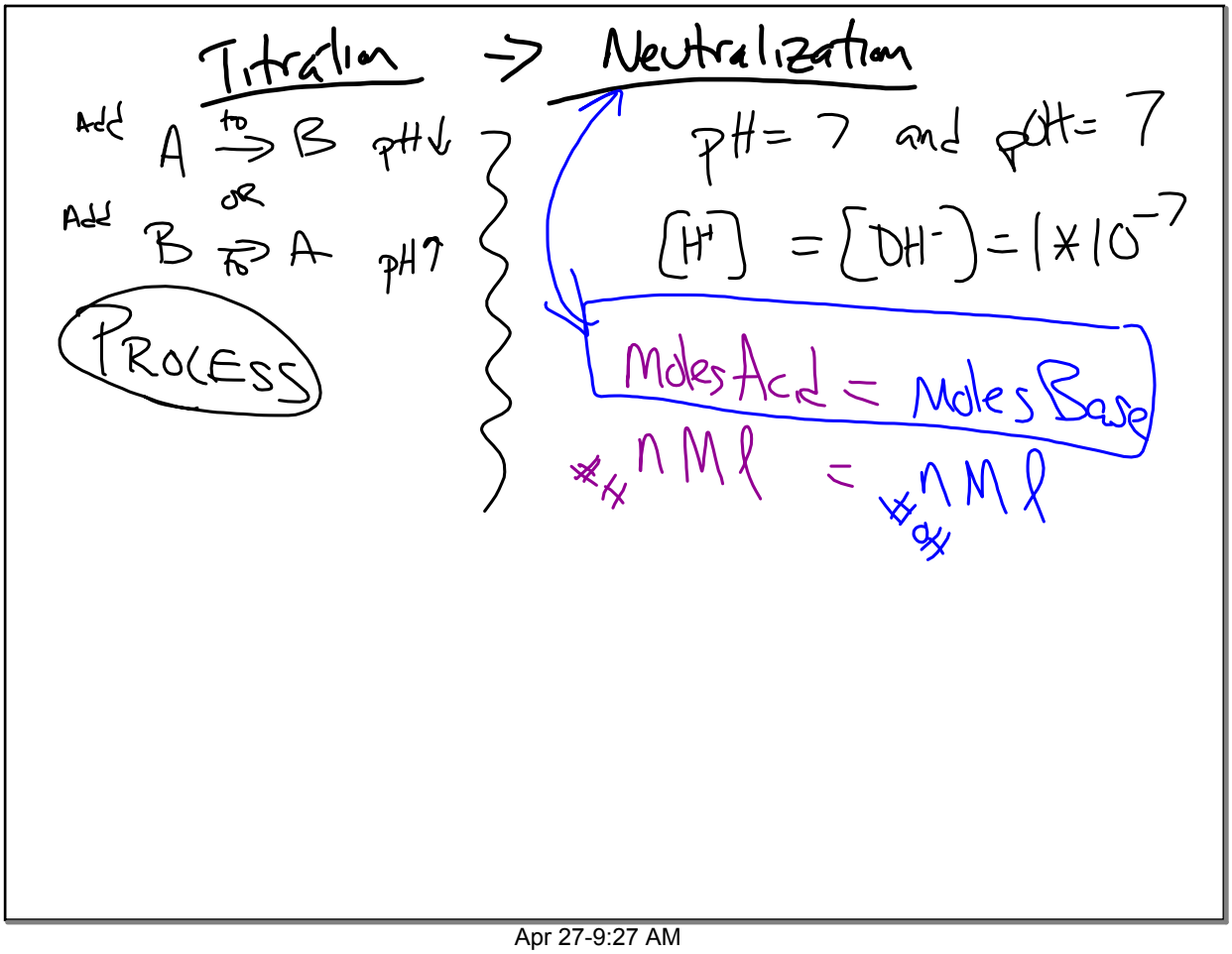
