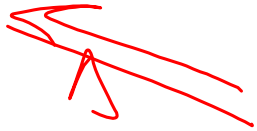
15 $HA \rightleftharpoons H^+ + A^-$



Mar 13-7:36 AM

17 $NaOH + HF \rightarrow NaF + HOH$

50ml 0.1M 30ml 0.25M

$K_a = 6.8 \times 10^{-4}$

Next Moves ①

	OH^-	+ HF	⇒ F^-	+ HOH
I	0.005	0.0075	∅	
Δ	-0.005	-0.0050	+0.005	
E	∅	0.0025 moles	0.005 moles	

② ReactM

80ml

0.08 0.08

0.03125M HF

0.0625M F^-

$pH = pK_a + \log \frac{\text{base}}{\text{acid}}$

$= -\log(6.8 \times 10^{-4}) + \log \frac{0.0625}{0.03125}$

$pH = 3.1685$

Mar 13-7:50 AM

(21) $Ba(OH)_2 \rightarrow Ba^{+2} + 2OH^-$

0.0012 M 0.0012 0.0024

SB ⇒ Dissociates ~ 100%

Grp 2 $C_a B_a S_r$

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(20) HB_z + NaB_z B_z^- Buffer

WA Soluble Salt of WA Conj Base

Mar 13-8:01 AM

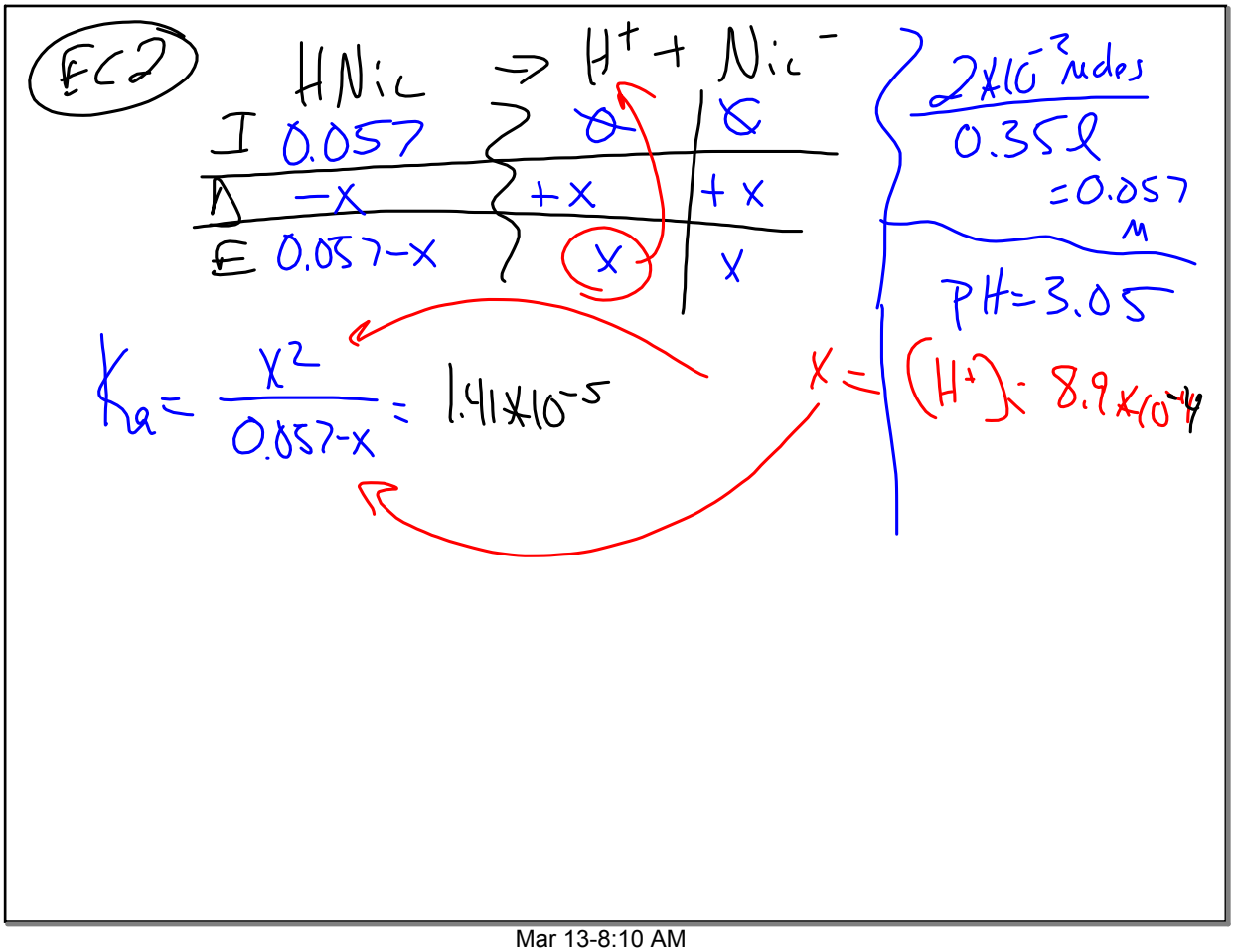
(22) $PbI_{2(s)} \Rightarrow Pb^{+2}_{(aq)} + 2I^-_{(aq)}$

