

11.38 CCl_2F_2 $H_v = \frac{289\text{J}}{\text{g}}$ H_2O (15°C)
 $200\text{g}, 288\text{K}$

$mC_{\text{ST}} + mH_f = \text{Total energy to freeze } 200\text{g H}_2\text{O at } 15^\circ\text{C}$

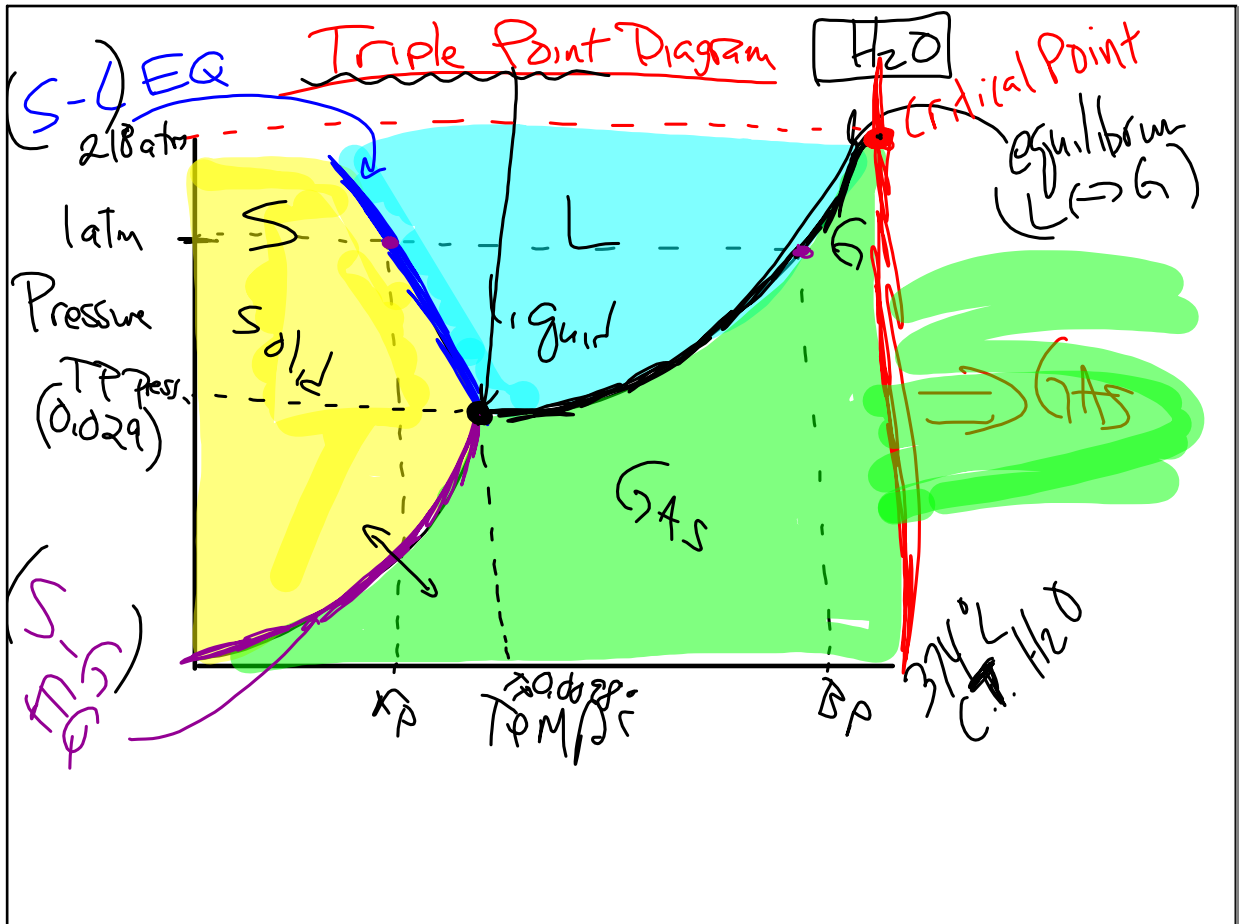
$(200)(4.18)(15) + 200(334) = 79340\text{J}$

