

⑩ SO_3 Formal Charge on S

$$6 + 3(6) = 24$$

$$\begin{array}{r} -6 \\ \hline 18 \\ -18 \\ \hline 0 \end{array}$$

FC = value - (all $\frac{a}{nb}$ + $\frac{1}{2}b$)

$$S = 6 - (2 + 4) = +2$$

Resonance

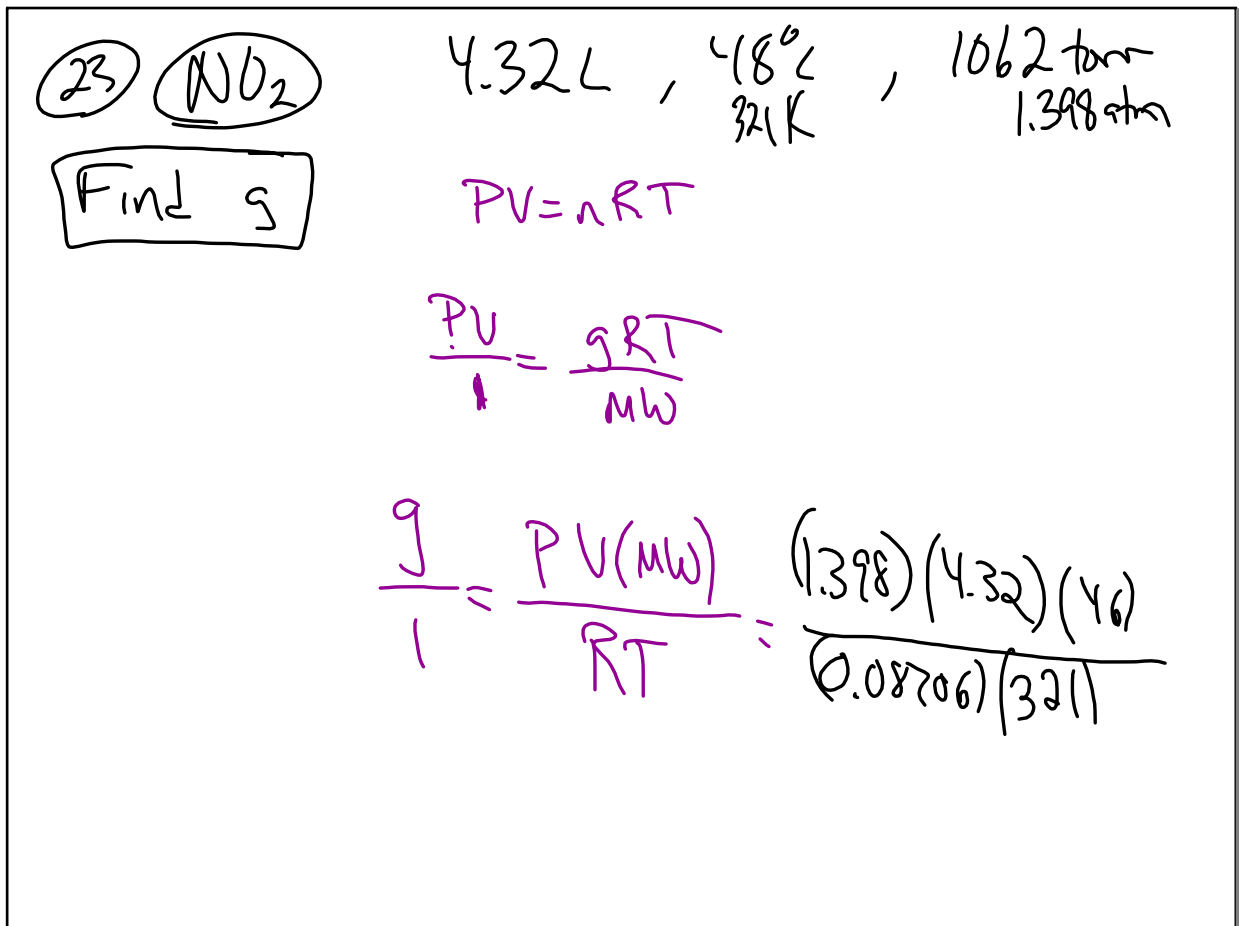
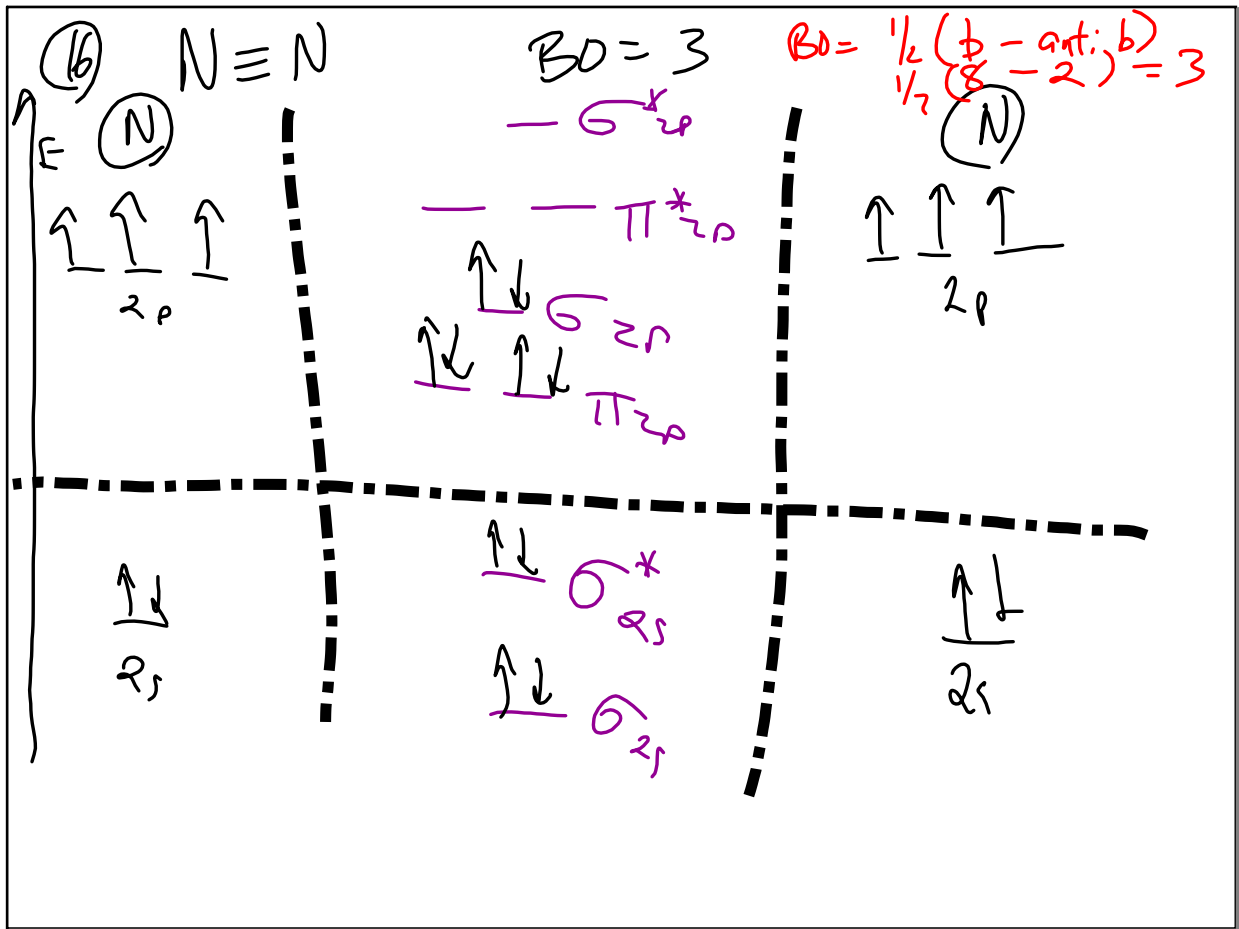
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⑬ XeF_4

$$8 + 4(7) = 36 - 8 = 28 - 24 = 4$$

4b, 2nb

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(26) $\text{NH}_4\text{NO}_2 (\text{s}) \rightarrow \text{N}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{g})$ 3 mole GAS

35g

35g NH₄NO₂	1 mole NH ₄ NO ₂	3 mole GAS
64g NH ₄ NO ₂	1 mole NH ₄ NO ₂	1 mole NH ₄ NO ₂

1.64 mole gas

525°C = 798K

1.5 atm

$PV = nRT$

$V = \frac{nRT}{P} = \frac{1.6g(0.08206)(798)}{1.5 \text{ atm}}$

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(27) 0.924g N₂O + 0.825g NO

$P_{\text{N}_2\text{O}} = ?$ $P_{\text{NO}} = ?$ $P_T = 1.32 \text{ atm}$

$P_{\text{N}_2\text{O}} = X_{\text{N}_2\text{O}} P_T$

$= \frac{\text{Mole N}_2\text{O}}{\text{Total Mole}} (1.32)$

$P_{\text{N}_2\text{O}} = 0.75 \text{ atm}$ + 0.57 atm →

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