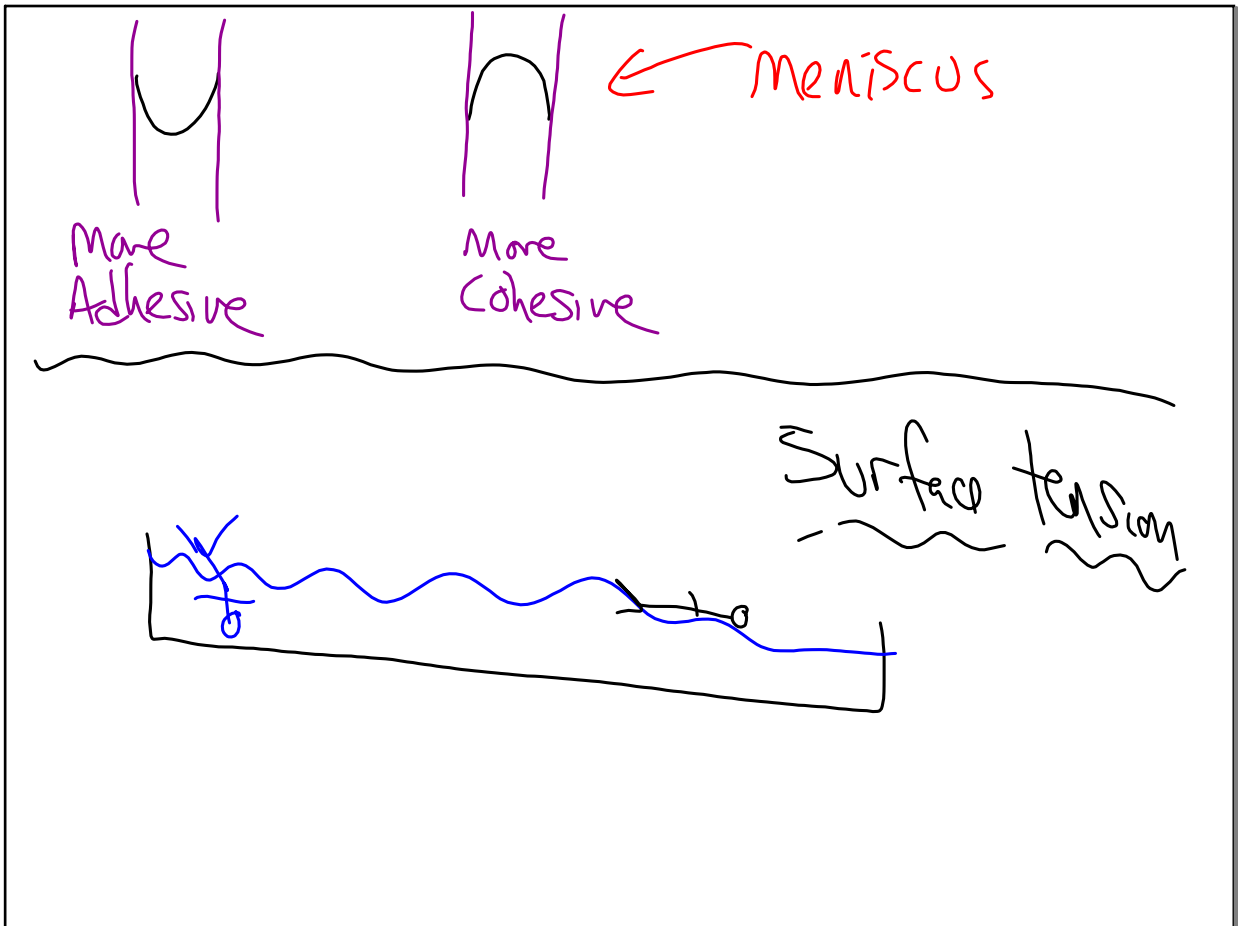
