Chap 5 Thermochemistry

movement heat

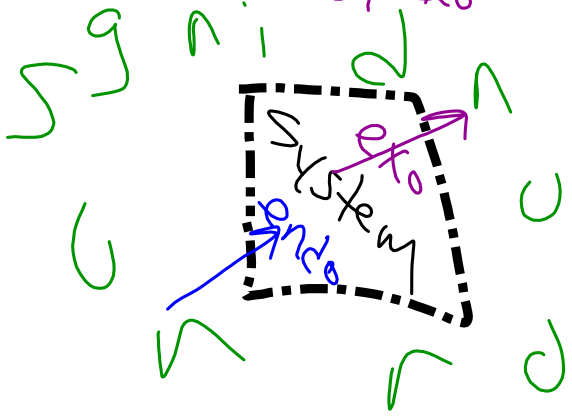
ENDOTHERMIC - heat enters (absorbed)
→ gets hotter

EXOTHERMIC → heat exits (released)
→ gets colder

Oct 14-7:40 AM

System vs Surroundings

Point of view for Endo/Exo



Oct 14-7:59 AM

hot vs. cold

molecular motion

Exo - Spontaneous $\ominus\Delta H$

Endo - requires work on the system. $\oplus\Delta H$

Oct 14-8:03 AM

Heat → Joules (J) $\frac{Kg*m^2}{sec^2}$

↳ ENERGY

$KE = \frac{1}{2} m v^2$

