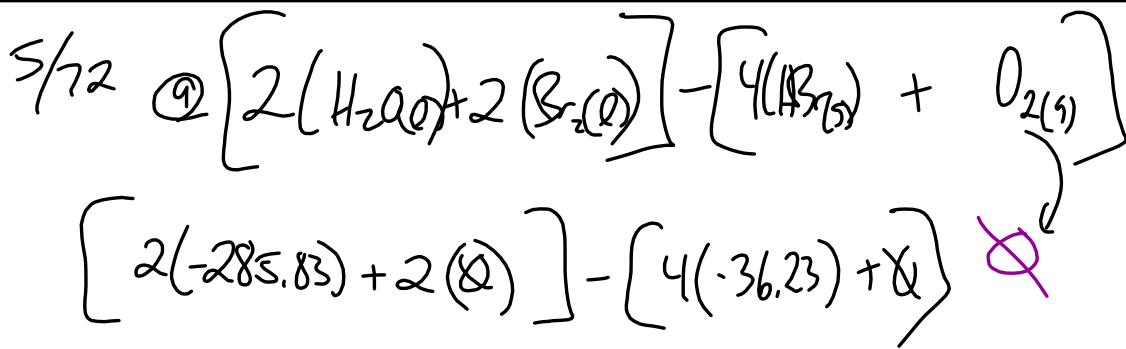


Oct 22-8:06 AM



PS 5.2

①

<del>12g KO</del>	<del>1 mole KO</del>	<del>-482 kJ</del>
	28g KO	2 mole KO

Oct 22-8:20 AM

$$\begin{aligned}
 \textcircled{13} \quad Q &= mc \Delta T \\
 &= 100(4.18)(6.8) \\
 Q &= 2842.4 \text{ J} \\
 &= \underline{\underline{2.84 \text{ kJ}}}
 \end{aligned}$$

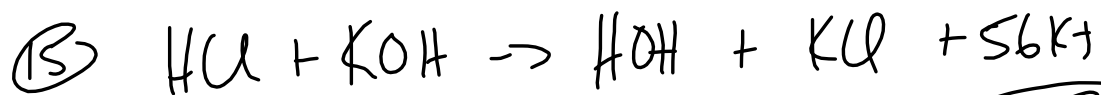
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$$\textcircled{14} \quad 1.96 \text{ g Ti}, \quad \frac{9.84 \text{ kJ}}{^\circ\text{C}}, \quad \Delta T = 61.98^\circ\text{C}$$

$$\text{Find } \frac{\text{kJ}}{\text{mole Ti}}$$

$$\begin{array}{c|c|c|c}
 \cancel{9.84 \text{ kJ}} & \cancel{61.98^\circ\text{C}} & & \cancel{48 \text{ g Ti}} \\
 \hline
 & & \cancel{1.96 \text{ g Ti}} & \text{Mole Ti} =
 \end{array}$$

Oct 22-8:41 AM



50ml      0.4M

0.22M

② Next.

mole A = mole B

$nMl = nMl$

$(1)(0.22)(50) = (1)$

$$\Delta H = -56\text{kJ}$$

$$C = \frac{4.18\text{J}}{90\text{C}}$$

$$D = \frac{1.01\text{g}}{\text{ml}}$$

$$T_f = ?$$

Oct 22-8:45 AM