

F2017 (19)

Moles Acid = Moles Base

$$n \times M \times l = n \times M \times l$$

$$\frac{(2) (0.0875) (\text{Some})}{0.115} = \frac{(1) (\cancel{0.115}) (\text{ml})}{0.115}$$

Moles = $n \times l$

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(22)

0.4445	500g	36%
g		
C	M	DT

(21) $\text{Na}_3\text{PO}_4 \rightarrow 3\text{Na}^+ + \text{PO}_4^{-3}$

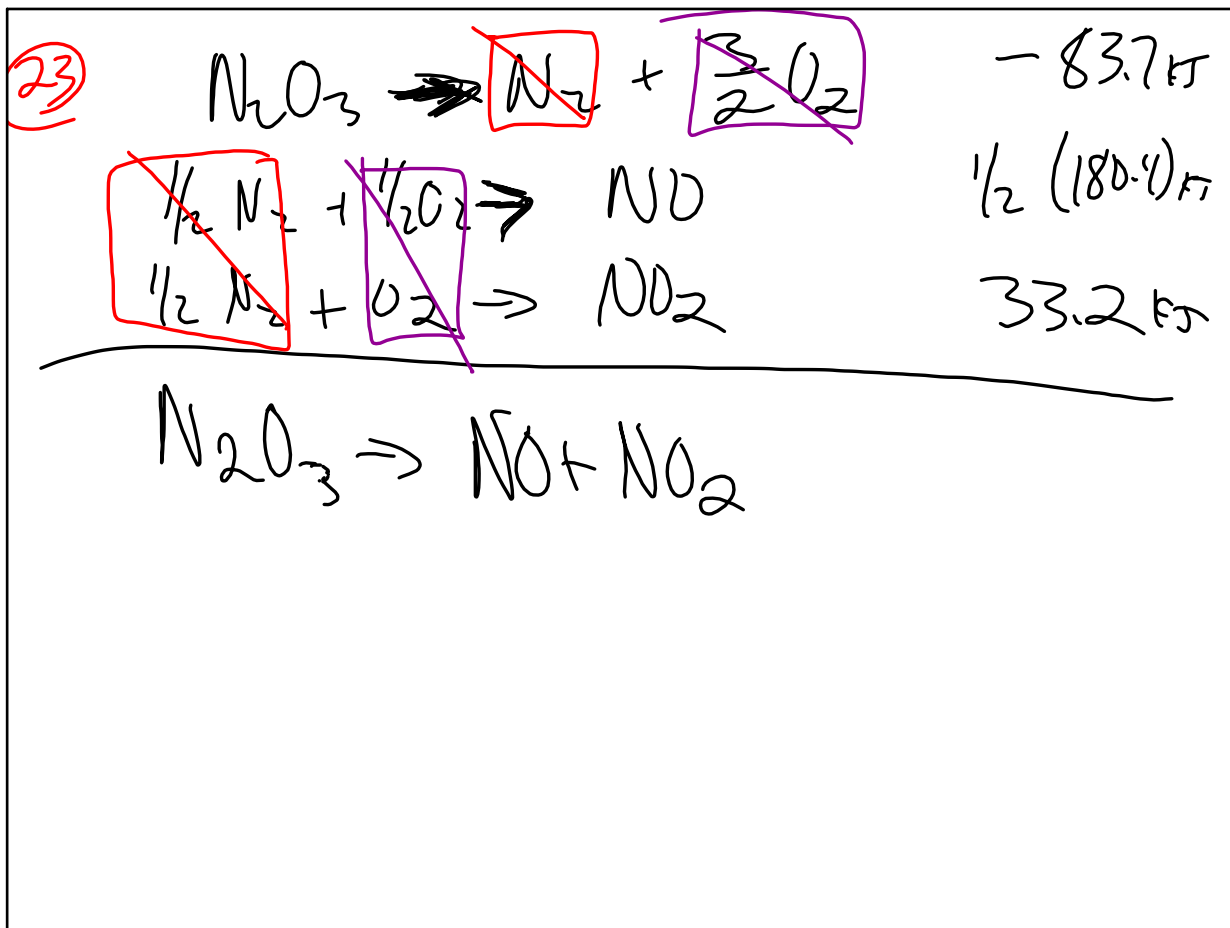
4.57 L
0.847 M

3(0.847)

