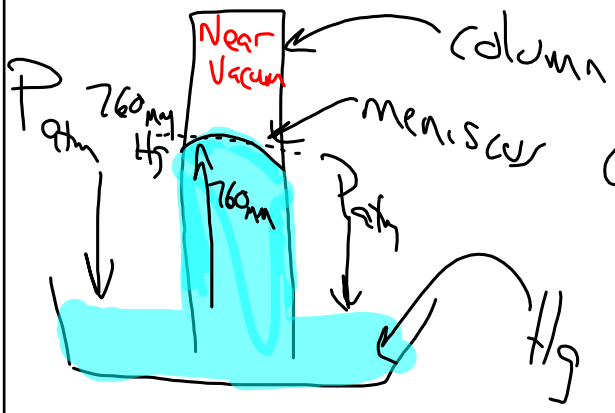
# Barometer - Torricelli (Torr) (mmHg)

↳ Measure ATMOSPHERIC Pressure




Cohesive > Adhesive

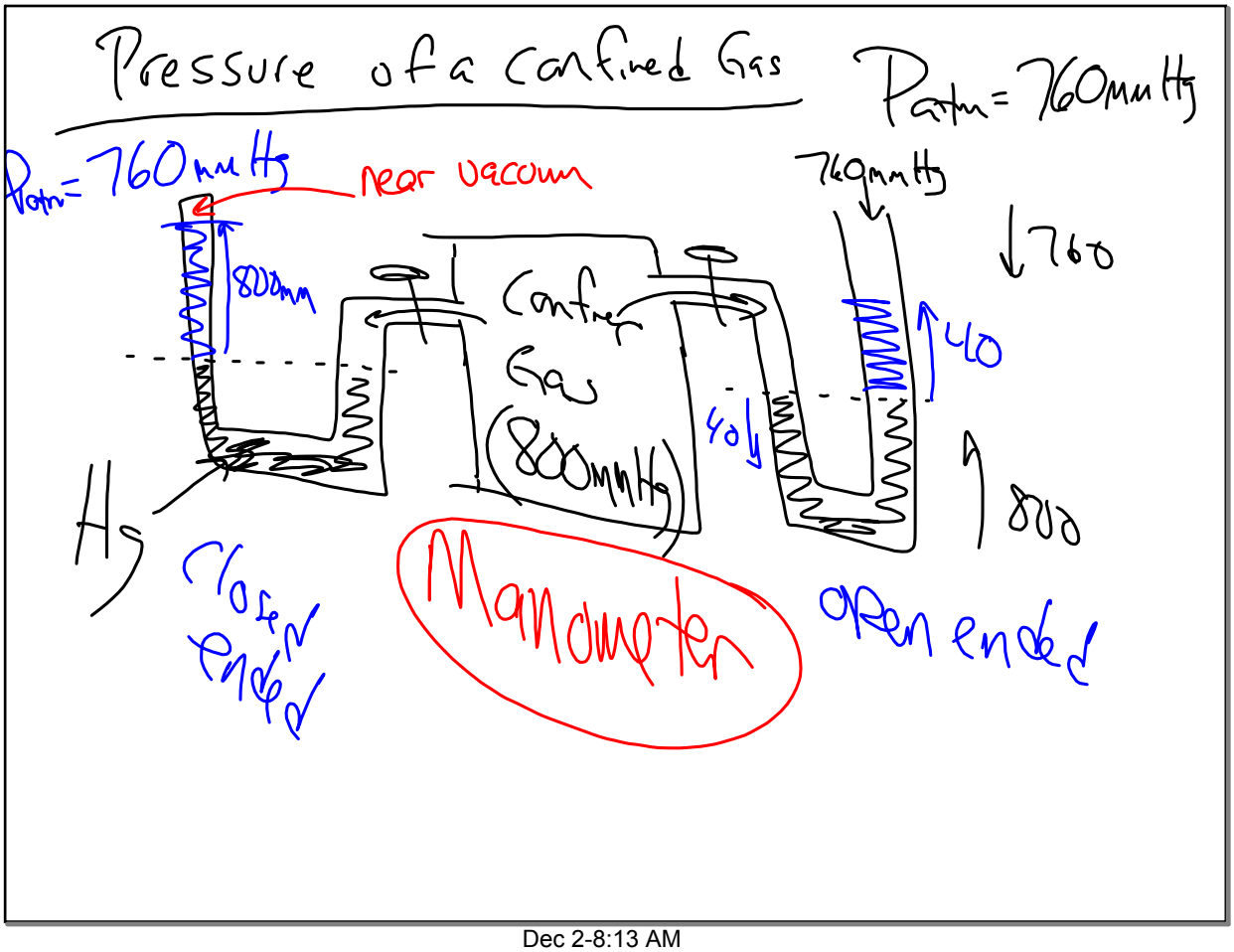
Dec 2-7:44 AM

## Conversion

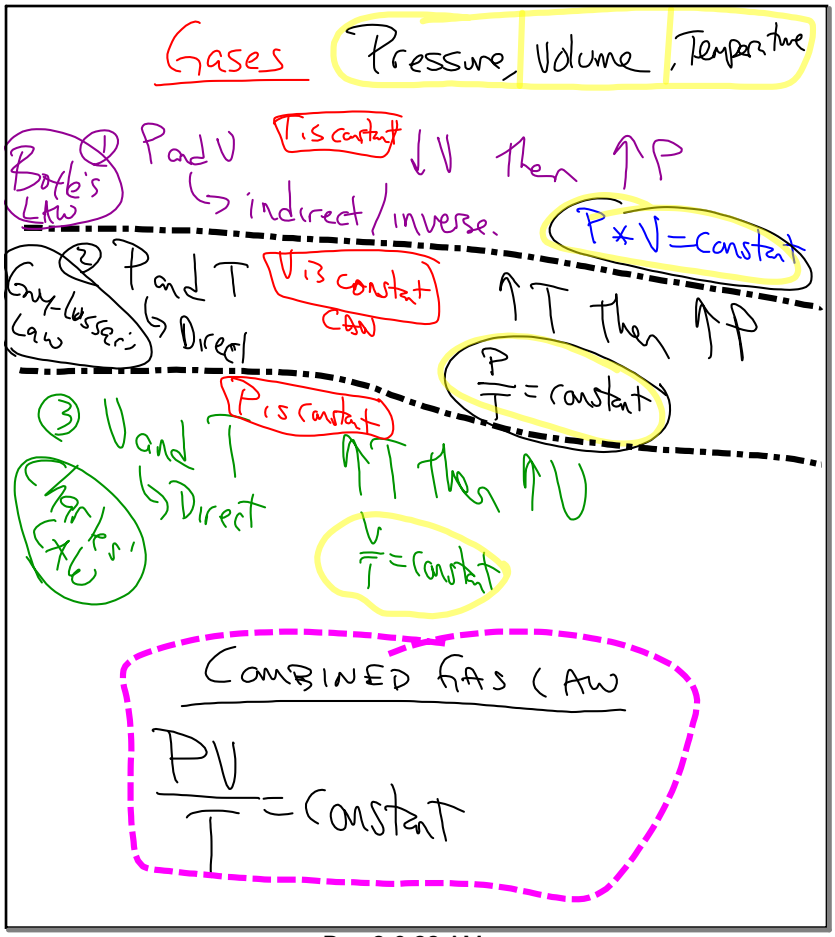
