(11.21) @ Buffer 0.12M HLac 0.11M NaLac conj.

WA salt of WA

STOPS any drastic Changes in pH

NaLac → Na^+ + OH^- (SB) + $HLac^-$ (WA)

Basic salt.

Start with adding Base

Start adding Acid to my salt

$pH = pKa + \log \frac{b}{a}$

$pH = -\log(1.4\times10^{-4}) + \log \frac{0.11}{0.12} = 3.816\ pH$

Mar 1-7:36 AM

(b) 85ml 0.13M HLac + 95ml 0.15M NaLac

180ml total = 0.180 l

Need to Recalc new M bc once Mixed Volume increase changing M.

① Convert to moles.

$\frac{1.1\times10^{-2}\ mole\ HLac}{0.18\ l}$ $\frac{1.42\times10^{-2}\ mole\ NaLac}{0.18\ l}$

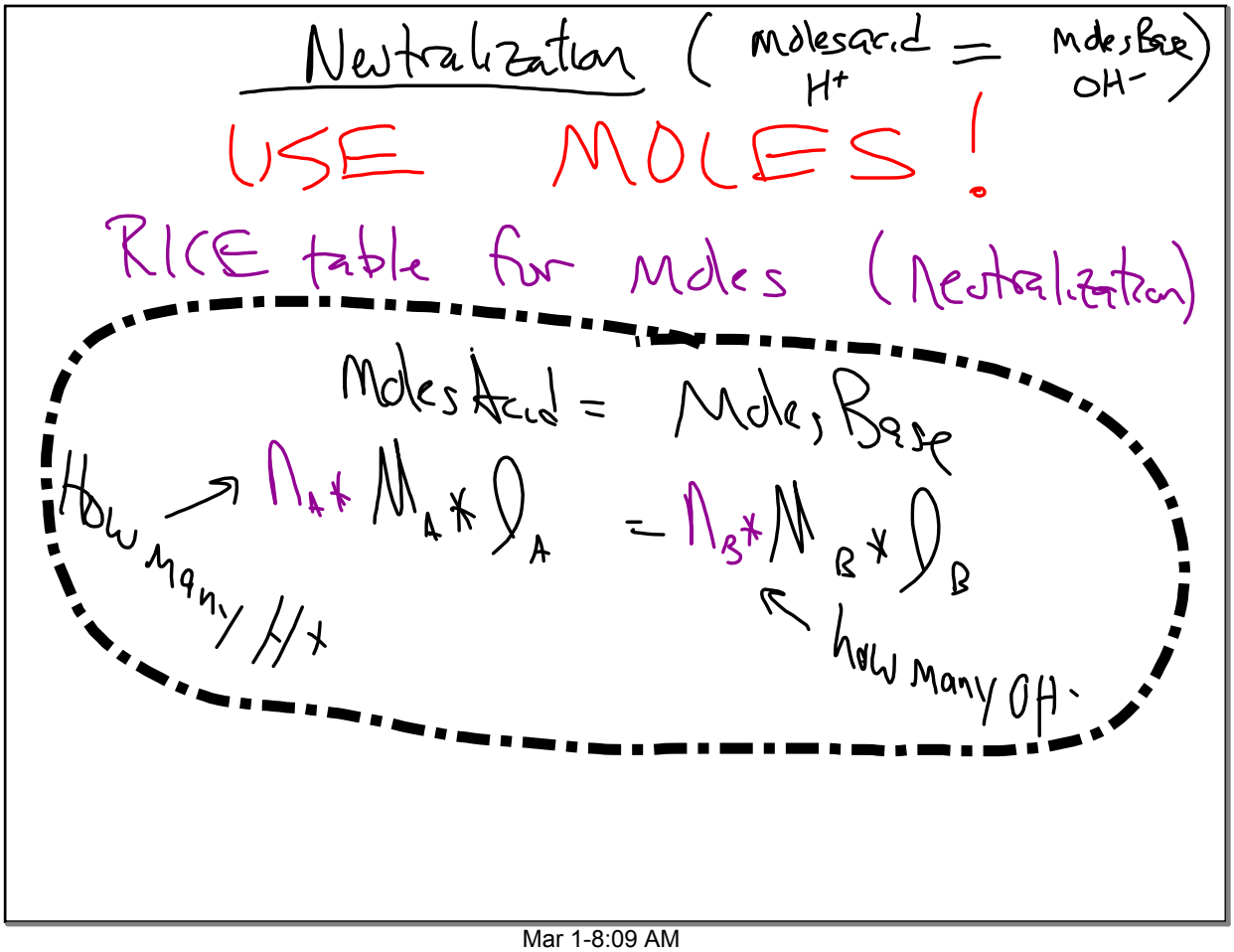
② Recalc M dividing by new Total Vol

6.11×10^{-2} M HLac 7.89×10^{-2} M NaLac

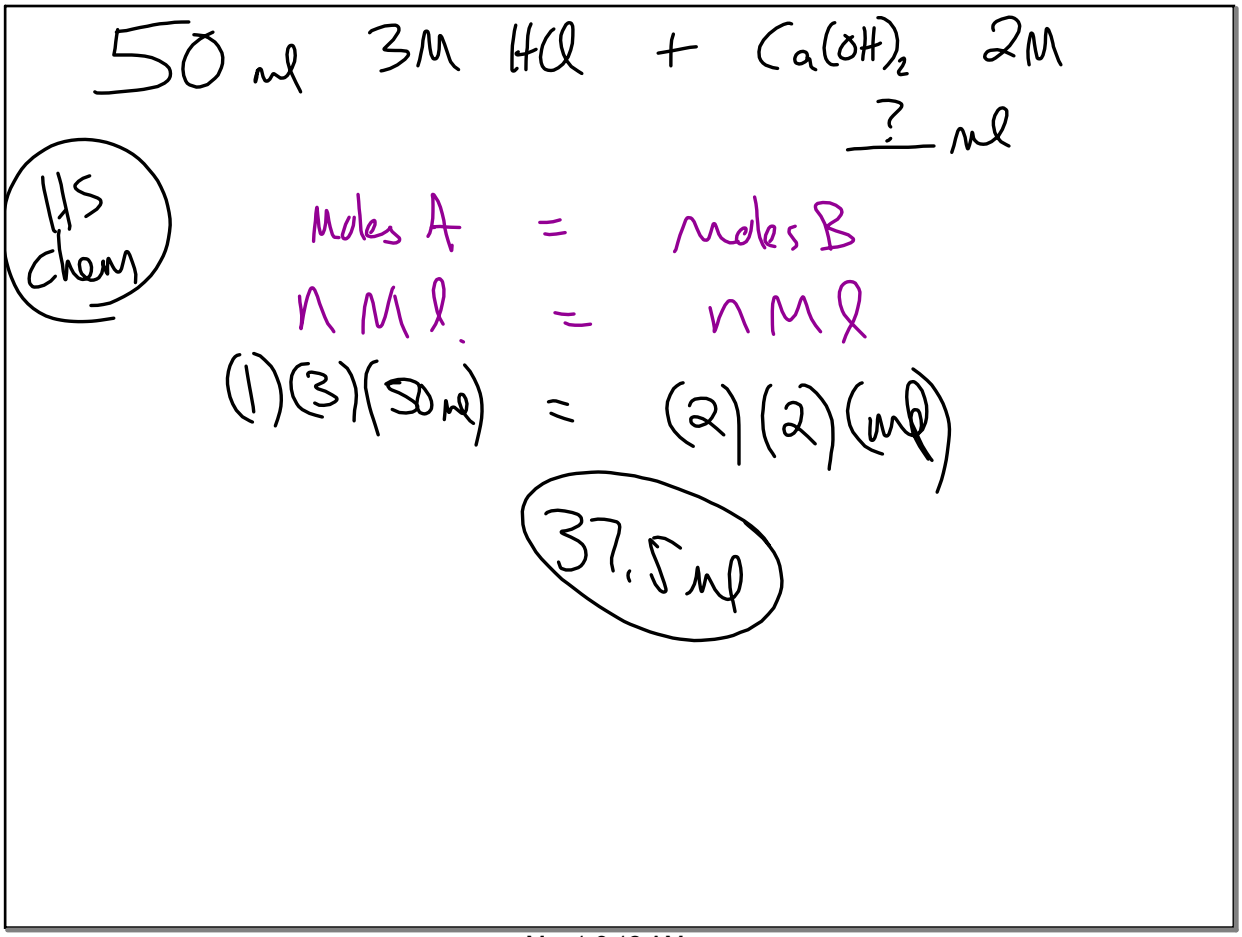
$pH = -\log(1.4\times10^{-4}) + \log \frac{7.89\times10^{-2}}{6.11\times10^{-2}}$ (pKa, b/a)

