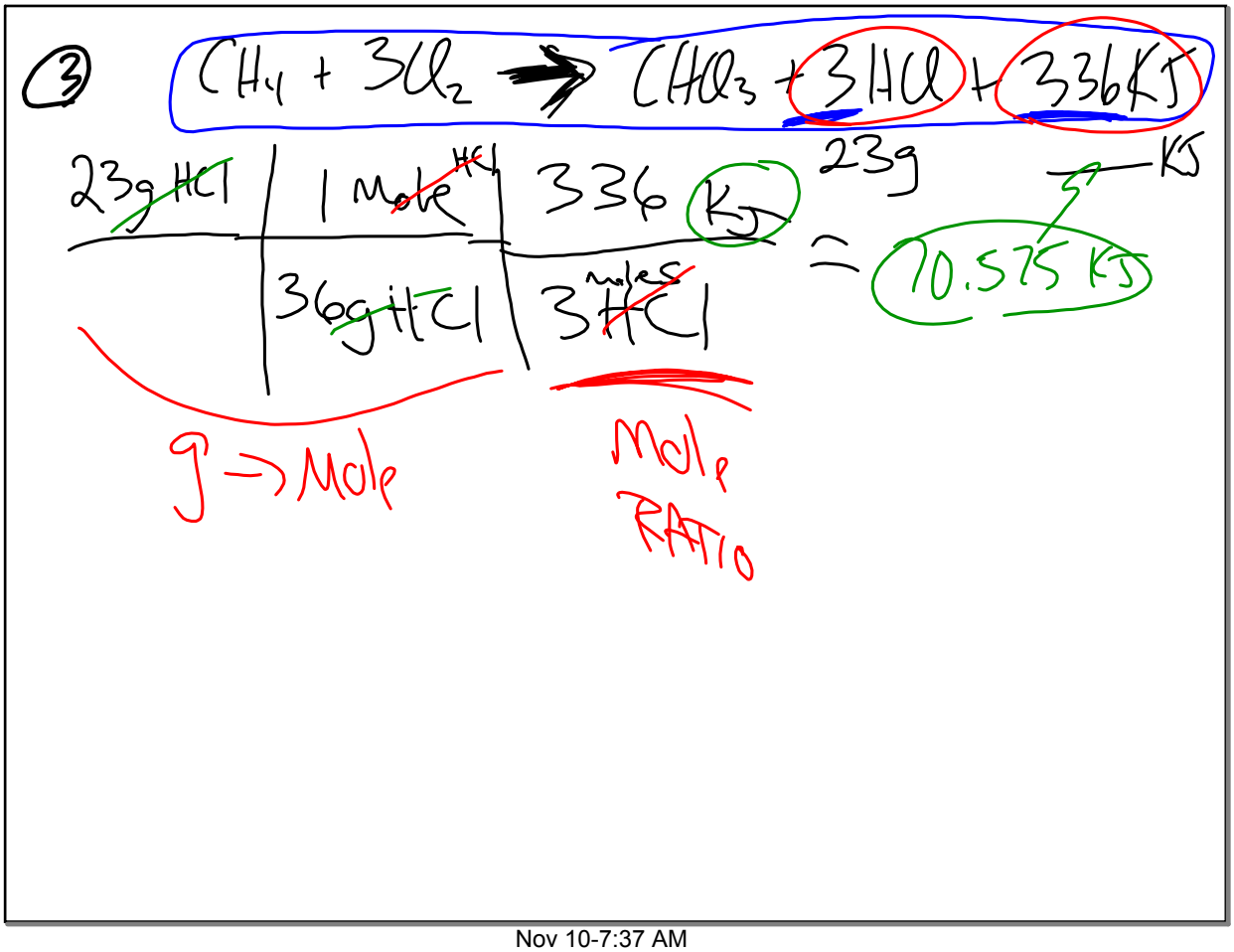
