(13) $\Delta G = \Delta H - T\Delta S$

$104.7 = \Delta H - (298)(0.173)$

$\Delta G = [3(86.7)] - [1(103.6) + 1(51.8)]$

(14) $\Delta G = \Delta H - T\Delta S$

$= 137 - (0.12)$

Spont ⊖ ↑T

Non Spont ⊕ ↓T

Mar 25-7:20 AM

(19) $\Delta G = -RT\ln K$

$26 = -(8.314 \times 10^{-3})(298)\ln K$

$\Delta G = [77 + 2(-237)] - [-943]$

$\Delta G = +26$ kJ

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(20) $\Delta G = \Delta H - T\Delta S$

$-121 = -151 - 298\,\Delta S$

$\Delta S = 0.101\ kJ$

$101\ J$

Spont ↔ non spont

$\Delta G = 0$

$0 = -151 - T(0.101)$

$T = \frac{151}{0.101} \approx 1500\ K$

Mar 25-8:13 AM

Chap 20 Redox + electrochem.

↳ flow of e⁻

Reduction

GER

(RIG)

FAT RED CAT

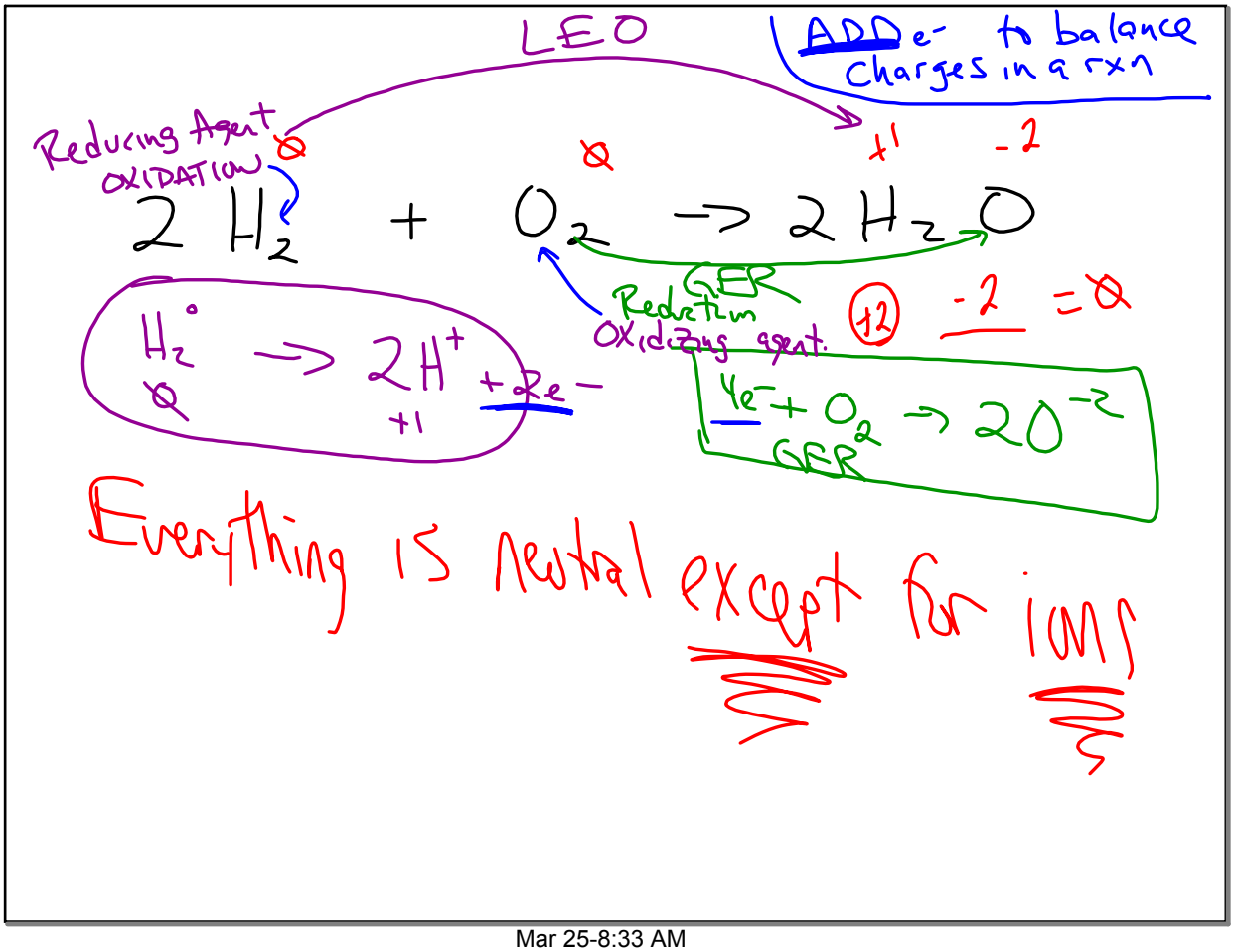
oxidation

LEO

(OIL)

→ Reduction occurs at cathode and gains mass

Mar 25-8:26 AM



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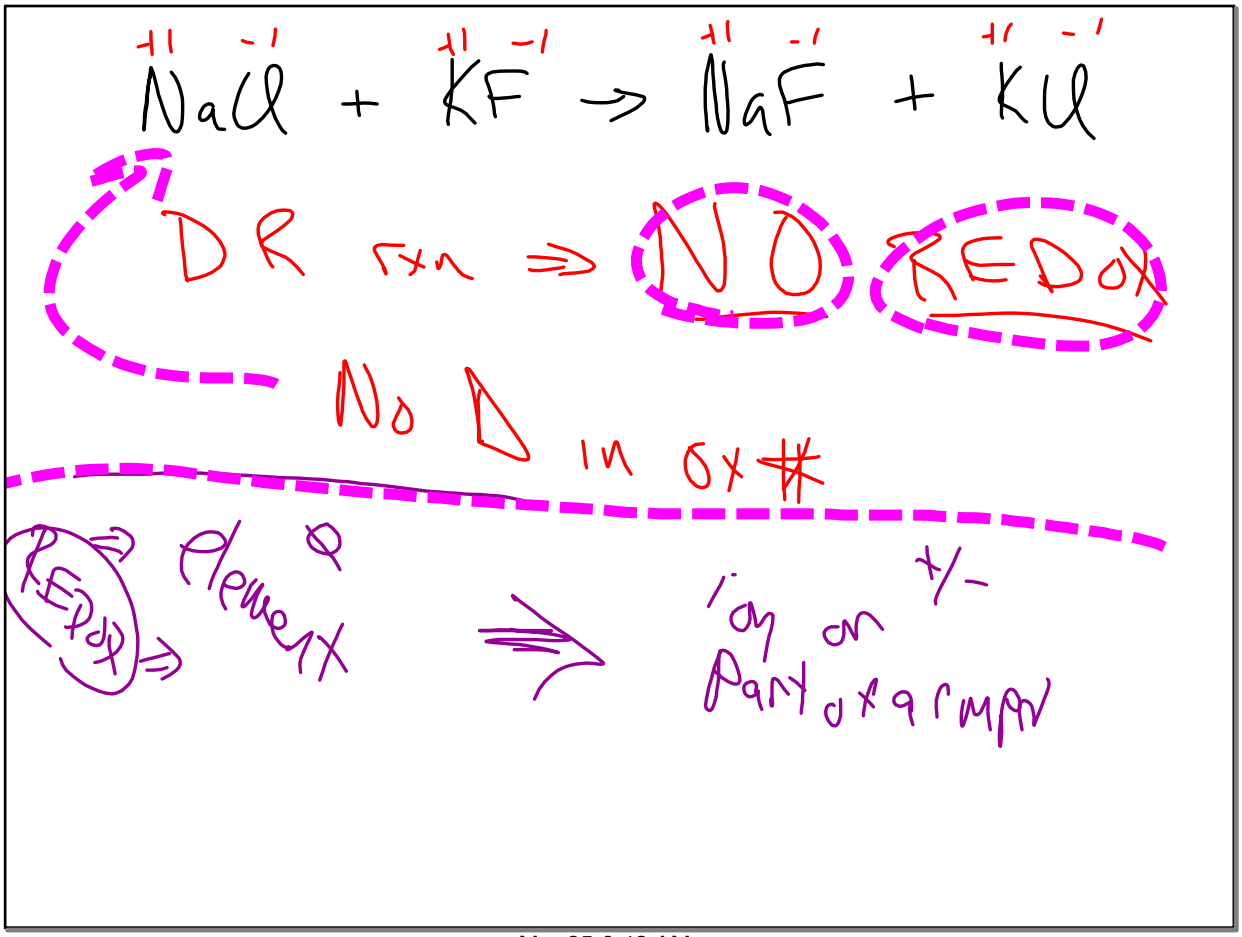

