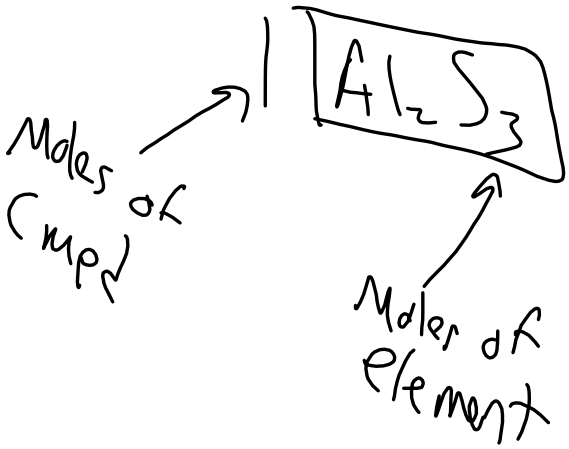
(17) $\frac{25\ g\ Al_2S_3}{150\ g\ Al_2S_3} \left| \frac{1\ mole\ Al_2S_3}{1\ mole\ Al_2S_3} \right| \frac{3\ mole\ S}{1\ mole\ S} \left| \frac{6.02 \times 10^{23}\ atoms\ S}{1\ mole\ S} \right.$



Oct 5-8:03 AM

(20) $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

$8.8\ g \quad \_\_ g$

$\frac{8.8\ g\ C_3H_8}{44\ g\ C_3H_8} \left| \frac{1\ mole\ C_3H_8}{1\ mole\ C_3H_8} \right| \frac{5\ mole\ O_2}{1\ mole\ O_2} \left| \frac{32\ g\ O_2}{1\ mole\ O_2} \right. =$

Oct 5-8:22 AM

(14) $2Li + 2H_2O \rightarrow 2LiOH + H_2$

14g ——— g

$$\left|\frac{14g\ Li}{}\right|\frac{1\ mole\ Li}{7g\ Li}\left|\frac{1\ mole\ H_2}{2\ mole\ Li}\right|\frac{2g\ H_2}{1\ mole\ H_2} =$$

Oct 5-8:25 AM

(23) $2\boxed{NaOH} + \boxed{H_2SO_4} \rightarrow 2H_2O + Na_2SO_4$

H_2SO_4 Moles A = Moles B $NaOH$

$n_A * M_A * l_A = n_B * M_B * l_B$

$(2)(0.0875)(50ml) = (1)(0.115)(ml)$

Oct 5-8:27 AM

(21) $4NH_3 + 5O_2 \rightarrow 4NO + 6H_2O$

8.5g 12g

% yield = $\frac{12}{15} \times 100$

80%

$$\frac{8.5\text{g}\,NH_3}{} \left| \frac{1\text{mole}\,NH_3}{17\text{g}\,NH_3} \right| \frac{4\text{mole}\,NO}{4\text{mole}\,NH_3} \left| \frac{30\text{g}\,NO}{1\text{mole}\,NO} \right. = 15\text{g}\,NO$$

Supposed to get

Mathematically

Oct 5-8:30 AM

(26) $3Ba(OH)_2 + 2H_3PO_4 \rightarrow Ba_3(PO_4)_2 + 6HOH$

SB WA (s) (l)

$3Ba^{+2} + 6OH^- + 2H_3PO_4 \rightarrow Ba_3(PO_4)_2 + 6HOH$

(s) (l)

Oct 5-8:34 AM

EC1 0.353 g KOH, — ml, ~~0.0176 M KOH~~

$\frac{0.0176 \text{ mole KOH}}{1000 \text{ ml}}$

$\frac{1000 \text{ ml}}{0.0176 \text{ mole KOH}} \left| \frac{1 \text{ mole KOH}}{56.10 \text{ g KOH}} \right| \frac{0.353 \text{ g KOH}}{} =$

