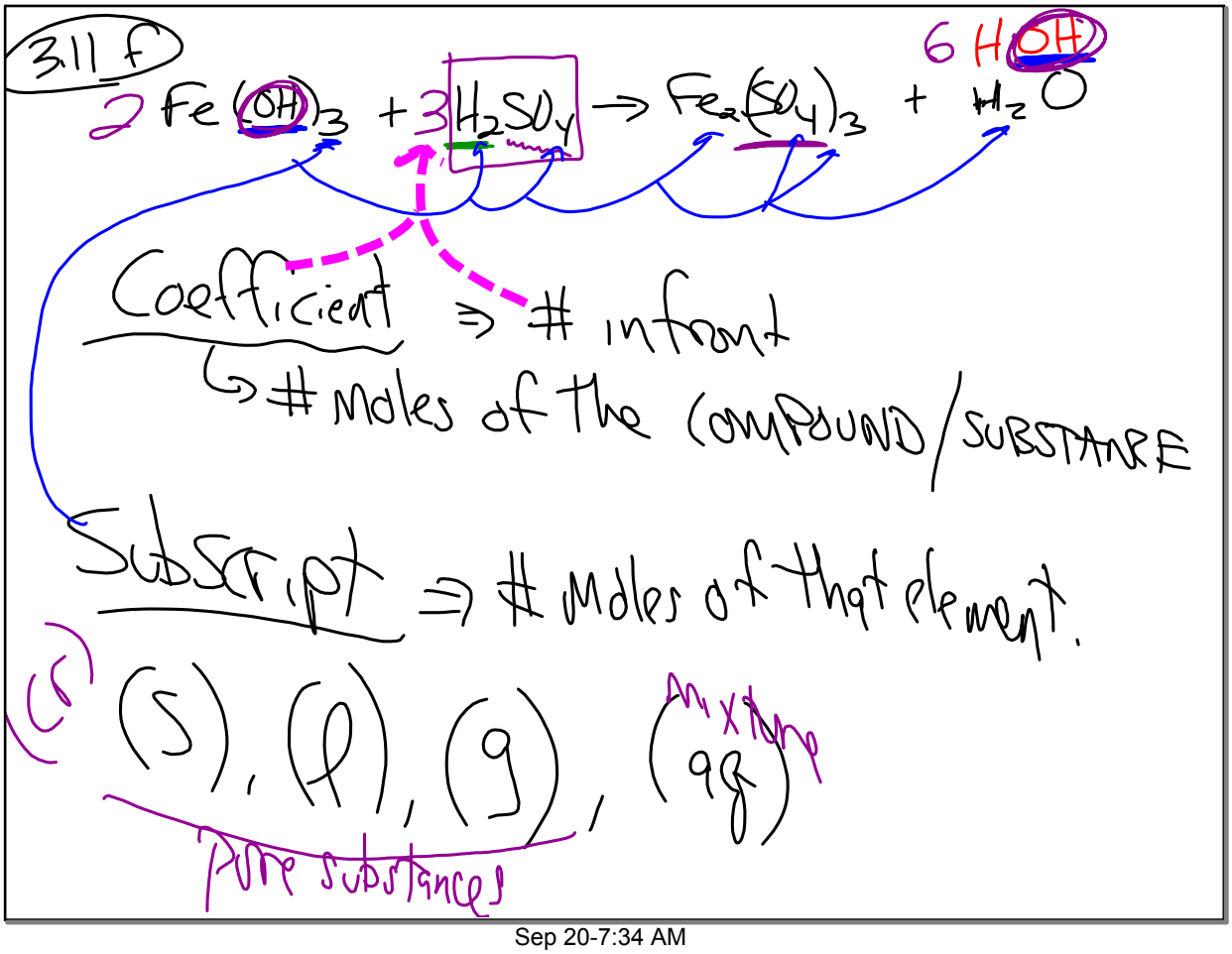
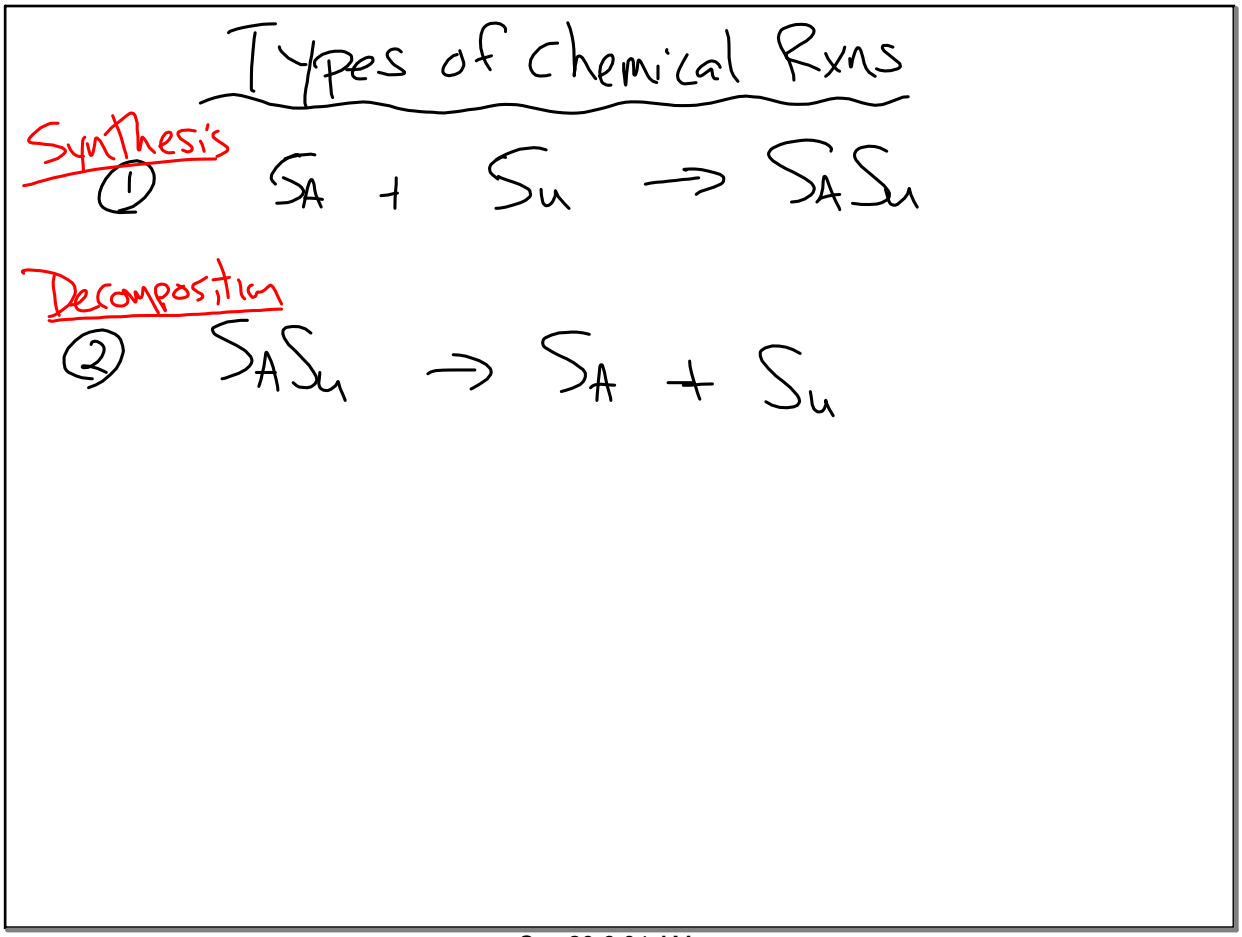


