

(Eci) 0.353g KOH / — ml, ~~0.0176M KOH~~

0.0176 mole KOH  
1 l

1000ml	<del>0.353g KOH</del>	<del>1 mole KOH</del>	= 358.16ml
<del>0.0176 mole KOH</del>		<del>55g KOH</del>	

<del>1 l</del>	<del>1 mole <math>\text{Al}_2(\text{O}_3)</math></del>	<del>53g <math>\text{Na}_2\text{CO}_3</math></del>	2.5 l 2500ml
<del>0.2 mole <math>\text{Al}_2(\text{O}_3)</math></del>	<del>106g <math>\text{Al}_2(\text{O}_3)</math></del>		

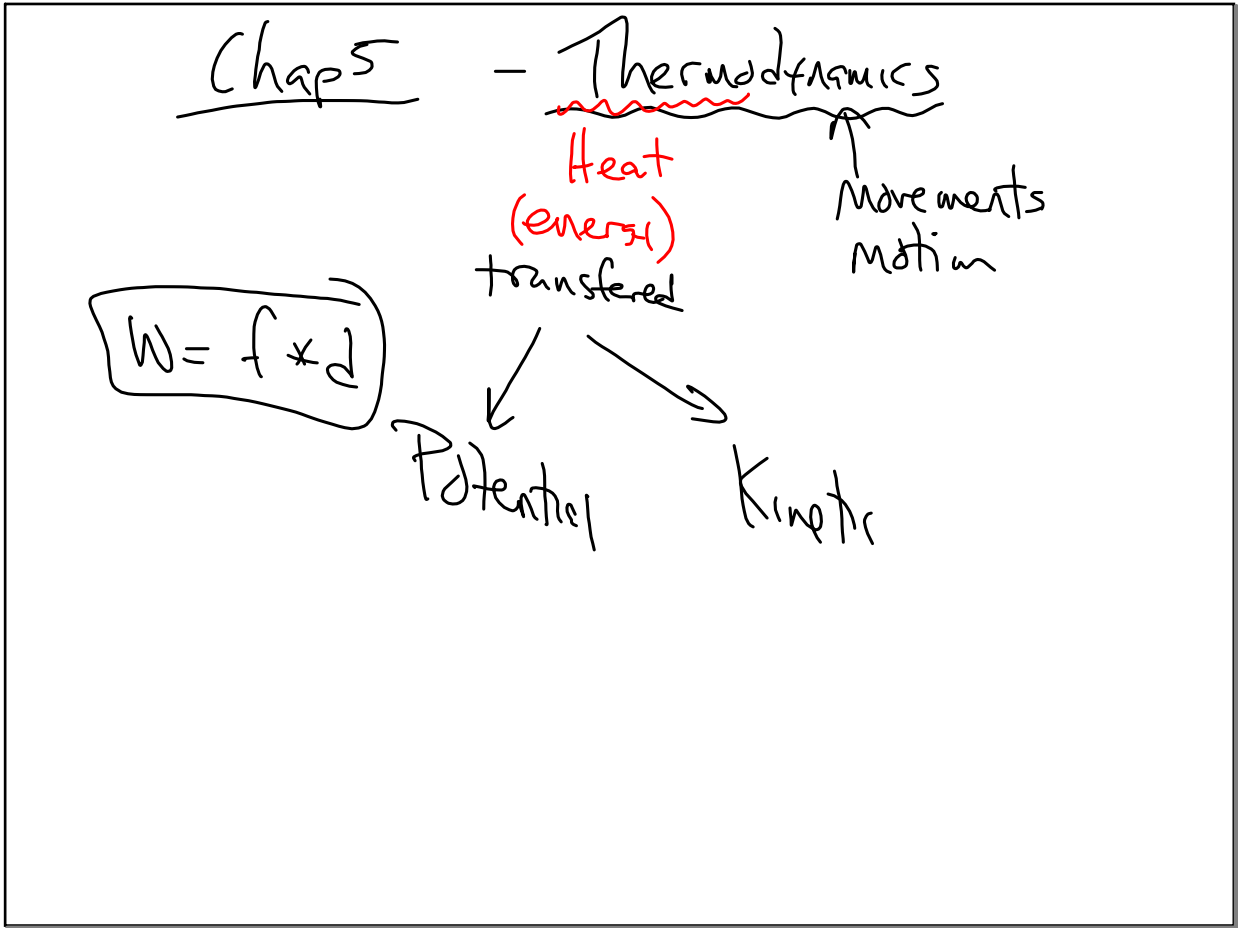
Oct 12-8:40 AM

(24)  $\frac{4g \text{ NaOH}}{100ml}$

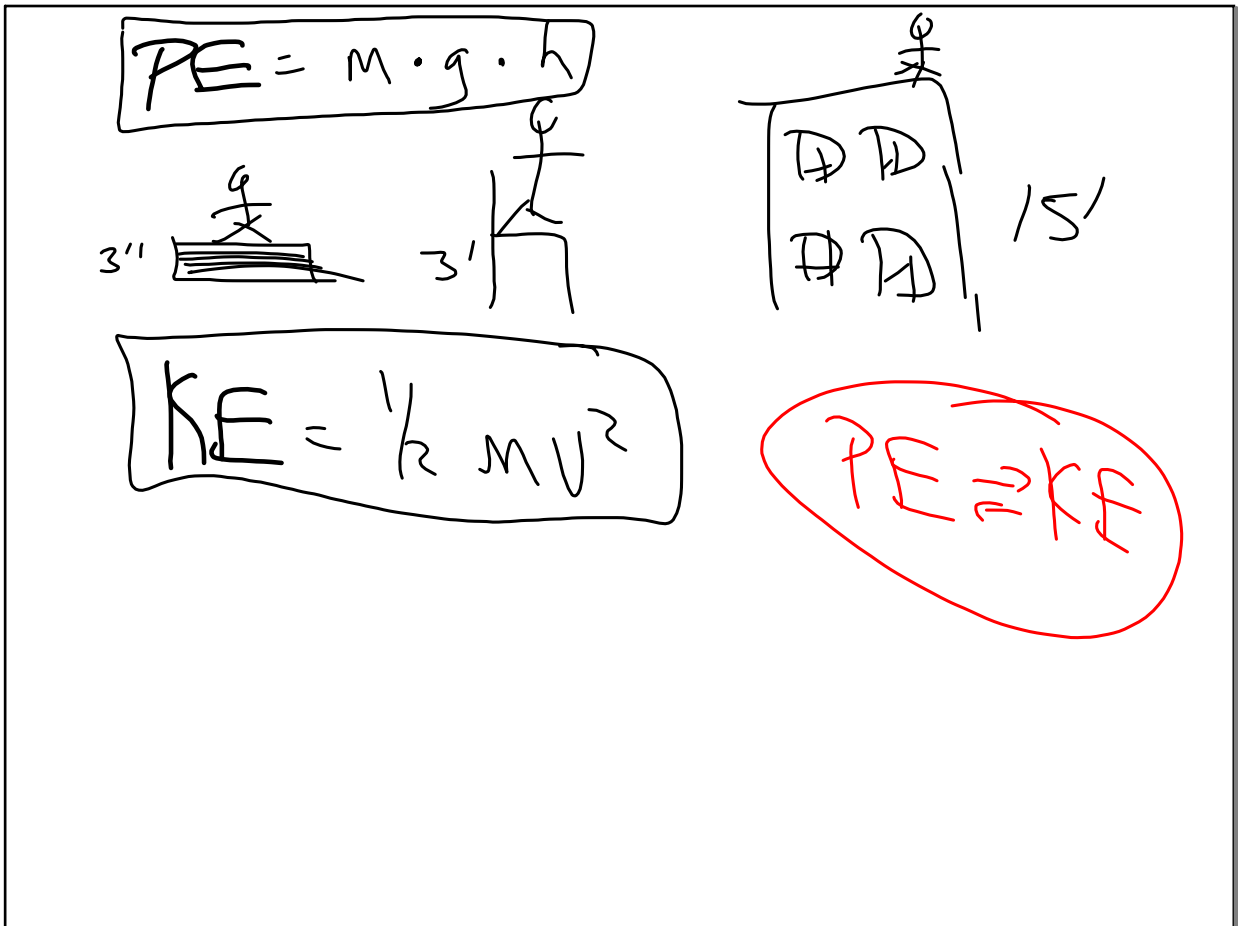
$M = \frac{\text{moles}}{l}$

4g NaOH	1 mole NaOH	= 1 M
0.1 l	40g NaOH	

Oct 12-8:46 AM



Oct 12-8:49 AM



Oct 12-9:05 AM

