13.28

MORE Polar

$CH_3(CH_2)_n COOH$

$\overset{O}{\overset{||}{C}} - OH$

POLAR END

$CH_3 \; CH_2CH_2CH_2CH_2CH_2 \; COOH$

Less Polar.

Jan 3-8:28 AM

SAT - Full

UNSAT - Vacancy

"Supa" saturated - Temporarily find 'extra space'
Usually heat

Jan 3-8:48 AM

M = Molarity m = Molality

$$M = \frac{\text{Moles solute}}{l \text{ of solution}}$$

Solute + Solvent

$$m = \frac{\text{Moles of solute}}{Kg \text{ solvent}}$$

Jan 3-8:58 AM

PPM — parts per Million

% (percent) = parts per hundred (pph)

$$PPM = \frac{\text{Part}}{\text{Whole}} * 1,000,000 \quad ^{10^6}$$

Jan 3-9:00 AM

Dilute → less solute in solvent

US

Concentrated → More solute in solvent

→ Both UN saturated.

Jan 3-9:01 AM

Gases → ↑ [ ] of a gas

increase Pressure.

Concentration

$[NaCl_{(aq)}] = 3M$

Henry's Law   ↑ P of a gas in a liquid, you ↑ [ ]

$S_{(g)} = K C_{(g)}$

Solubility
Pressure

Jan 3-9:04 AM

Find   % (pph) and Find Molality

$\boxed{0.75\,m}$ Molal

$\% = \dfrac{Part}{Whole} \times 100$

$$\dfrac{13.5}{13.5 + 100} = 11.89\ \%$$

Jan 3-9:07 AM

$M = \dfrac{Moles}{L\,soln}$ Volume.

$m = \dfrac{Moles}{kg\ solvent}$ Mass

?

USE DENSITY to Convert

Jan 3-9:13 AM

13/32, 44, 50a

Jan 3-9:16 AM