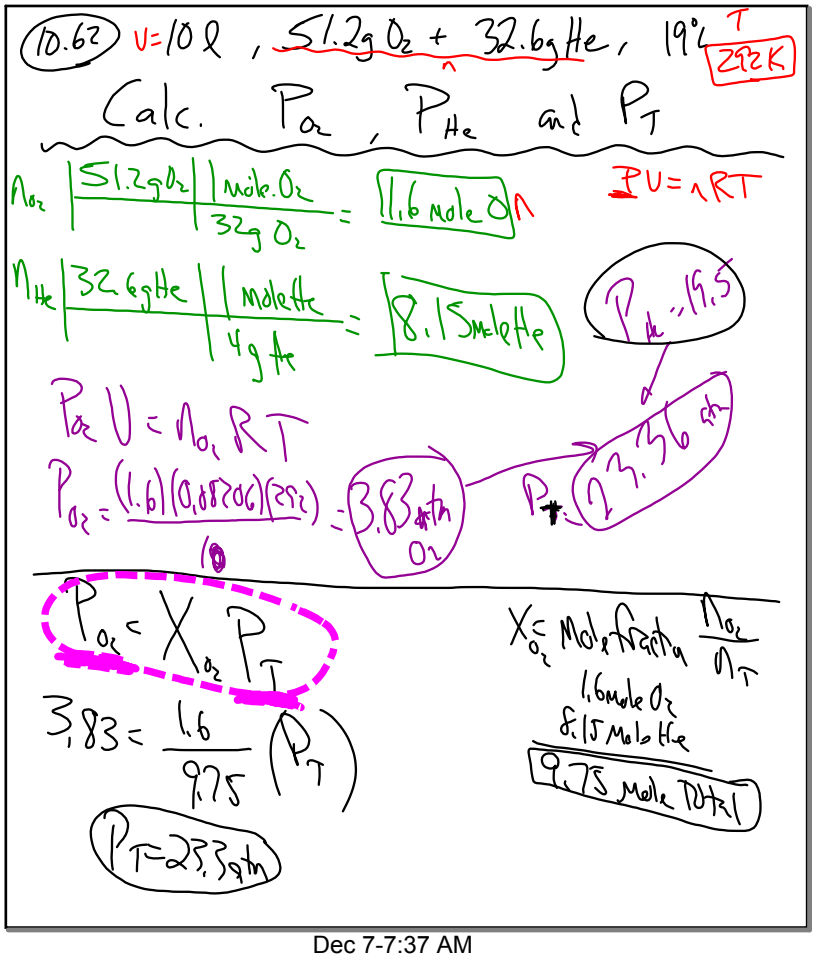
