Unsat (Dilute vs Concentrated) | SAT Full. | SUPA sat. Super → only under Special Conditions!

$$\text{Molarity} = \frac{\text{Moles solute}}{\text{L of Solution}}$$

Big (M)

Solute + Solvent

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$$\text{Mass \%} = \frac{\text{Part}}{\text{Whole}} * 100$$

ppH (parts per hundred)

Solute + Solvent.

$$\text{PPM} = \frac{\text{Part}}{\text{Whole}} \times 1{,}000{,}000$$

10^6

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Mole fraction

$$X = \frac{\text{part moles}}{\text{whole moles (TOTAL)}}$$

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Molality (m) = $\frac{\text{moles of solute}}{\text{Kg of solvent}}$

(1.HL2)

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Dissolve Quickerer Beterer

1. Heat
2. ↑ P on gas (Flat subs i")
3. Stir
4. ↑ Surface area
5. Bond type.
6. How much already dissolved

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Henry's Law

$C_g \approx P_g$

$C_g = k \cdot P_g$

- C_g: Solubility of gas in solution
- k: constant
- P_g: Pressure on gas

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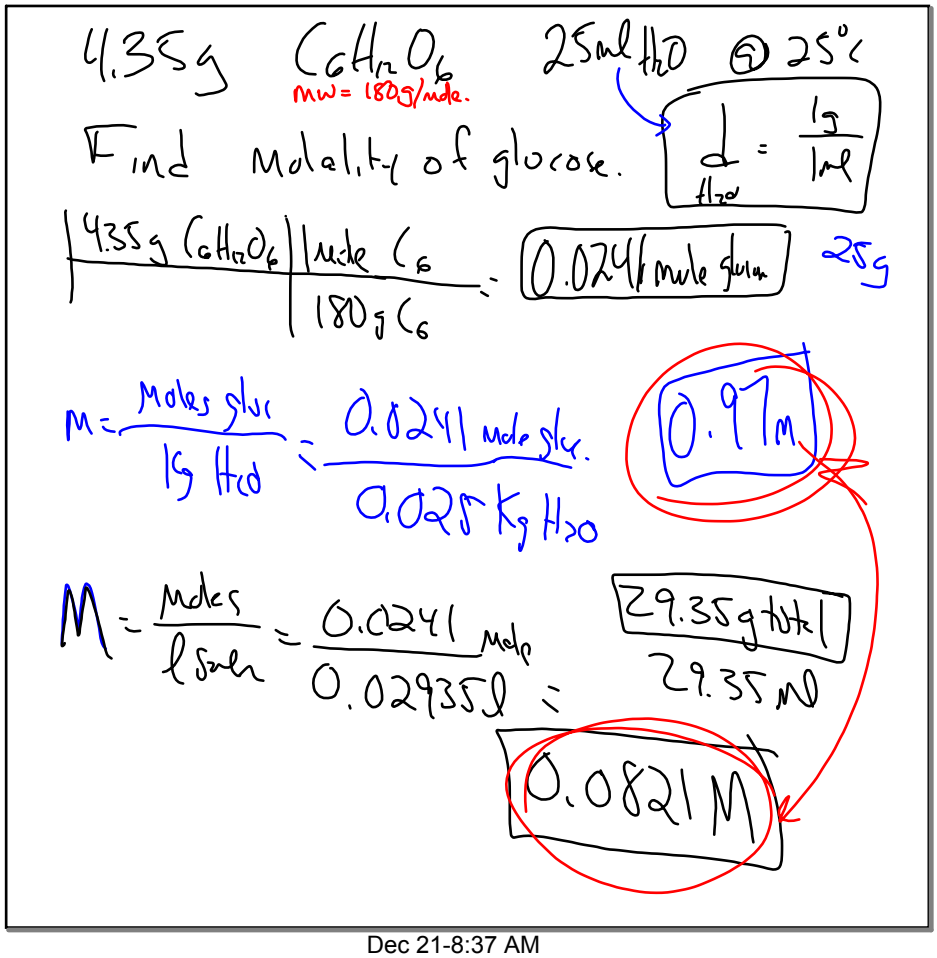

