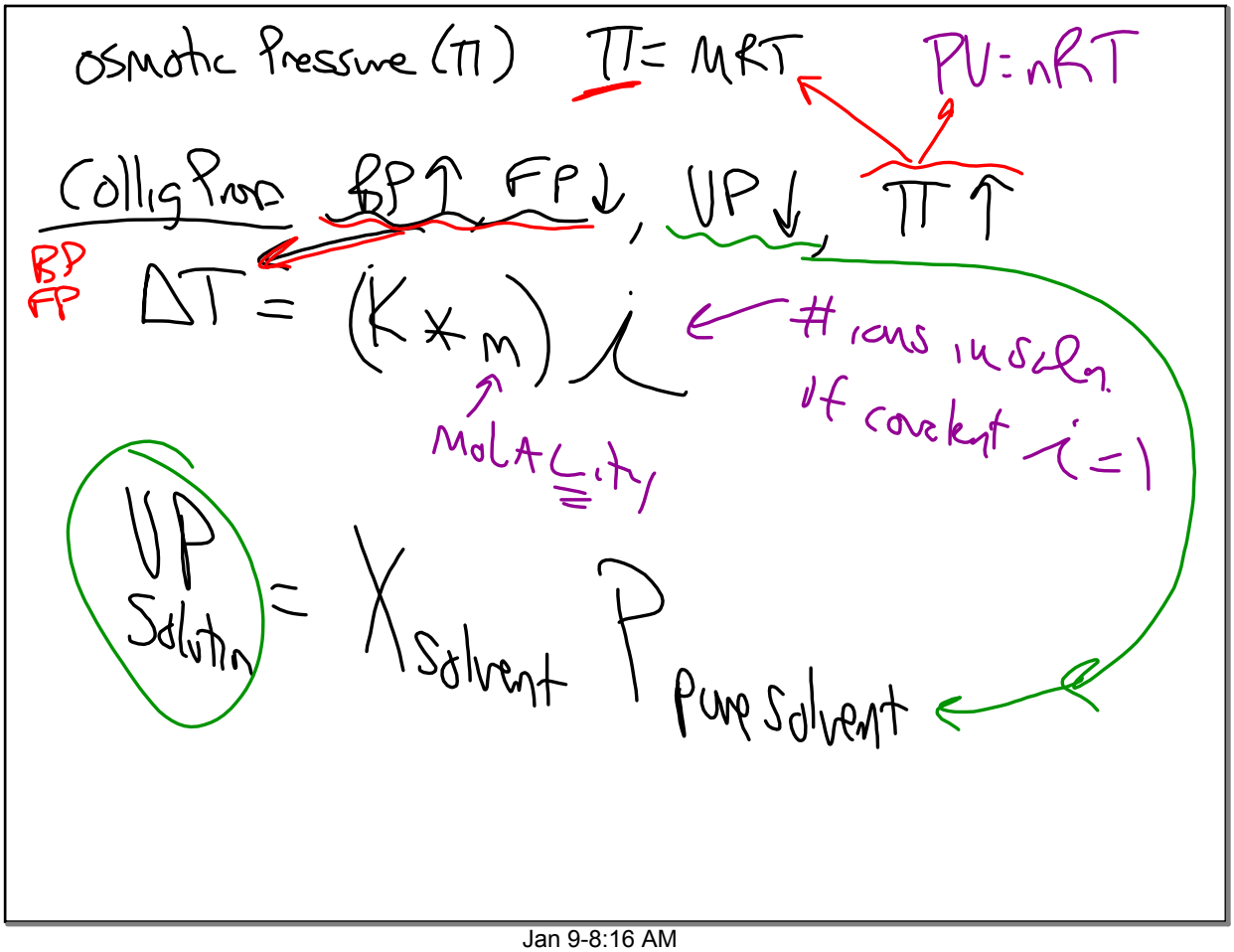
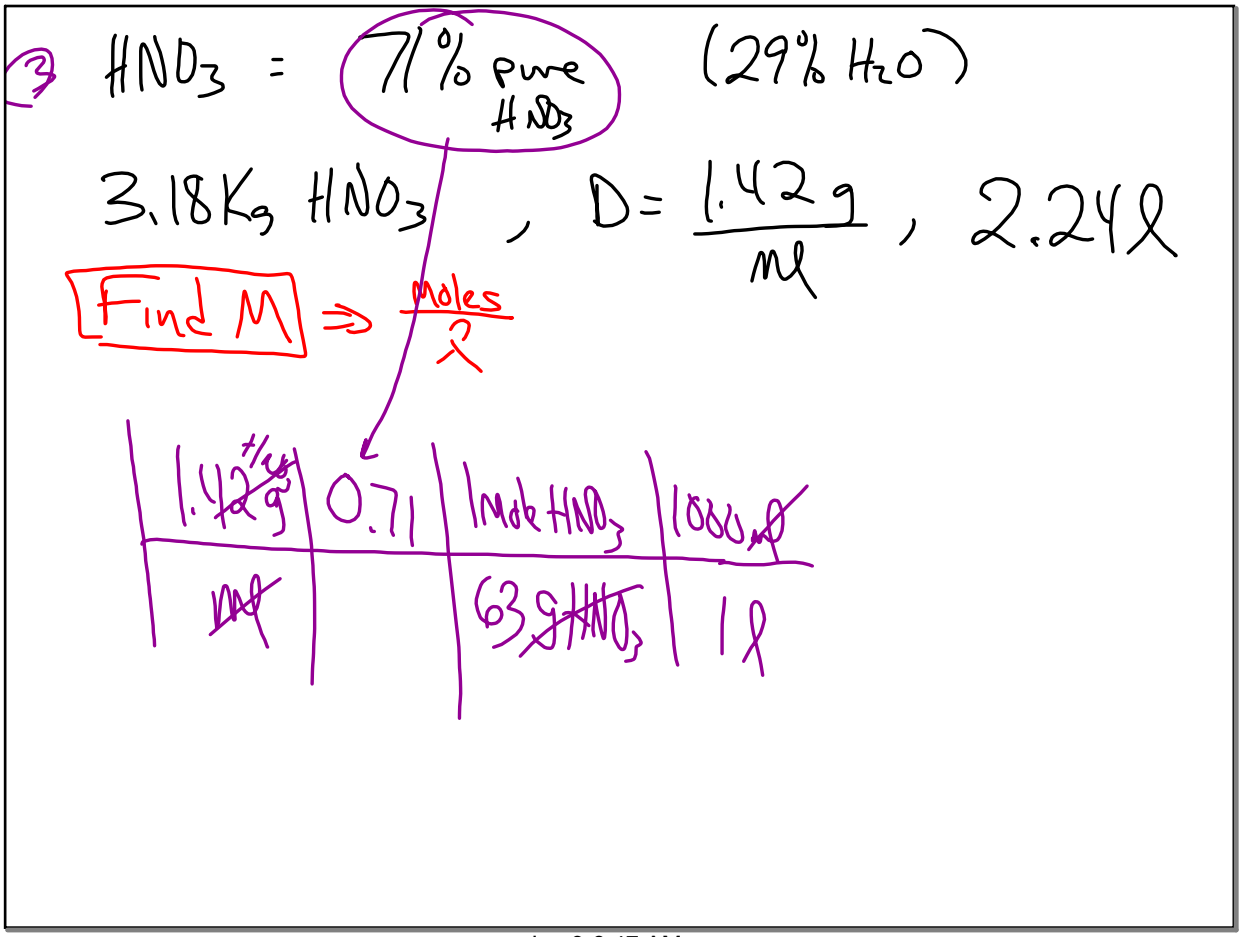