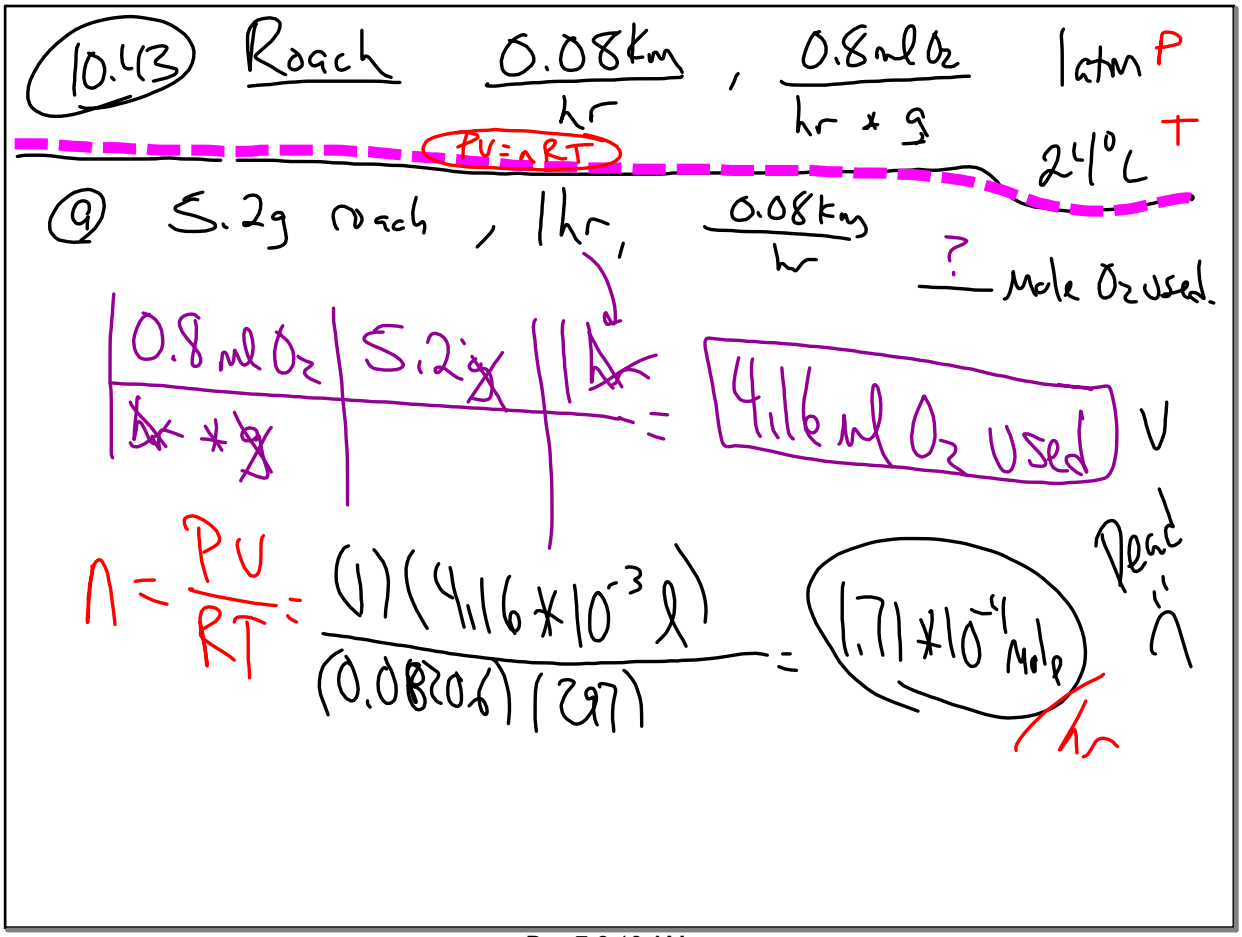