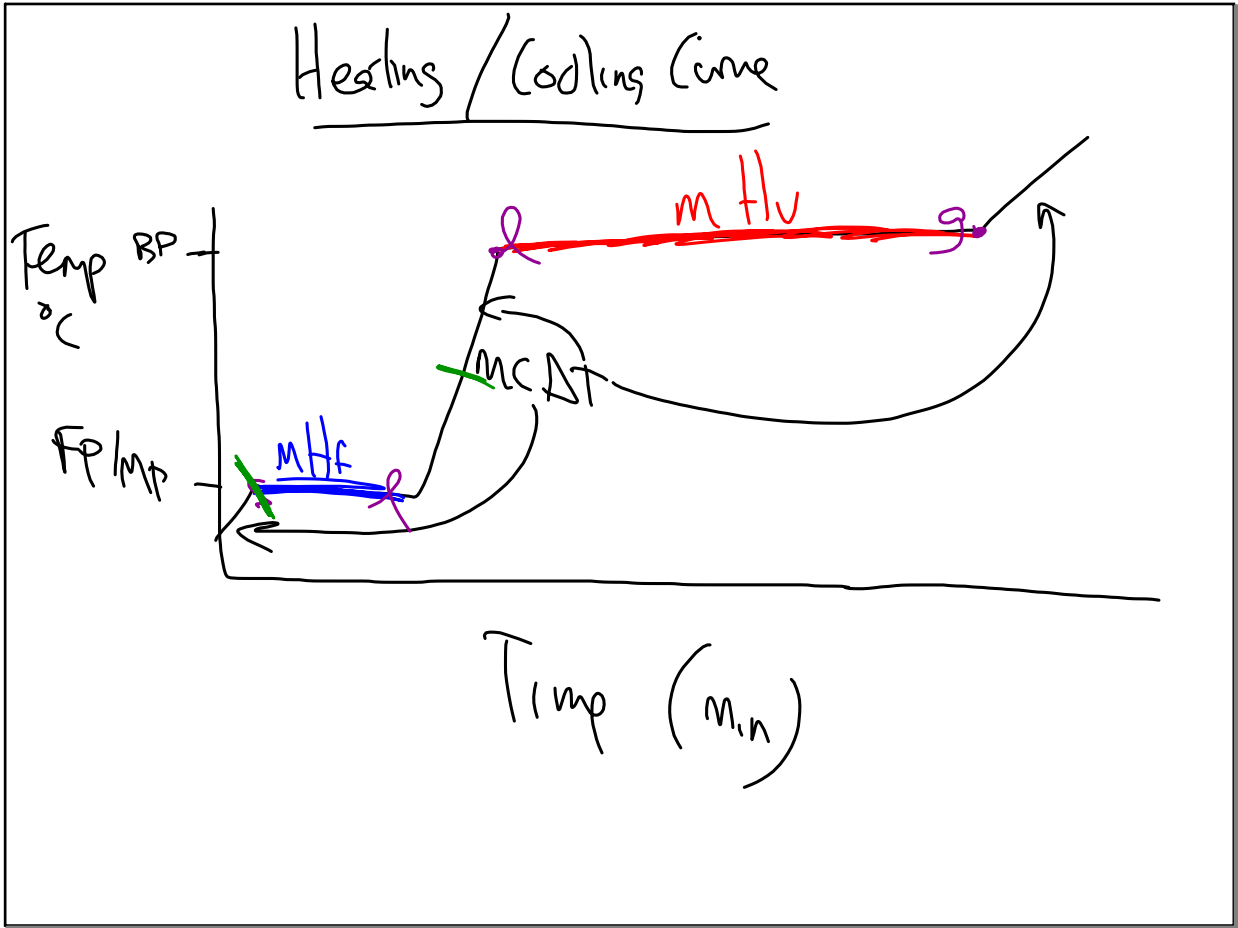