## Std. Pressure

760 Torr
760 mmHg = 1 atm = 101.35 KPa

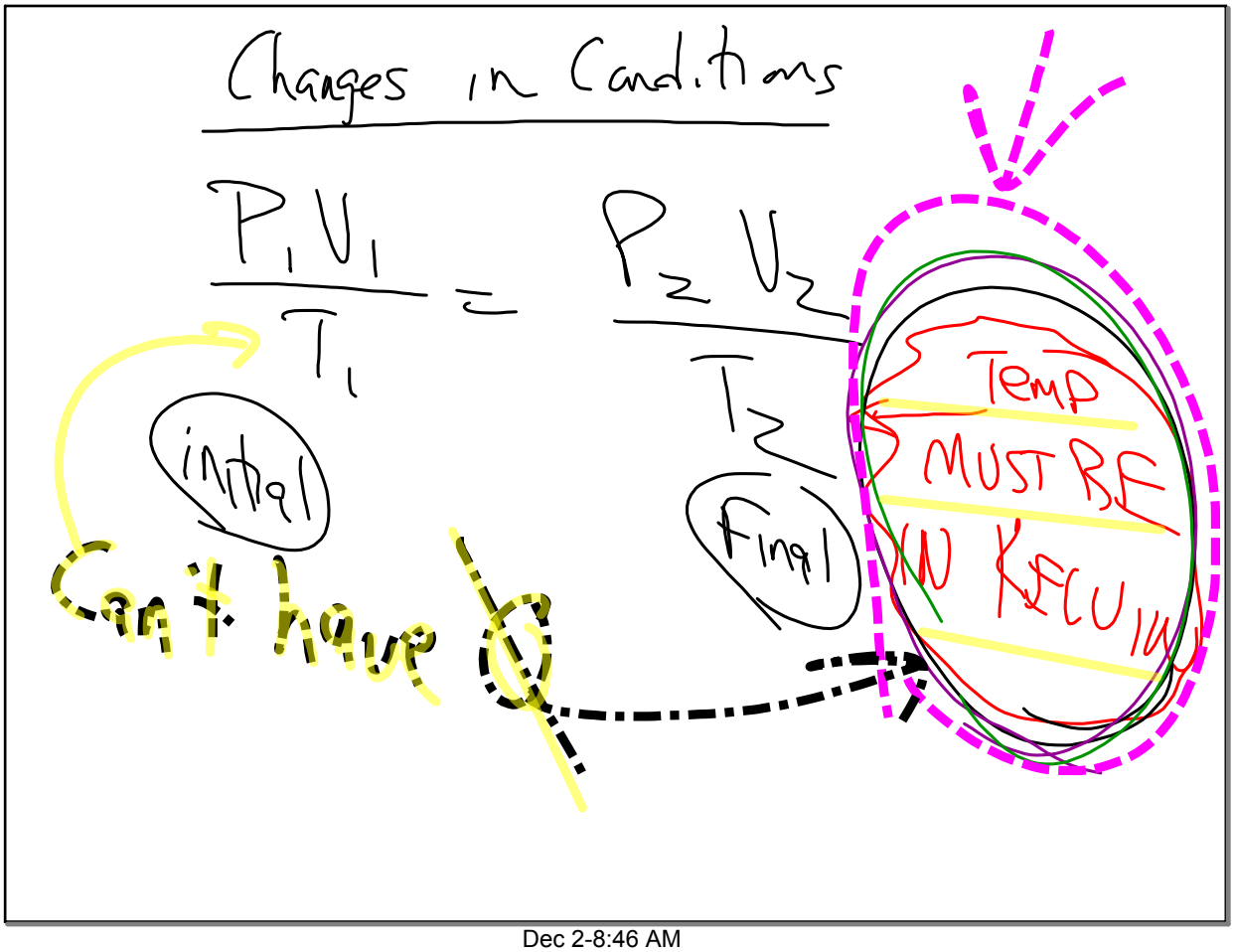
Dec 2-8:09 AM



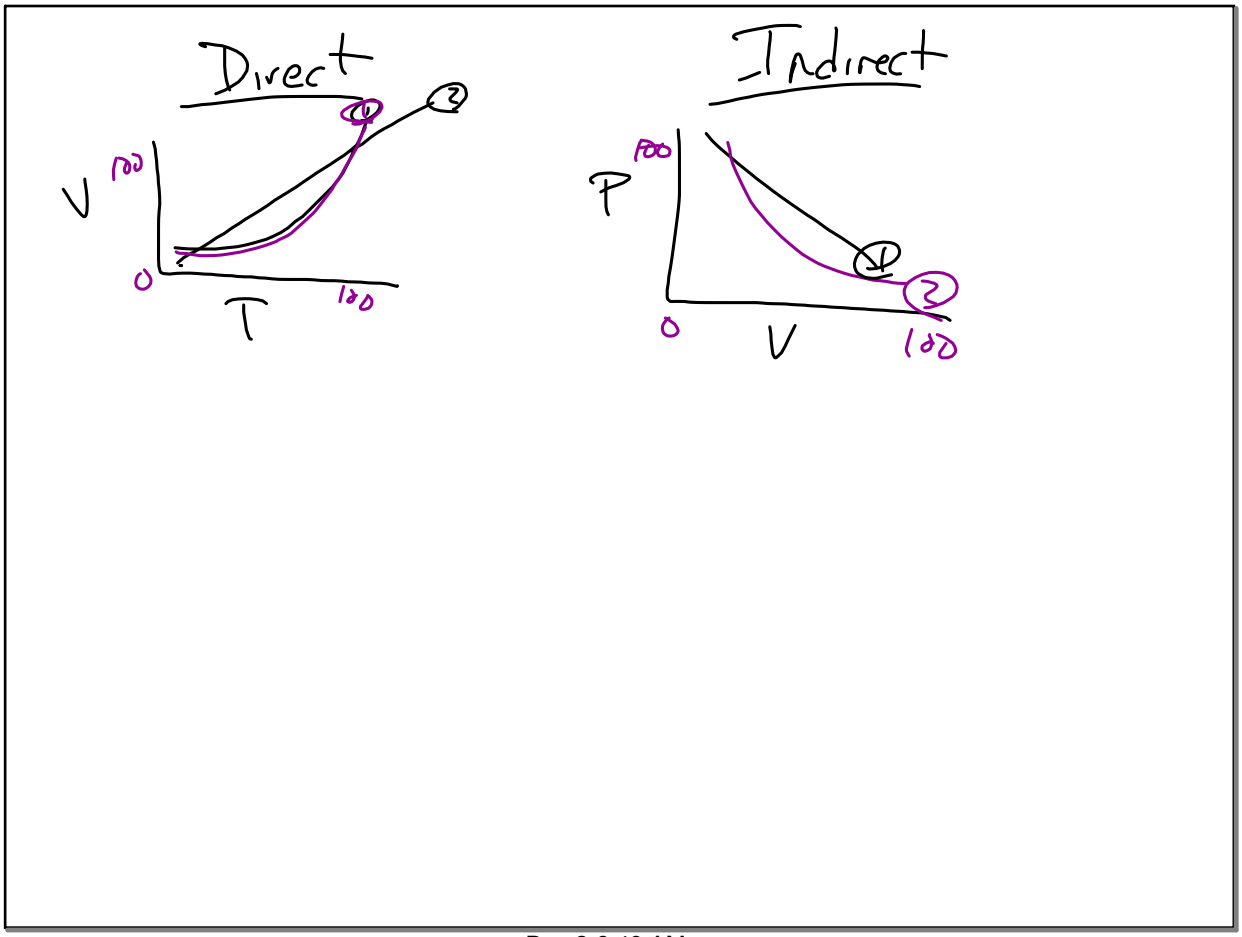

Dec 2-8:13 AM




Dec 2-8:22 AM




Dec 2-8:46 AM




Dec 2-8:49 AM

