

FLM PS

② $8.55 \text{ mi}^3 = \underline{\quad\quad\quad} \text{ l}$

5280 ft = 1 mile
 1 ft = 12 in
 1 in = 2.54 cm

$1000 \text{ ml} = \text{cm}^3$
 $1 \text{ l} = 10^{-3} \text{ m}^3$

8.55 mi^3	5280^3 ft^3	12^3 in^3	2.54^3 cm^3	1 l	5.28×10^3
	1^3 mi^3	1^3 ft^3	1^3 in^3	1000 ml	

Sep 17-7:19 AM

⑥ $24' \times 51' \times 10'$ room } Ceiling — yd^2
 $1224 \text{ ft}^2 = \underline{\quad\quad\quad} \text{ yd}^2$

1224 ft^2	1 yd^2	$= 136 \text{ yd}^2$
	3^2 ft^2	

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⑦ $\frac{5.83 \text{ pounds Gasoline}}{1 \text{ gal}} = \frac{3 \text{ tons.}}{\text{Volume!}}$

Unit to find \Rightarrow ALWAYS on TOP (numerator)

1 gallon	2000 pounds	3 tons	= 1029.16 gal
5.83 pounds	1 ton		

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⑧ 370 pounds. $4' \times 6' \times 3'$

$72 \text{ ft}^3 \rightarrow \text{in}^3 \rightarrow \text{gal}$

$\frac{\text{gal H}_2\text{O}}{1 \text{ gal} = 231 \text{ in}^3}$

Unit

1 gal	231 in³	ft³	
1 gal	12³ in³	72 ft³	

$\frac{1}{10 \text{ cm}} =$

Sep 17-7:51 AM

Chap 2 Atomic Theory

empirical → reduced
us

Molecular → Not reduced.

Ex: Lead (IV) oxide

$$\begin{matrix} +4 & & -2 \\ \text{Pb} & \text{---} & \text{O} \\ \text{2} & & \text{2} \end{matrix}$$

Pb₂O₄ Molecular

PbO₂ empirical

← Neutral
need 2 O's

PbO₂

Sep 17-7:59 AM

Thompson's Cathod Ray tube (CRT)

Mach's CRT (Goldstein)

Cathode

Cations ⊕ ions

⊖ charges
Subatomic particles
Green glow.

Anode ⊕

Anions ⊖ ions

Sep 17-8:33 AM

$$2 / 22 + 26$$

Sep 17-8:47 AM