Jan 3-7:53 AM



Jan 3-8:20 AM



Jan 3-8:23 AM

R-chem Specific Heat Lab

$T_f = 23^\circ\text{C}$

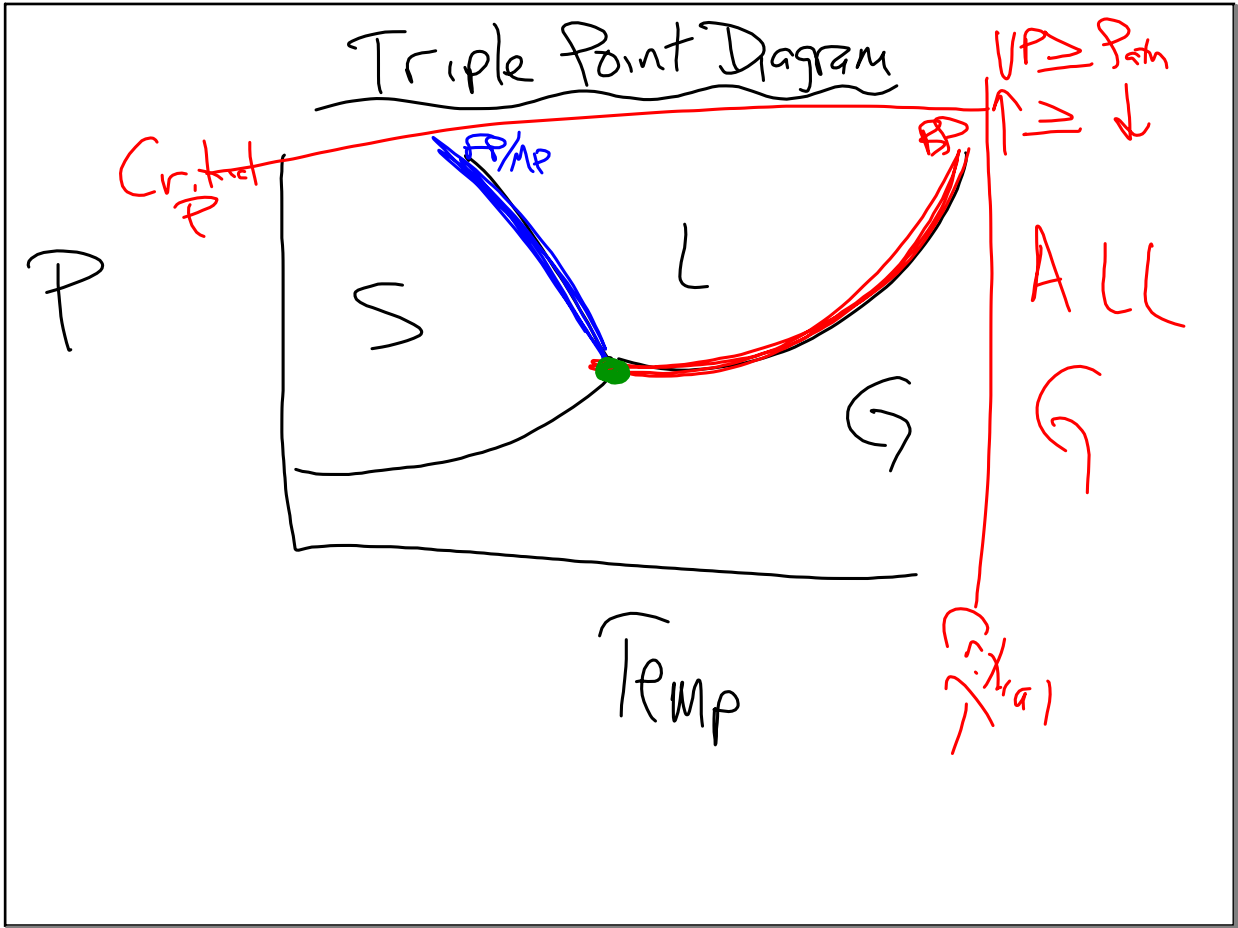
Heat Energy →

Hot METALS (Boiling water to heat) 100°C
metal
 $MC\Delta T$

↔ Room temp (top of water)
Styrofoam $\sim 22^\circ\text{C}$
H₂O
 $MC\Delta T$

$MC (B_{\text{ign}} - S_{\text{melt}}) = M (23 - 22)$
 $100 - 23$

Jan 3-8:37 AM



Jan 3-8:45 AM

The figure contains handwritten notes for "Chap 13 - SOLUTIONS CHEM.". The main title "Solute in a Solvent" is enclosed in a rectangular box. An arrow points from the box to the text "Gets dissolved" on the left and "does the dissolving." on the right. Below the box, there is a purple note: "Bort - 'Like Dissolve, like'" followed by "Bonding types." on the next line.

Jan 3-8:55 AM

Solution Examples

① $KCl(aq) = \underset{\substack{\text{Ionic} \\ \text{(Ion-Dipole)}}}{KCl(s)} + \underset{\text{POLAR}}{H_2O(l)}$
Solute Solvent

② Arnold Palmer: $Ice-T(l) + \text{Lemon Aid}(l)$
 Solvent = greater quantity.

