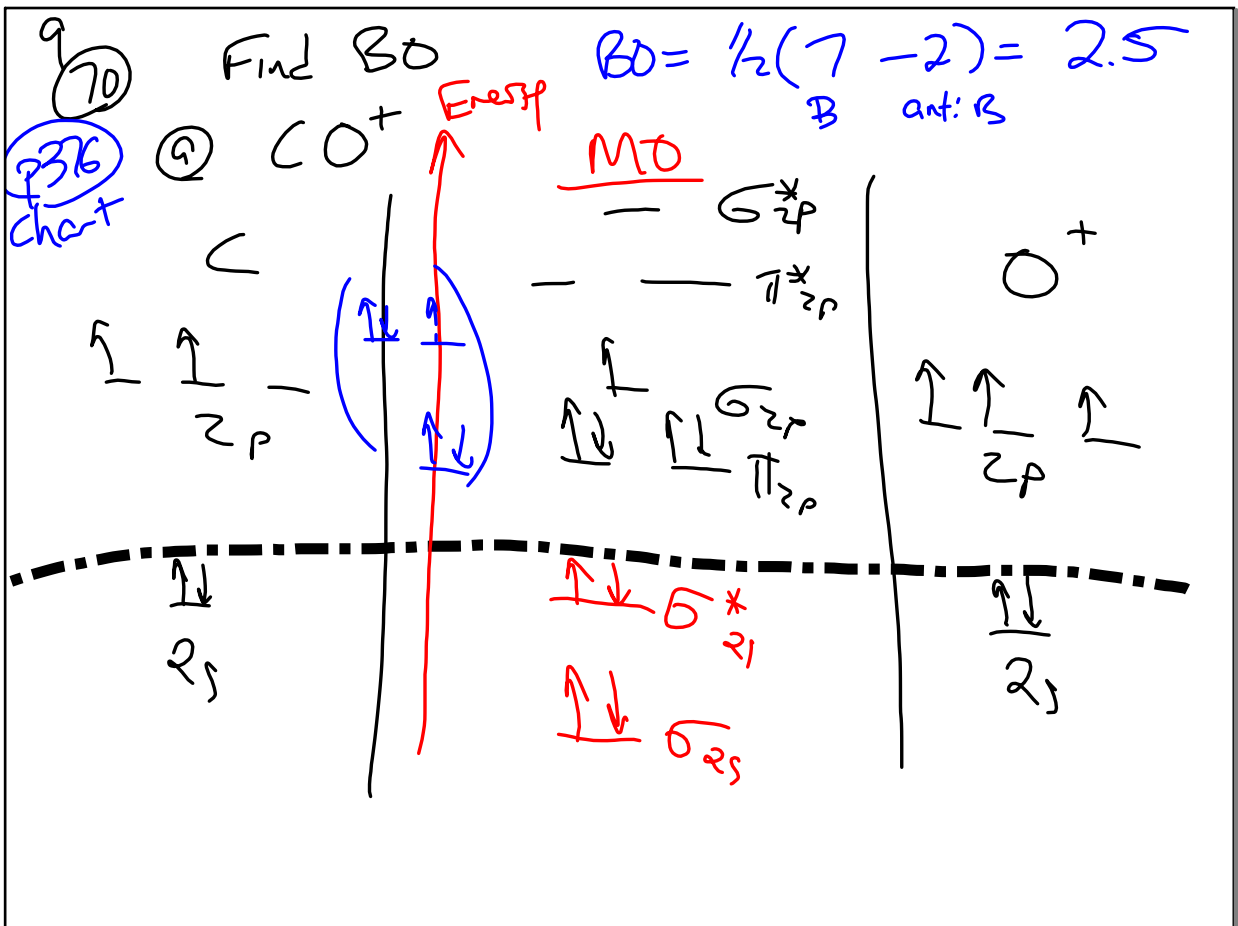
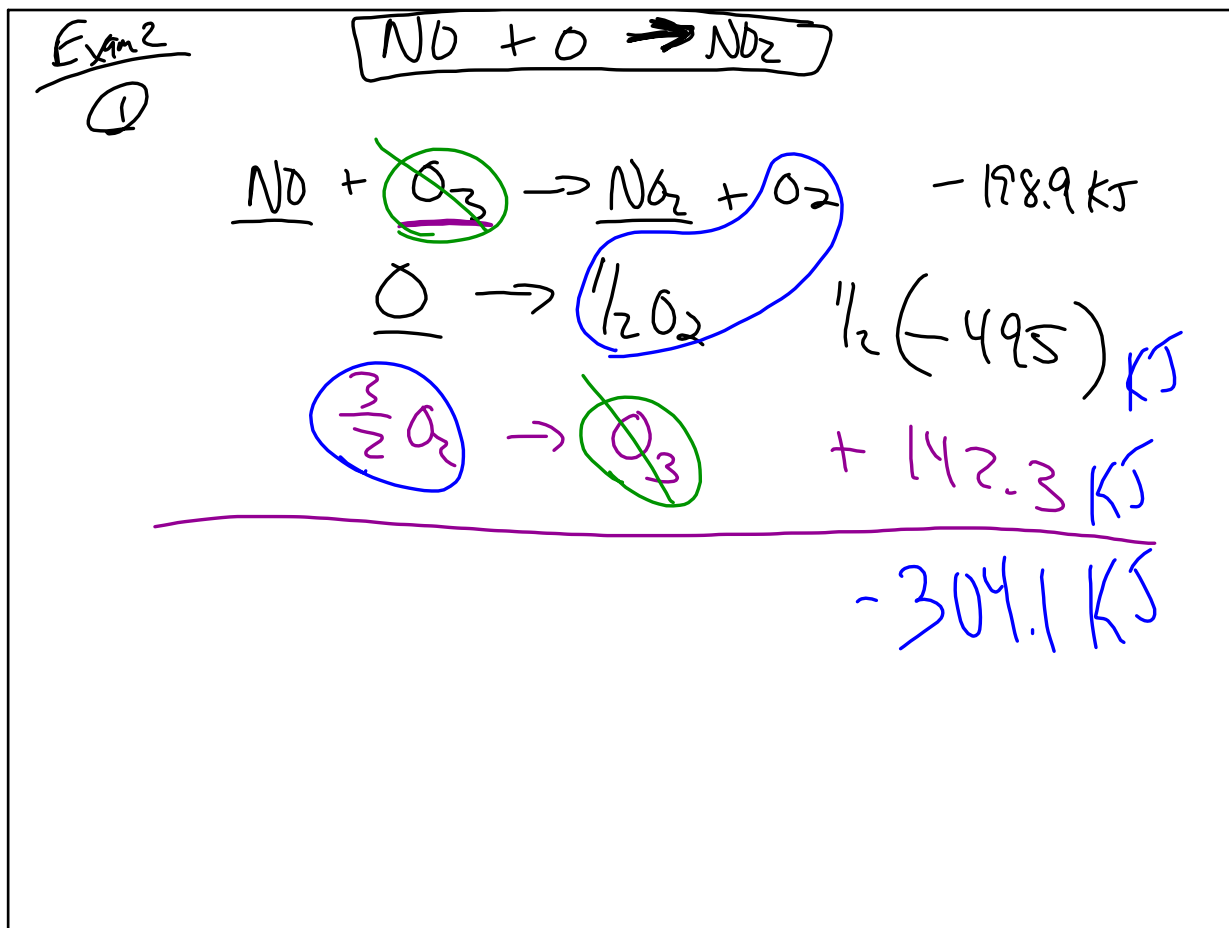