Dec 7-7:37 AM



Dec 7-8:10 AM

(b) 1 gt airtight jar

1 gt | 1 L / 1.0567 L = 0.95 L air

Consume > 20% O_2 in 48 hrs:

Air is ~21% O_2

$0.21(0.95) = 0.19\ L\ O_2$ have

g_{O_2}

Consume 1.71×10^{-4} mole / hr * 48 hrs / 1 = 0.0082 mole O_2 needed in 48 hrs.

$PV = nRT$

$(1)V = (0.0082)(0.08206)(298)$

Use $V = 0.2\ L$

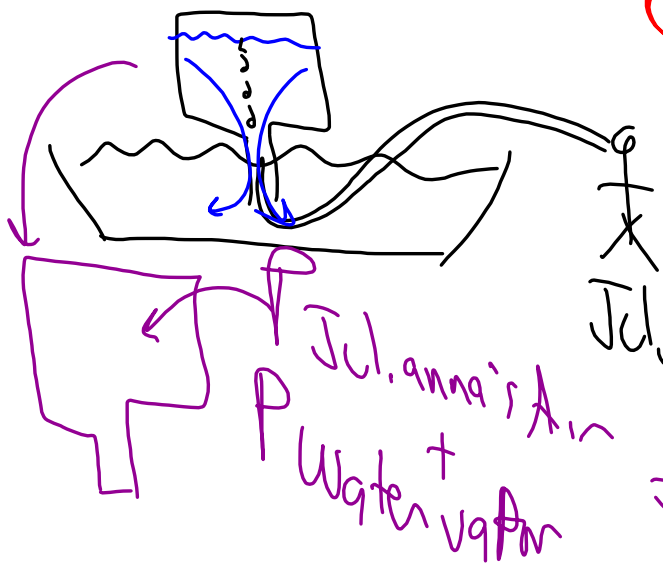
Consume 100% of O_2