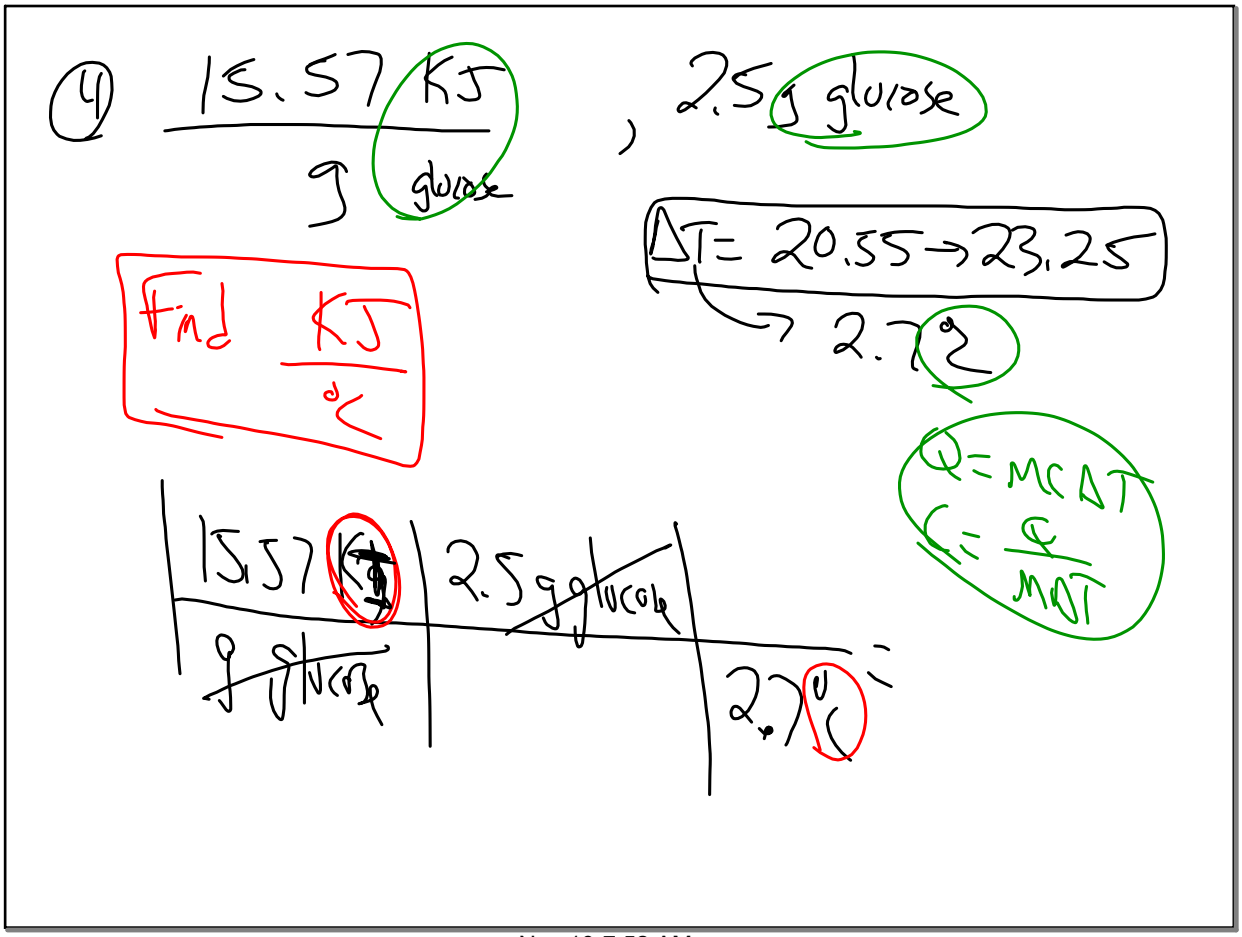
