

Ammonium Carbonate

NH_4^{+1} CO_3^{-2} $(\text{NH}_4)_2\text{CO}_3$

ACIDS

Polyatomic ending ↓
 ATE - ic
 ITE - ous.

Sep 8-7:37 AM

Acids

① H + Polyatomic ion

Ate - ic $\text{H}_2\text{SO}_3 \rightarrow$ Sulfurous acid.
 ite - ous. H_2SO_4 - Sulfuric acid

H_3PO_4 - phosphoric acid
 H_3PO_3 ← phosphorous acid

Sep 8-7:53 AM

② H + element.

HCl hydro____-ic acid
hydrochloric acid

H₂S - hydro Sulfuric acid

Sep 8-7:56 AM

CO

(+2)

FeCl₂

Iron (II) chloride

Transition elements

- ① Can have multiple ox #'s.
- ② Colors. Colorful compounds.

vs

(+3)

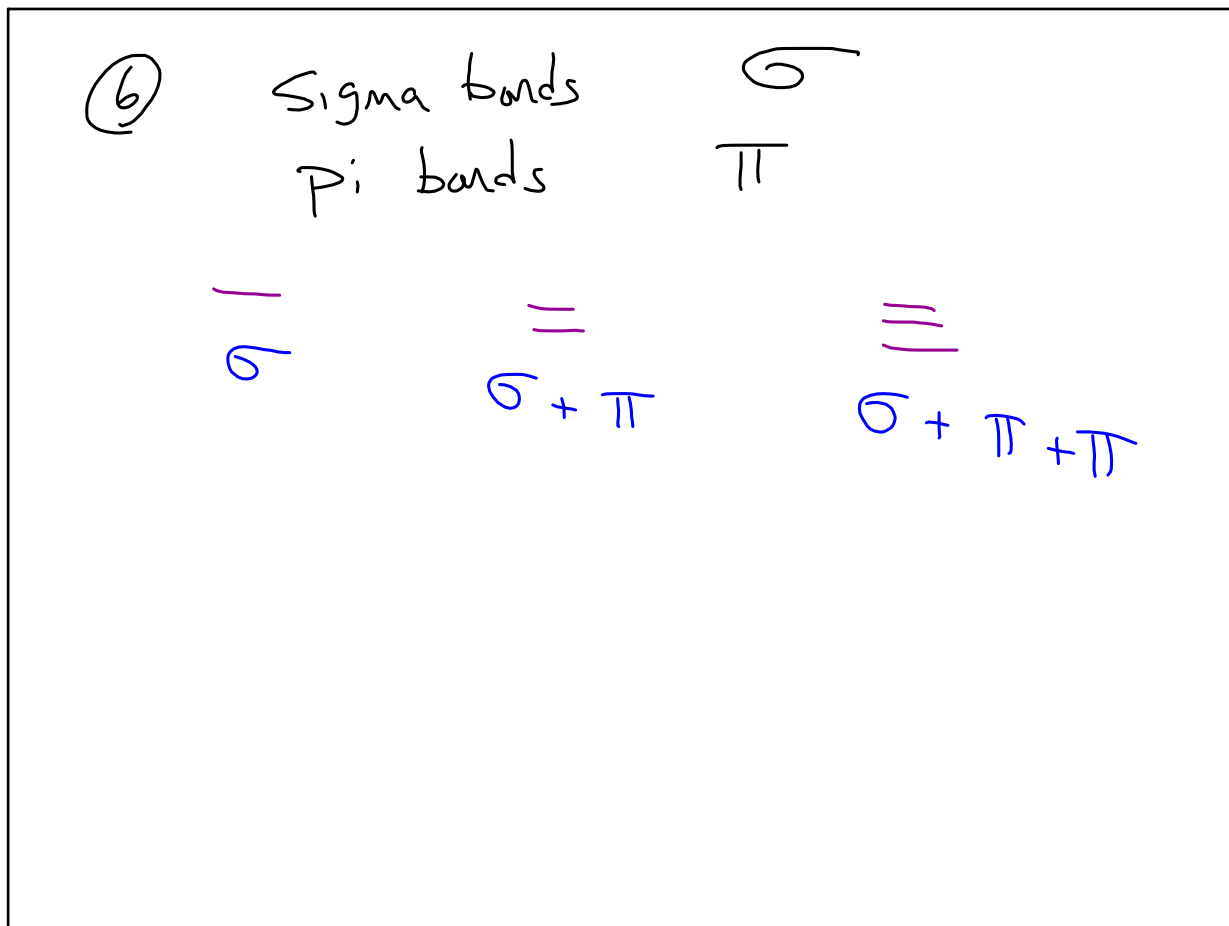
FeCl₃

Iron (III) chloride

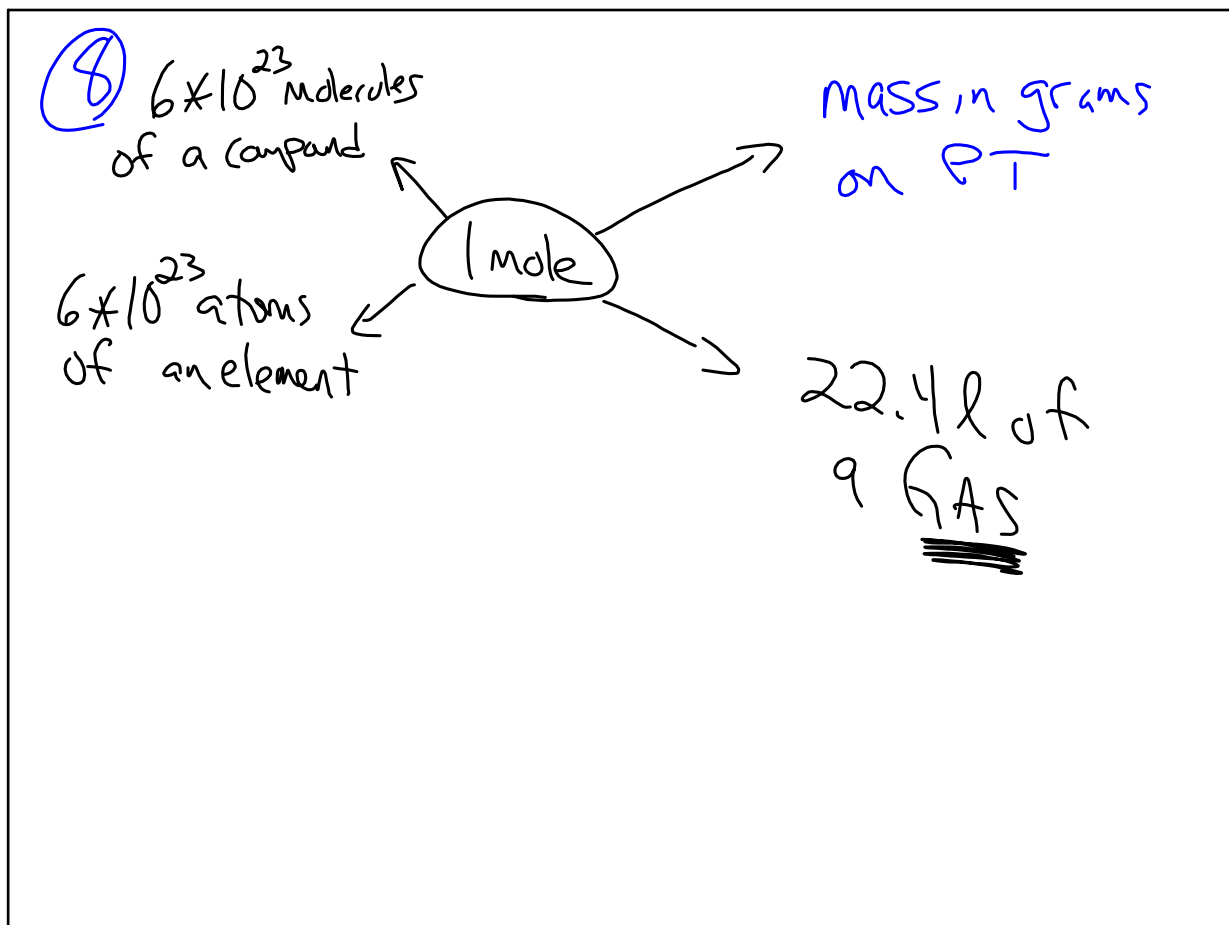
Iron Chloride

Stock system - Roman # IS the ox #

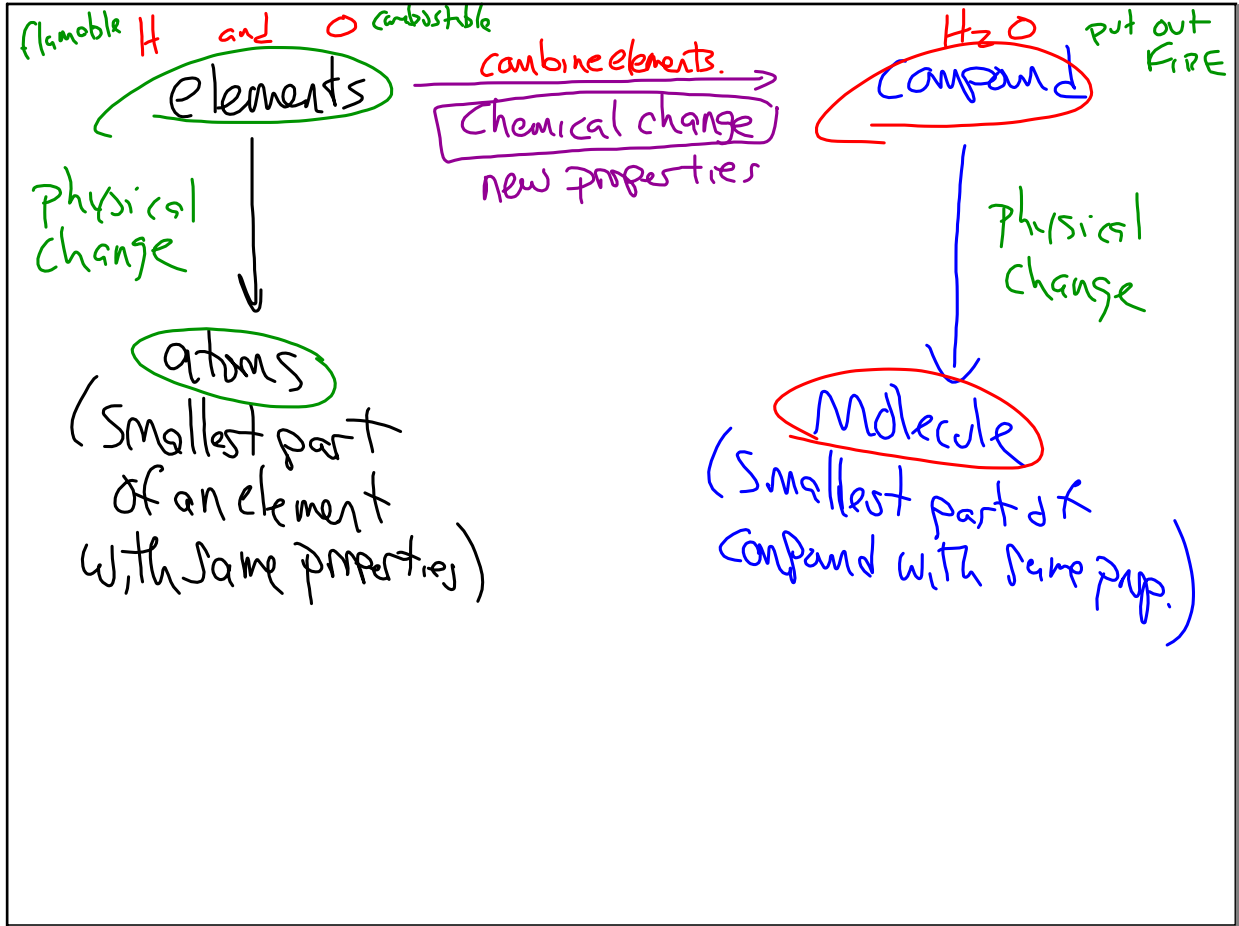
Sep 8-7:58 AM



Sep 8-8:39 AM



Sep 8-8:45 AM



Sep 8-8:49 AM

④ how many mole oxygen are in 1×10^{22} molecules SO_3 .

$\frac{1 \text{ mole } SO_3}{6 \times 10^{23} \text{ molecules } SO_3}$ ← conversion factor.

1 mole SO_3	1×10^{22} molecules SO_3	3 mole O	= 0.05
6×10^{23} molecules SO_3		1 mole SO_3	

Sep 8-8:56 AM

$$\frac{40 \text{ miles}}{\text{hour}} \Rightarrow \frac{\text{Km}}{\text{hr}}$$

$$(1 \text{ in} = 2.54 \text{ cm})$$

40 miles	5280 ft	12 in	2.54 cm	1 m	1 Km
1 hr	1 mile	1 ft	1 in	100 cm	1000 m

~~64 Km~~
hr

Sep 8-9:03 AM

$$308 \text{ ft}^2 = \text{_____} \text{ in}^2$$

308 ft²	(12) ² in ²	=
	(1) ² ft²	

Sep 8-9:10 AM