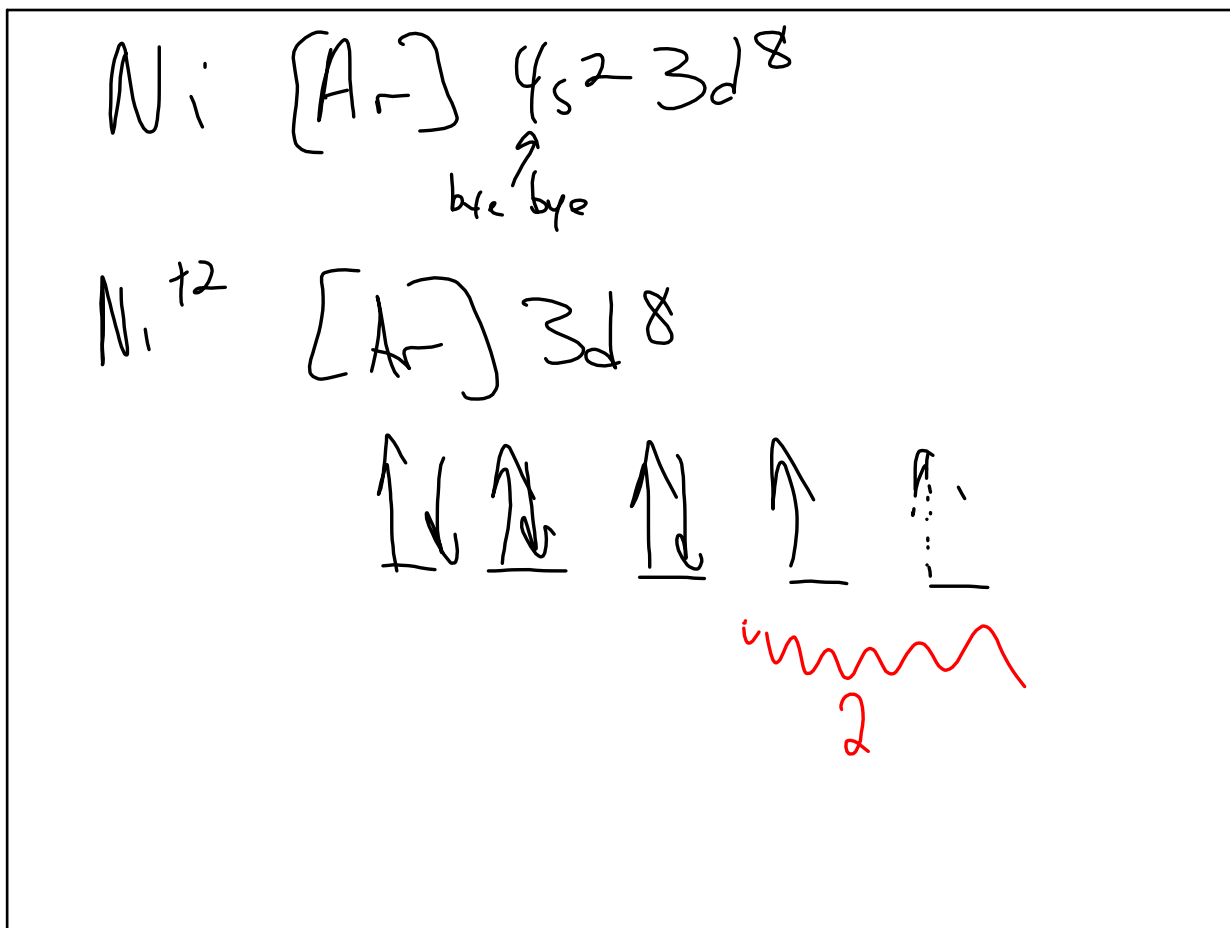
0.847

Mole Ratio \times $3 \times$ ~~\times~~

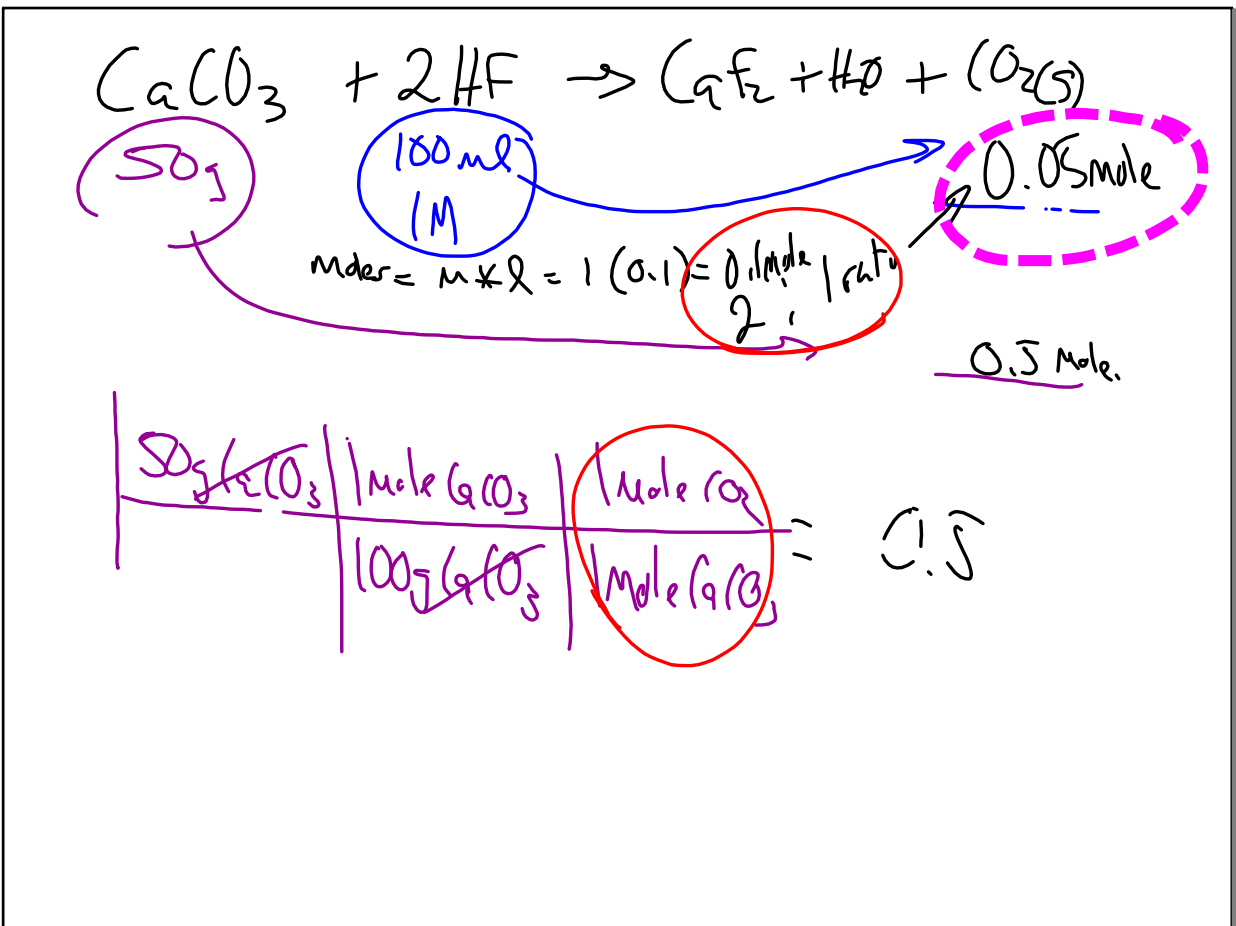
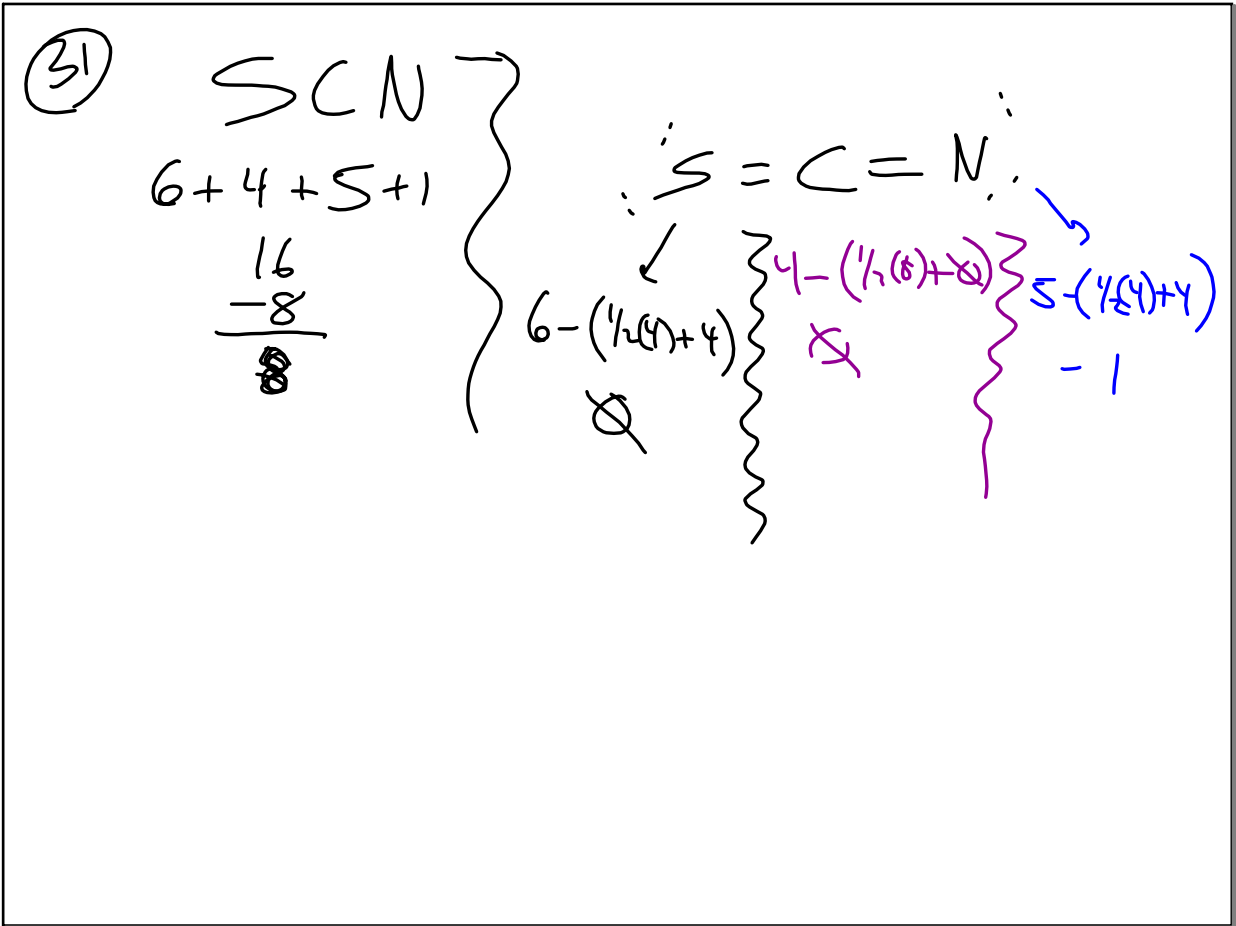
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$$\text{Water Heat gained} = \text{Heat lost MeCl}$$

$$mC\Delta T = mC\Delta T$$

$$(50)(4.18)(23-18) = (95)c(75-23)$$

$$c = 0.211 \text{ J/g}^\circ\text{C}$$

Jan 11-8:18 AM

Exm 4

$$\text{(15)} \quad \frac{1.12\text{g}}{\text{mL}} * \frac{100\text{mL}}{1} = 112\text{g solvent}$$

$$\text{Molality} = \frac{\text{Moles solute}}{\text{kg solvent}} = \frac{\frac{0.195\text{g}}{286\text{g/mol}}}{0.112\text{kg}}$$

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(16) $\Delta T = (K \times M) i$
 (9) $2.56 = (1.86 \times M) i$
 $m = 1.376 m$

(6) $M = \frac{\text{Moles solute}}{\text{Kg Solvent}}$
 $\frac{1.376}{1} = \frac{\text{Moles solute}}{0.045 \text{ Kg}}$
 0.0619 mole

$\frac{\text{Moles}}{1} = \frac{g}{\text{MW}}$ $\text{MW} = \frac{g}{\text{Mole}} = \frac{7}{0.0619}$
 113

Jan 11-8:41 AM

HW 2016 # 1-20

Jan 11-8:45 AM