3.96

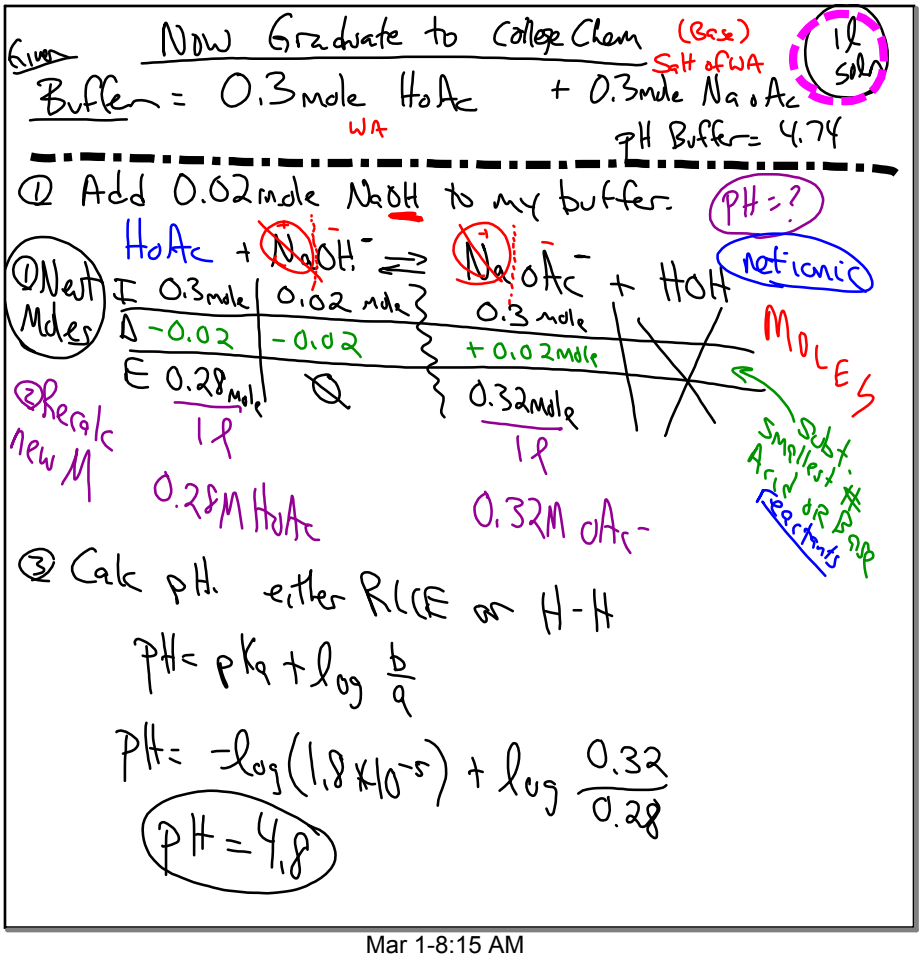
Mar 1-7:57 AM