Dec 7-8:22 AM

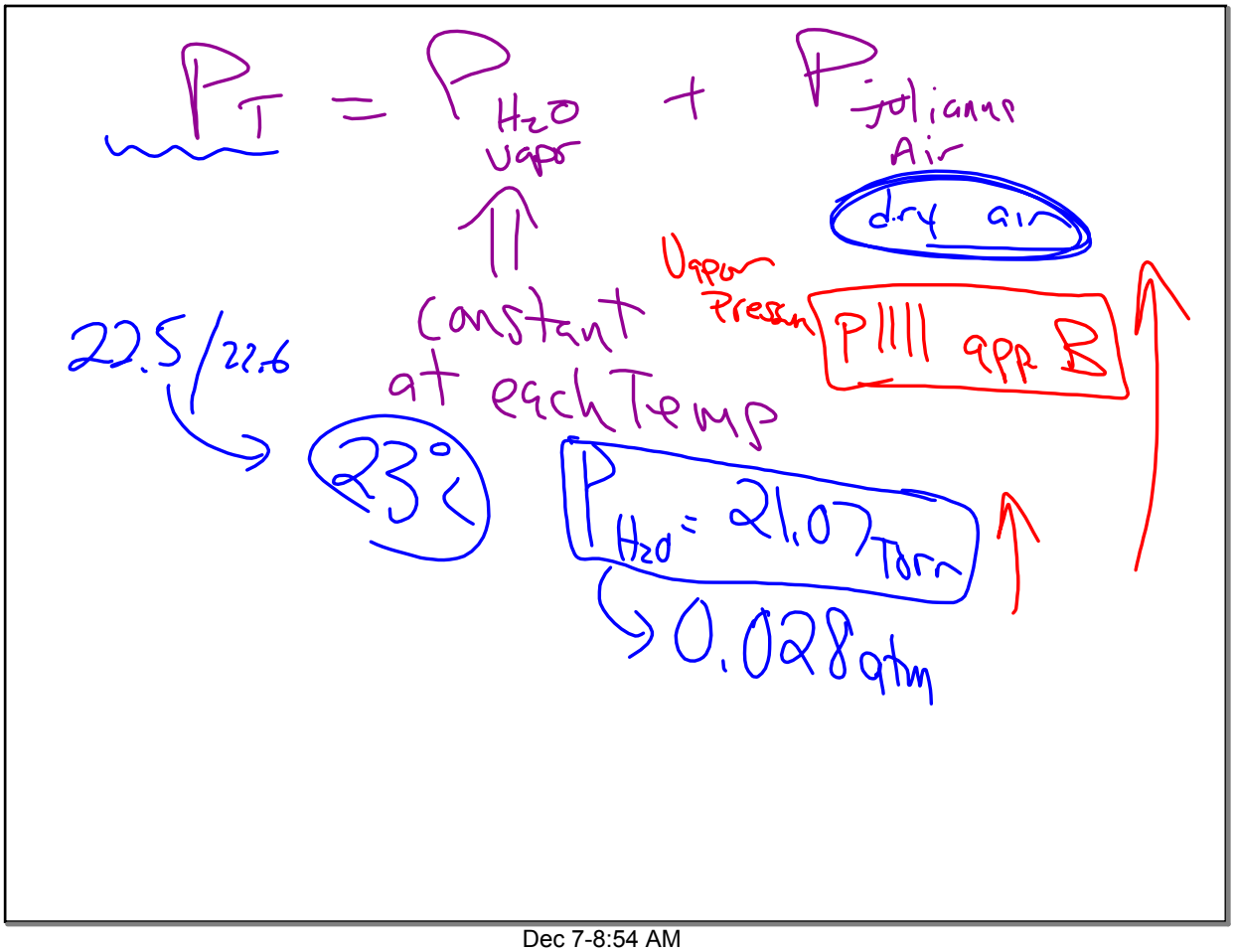
Collecting a Gas

Collecting a gas "over" water

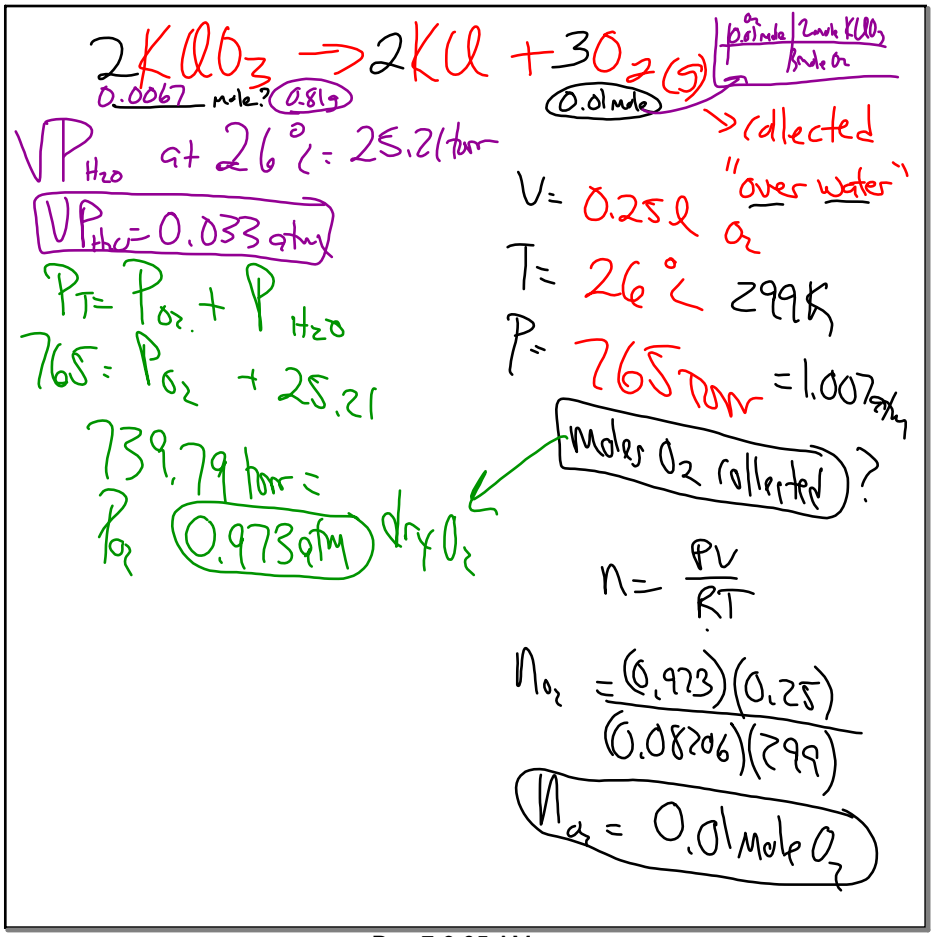
under



Dec 7-8:42 AM



Dec 7-8:54 AM



Dec 7-9:05 AM

KMT Kinetic Molecular Theory

① Fat guys run slow

② gas go in straight lines until they collide → hit → bounce

FAST → light (least mass)

Dec 7-9:14 AM

10 / 54 + 58

Dec 7-9:16 AM