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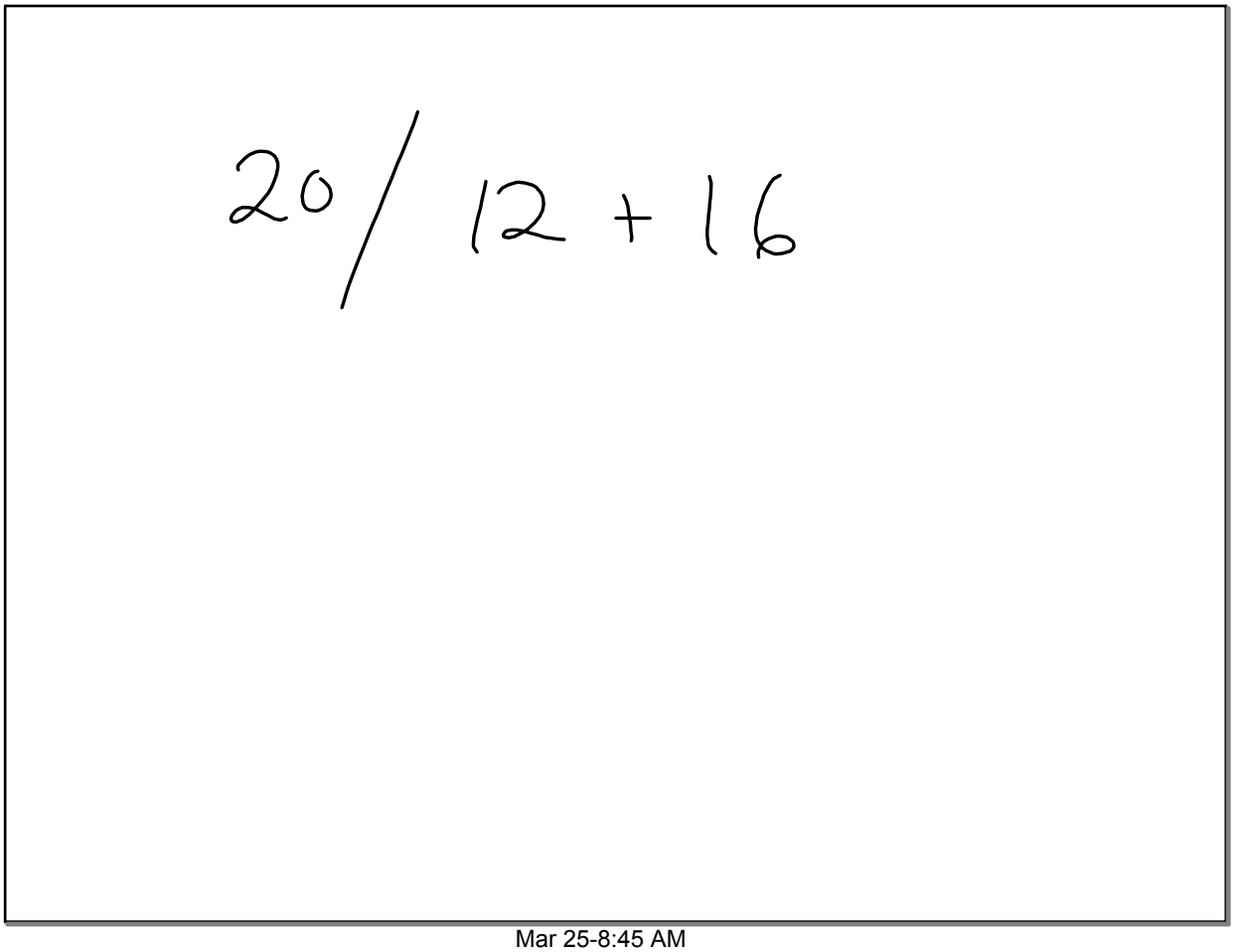

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