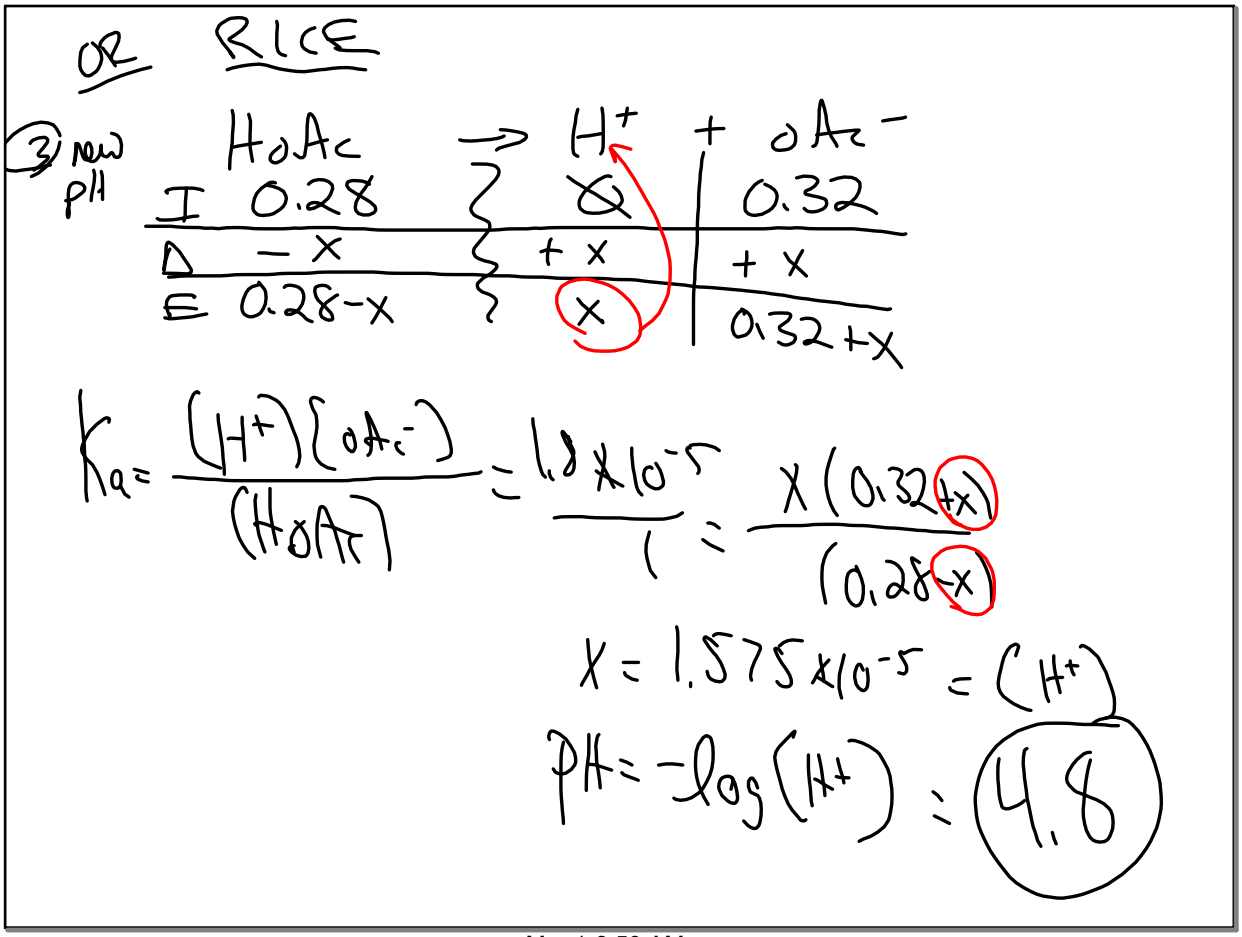
Mar 1-8:09 AM



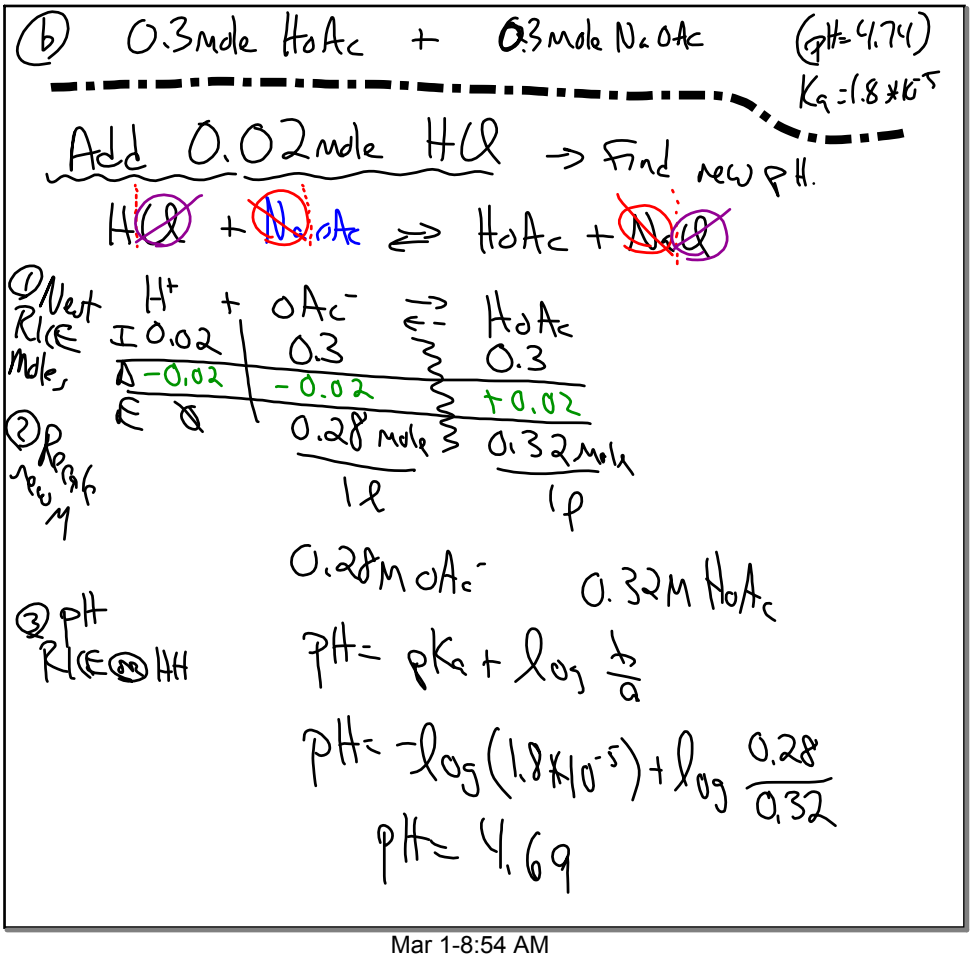
