

$$\textcircled{10} \quad 0.2740(300.991) + 0.2790(303.944) + \\ + 0.4470(308.963) =$$

Weighted Avg

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$$\textcircled{12} \quad 2 \text{ g Al} = \text{_____ atoms Al}$$

$$\frac{2 \text{ g Al} \quad | \quad 1 \text{ mole Al} \quad | \quad 6 \times 10^{23} \text{ atoms Al}}{26.98 \text{ g Al} \quad | \quad 1 \text{ mole Al}} =$$

Sep 25-9:02 AM

Atoms N

(16)

10g NH₄NO₃	1mole NH₄NO₃	2mole N
80	1mole NH₄NO₃	1mole N

mole ratio!

6×10^{23} atoms N

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(19)

C $\frac{40g}{12g} = \frac{3.33 \text{ mole C}}{3.33} = 1$

H $\frac{6.7gH}{1gH} = \frac{6.7 \text{ mole H}}{3.33} = 2$

O $\frac{53.3gO}{16gO} = \frac{3.33 \text{ mole O}}{3.33} = 1$

emp

C H₂ O

Sep 25-9:13 AM

PS 3-2
(1996)

8, 12, 14, 19, 24

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