$\frac{79340\text{J}}{289\text{J}} = 274.53\text{g CCl}_2\text{F}_2$

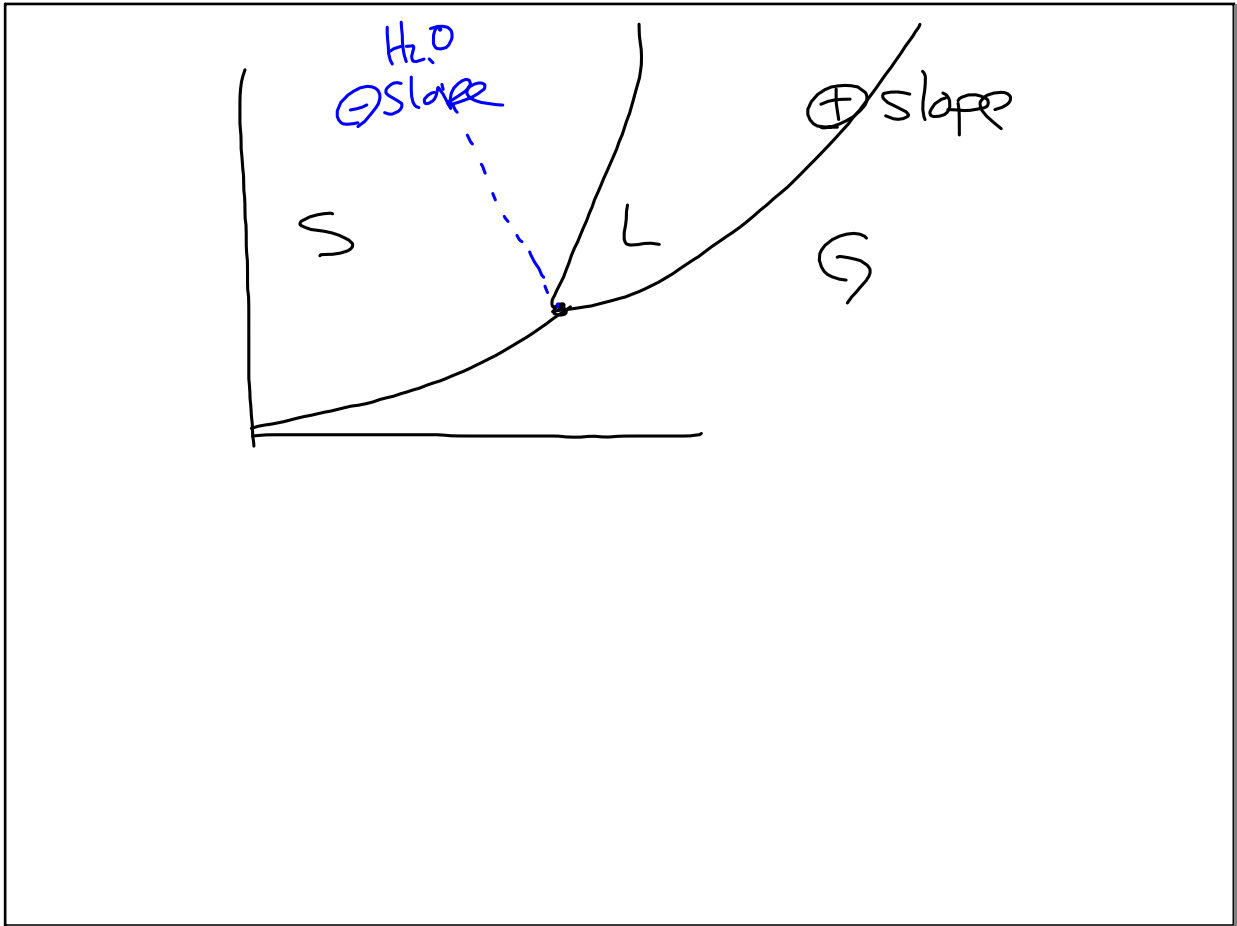
$H_f = \frac{334\text{J}}{\text{g}}$
 $C = \frac{4.18\text{J}}{\text{g}\cdot\text{K}}$

Freeze

Dec 20-7:40 AM



Dec 20-8:51 AM



Dec 20-9:03 AM

<u>tlw</u>	11-		
PS 1		1-15	skip #2
PS ¹¹⁻ 2		1-20	

Dec 20-9:13 AM