

$$\textcircled{7} \quad \Delta H = n \sum \Delta H_{\text{prod}} - n \sum \Delta H_{\text{react}} \quad -196 \text{ kJ}$$

$$\textcircled{8} \quad \Delta S \quad \text{kJ} \quad \text{''} \quad \text{''} \quad -191 \text{ J}$$

$$\textcircled{6} \quad \Delta G = \Delta H - T \Delta S \quad -200,172$$

$$\textcircled{9} \quad \Delta G = -RT \ln K$$

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$$\textcircled{8} \quad E = E^{\circ} - \frac{RT}{nF} \ln Q$$

Volts

$$R = 8.314 \text{ J}$$

$$\Delta G = \Delta G^{\circ} - RT \ln Q$$

$$\downarrow$$

$$\text{kJ}$$

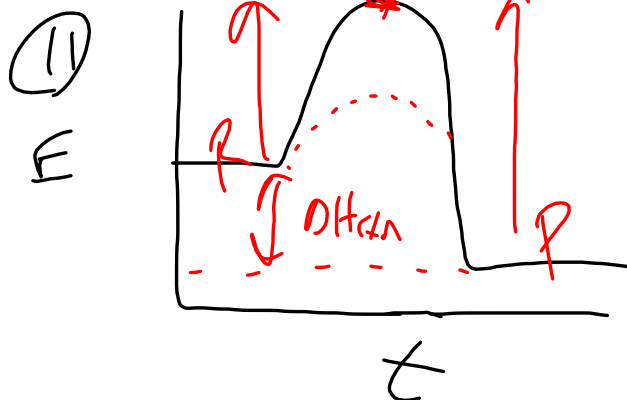
$$\downarrow$$

$$8.314 \times 10^{-3} \text{ kJ}$$

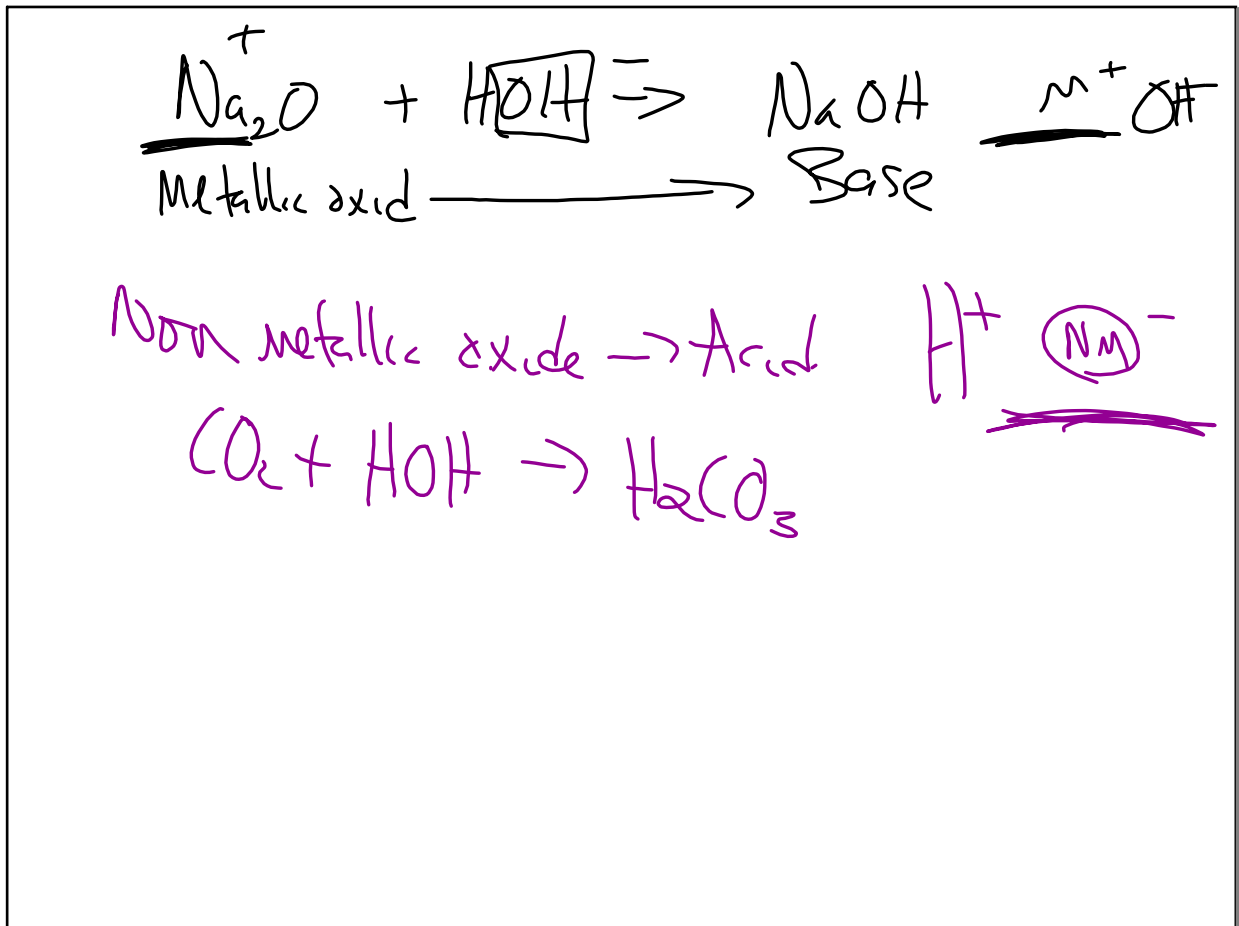
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$$\textcircled{10} \quad \Delta G = \Delta H - T\Delta S$$
$$0 = \Delta H - T\Delta S$$
$$\Delta H = T\Delta S$$
$$T = \frac{\Delta H}{\Delta S}$$