Mar 1-8:12 AM



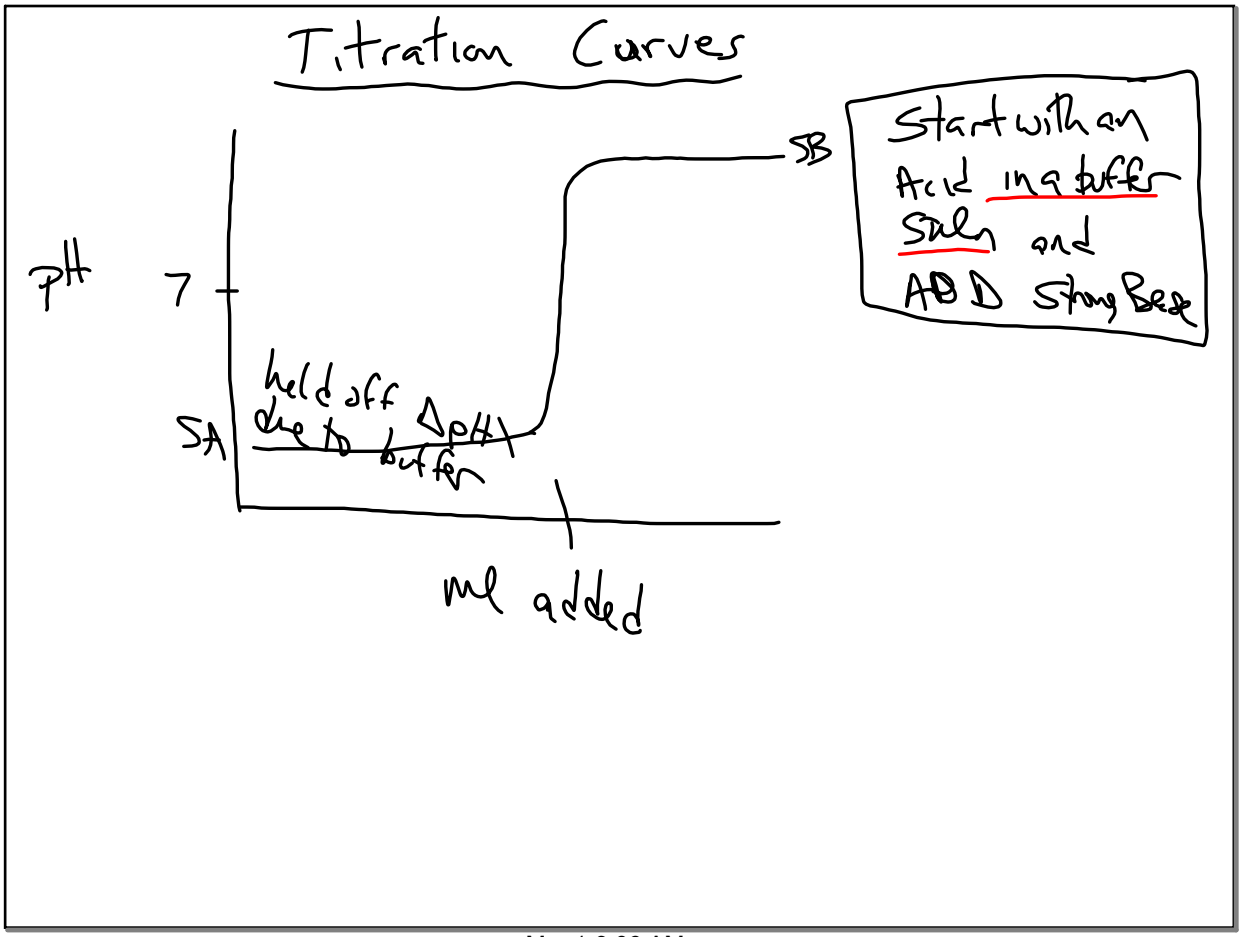
Mar 1-8:15 AM



