

① Low BP \Rightarrow Weakest IMF.
 \hookrightarrow Weak. Large Radius

③ Ar
 molecular (elemental)
 LDF only.

HI
 Polar
 More Dipole Dipole
 \hookrightarrow Bigger ($>$ Mass)
 More LDF. More inter.

HCl
 Polar
 less Dipole Dipole
 more intra.

Increasing BP.

Low IMF \longrightarrow High IMF

Ar < HCl < HI

Intermolecular

Dec 20-8:05 AM

⑤ Ion - Dipole

Ionic: Na^+ Cl^-

Polar: H_2O

N_3^-

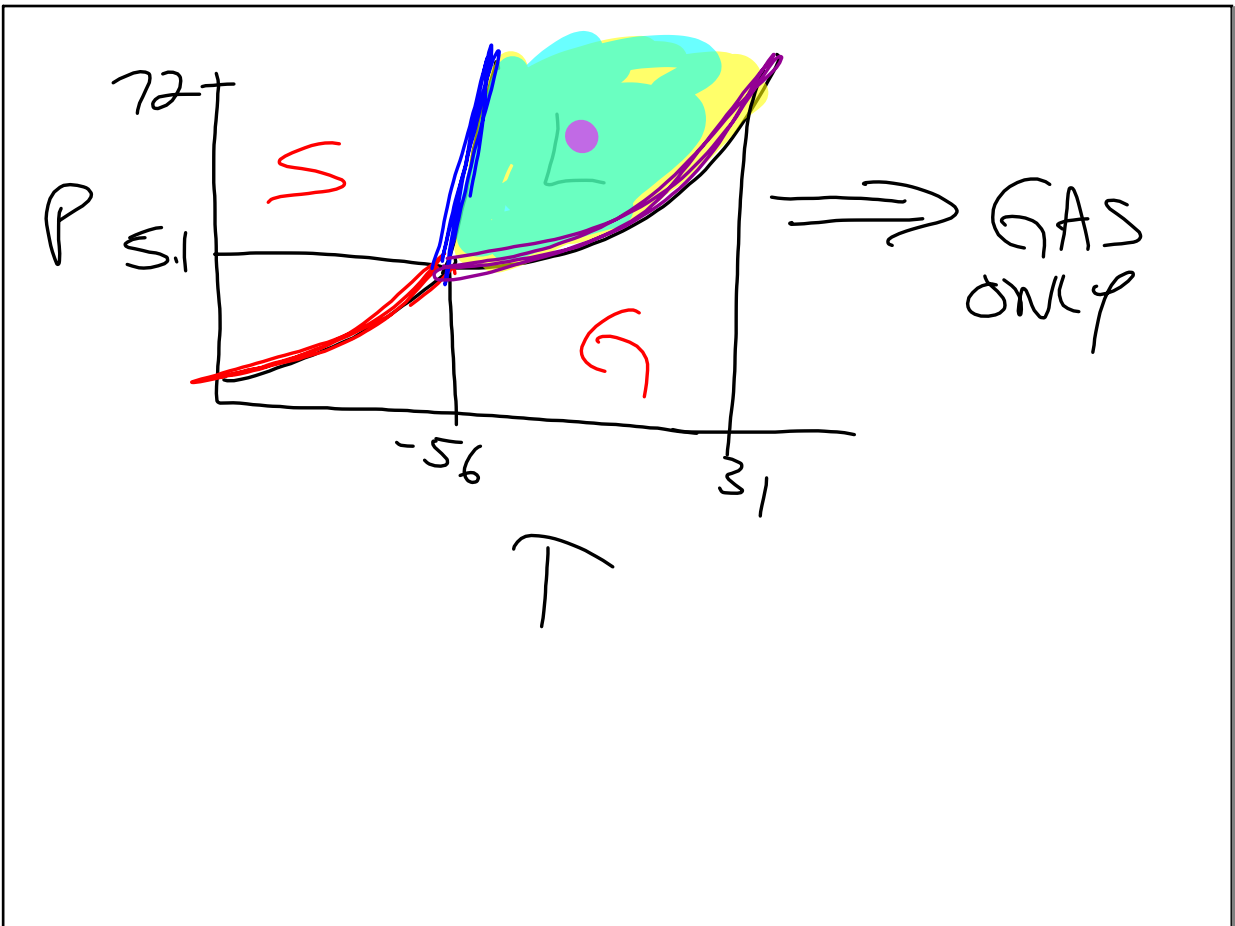
Cl^-

Dec 20-8:24 AM

① CO_2 $\frac{BP}{77^\circ} = \frac{H_v}{?}$ $\frac{31.32}{40.68} \text{ KJ/mole}$

② $\Rightarrow H_2O$ $\frac{100^\circ}{2260 \cancel{J}} \cdot \frac{1 \text{ KJ}}{1000 \cancel{J}} \cdot \frac{1.8 \cancel{g} H_2O}{1 \text{ mole } H_2O} = 40.68 \text{ KJ/mole } H_2O$

Dec 20-8:33 AM



Dec 20-8:38 AM

Chap 13 → Solution chemistry

Solute - gets dissolved

Solvent - does the dissolving

← MORE present

Solvation Army

Dec 20-8:41 AM