⑤ H_3PO_4, 175mL, 3.5M — g

0.175L | 3.5 mole H_3PO_4 / L

M

$$\frac{3.5\ \text{mole}\ H_3PO_4}{1\ \text{L}} \times 0.175\ \text{L} \times \frac{98\ \text{g}\ H_3PO_4}{1\ \text{mole}\ H_3PO_4}$$

Oct 5-7:40 AM

③ % = $\frac{\text{Part}}{\text{Whole}} \times 100$

$$\%\ NaCl = \frac{50}{50+500} = \frac{50}{550}$$

⑦

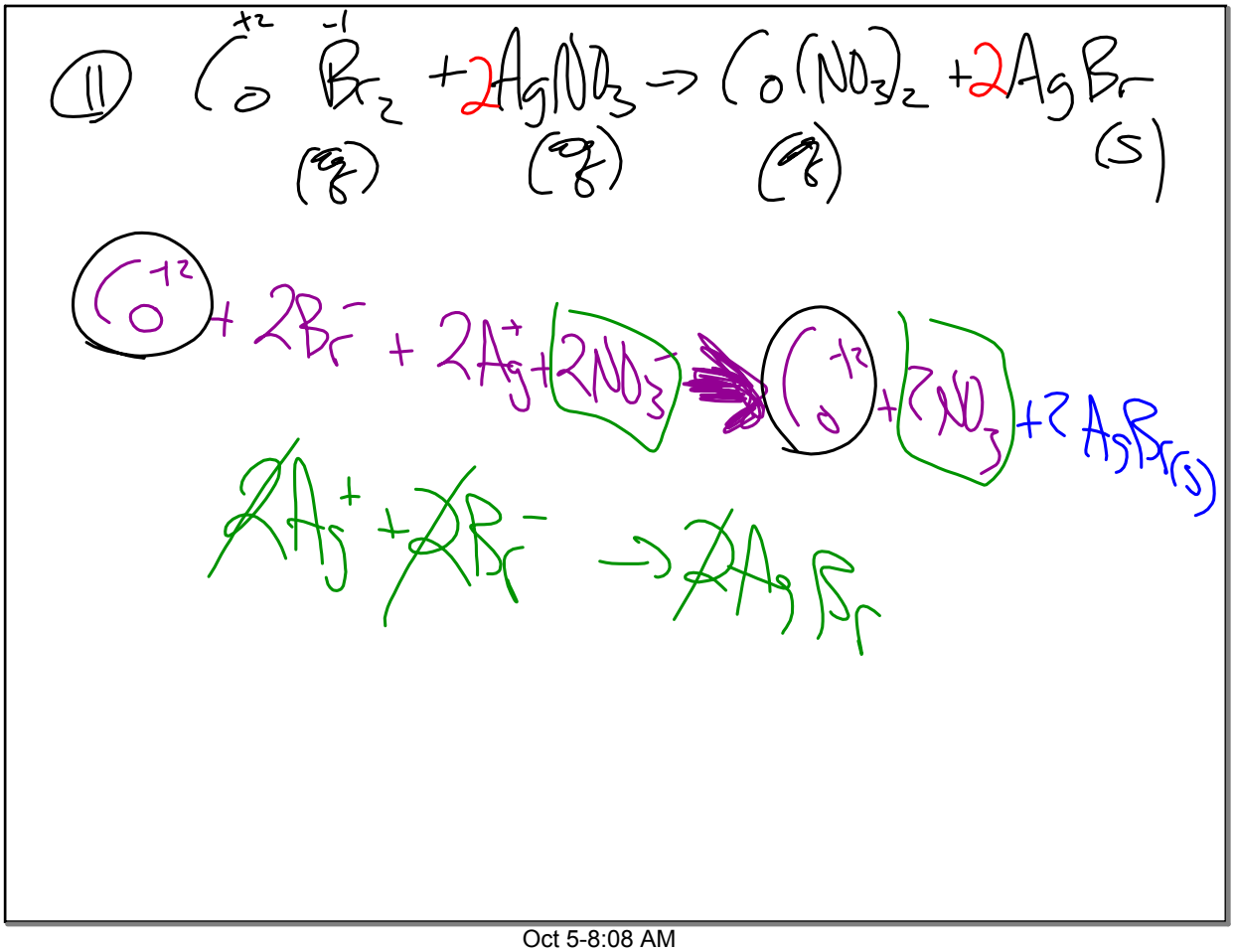
1 Na_3PO_4 — 3 Mole Na^+ — 0.2M

$[Na^+] = 0.6M$

