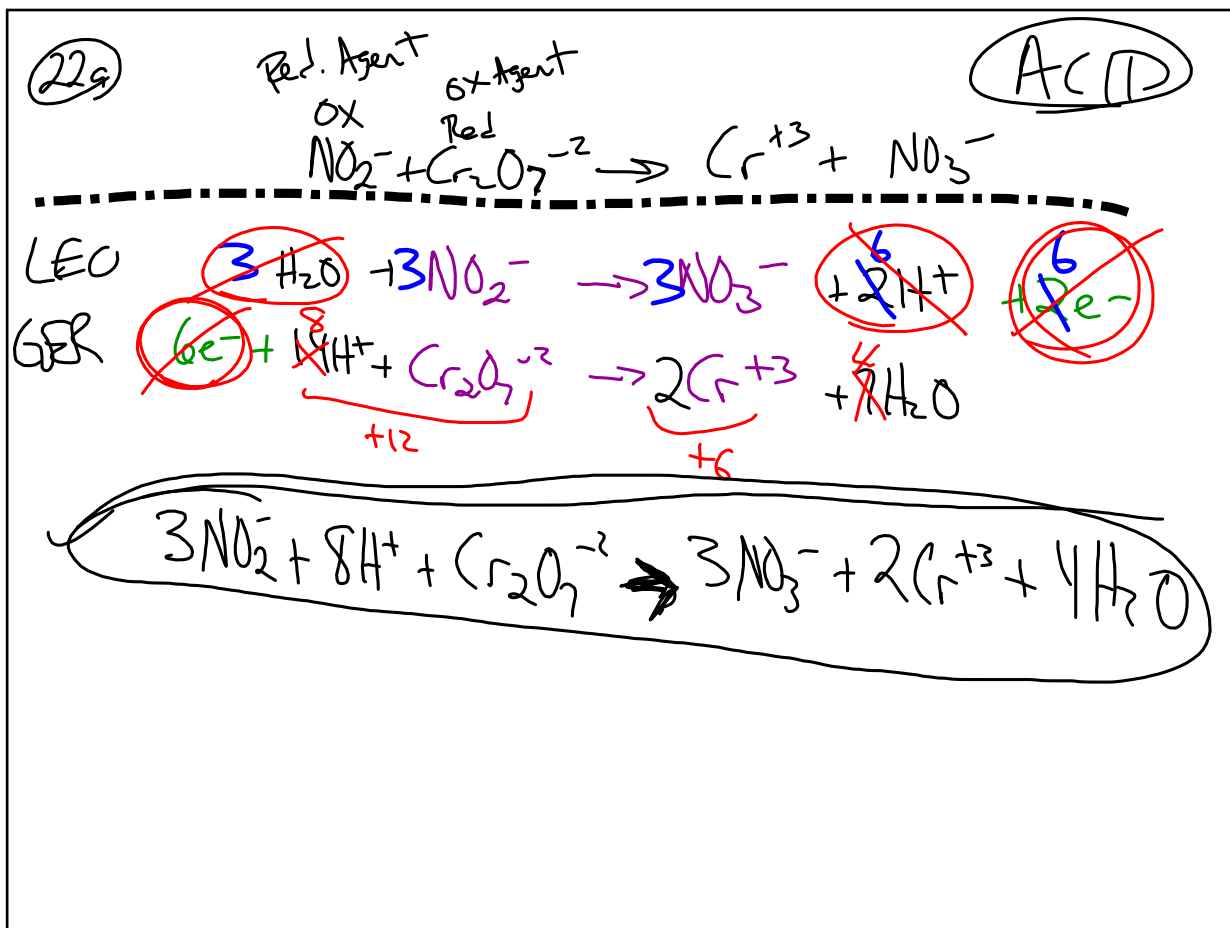


Mar 28-7:38 AM



Mar 28-7:59 AM

BASIC Soln

NO extra H^+ ← ☹️

OH^- ☺️

$H^+ + OH^- \rightarrow H_2O$

① BALANCE e^- if in acidic soln

② THEN ⇒ Add OH^- to the H^+ 's & other side of eqn. AND Add same # OH^-

Mar 28-8:08 AM

ACID $CN^- + MnO_4^- \rightarrow CNO^- + MnO_2$ (5)