$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

$K = °C + 273$

$\frac{252}{0.45}\frac{(1)(6)}{295} = \frac{(0.45)(V)}{252}$ $\frac{252}{0.45}$

$V = 11.39\ l$

Dec 2-8:54 AM

Regnits Chem

1 Mole of Any Gas = 22.4 l

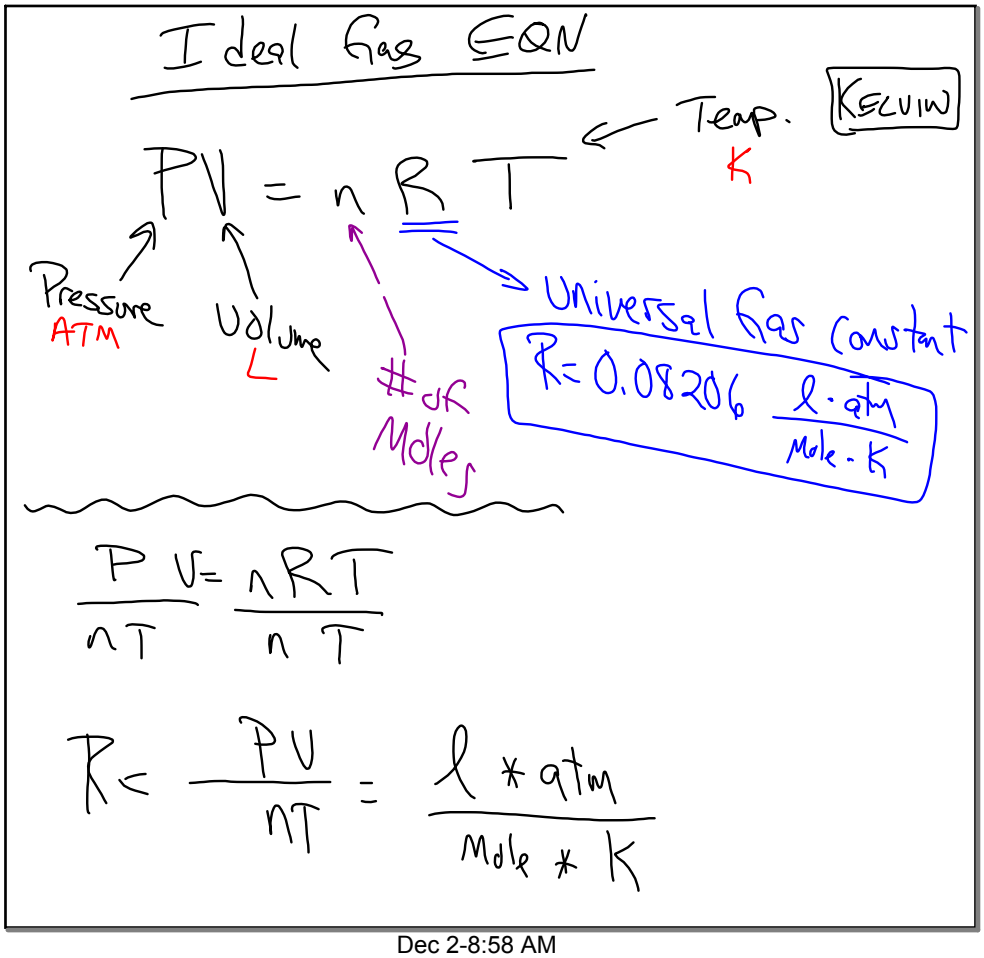