Jan 9-8:16 AM



Jan 9-8:47 AM

(9) $X_{Solute} = \frac{\text{Moles Solute}}{\text{Moles solute} + \text{Moles solvent}} = 1$

(8) CH_3OH 32g/mole H_2O 18g/mole.

32g

$\frac{32}{18} = 1.78$

$X_{CH_3OH} = \frac{\text{Moles } CH_3OH}{\text{Moles } CH_3OH + \text{Moles } H_2O} = \frac{1}{1 + 1.78}$

Jan 9-8:53 AM

(7) 1750g H_2O + ~~0.05M Na_2SO_4~~

? Moles

$\frac{0.05 \text{ Moles } Na_2SO_4}{\text{Kg } H_2O}$

$\left| \frac{0.05 \text{ Mole } Na_2SO_4}{1 \text{ Kg } H_2O} \right| \left| 1.75 \text{ Kg } H_2O \right| =$

Jan 9-9:00 AM

(13) $\Delta T = (K_b * m)\, i$

$\Delta T = (0.5 * 1)\, 5$

$\Delta T = 2.6$

$1m\ Al_2(SO_4)_3$

$2Al^{+3} + 3SO_4^{-2}$

5 ions

Old

Insulin BP 100.51 $\Delta T = 0.51$

$\Delta T = (K_b * m)\, i$

$0.51 = (K_b * 1)\, 1$

$K_b = 0.51$

Jan 9-9:05 AM

(19) EG (62 g/mole) + H_2O (18 g/mole) $\frac{62}{18}$

$$X_{EG} = \frac{\text{Mole EG}}{\text{Mole EG} + \text{Mole } H_2O} = \frac{1}{1 + \frac{62}{18}}$$

Jan 9-9:10 AM

(16) 0.25g cytc , 50ml soln , 1.52kPa

50×10^{-3} L 0.015 atm

MW $\frac{g}{mole}$

25°C

Kelvin 298

$PV = nRT$

$\frac{PV}{1} = \frac{gRT}{MW}$

$\frac{MW}{1} = \frac{gRT}{PV} = \frac{(0.25)(0.08206)(298)}{(0.015)(50 \times 10^{-3})}$

Jan 9-9:13 AM