Oct 5-8:38 AM

Q5 Ca acetate

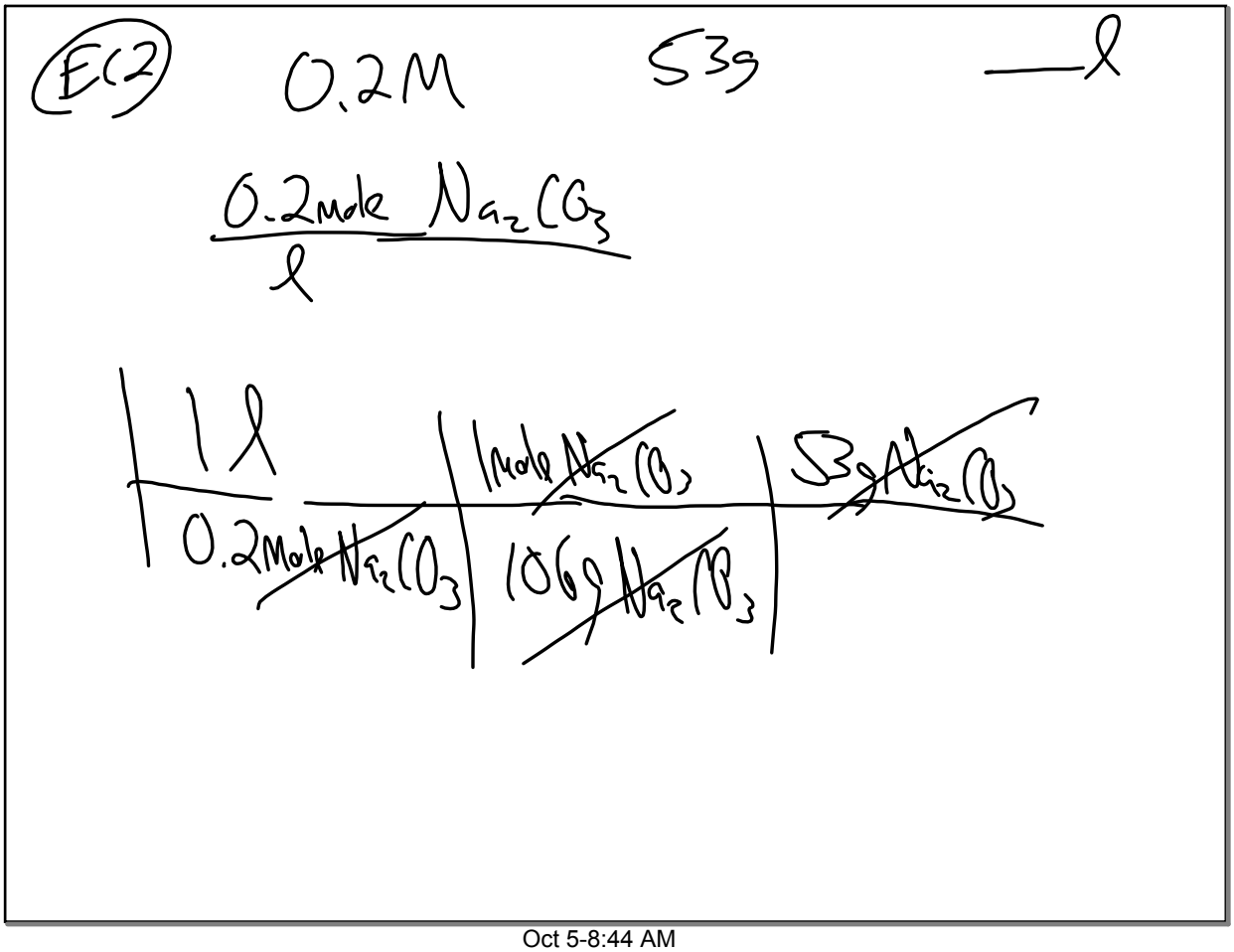
a) $Ca(C_2H_3O_2)_2 + NaOH \rightarrow Na\,Act + Ca(OH)_2$ (aq) (aq)

b) $Ca(C_2H_3O_2)_2 + Pb(NO_3)_2$

c) $Ca(C_2H_3O_2)_2 + FeCl_3$

$Ca(C_2H_3O_2)_2 + NH_4PO_4$ $Ca_3(PO_4)_2$ (s)

Oct 5-8:40 AM


E(2)
0.2M
53g
—l
0.2Mole Na_2CO_3
l
1 l
0.2Mole Na_2CO_3
1Mole Na_2CO_3
106g Na_2CO_3
53g Na_2CO_3

Oct 5-8:44 AM