* AT STP

273K 0°C

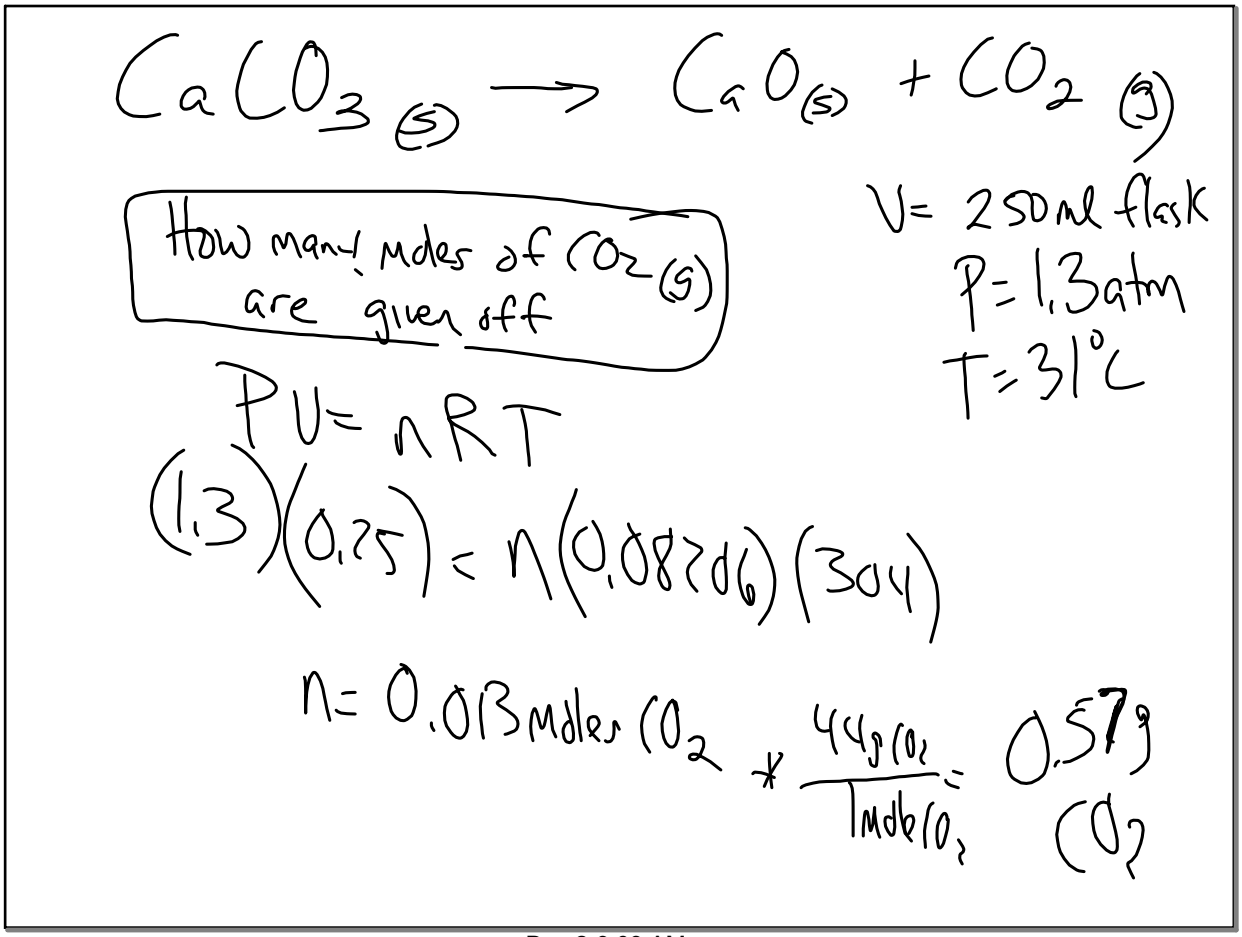
ONLY

1 atm 760 mm Hg 101.3 kPa

Dec 2-8:57 AM




Dec 2-8:58 AM




Dec 2-9:03 AM

Density of a gas $= \frac{\text{mass}}{\text{volume}} = \frac{g}{l}$ $\frac{g}{V}$

$PV = nRT$

$\frac{PV}{1} = \frac{gRT}{MW}$

$\frac{P(MW)}{1} = \frac{gRT}{V}$

$\frac{P(MW)}{RT} = \frac{g}{V}$

$D = \frac{g}{V} = \frac{P(MW)}{RT}$

Density of a gas

Dec 2-9:09 AM

Molarity of soln $\frac{M}{1} = \frac{\text{moles}}{l}$ $\frac{n}{V}$

$PV = nRT$

$\frac{n}{V} = \frac{P}{RT} = \text{MOLARITY}$

Dec 2-9:12 AM

$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

$PV = nRT$

Kelvin

$density = \frac{P(mw)}{RT}$

$M = \frac{P}{RT}$

Dec 2-9:14 AM

10/ 24, 38, 50a

Molec. Geom | AB

Dec 2-9:16 AM