

(38b) $\frac{[\text{OH}^-]}{[\text{H}^+]} = 0.014 = 1.4 \times 10^{-2}$

$$[\text{H}^+] [\text{OH}^-] = 1 \times 10^{-14}$$

$$[\text{H}^+] = 7.14 \times 10^{-13}$$

$$\text{pH} = -\log([\text{H}^+]) = 12.146 \quad \text{Base}$$

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(38c) $\text{pH} = 6.6$ $\text{pOH} = 7.4$ Find $[\text{H}^+]$ and $[\text{OH}^-]$

$$\text{pH} = -\log([\text{H}^+])$$

$$-6.6 = \log([\text{H}^+])$$

$$[\text{H}^+] = 2.5 \times 10^{-7}$$

$$[\text{H}^+] [\text{OH}^-] = 1 \times 10^{-14}$$

$$3.98 \times 10^{-8}$$

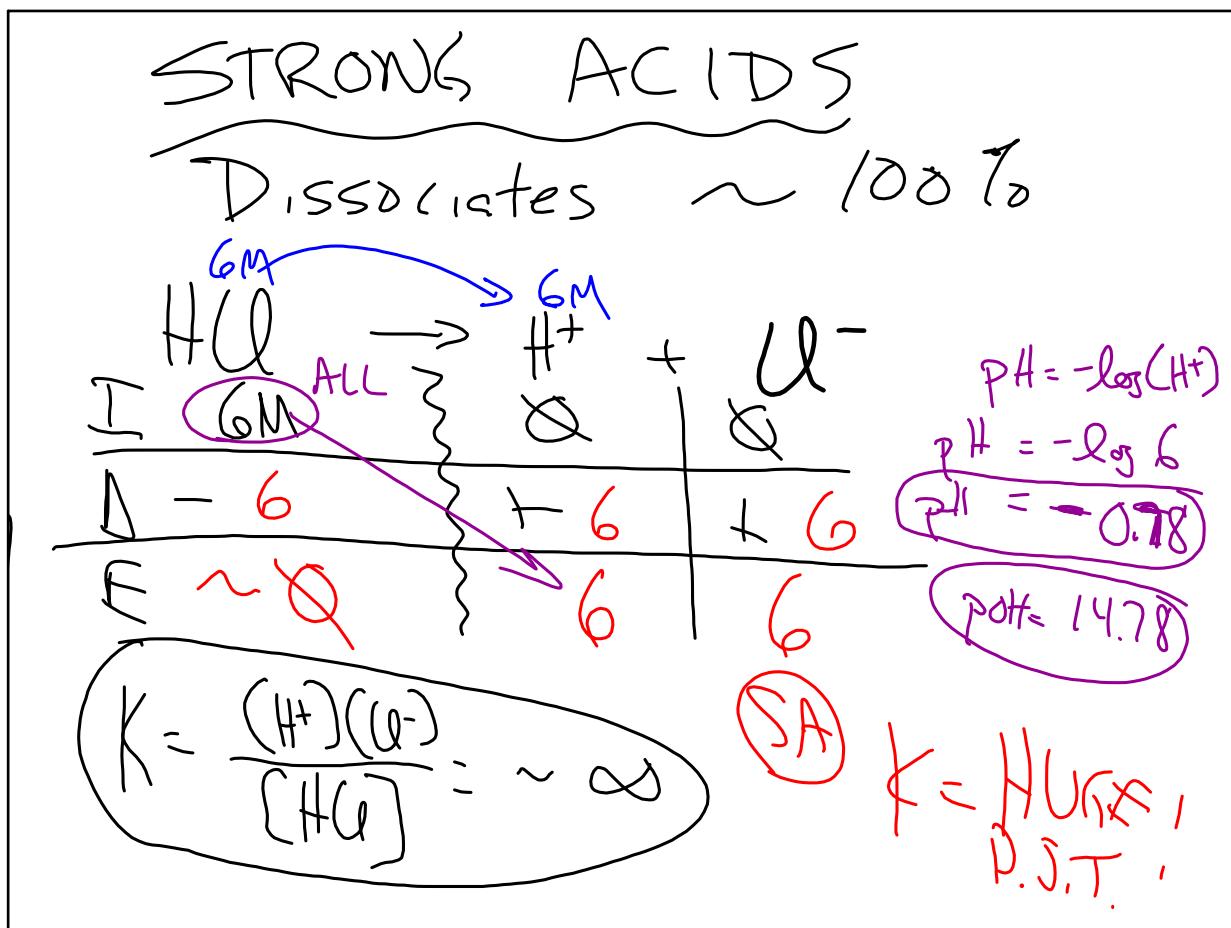
$$-7.4 = \log([\text{OH}^-])$$

$$\text{pH} + \text{pOH} = 14$$

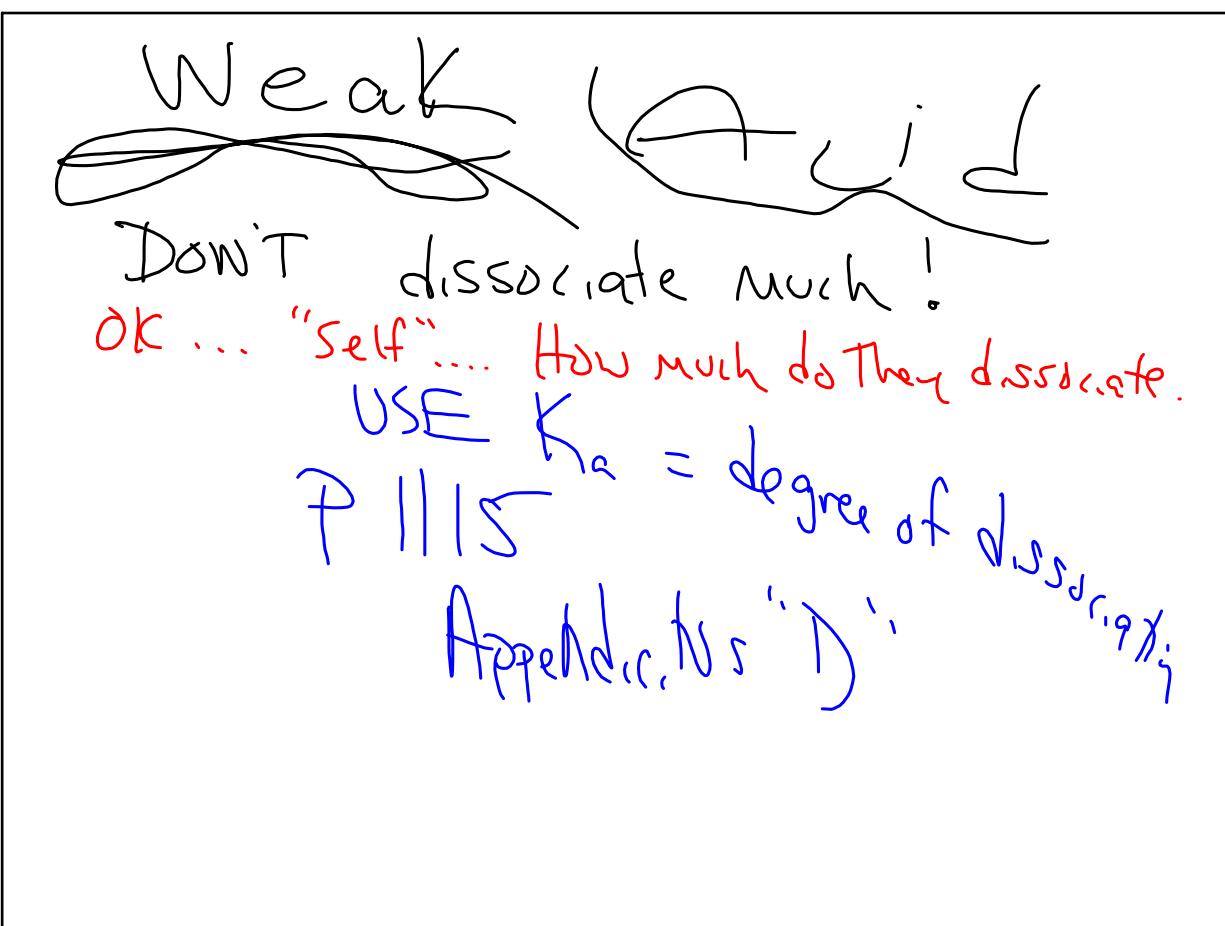
$$\text{pOH} = -\log([\text{OH}^-])$$

AT 25 °C

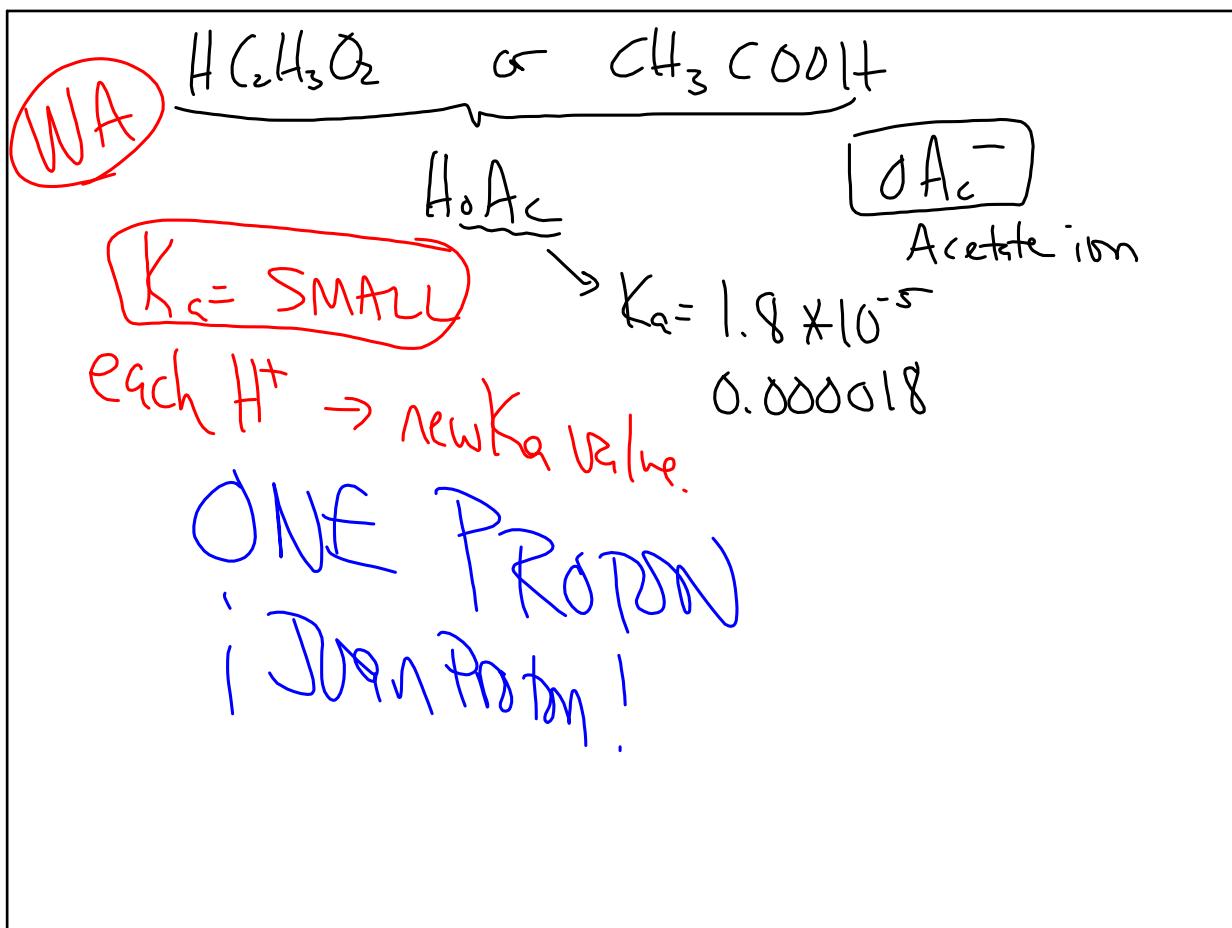
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