⑧ $Rate = K[NO]^2[Br_2]^1$

$\frac{1}{2}\frac{\Delta[NO]}{\Delta t} = \frac{\Delta[Br_2]}{\Delta t}$

$2NO + Br_2 \rightarrow 2NOBr$

25% (1/4) 50% (1/2)

Feb 6-8:07 AM

⑪ Δ calc Δ rate

$1.5^{\boxed{x}} = 2.3$

$x\frac{\ln 1.5}{\ln 1.5} = \frac{\ln 2.3}{\ln 1.5}$

Feb 6-8:21 AM

(R) $\ln A_t = -kt + \ln A_0$

$\ln A_t = -(3.4)(1) + \ln 100$

1°: $\ln A_t = -kt + \ln A_0$

$t_{1/2} = \frac{0.693}{k}$ (1°)

2°: $\frac{1}{A_t} = kt + \frac{1}{A_0}$

$\frac{t_{1/2}}{1} = \frac{1}{k[A_0]}$ (2°)

$\frac{k}{1} = \frac{1}{t_{1/2}[A_0]}$

$k = \frac{1}{185(0.01)}$

Feb 6-8:25 AM

(25) Rate $= k[O_2][NO]^2$

(1) (2)

(16)

100 $\xrightarrow{1}$ 50 $\xrightarrow{2}$ 25 $\xrightarrow{3}$ 12.5 $\xrightarrow{4}$ 6.25 $\xrightarrow{5}$ 3.12 ✓

Feb 6-8:37 AM

(7) $\frac{1}{A_t} = kt + \frac{1}{A_0}$

90% reacted
10% left am
0.1 (0.1) = 0.01 left over

$\frac{1}{0.01} = (0.014)t + \frac{1}{0.1}$

10% of A_0

Feb 6-8:41 AM

$K = \frac{(Prod)^{coeff}}{(React)^{coeff}}$

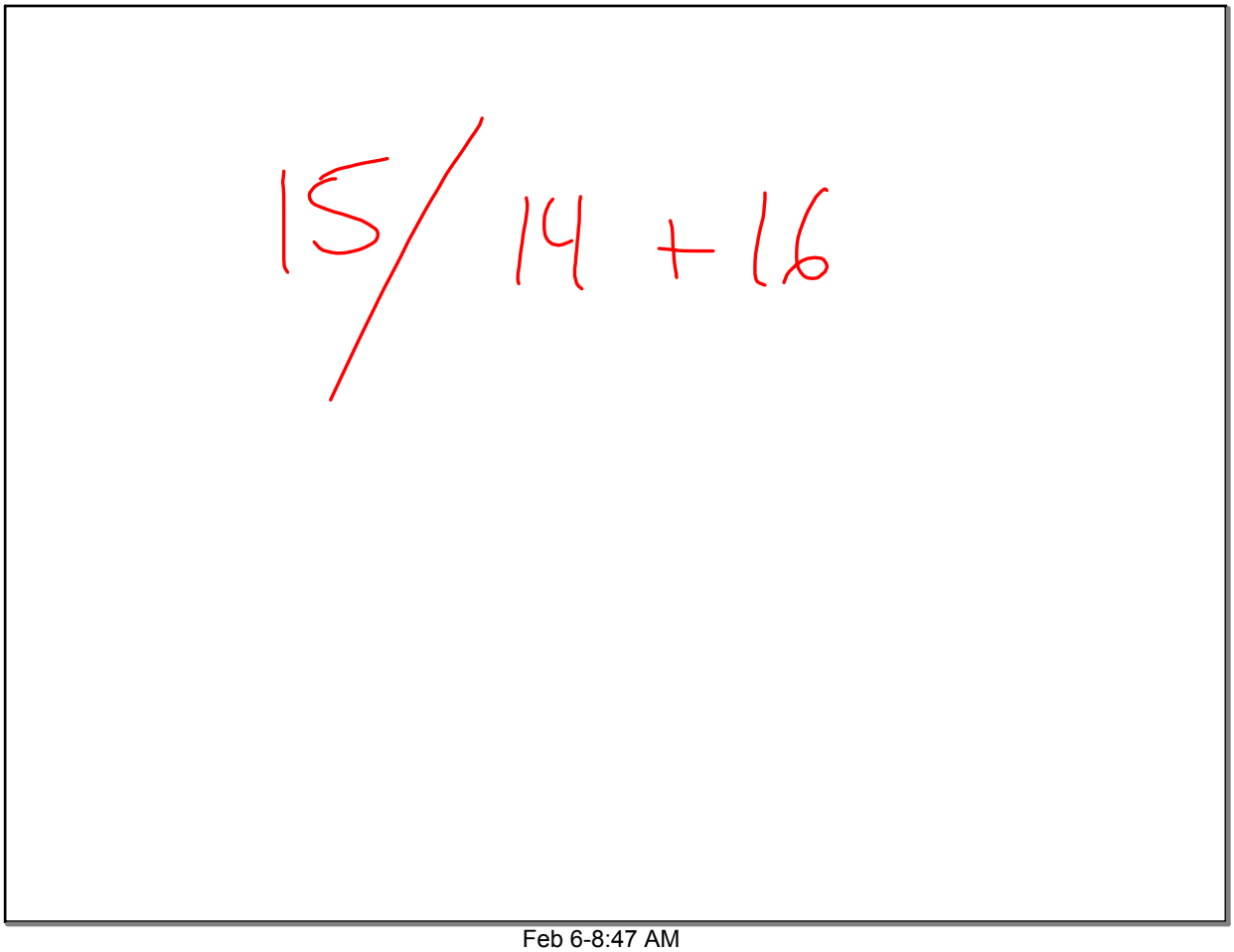
(s) or (l) = 1

K_c g

K_p P

Only (aq) and (g)

Feb 6-8:44 AM



Feb 6-8:47 AM