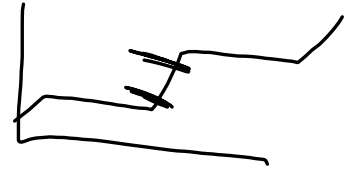


250g H₂O at 22°C → 98°C ΔT 76°C

$C = 4.18 \text{ J/g}^\circ\text{C}$

Find J

4.18 J	250g H ₂ O	76°C	=
g °C			



C M ΔT

Oct 16-8:38 AM

Some 1M HCl + Some 1M NaOH

ΔT = 21 → 27.5°C

4.18 J/g°C

Total = 100ml → 100g

d = 1g/ml

4.18 J	100g	6.5K	=	2700 J/g
g K				

Oct 16-8:41 AM

$$\frac{2700\text{J}}{\cancel{\text{g H}_2\text{O}}} \times \frac{\cancel{18\text{g H}_2\text{O}}}{1\text{mol H}_2\text{O}} =$$

Oct 16-8:44 AM

Hess's Law

5/52154

Oct 16-8:46 AM