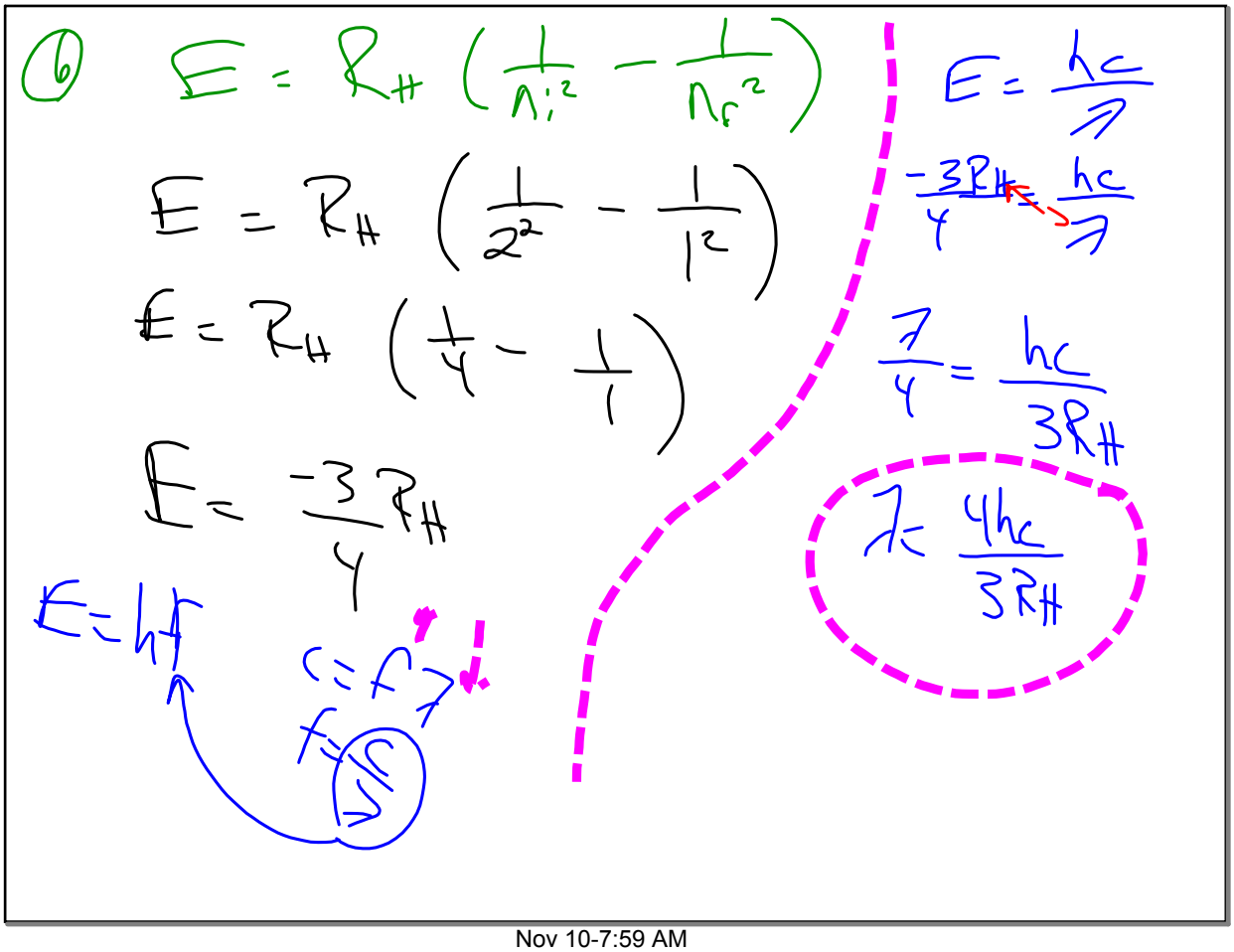