$$KE = \frac{1}{2} m v^2$$

$$= \frac{(Kg)}{1} \frac{M^2}{Sec^2}$$

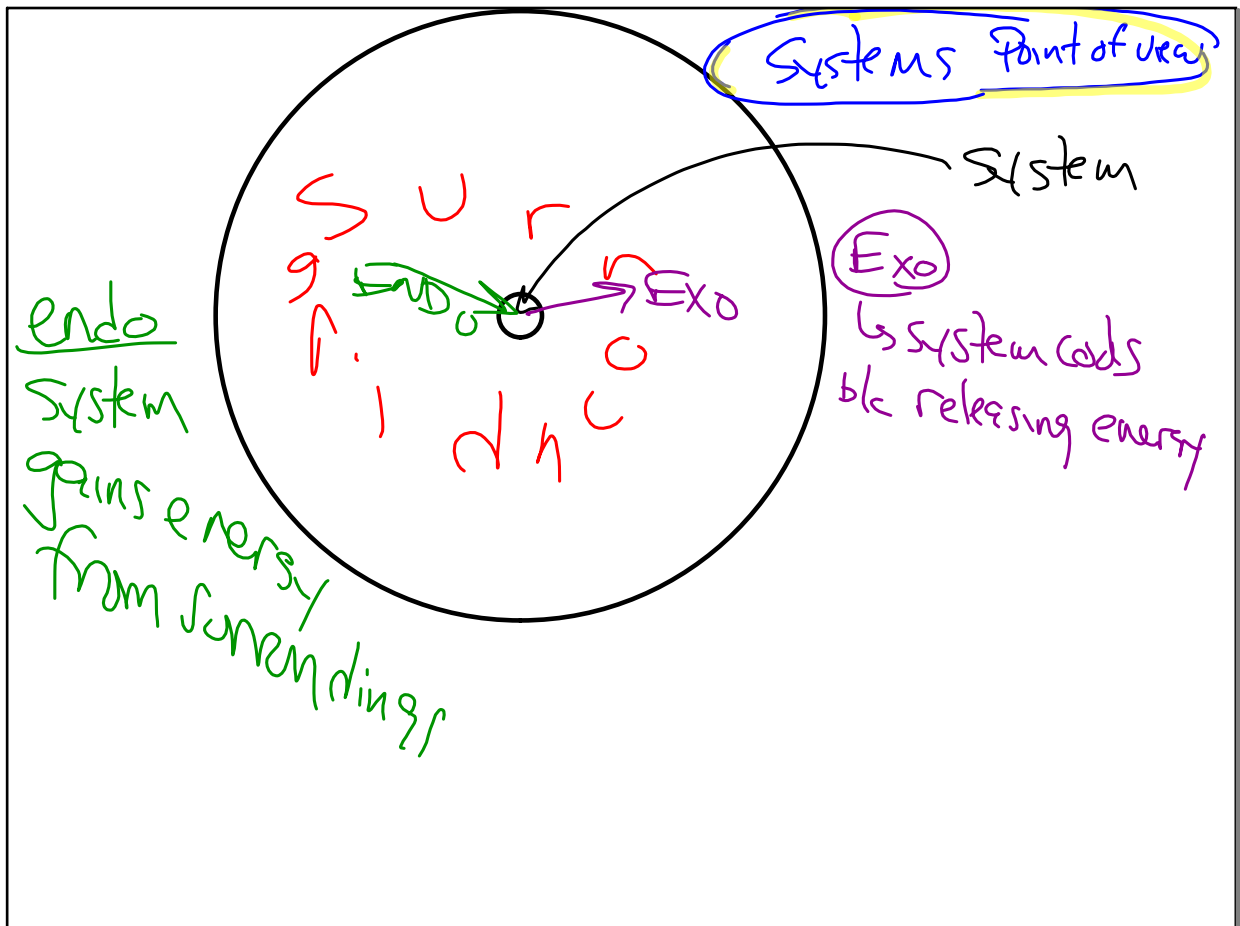
J =  $\frac{Kg \times M^2}{Sec^2}$

$$PE = mgh$$

$$= \frac{Kg}{1} * \frac{M}{Sec^2} * \frac{m}{1}$$

J =  $\frac{Kg \times m^2}{Sec^2}$

Oct 12-9:08 AM



Oct 12-9:10 AM

Spontaneous

Cool down  $\rightarrow$  EXO

Oct 12-9:13 AM