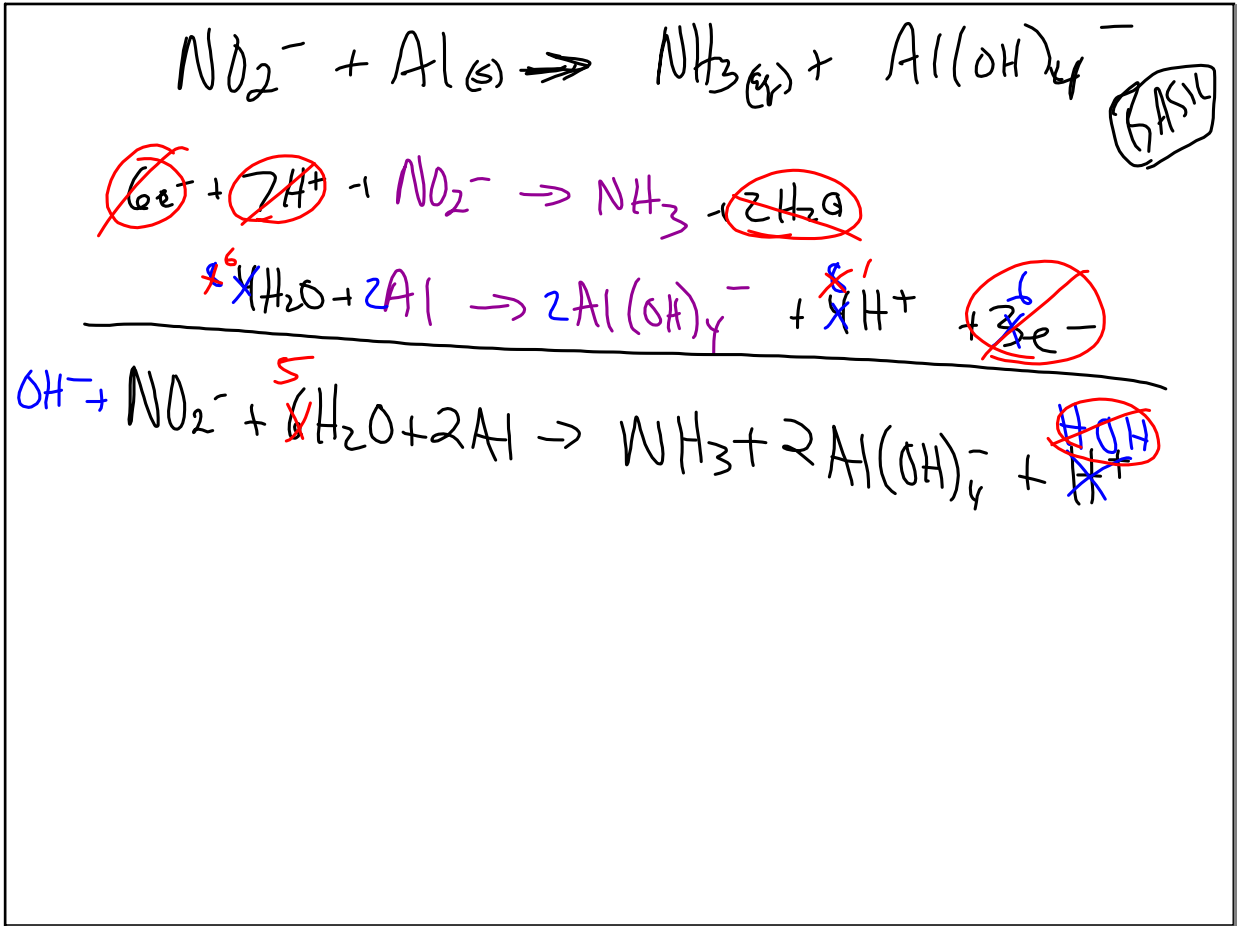
~~$3H_2O + 3CN^- \rightarrow 3CNO^- + 2H^+ + 2e^-$~~ BASIC