$J = Kg \left(\frac{M}{Sec}\right)^2$

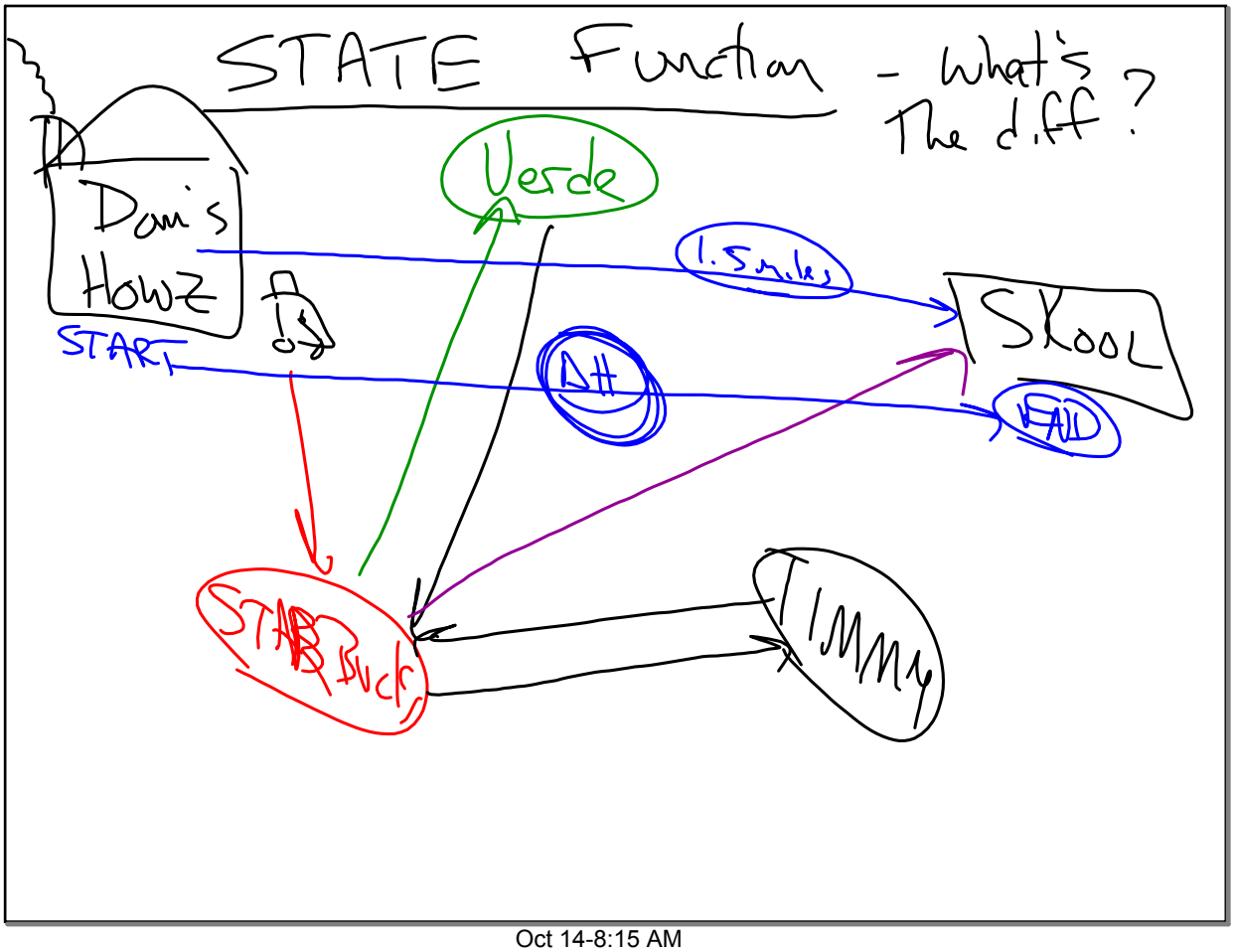
$J = \frac{Kg*m^2}{sec^2}$

$PE = mgh$

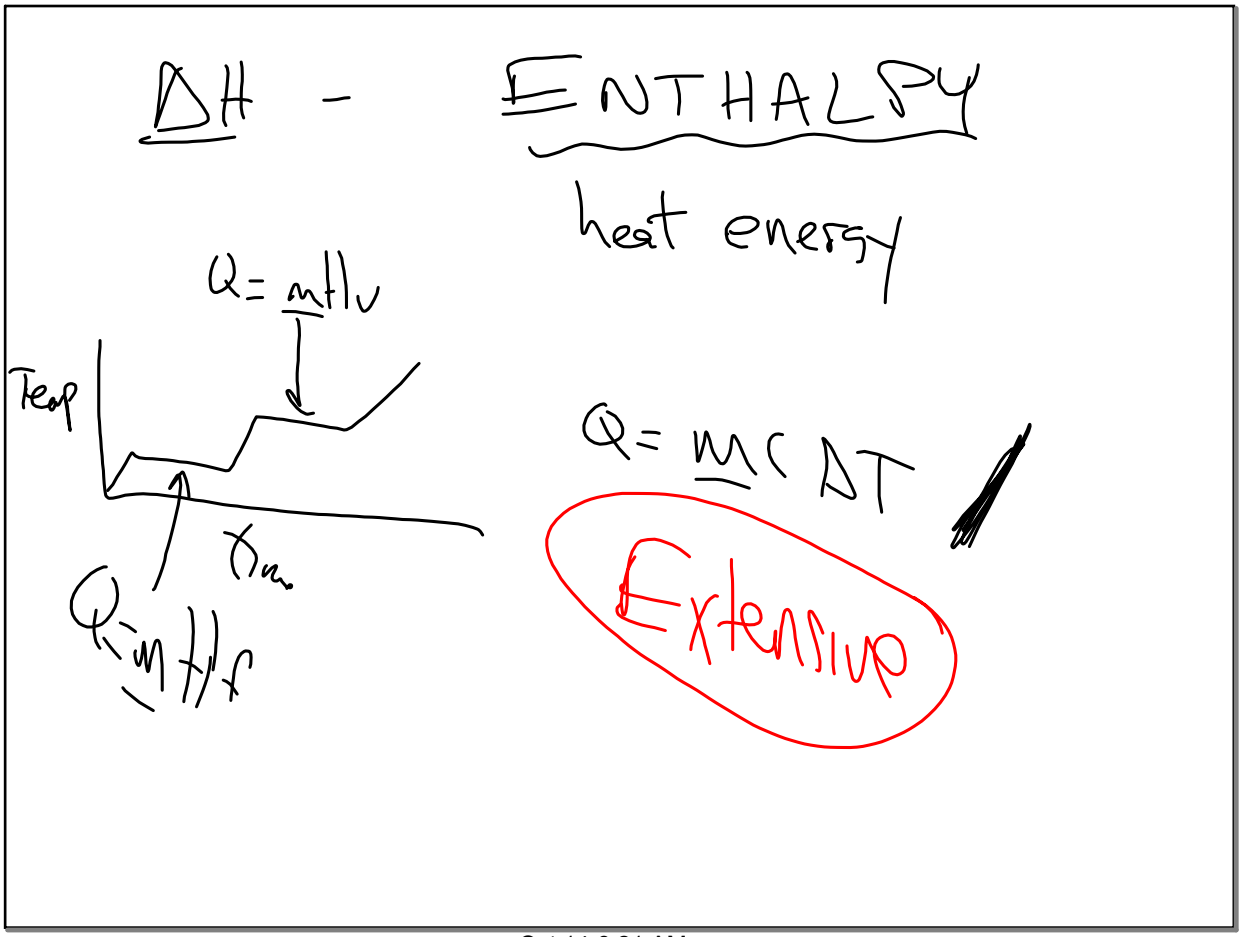
$PE = \frac{Kg}{1} * \frac{M}{Sec^2} * \frac{M}{1}$

$PE = \frac{Kg*m^2}{Sec^2}$

Oct 14-8:10 AM

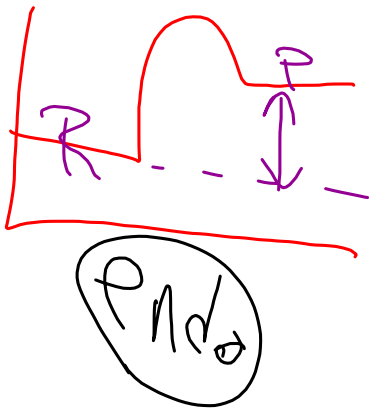


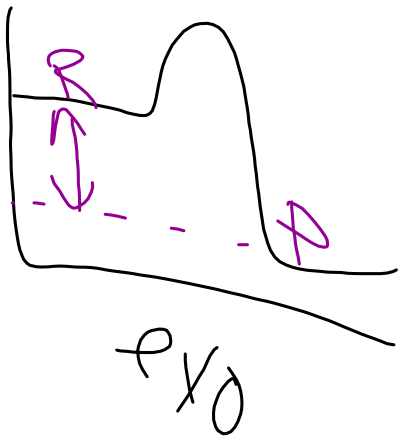
Oct 14-8:15 AM



Oct 14-8:21 AM

$\Delta H = H_{products} - H_{reactants}$





Oct 14-8:23 AM

HW p201 ish

5 / 4, 12, 14, 20, 26

Oct 14-8:27 AM