$$\frac{M}{1} = \frac{\text{moles solute}}{\text{l of solution}}$$

solute + solvent.

$$6M = \frac{6 \text{ moles } (C_6H_{12}O_6}{l}$$

Oct 1-7:44 AM

Prepare 250ml of 1.5M Copper(II) sulfate.

0.250l

Cu^{+2} SO_4^{-2}

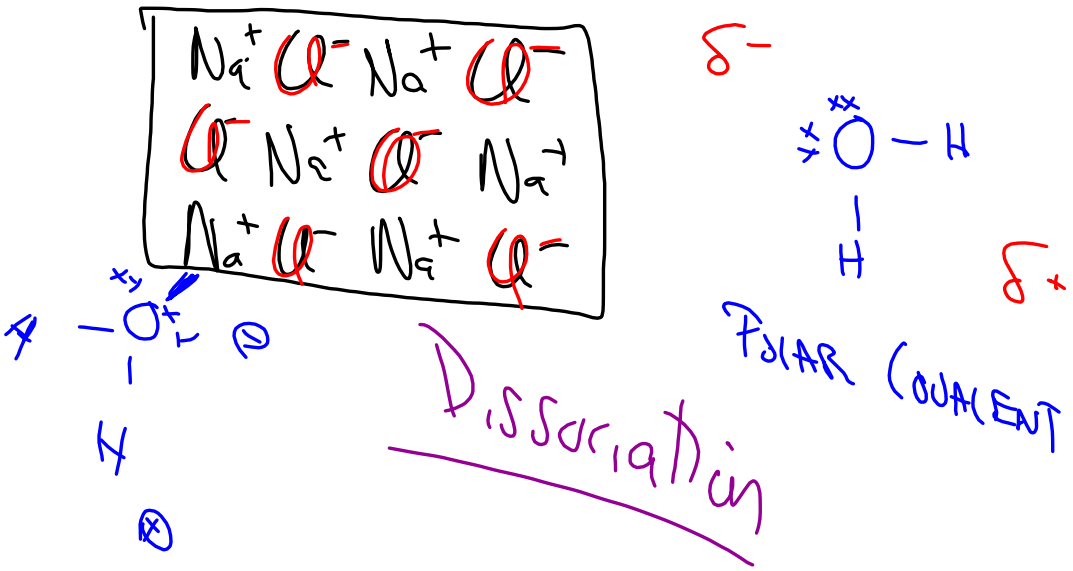
60 grams

1.5 moles $CuSO_4$ / 1l

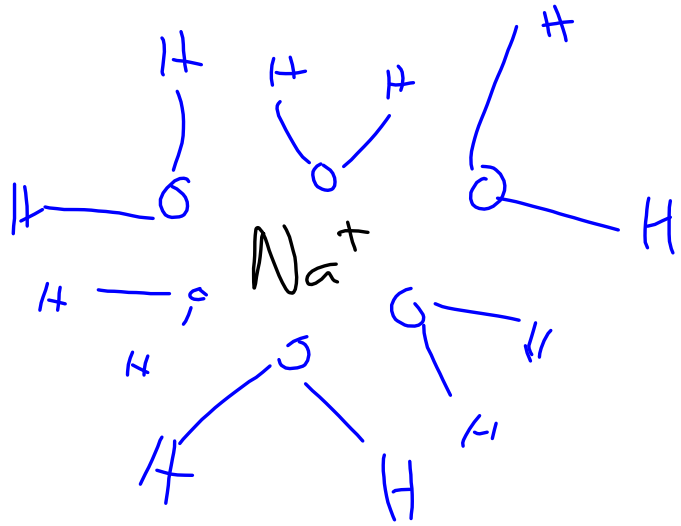
$$\frac{1.5 \text{ mole } CuSO_4}{1\,l} \left| \frac{160\,g\,(CuSO_4}{1 \text{ mole } CuSO_4} \right| 0.250l = 60g\,(CuSO_4$$