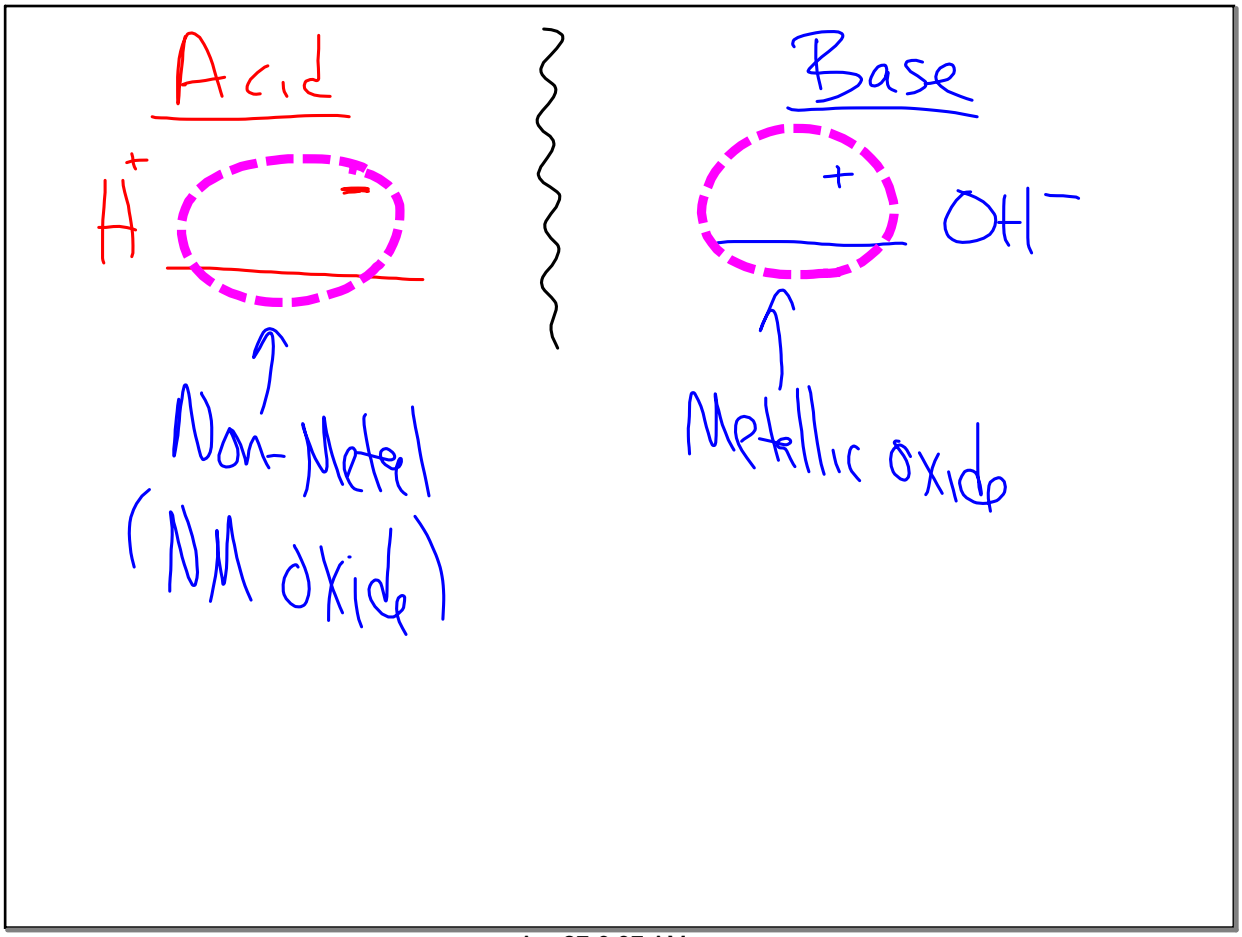


