

① $N_2 + 3H_2 \rightarrow 2NH_3$
 1g N_2 + 3g H_2 - 0.214g used =
 2.786g left over

1g N_2	1 mole N_2	3 mole H_2	2g H_2	= 0.214g H_2 used
28g N_2	1 mole N_2	1 mole H_2	1 mole H_2	

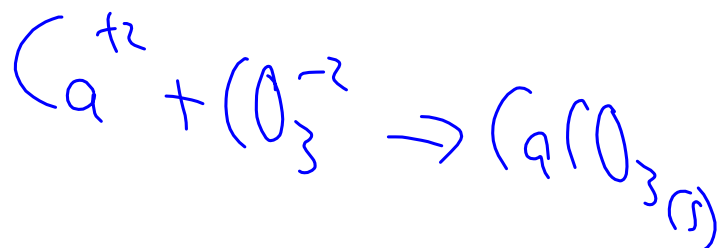
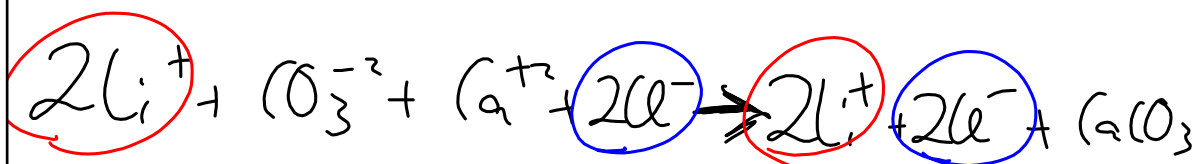
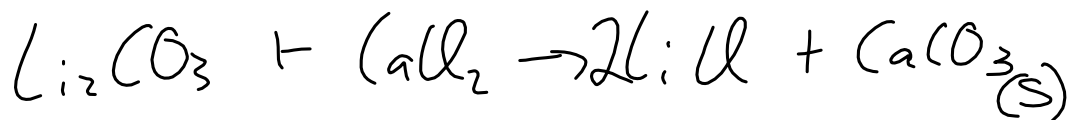
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②

25g Al_2S_3	1 mole Al_2S_3	3 mole S	6×10^{23} atoms S
150g Al_2S_3	1 mole Al_2S_3 (3)	1 mole S	element

↑
Compound
↓
Molecule

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②⑥ moles ^{Acid} A = moles ^{Base} B

$$\underline{n} \text{ M L} = \underline{n} \text{ M L}$$

$$(2)(0.737)(35 \text{ ml}) = (1)(0.827) \text{ ______ ml}$$

②⑨ Mole start = mole end

$$M \times l = M \times l$$

$$(0.75 \text{ M})(85.6 \text{ ml}) = M(6200 \text{ ml})$$

Oct 7-8:02 AM