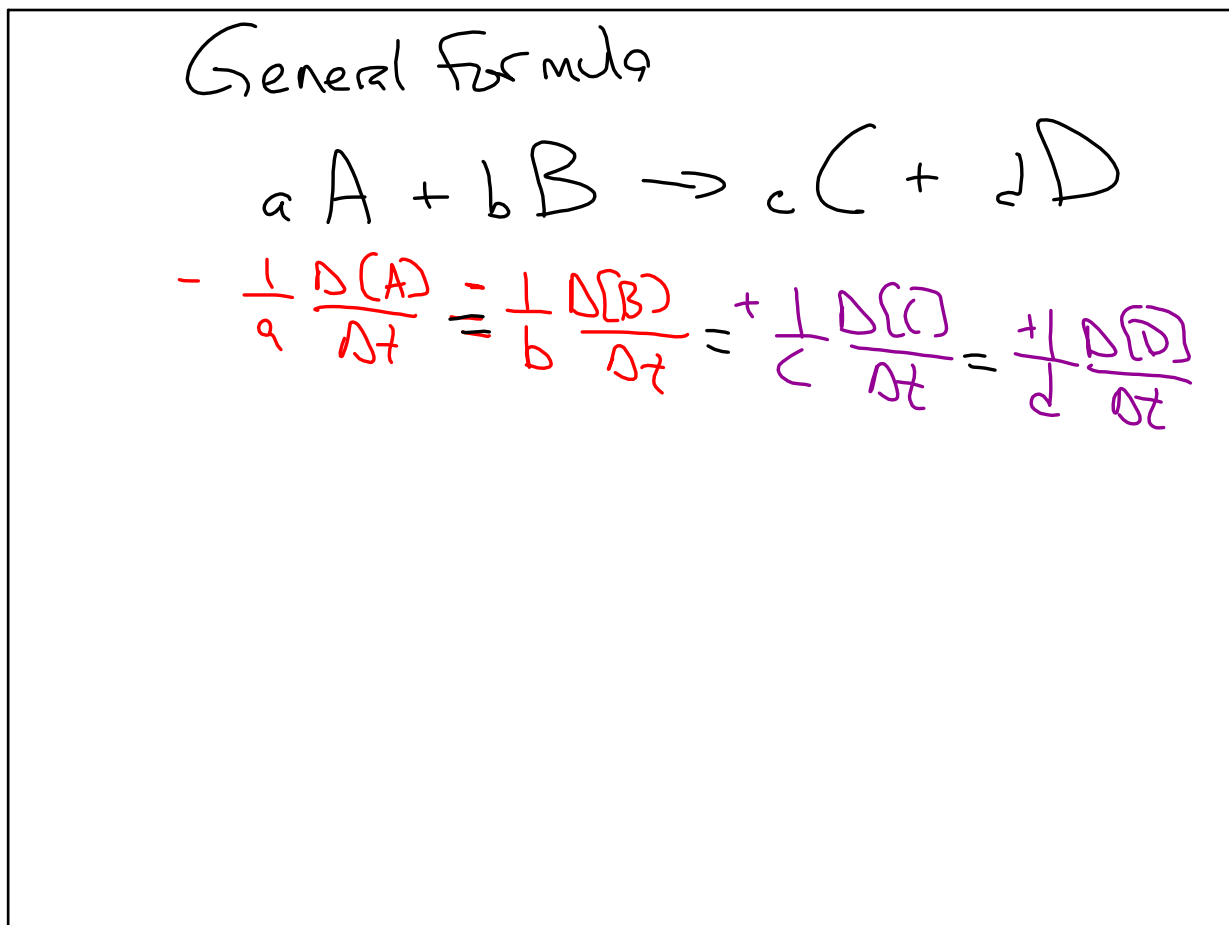


Jan 30-10:03 AM



Jan 30-10:27 AM

$2H_2 + O_2 \rightarrow 2H_2O$

$$-\frac{1}{2} \frac{\Delta[H_2]}{\Delta t} = -\frac{\Delta[O_2]}{\Delta t} = +\frac{1}{2} \frac{\Delta[H_2O]}{\Delta t}$$

Interms of O_2 , how quickly is H_2 disappearing

(*) $-\frac{1}{2} \frac{\Delta[H_2]}{\Delta t} = -\frac{\Delta[O_2]}{\Delta t}$

$\frac{\Delta[H_2]}{\Delta t} = \frac{2\Delta[O_2]}{\Delta t}$

$2x = y$
 $x = \frac{1}{2}y$

Jan 30-10:29 AM

③ Rate Law

Reaction orders
Rate orders

DATA TABLE

$$\text{Rate} = k [reactant 1]^n [reactant 2]^m \dots$$

$\frac{\text{Molarity}}{\text{Sec}}$

↑
 Rate constant

↑
 concentration or at that instant in time

Jan 30-10:35 AM

$$14 / 16 + 22$$

Jan 30-10:46 AM