(52) $\frac{242\,J}{g*K}$, ? T, 62g EG, 13.1°C → 40.5°C

242 J	62 g	27.4 K
g * K		

=

	°C	K
FP	0	273
BP	100	373

Δ = 100 = 100

ΔT°C = ΔT K

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(54) 3.88g $NH_4NO_3(s)$, 60g $H_2O(l)$, 23 → 18.4°C

solute, solvent

ΔT = 4.6°C

Find $\frac{KJ}{Mole\ NH_4NO_3}$

$\frac{4.18\,J}{g*°C}$ Solution

Solution

4.18 J	63.88 g	4.6°C
g*°C		

= 1228.28 J solution

1.23 KJ

Water gain heat (endo) lose heat

Find $\frac{KJ}{mole\ NH_4NO_3}$

1.23 KJ	80 g NH_4NO_3
3.88 g NH_4NO_3	1 Mole NH_4NO_3

= 25.33 KJ / Mole NH_4NO_3

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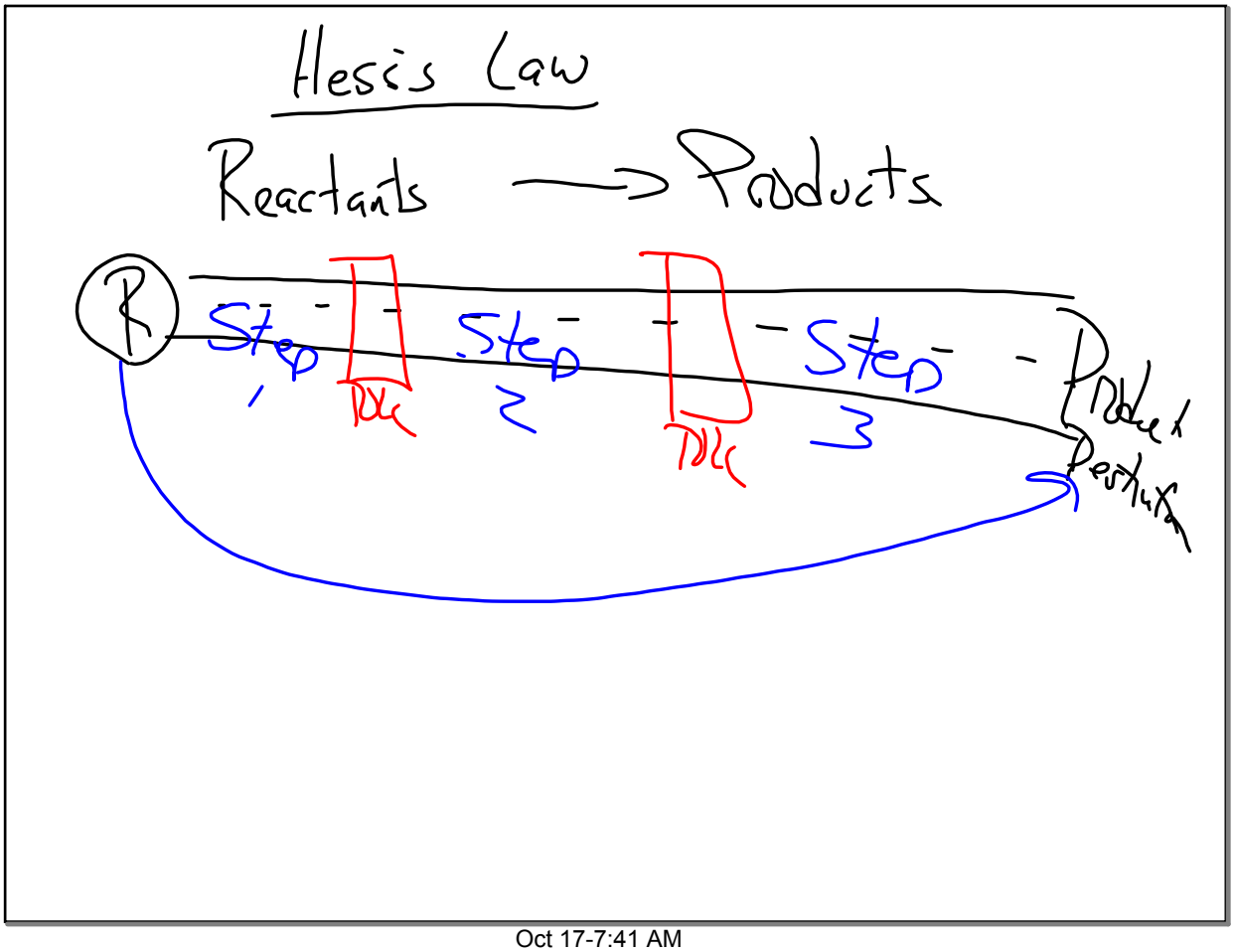