Apr 27-9:27 AM



Apr 27-9:37 AM

$H_2O + CO_2 \rightarrow H_2CO_3$ (CO_3^{-2})

$H_2O + SO_2 \rightarrow H_2SO_3$

$H_2O + NO_2 \rightarrow 2HNO_3$ (H^{+1}, NO_3^{-1})

Non-Metallic Oxides That produce an Acid

Apr 27-9:40 AM

$Na_2^{+1}O^{-2} + H_2O \rightarrow 2NaOH$

Metallic anhydride HOH

$CaO + HOH \Rightarrow Ca(OH)_2$

Metallic oxide + Water → Base

(+) (+) OH^-

Apr 27-9:51 AM

A B S

No H^+ ___

No ___ OH^-

'inorganic'

NO "C" as main element.

$NaCl$

CARBON

C_2H_5OH ⇒ Organic alcohol

Apr 27-9:55 AM

SA

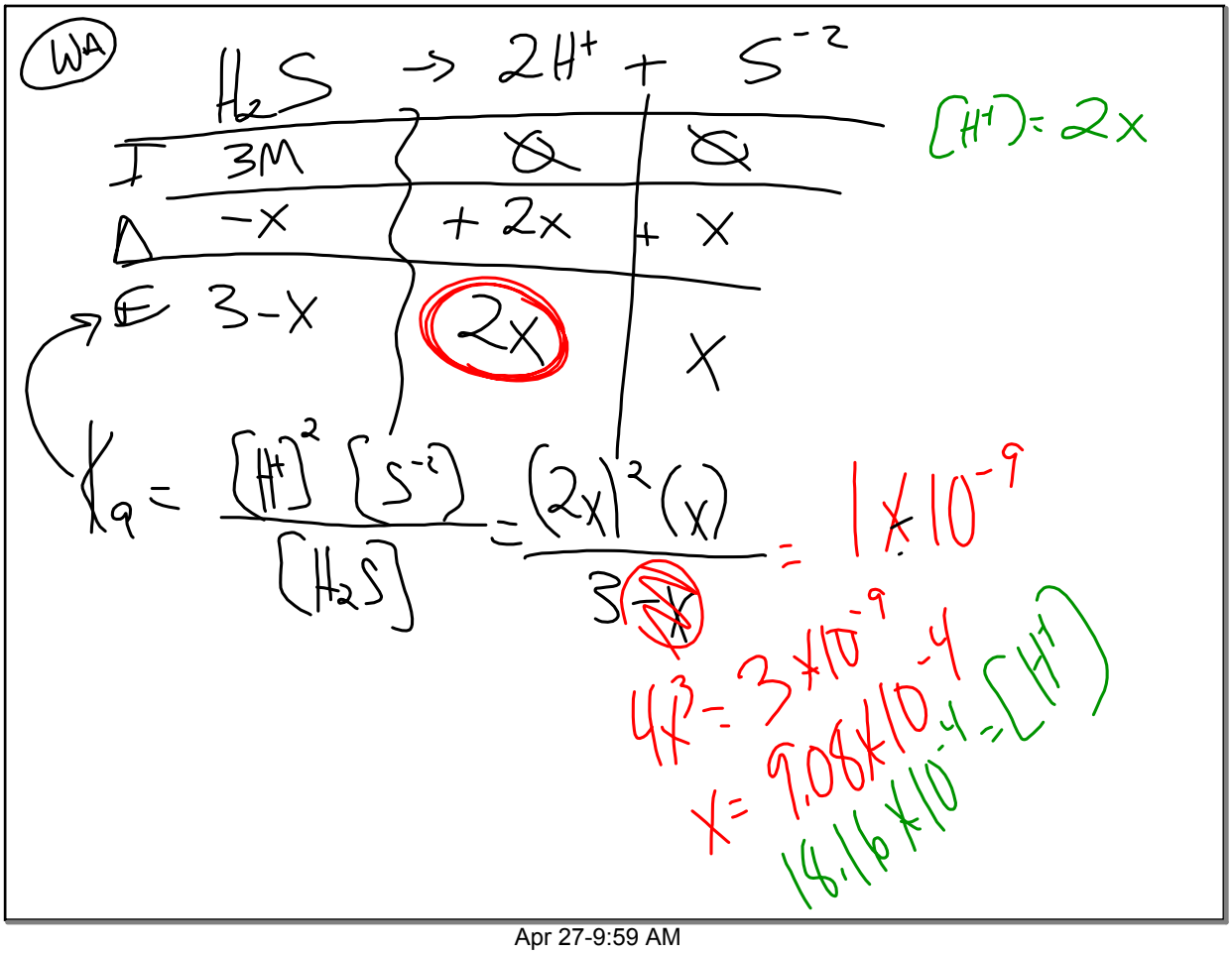
$H_2SO_4 \rightarrow 2H^+ + SO_4^{-2}$

3M 6M 3M

$pH = -\log[H^+]$

$1 * 10^{-6}$

Apr 27-9:58 AM


WA
$H_2S \rightarrow 2H^+ + S^{-2}$
I 3M 0 0
Δ −x +2x +x
E 3−x 2x x
$[H^+] = 2x$
$K_a = \frac{[H^+]^2[S^{-2}]}{[H_2S]} = \frac{(2x)^2(x)}{3-x} = 1 \times 10^{-9}$
$4x^3 = 3 \times 10^{-9}$
$x = 9.08 \times 10^{-4}$
$18.16 \times 10^{-4} = [H^+]$

Apr 27-9:59 AM