~~$2e^- + 2H^+ + 2MnO_4^- \rightarrow 2MnO_2 + 2H_2O$~~

ACID ⇒ BASIC

$3CN^- + 2H^+ + 2MnO_4^- \rightarrow 3CNO^- + 2MnO_2 + H_2O + 2OH^-$

Mar 28-8:12 AM



Mar 28-8:19 AM

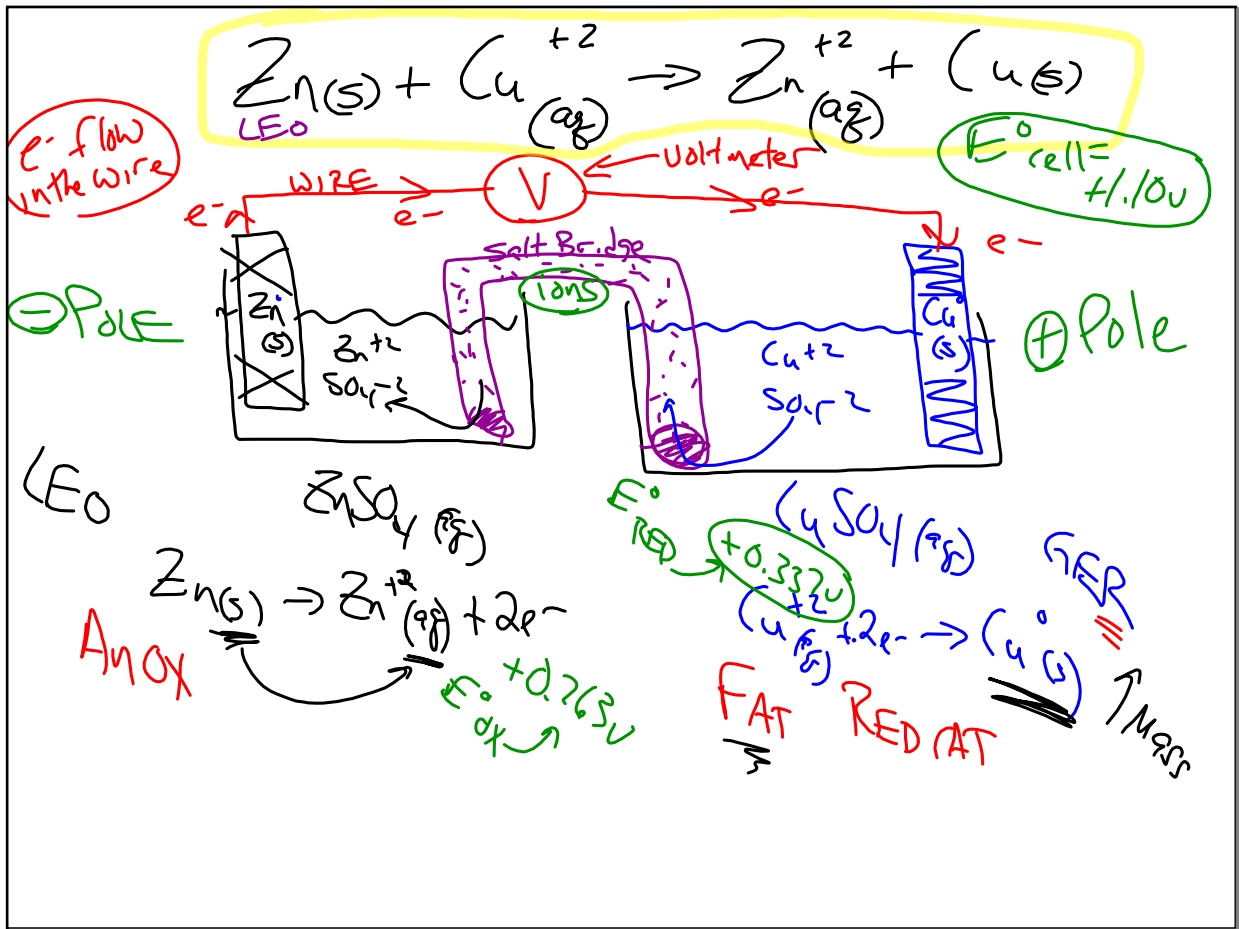
Electrochemistry

Electrochemical cell - Spontaneous redox rxn produces a current (electricity) (flow of e⁻) (LEO → GER)

↳ Voltaic cell

↳ Galvanic cell

Mar 28-8:38 AM



(E) 9/11/17 STD Reduction Potent. 9/5

$\oplus E^{\circ}$ SPAWT

$1V = \text{VOLT} = \frac{1 \text{ Joule}}{1 \text{ coulomb}} \quad \frac{J}{\text{Coul.}}$

Zn / Zn^{+2}

$E^{\circ}_{red} = -0.763V$

$E^{\circ}_{ox} = +0.763V$

Cu / Cu^{+2}

$E^{\circ}_{red} = +0.337V$

Greater Reduction Potential Ox Agent.

Mar 28-9:01 AM

EMF E° (Volts)

↳ electro Motive force

STP, IM, Ictm E°

Non-Std E

Mar 28-9:13 AM

20/26 + 34

Mar 28-9:15 AM