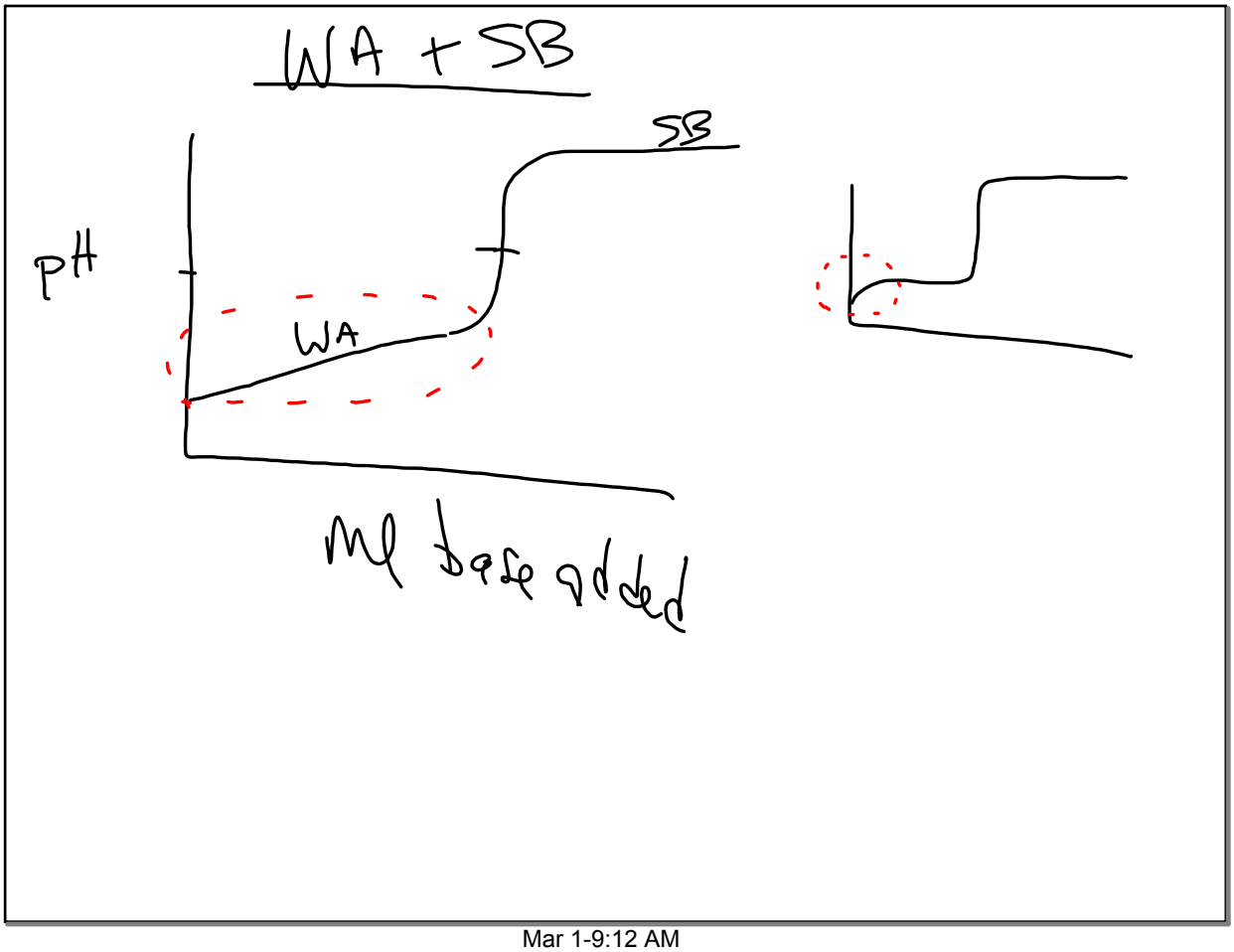
Mar 1-8:50 AM



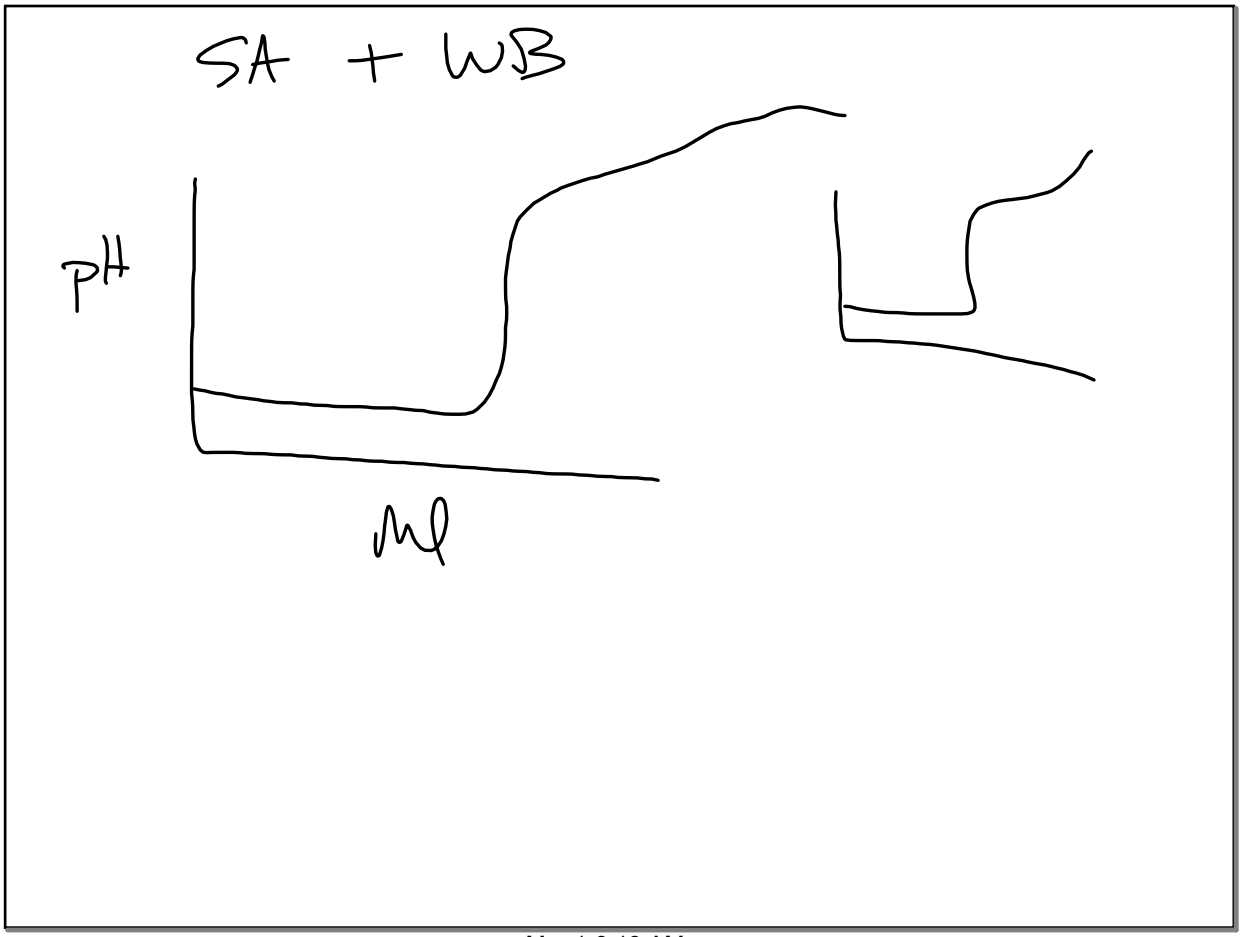