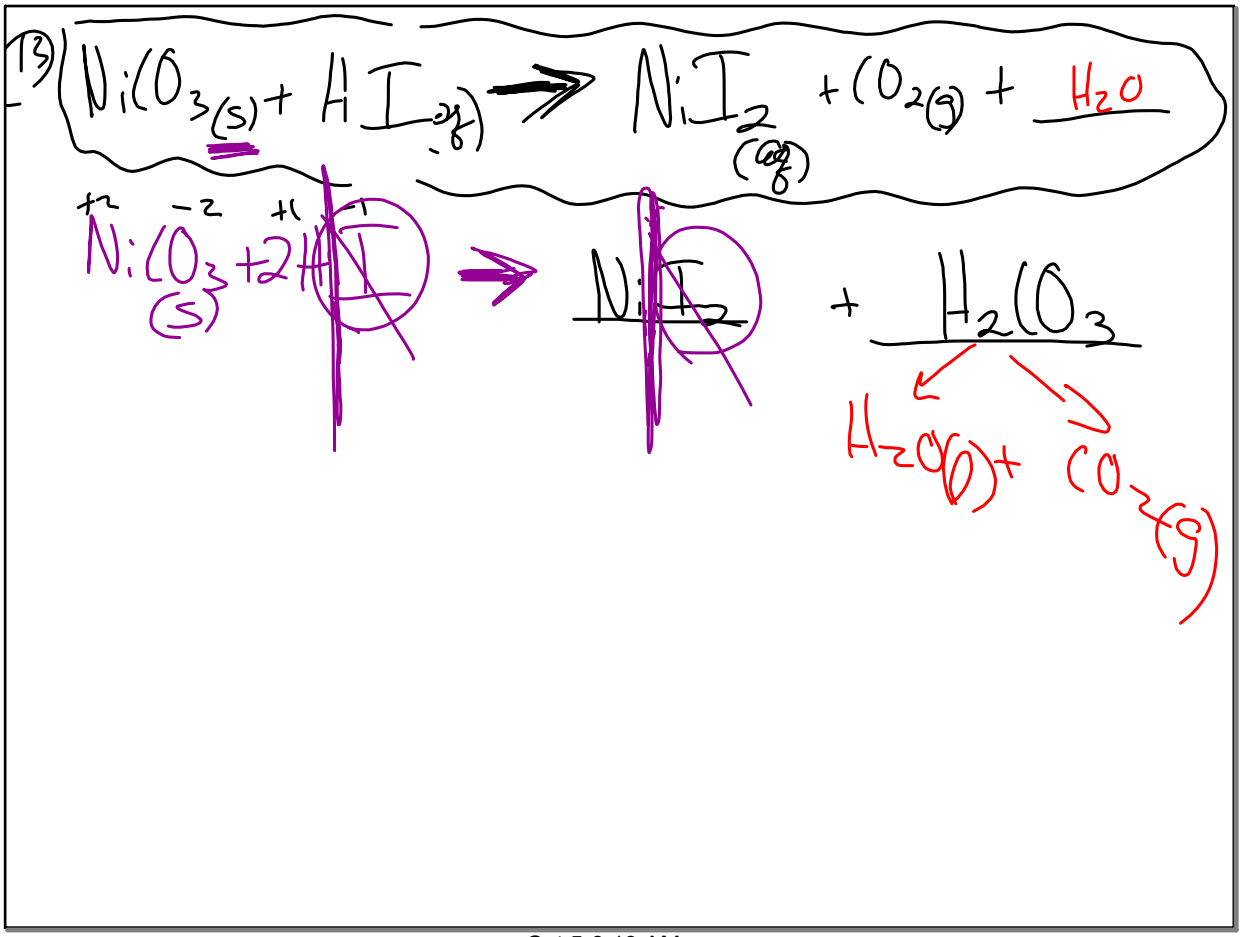
1 Mole Na_3PO_4 : $\frac{3\ \text{Mole}\ Na^+}{0.6M}$: $\frac{1\ \text{Mole}\ PO_4^{-3}}{0.2M}$

0.2M

Oct 5-8:01 AM



Oct 5-8:08 AM



Oct 5-8:13 AM

(21) moles A = moles B

$n_A * M_A * l_A = n_B * M_B * l_B$

(1) M (17.5 mL) = (1) (0.25) (29.6 mL)

Oct 5-8:22 AM

(23) 200mL + 500mL 0.35M HCl

moles start = moles end

M * l

(0.35)(200) = (700)

Oct 5-8:26 AM

PS 4-1 even #'s

Oct 5-8:28 AM