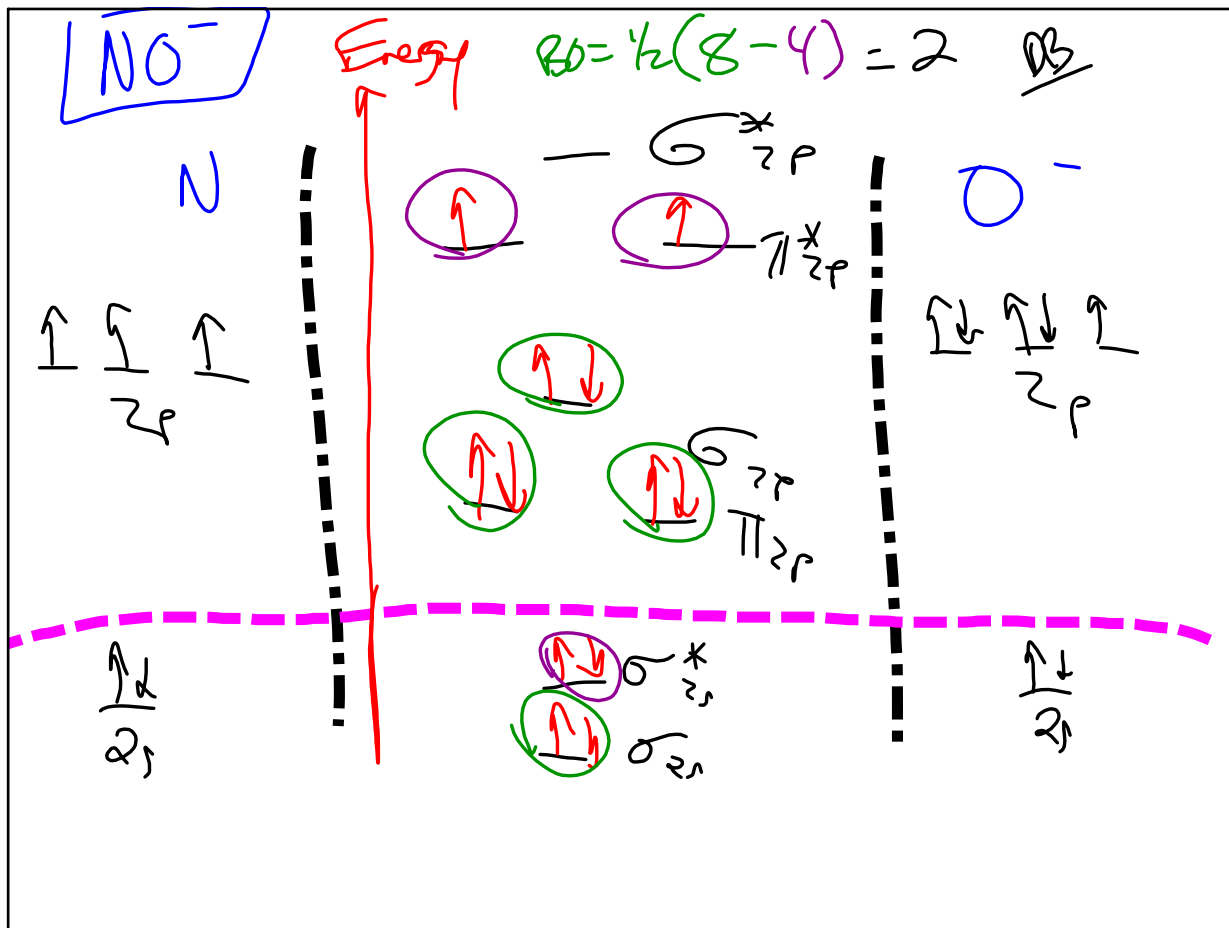