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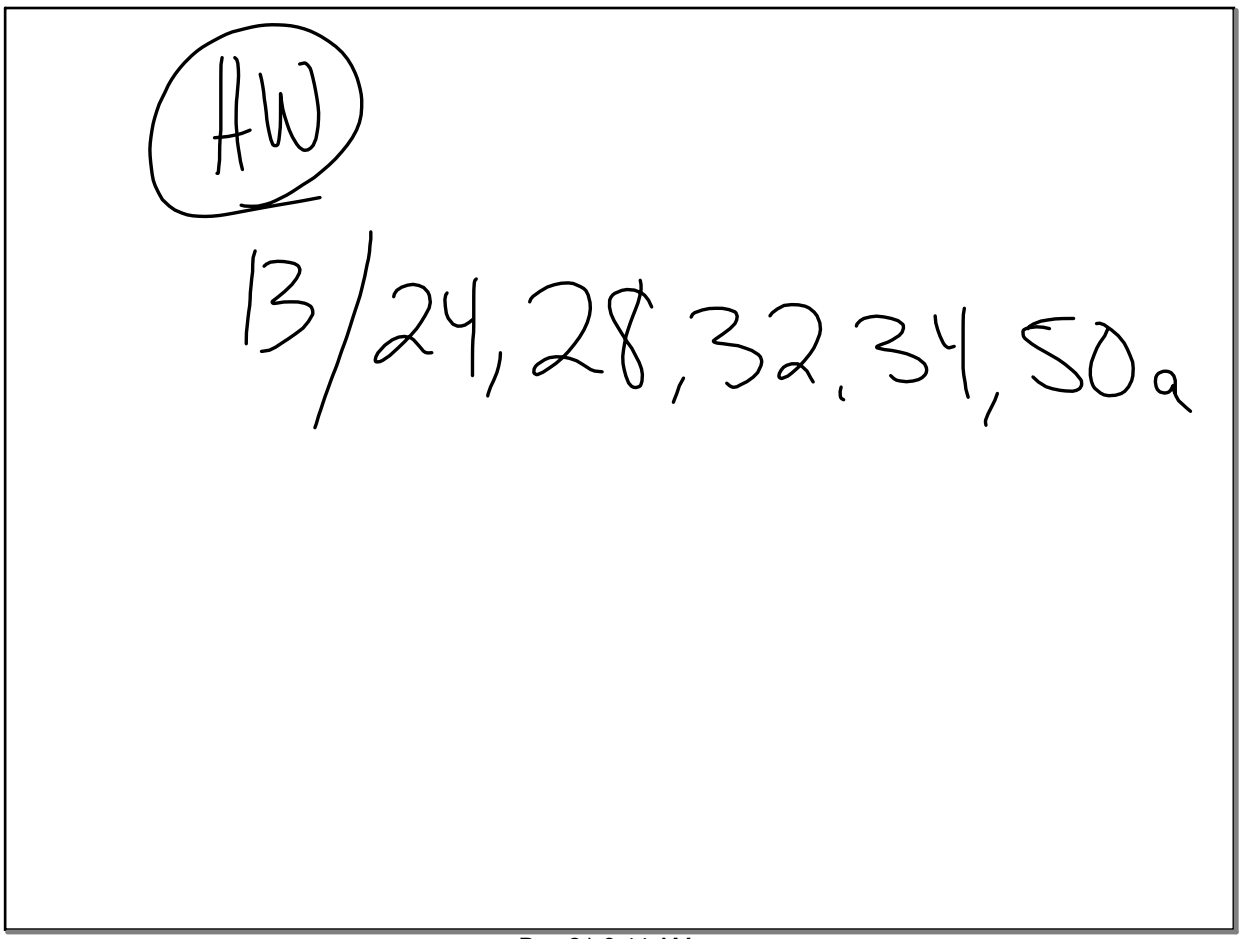

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