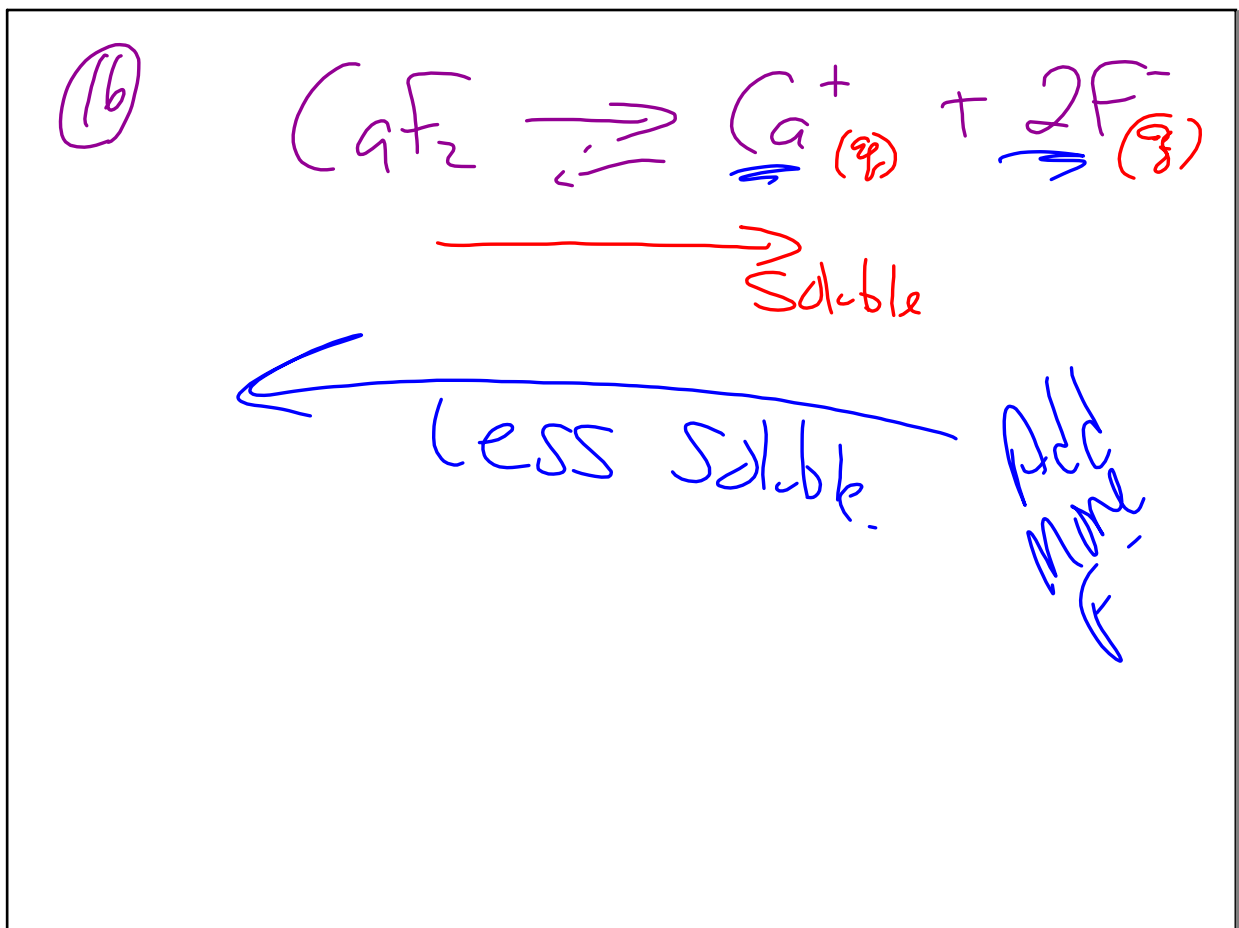
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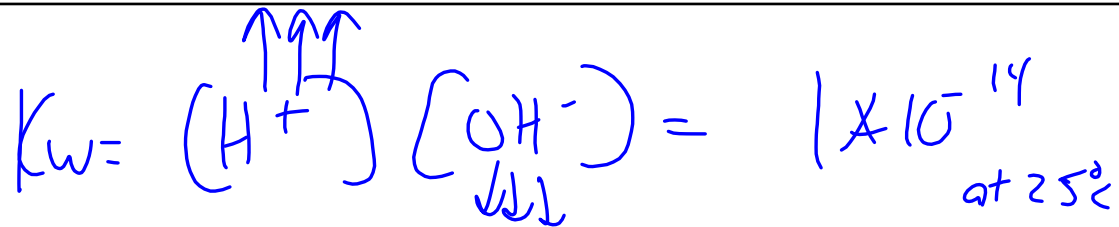
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$$pH + pOH = 14$$

$$pH = -\log(H^+)$$

$$-pH = \log(H^+) \quad \text{Move } \ominus \text{ to } H^+$$

Then anti-log.


Apr 30-9:01 AM

$$\textcircled{29} \frac{\Delta M}{\Delta \text{Sec}} = \frac{(0.0143 - 0.00701)}{26 \text{ sec}}$$

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28  $\Delta G^\circ = -nFE^\circ$

$$\Delta G_{(J)} = -6 \left( \overset{J}{96500} \right) \left( \overset{V}{1.34} \right)$$

$$\therefore -775860 \text{ J}$$


Apr 30-9:13 AM

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$$\begin{array}{l} 1 \rightarrow 3 \\ \text{(A)}^2 \quad 3^2 = 9 \\ 1 \rightarrow 2 \text{(B)} \quad 2^2 = 1 \end{array} \quad \text{Psk}$$

Apr 30-9:16 AM