

5/54a

Soln

$\text{NH}_4\text{NO}_3(s) \rightarrow \text{NH}_4^+ + \text{NO}_3^-$

3.88g
60g H₂O

4.18 J	63.88g soln	4.6°C	80%
9.°C			$\frac{3.88\text{g}}{\text{Molar Mass}} \text{ mole}$

$\Delta T = 23^\circ - 18.4^\circ$

$Q = mC\Delta T$

$C = \frac{4.18 \text{ J}}{9.^\circ\text{C}}$

25325.46 J/mole

$+25.3 \text{ kJ/mole}$

~~kJ/mole NH₄NO₃~~

Oct 18-7:38 AM

5.43

$\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \quad \Delta H = -65.5 \text{ kJ}$

$1 \text{ Ag}^+ + 1 \text{ Cl}^- \rightarrow 1 \text{ AgCl} + 65.5 \text{ kJ}$

Reactant
Go into
ENDO

Product → Produced
Goes out - leaves
EXO

① $\frac{0.2 \text{ mol AgCl}}{1 \text{ mol AgCl}} \times 65.5 \text{ kJ} = 13.1 \text{ kJ}$

② $\frac{2.5 \text{ g AgCl}}{143 \text{ g AgCl}} \times 65.5 \text{ kJ} = 1.15 \text{ kJ}$

How much of a mole.

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②

21.4g	1 kg	100 ³ cm ³
cm³	1000g	1 ³ m ³

+
NaHCO₃⁻

+2 Mg ₃	-3 N ₂
Ca	CO

Oct 18-8:19 AM

S / 44 + SS

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