

15.22 $2\text{Cl}_2(g) + 2\text{H}_2\text{O}(g) \rightleftharpoons 4\text{HCl}(g) + \text{O}_2(g)$
 $K_p = 0.0752$

Q) Full eqn $\frac{1}{K_p}$

Q) $1/2$ eqn $K^{1/2}$ or $\sqrt{K} = 0.274 \sqrt{K_p}$

Q) $K_p = K_c (RT)^{\Delta n}$
 $0.274 = K_c [(0.05206)(753)]^{1/2}$

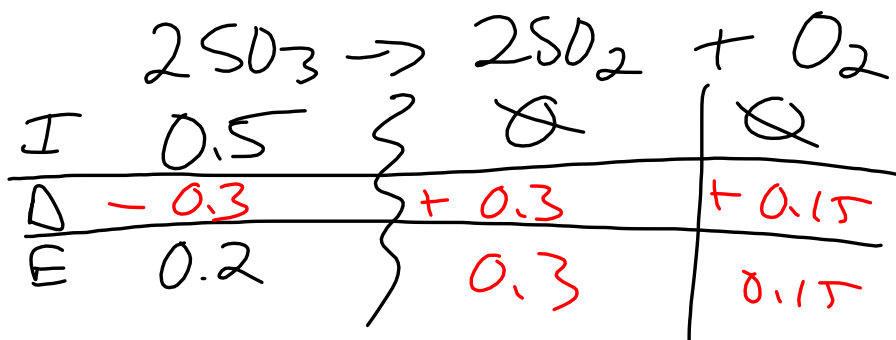
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Reaction: $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$

Initial (I)	$1 \times 10^{-3} \text{ M}$	$2 \times 10^{-3} \text{ M}$		time = ∞
Change (C)	-0.935×10^{-3}	-0.935×10^{-3}	$+1.87 \times 10^{-3}$	MOLE RATIO
Equilib (E)	0.065×10^{-3}	1.065×10^{-3}	1.87×10^{-3}	

Find $K_c = \frac{(\text{HI})^2}{(\text{H}_2)(\text{I}_2)} = \frac{(1.87 \times 10^{-3})^2}{(0.065 \times 10^{-3})(1.065 \times 10^{-3})}$
 $K_c = 50.5$

Feb 8-10:34 AM



$$K = \frac{(\text{SO}_2)^2 (\text{O}_2)}{(\text{SO}_3)^2} = \frac{(0.3)^2 (0.15)}{(0.2)^2}$$

Feb 8-10:41 AM

$$15 / 42 + 44$$

Feb 8-10:46 AM