Nov 28-8:33 AM



Nov 28-8:47 AM



③ $\text{CH}_4 + 3\text{Cl}_2 \rightarrow \text{CHCl}_3 + 3\text{HCl} + 336 \text{ kJ}$

23g HCl	1 mole HCl	336 kJ
36.5g HCl	3 mole HCl	=

④ $\frac{\text{kJ}}{\text{g}}$ $\frac{15.57 \text{ kJ}}{\text{g}}$, 2.5g, $\frac{20.55 \rightarrow 23.25^\circ\text{C}}{\Delta T = 2.7^\circ\text{C}}$

$\frac{15.57 \text{ kJ}}{\text{g}}$	$\frac{2.5 \text{ g}}{\text{g}}$	2.7°C
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Nov 28-9:11 AM

⑥ $E = R_H \left(\frac{1}{n_i^2} - \frac{1}{n_f^2} \right)$ | $E = hf$

$E = R_H \left(\frac{1}{2^2} - \frac{1}{\infty} \right)$ | $E = \frac{hc}{\lambda}$

$E = R_H \left(\frac{1}{4} - \frac{1}{\infty} \right)$ | $\lambda = \frac{hc}{E}$

$E = -\frac{3R_H}{4}$

$= \frac{hc}{-\frac{3R_H}{4}}$
 $\frac{4hc}{3R_H}$

Nov 28-9:15 AM