Oct 1-8:22 AM

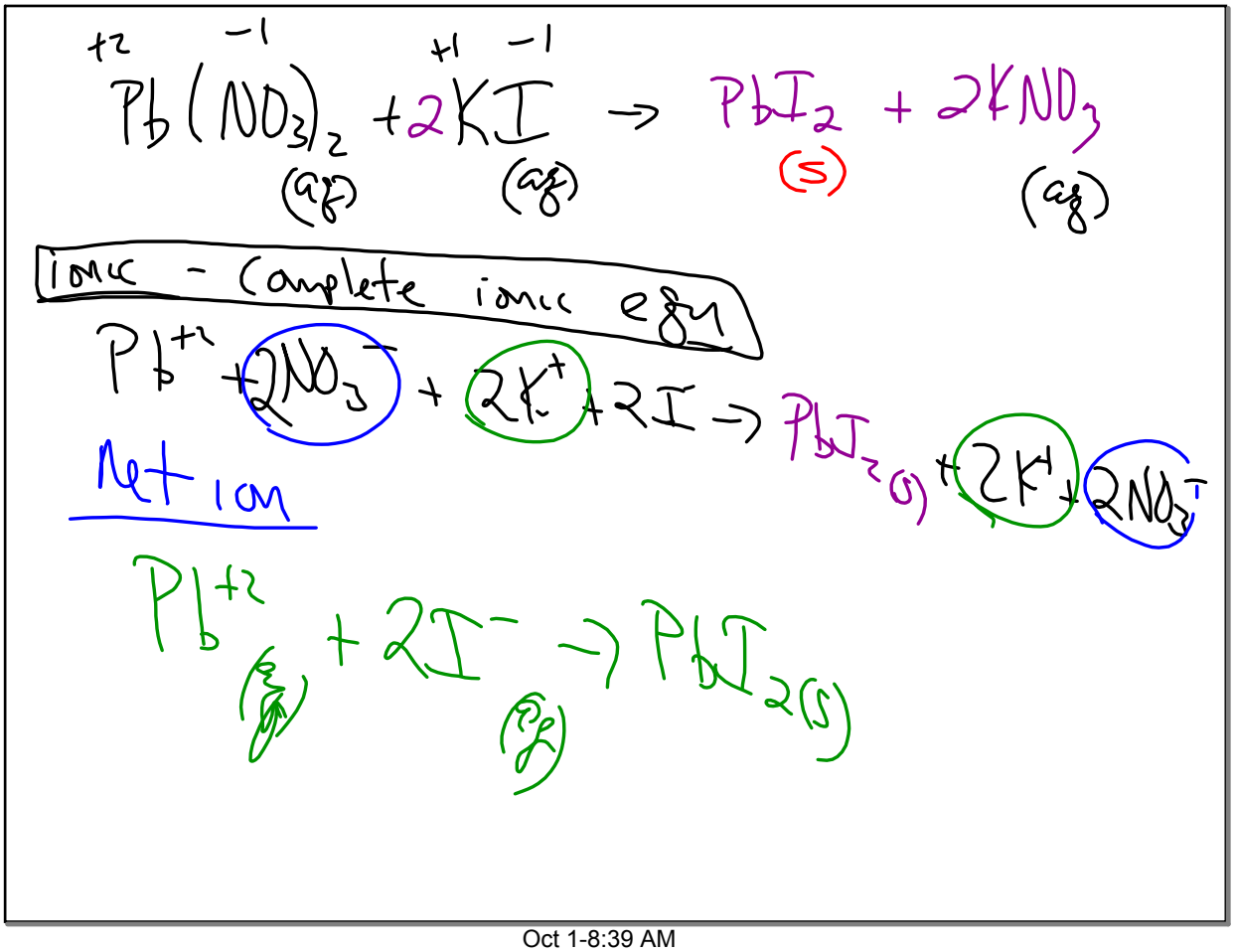
p125 Table 4.1 ⇒ Solubility



Oct 1-8:31 AM



Oct 1-8:37 AM



Oct 1-8:39 AM

4.24 a+c

4.62 a+c

Oct 1-8:47 AM