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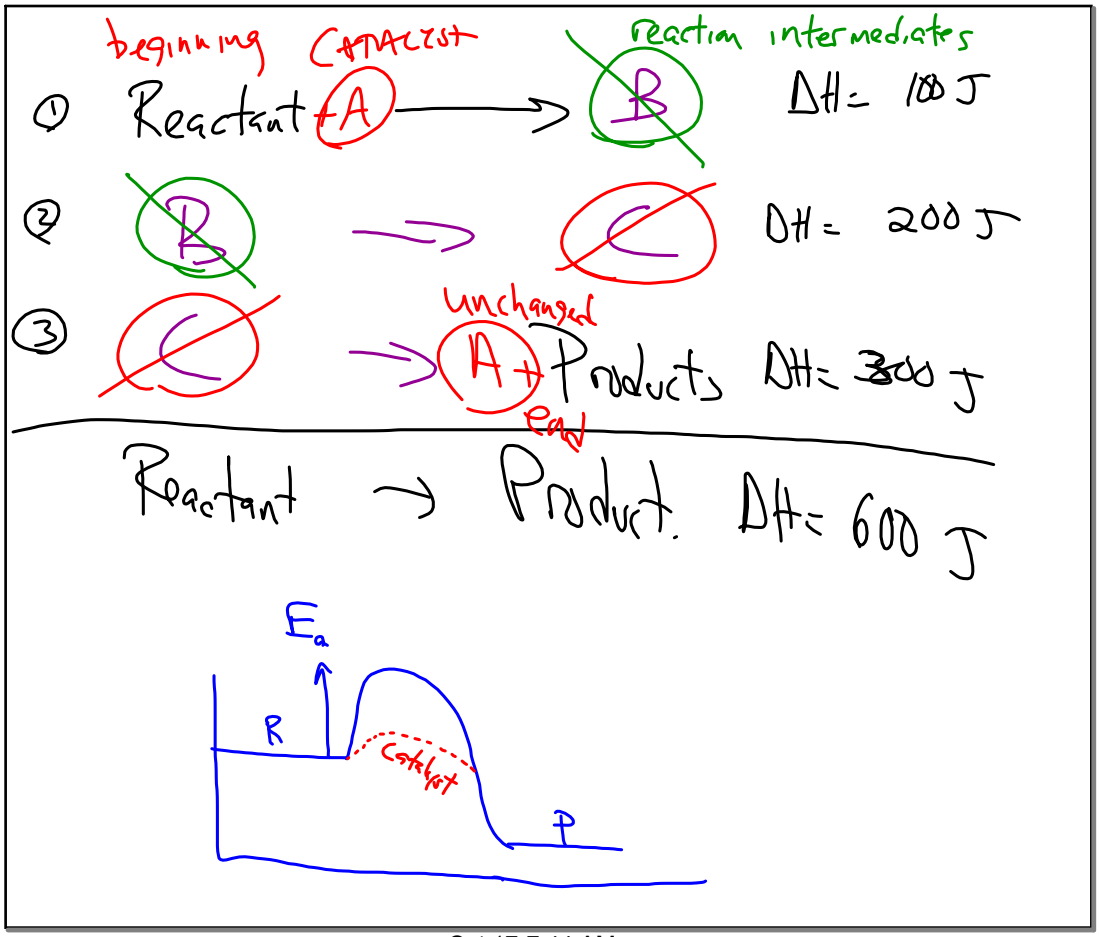