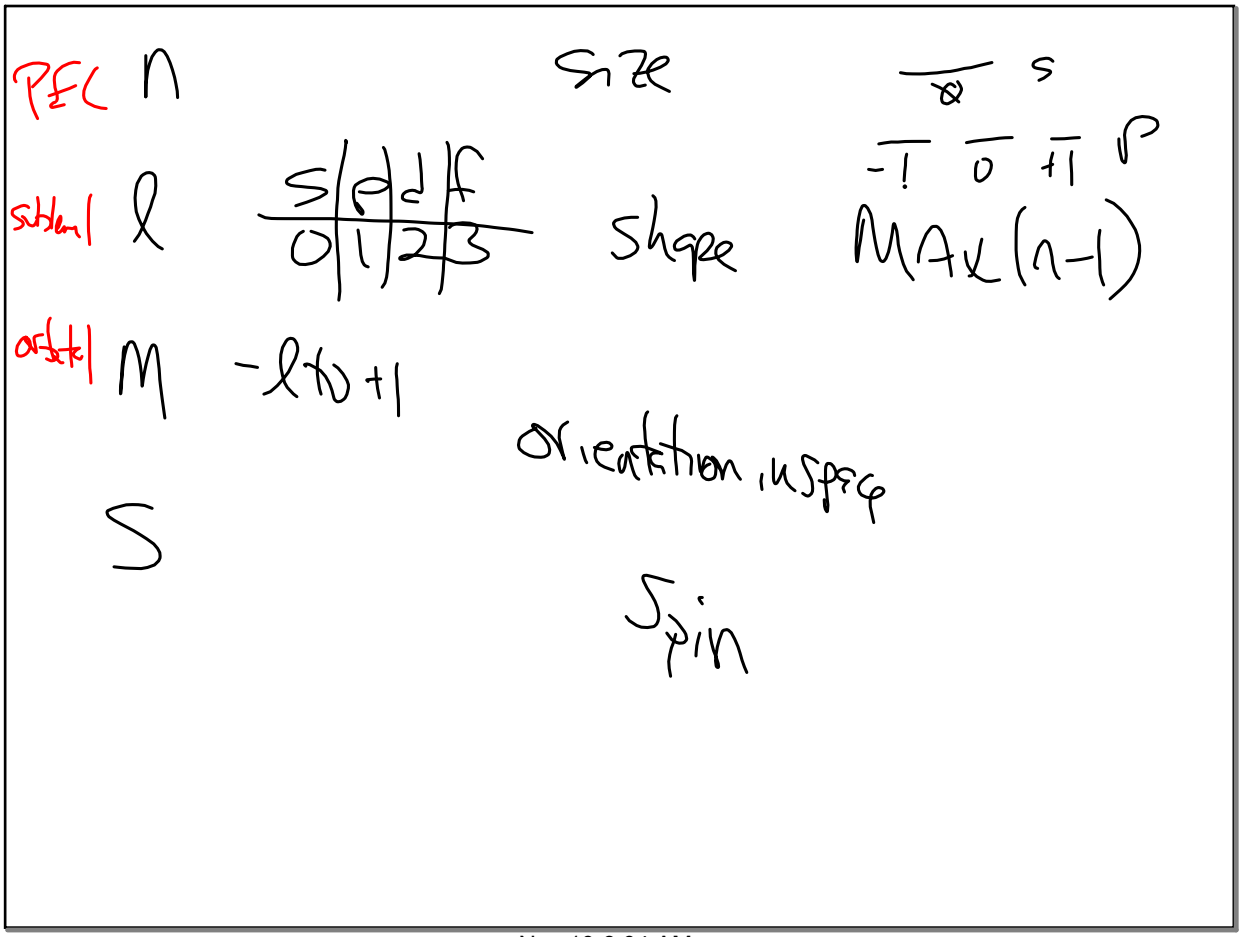
Nov 10-7:37 AM



Nov 10-7:52 AM



Nov 10-7:59 AM



Nov 10-8:04 AM

Co [Ar] $4s^2 3d^7$

Nov 10-8:09 AM

Fe [Ar] ~~$4s^2$~~ $3d^6$

Fe^{+3} [Ar] $3d^5$

$\rightarrow 1s^2 2s^2 2p^6 3s^2 3p^6 3d^5$

Nov 10-8:22 AM

15.3g $NaNO_3$, 3.44g , 1070J / 0L

1 mole $NaNO_3$ (1 mole / 85g)

Find ΔH KJ / 1 mole $NaNO_3$

1.070 KJ	3.44g	85g $NaNO_3$
	15.3g $NaNO_3$	1 mole $NaNO_3$ =

Nov 10-8:23 AM

Room Temp H_2O 20°C

heat gained = heat lost

hot metal 100°C

$MC\Delta T = MC\Delta T$

$MC(T_f - 20) = MC(100 - T_f)$

Nov 10-8:26 AM