$\underline{\frac{PS4-1}{\textcircled{3}}}$ ___ g H_3PO_4 , 175 ml , $\cancel{3.5M}$

$\underline{\frac{3.5 \text{ moles } H_3PO_4}{\ell}}$

$\left| \frac{3.5 \text{ moles } H_3PO_4}{\cancel{\ell}} \right| 0.175 \cancel{\ell} \left| \frac{98 \text{ g } H_3PO_4}{1 \text{ mole } H_3PO_4} \right| = 60.025 \text{ g } H_3PO_4$

① Mole   ② GFM

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⑦ $\boxed{Na_3 PO_4} \rightarrow 3 Na^+ + PO_4^{-3}$

          1   :   3   :   1

Mole
RATIO   0.2   :   0.6 :   0.2

$NH_4 Cl + HOH \rightarrow HCl + NH_4OH$

                              SA   WB

(13) $NiCO_{3(s)} + 2HI_{(g)} \rightarrow NiI_{2(g)} + CO_{2(g)} + H_2O_{(l)}$

$$H_2CO_3 \rightarrow H_2O + CO_2$$

$$H_2SO_3 \rightarrow H_2O + SO_2$$

$NiCO_{3(s)} + 2H^+ + 2I^- \rightarrow Ni^{+2} + 2I^- + CO_{2(g)} + H_2O_{(l)}$

Oct 3-8:58 AM

(19)

Oct 3-9:08 AM

㉑   Moles A = moles B

n $\underline{M\ell}$ = n $\underline{M\ell}$
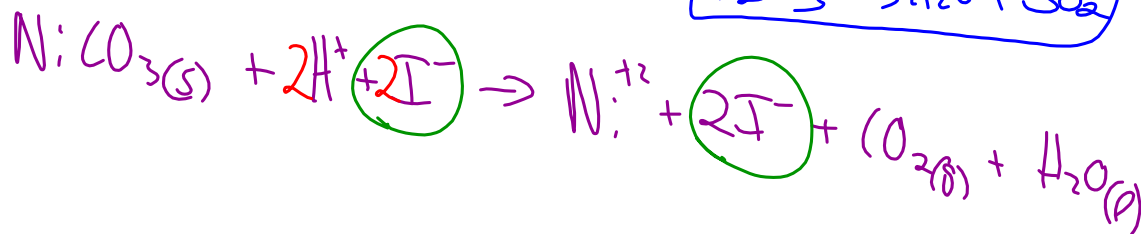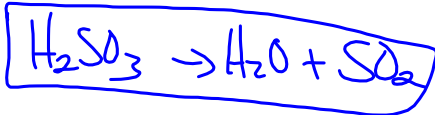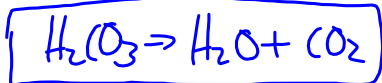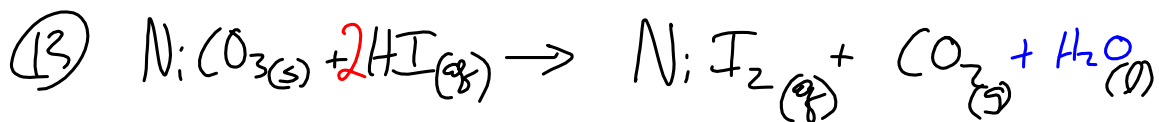
(1) $\underline{\underline{M}}$ (17.5mℓ) = (1)(0.25)(29.6 mℓ)

H $\cancel{|}$ (C$_2$H$_3$O$_2$)

Oct 3-9:10 AM

㉓   Moles start = Mole end

M * ℓ = M * ℓ

(0.35M)(500mℓ) = M (700mℓ)

Oct 3-9:12 AM