③ Air - gas in a gas.
 (g) in (g)

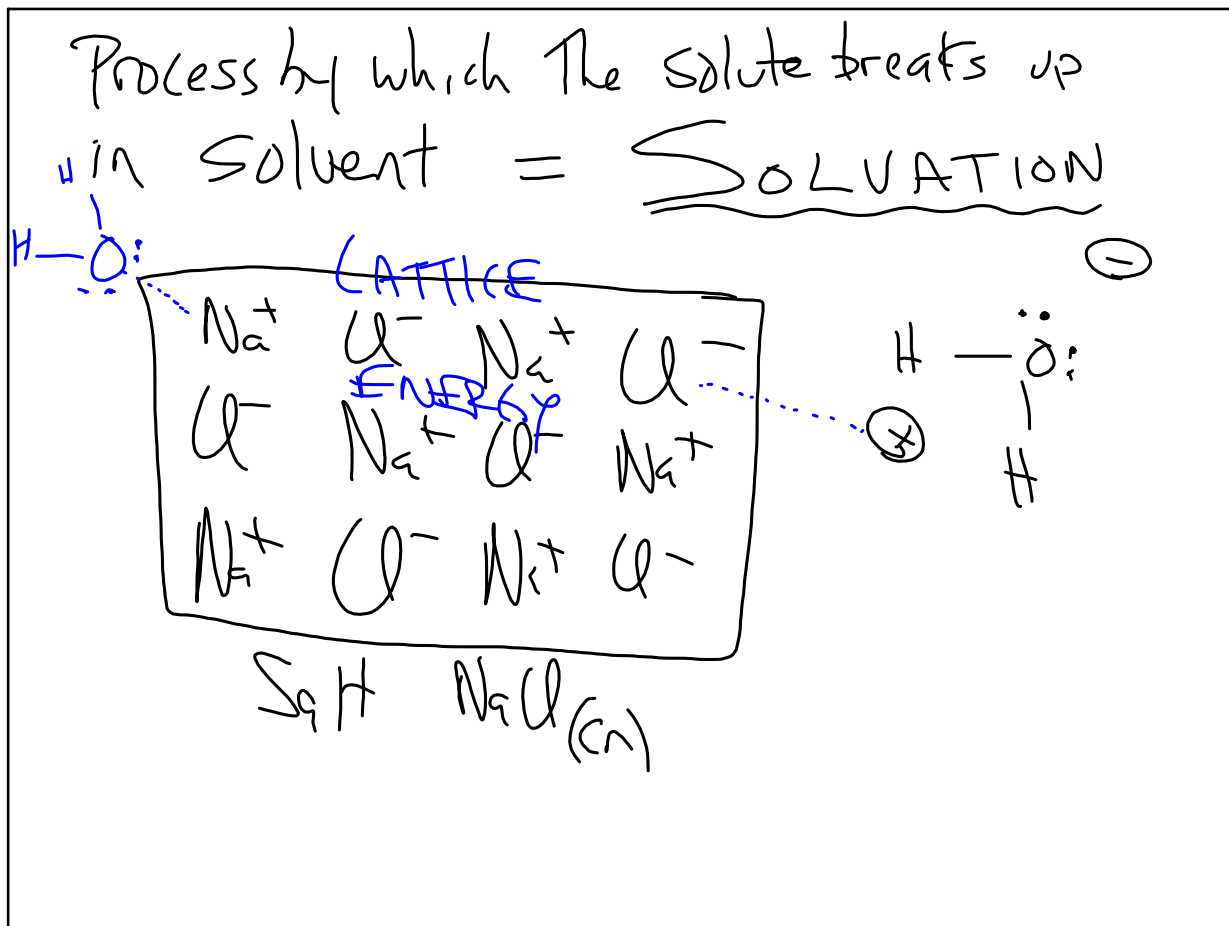
Jan 3-9:02 AM

④ Sparkling water
 AKA. Seltzer
 "Soda water" = $(O_2(g))$ in $H_2O(l)$

⑤ Clouds Water droplets (l) in a gas (g)

⑥ Brass (Cu(s) and Zn(s))
 Alloy → solid in solid.

Jan 3-9:08 AM



Jan 3-9:14 AM