

E1 2d6

21.4 g	1 kg	100 <sup>3</sup> cm <sup>3</sup>	<del>100 cm</del>   <del>100 cm</del>
<del>cm<sup>3</sup></del>	1000 g	1 <sup>3</sup> M <sup>3</sup>	1 m   1 m

④ [Na+].[HCO3-]

ionic bond

covalent

Compound = ionic

Oct 4-7:50 AM

⑩

50 g NaOH	100 ml	1 mole NaOH
750 ml	1 l	40 g NaOH

$M = \frac{\text{moles}}{l}$

⑪

HC2H3O2      KOH

moles A = moles B

1 M l = 1 M l

(1) M (200 ml) = (1) (0.2) (31.6 ml)

Oct 4-9:03 AM

⑱ 250ml 0.1 M  $\text{Na}_2\text{SO}_4$

$\frac{0.1 \text{ mole } \text{Na}_2\text{SO}_4}{1 \text{ l}}$   $\rightarrow$  ? g

0.1 mole $\text{Na}_2\text{SO}_4$ 1 l	0.250 l	112 g $\text{Na}_2\text{SO}_4$
		1 mole $\text{Na}_2\text{SO}_4$

Oct 4-9:08 AM

⑳  $\frac{0.126 \text{ mole } \text{HClO}_4}{1 \text{ l}}$        $\frac{0.102 \text{ mole } \text{HClO}_4}{1 \text{ l}}$

1000ml	<del>0.102 mole <math>\text{HClO}_4</math></del>
<del>0.126 mole <math>\text{HClO}_4</math></del>	

Oct 4-9:14 AM

22

Oct 4-9:16 AM