M ratio ⇒ x x 2x

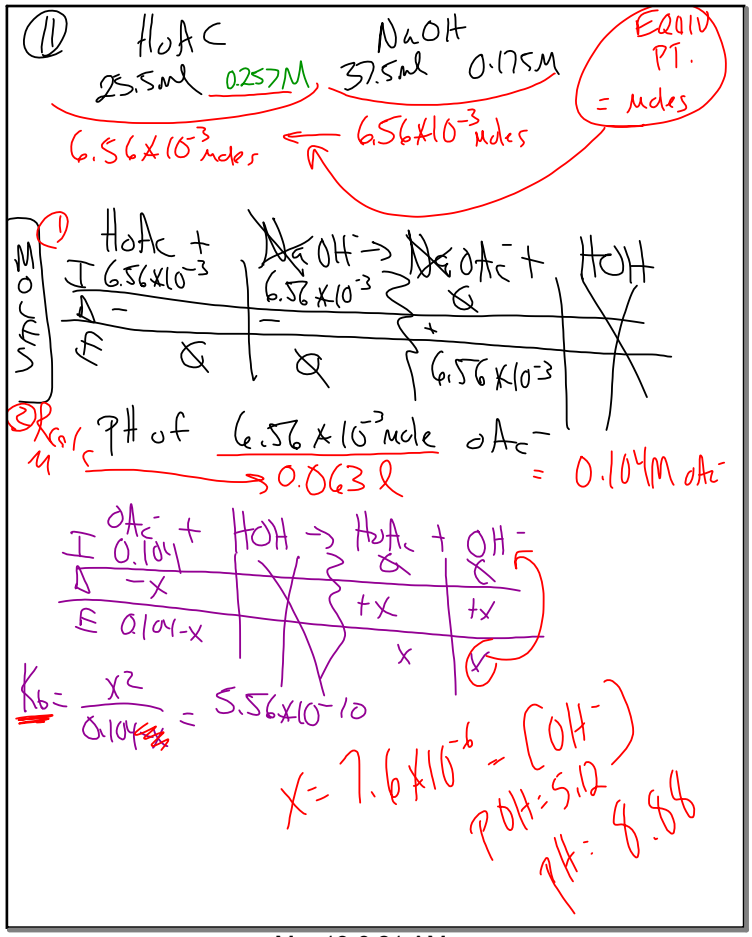
$$K_{sp} = [Pb^{+2}][I^-]^2$$

$(x)(2x)^2$

Mar 13-8:08 AM



Mar 13-8:10 AM



Mar 13-8:21 AM