

②  $N_2 + 3H_2 \rightarrow 2NH_3$

LAB  
0.0574g  
NH<sub>3</sub>

<del>0.6g H<sub>2</sub></del>	<del>1 mole H<sub>2</sub></del>	<u>2 mole NH<sub>3</sub></u>	<del>17g NH<sub>3</sub></del>
	2g H <sub>2</sub>	<del>3 mole H<sub>2</sub></del>	1 mole NH <sub>3</sub> = 3.4g NH <sub>3</sub>

% yield =  $\frac{0.0574g}{3.4} \times 100 = 1.69\%$

Oct 5-8:29 AM

③  $1 \text{ L } 0.025 \text{ M NaBr}$  contains  $2 \times 10^{-4} \text{ mole Br}^-$

~~0.025 mole NaBr~~  
1 L

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<u>1 L</u>	<del>1 mole NaBr</del>	<del>2 x 10<sup>-4</sup> mole Br<sup>-</sup></del>
<del>0.025 mole NaBr</del>	<del>1 mole Br<sup>-</sup></del>	<del>1</del>

Oct 5-9:01 AM

(FC2)

79.88% C, 20.12% H

Oct 5-9:05 AM