Sep 20-7:34 AM



Sep 20-8:04 AM

③ $V + RG \rightarrow G + RV$

Single Replacement

Regents Table ⑤ — V ← More active can replace what's below. G

④ $CD + AK \rightarrow CK + AD$

Double replacement

Sep 20-8:11 AM

⑤ Combustion

$C_xH_yO_z + O_2 \rightarrow CO_2 + H_2O$ — complete combustion.

SAME

Incomplete combustion

$\_\_ + O_2 \rightarrow CO + H_2O$

Sep 20-8:36 AM

NaN_3 - Sodium Azide

65 g / Mole

$NaN_{3(s)} \rightarrow Na_{(s)} + N_{2(g)}$

— g 72 L

Sep 20-8:46 AM

MOLE RATIO

$2\,NaN_{3(s)} \rightarrow 2Na_{(s)} + 3N_{2(g)}$

139.29 g 72 L

MOLE RATIO

$72\,L\,N_2$	$1\,mole\,N_2$	$2\,mole\,NaN_3$	$65\,g\,NaN_3$	$139.29\,g\,NaN_3$
	$22.4\,L\,N_2$	$3\,mole\,N_2$	$1\,mole\,NaN_3$	

① Convert to moles ② Moles A → Moles B ③ Mole → ans.

Sep 20-8:53 AM

How many ℓ CO_2 are produced by the combustion of 5g of Propane? C_3H_8

$$1\,C_3H_8 + 5\,O_2 \rightarrow 3\,CO_2 + 4\,H_2O$$

(5g) C_3H_8 → 7.64 ℓ CO_2 (g)

$$\frac{5\text{g}\,C_3H_8}{} \left|\frac{1\text{mole}\,C_3H_8}{44\text{g}\,C_3H_8}\right|\frac{3\text{ mole}\,CO_2}{1\text{ mole}\,C_3H_8}\left|\frac{22.4\,\ell}{1\text{ mole}\,O_2}\right. =$$

Sep 20-9:06 AM

Tarea para esta noche

3/ 38 + 58

Sep 20-9:15 AM