Mar 1-8:54 AM



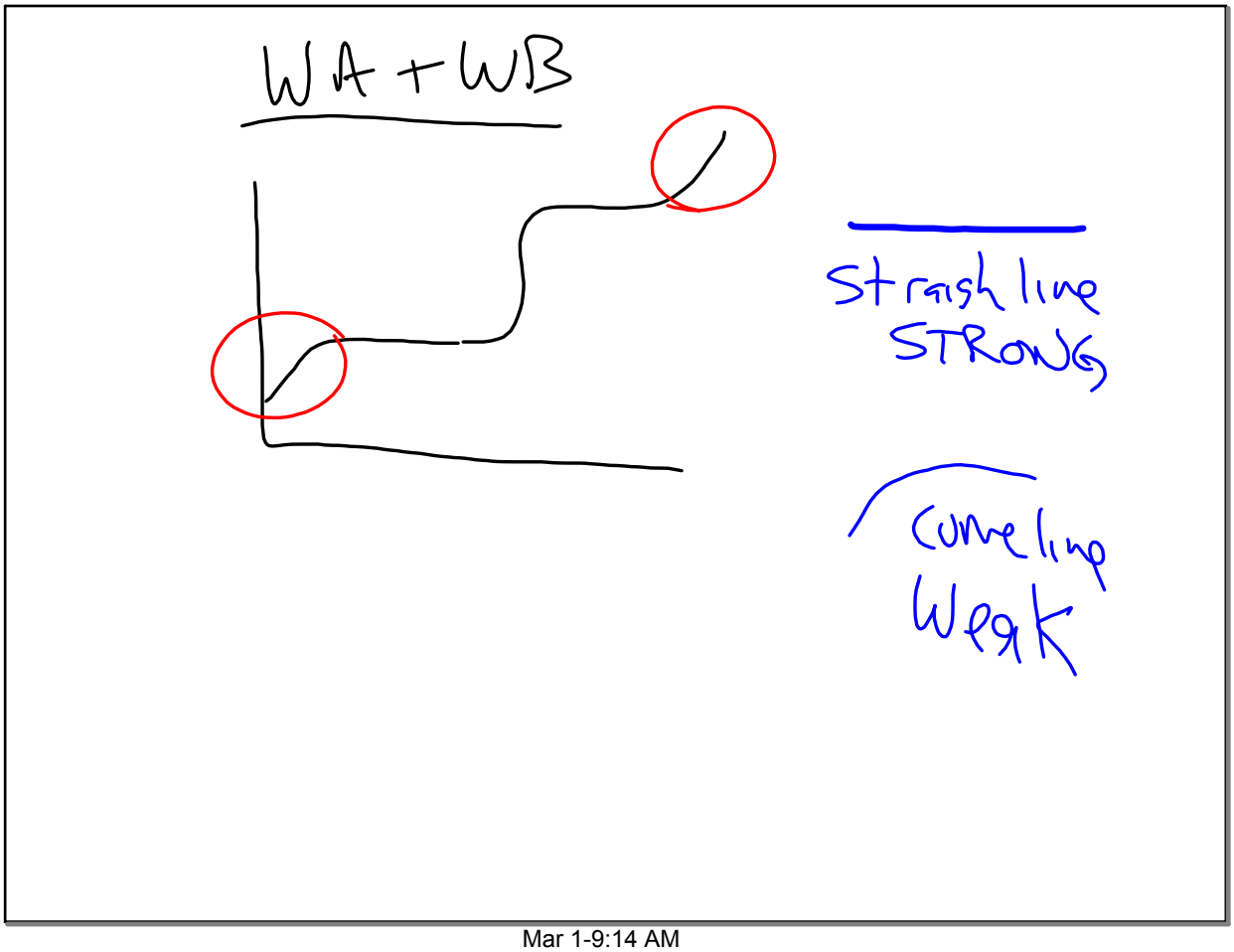
Mar 1-9:06 AM



Mar 1-9:12 AM



Mar 1-9:13 AM



Mar 1-9:14 AM

17 | 27 + 30

Mar 1-9:15 AM