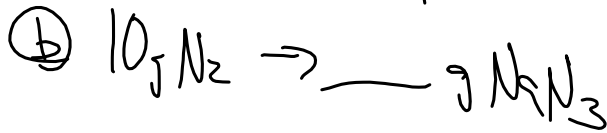
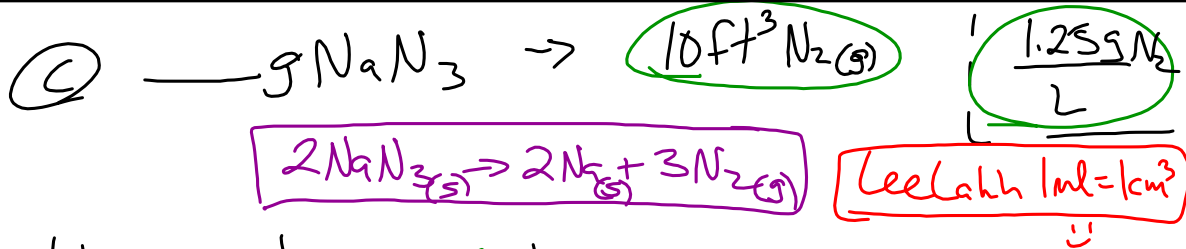


1.5 mole NaN_3	3 mole N_2	= 2.25 mole N_2
	2 mole NaN_3	



10g N_2	1 mole N_2	2 mole NaN_3	65g NaN_3	= 15.5g NaN_3
	28g N_2	3 mole N_2	1 mole NaN_3	

Sep 28-7:37 AM



1.25g N_2	10 ft ³ N_2	1L	1L	(2.5) ³ cm	(12) ³ in
N_2		1000	cm ³	1 inch ³	1 ft ³

1.25g N_2	1 mole N_2	2 mole NaN_3	65g NaN_3
	28g N_2	3 mole N_2	1 mole NaN_3

Sep 28-8:42 AM

Kool - Aid CAB (MOLARITY)

100 ml = 4th ribbed line from bottom

(M) Molarity - Measure of concentration of solute in the solution \leftrightarrow (Solute + Solvent)

$$M = \frac{\text{Moles of solute}}{\text{liters of solution}}$$

$$\text{Moles} = M * l$$

Sep 28-8:54 AM

Make 250ml 1M CuSO_4 (3 CuSO_4)

~~0.250l~~

$64 + 32 + 4(16) = 160 \text{ g/mole}$

$$M \Rightarrow \frac{1 \text{ mole } \text{CuSO}_4}{1 \text{ l}}$$

1 mole CuSO_4	1 l	$\frac{160 \text{ g } \text{CuSO}_4}{1 \text{ mole } \text{CuSO}_4} = 160 \text{ g } \text{CuSO}_4$
0.250l	0.250l	$40 \text{ g } \text{CuSO}_4$

Sep 28-9:05 AM

~~10ml~~ ~~0.5M NaCl~~ 0.5 mde NaCl
 ↓
 0.01 l 1 l

0.5 mde NaCl	0.01 l	505 NaCl
1 l		1 mde NaCl

Sep 28-9:13 AM

Dilution

Mdes start = mde end.

$M \times l = M \times l$

HW 4.62

~~$\frac{M}{l} = \frac{\text{Moles}}{l}$~~

Sep 28-9:15 AM