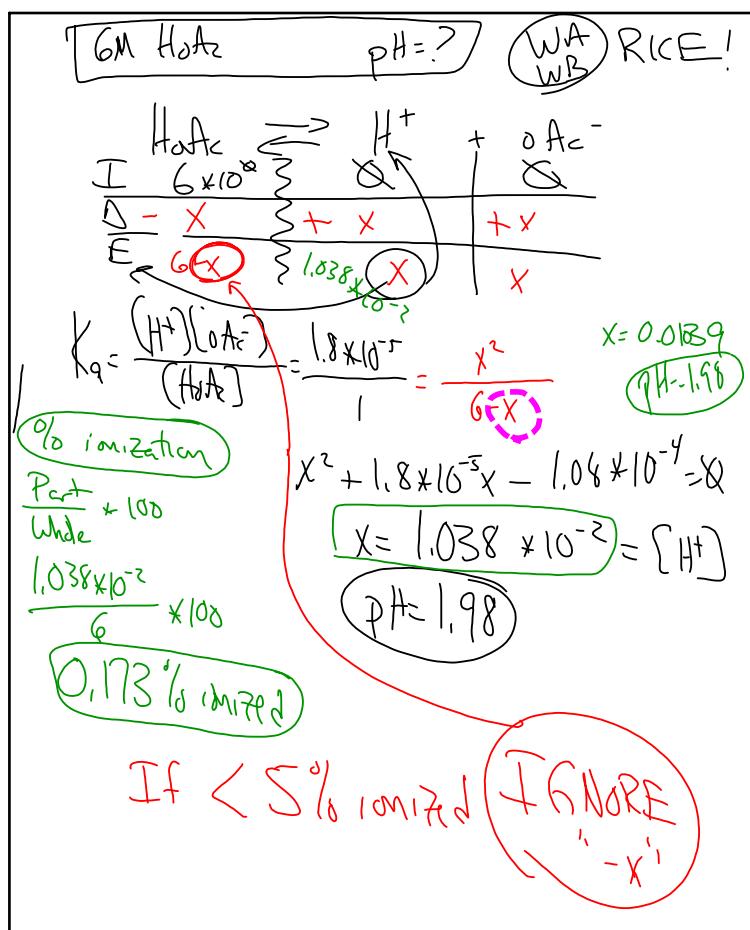
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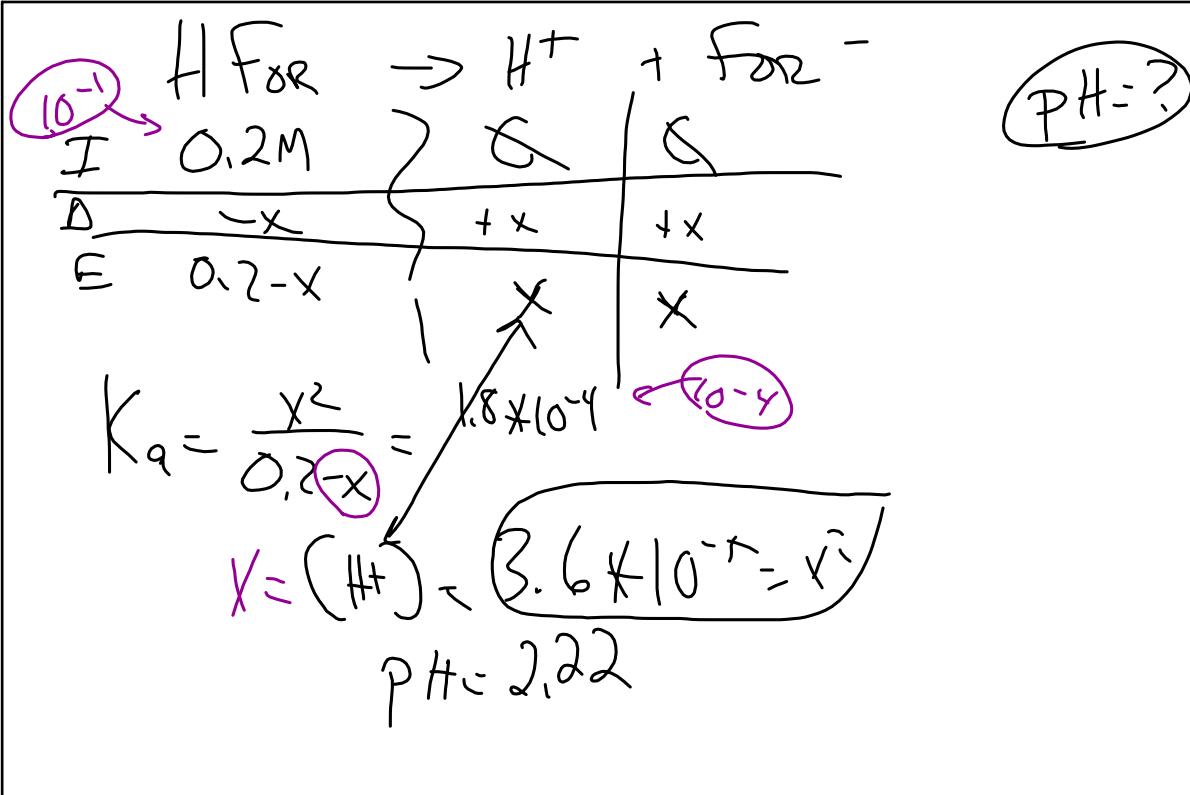
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Feb 23-8:06 AM



Feb 23-8:11 AM



Feb 23-8:23 AM

$$16 / 60 + 64$$

Feb 23-8:29 AM