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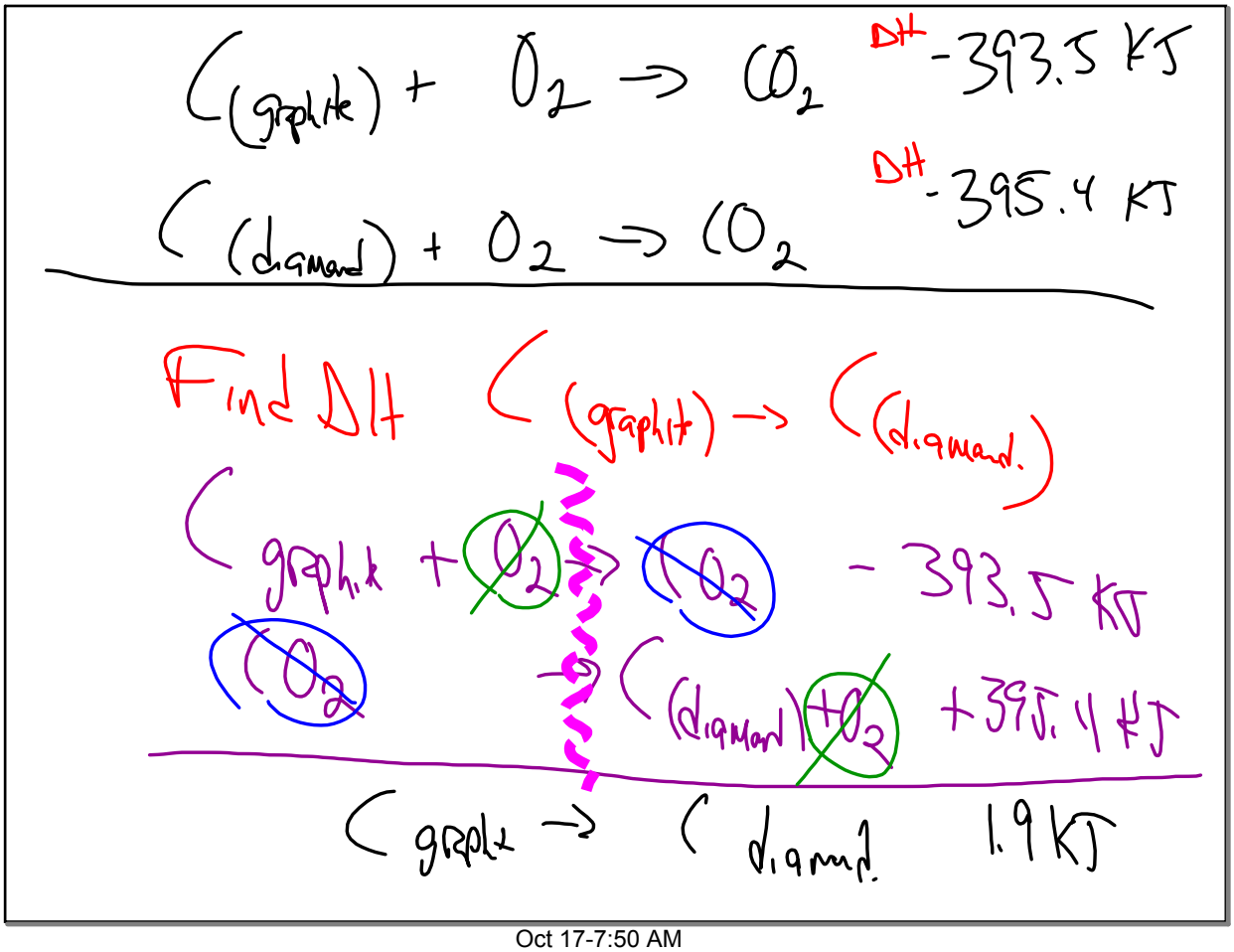

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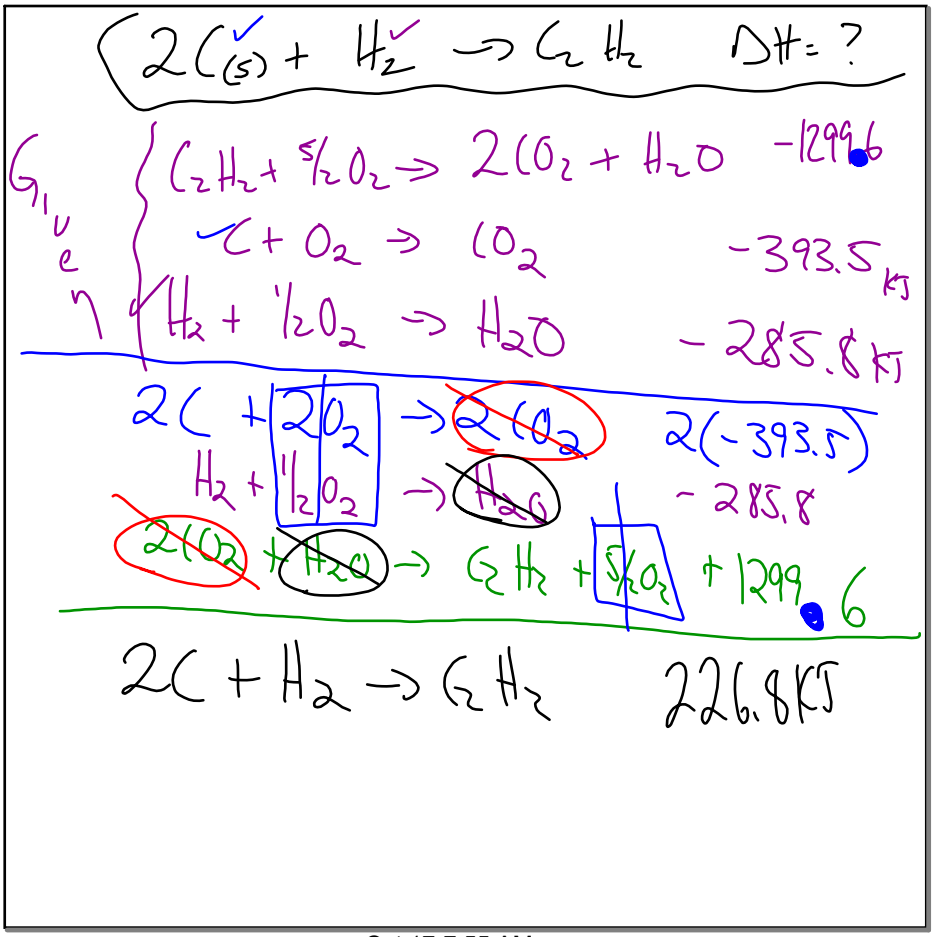

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