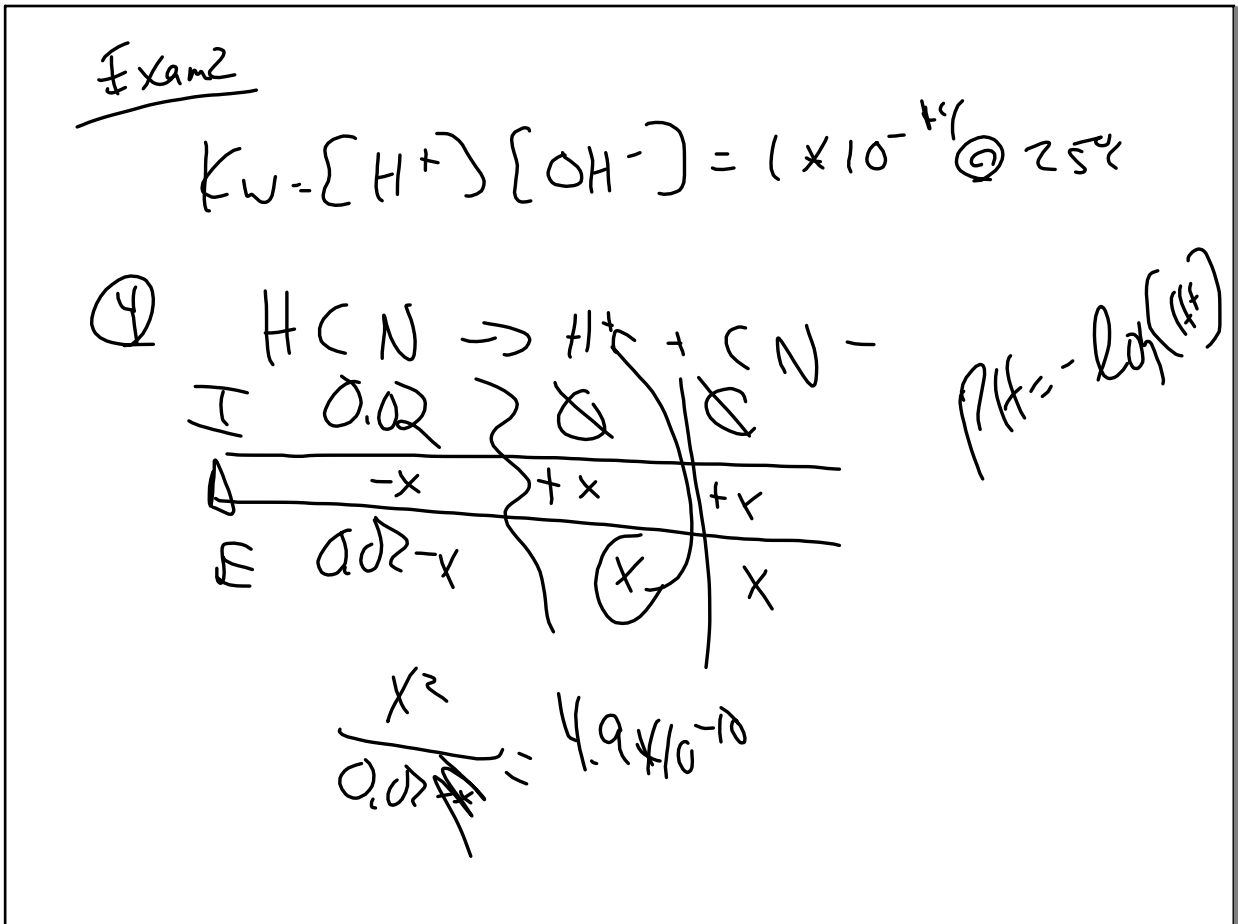
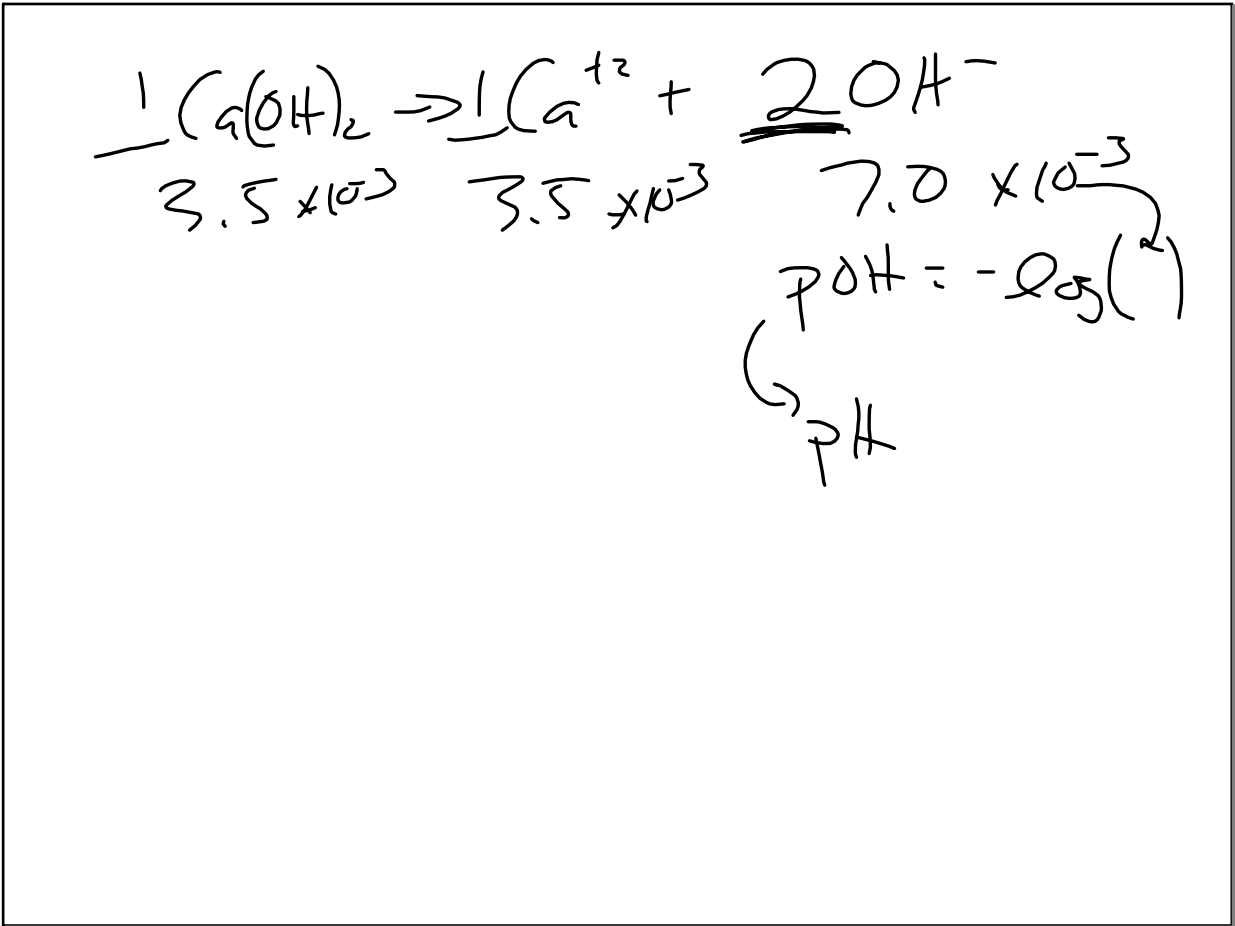


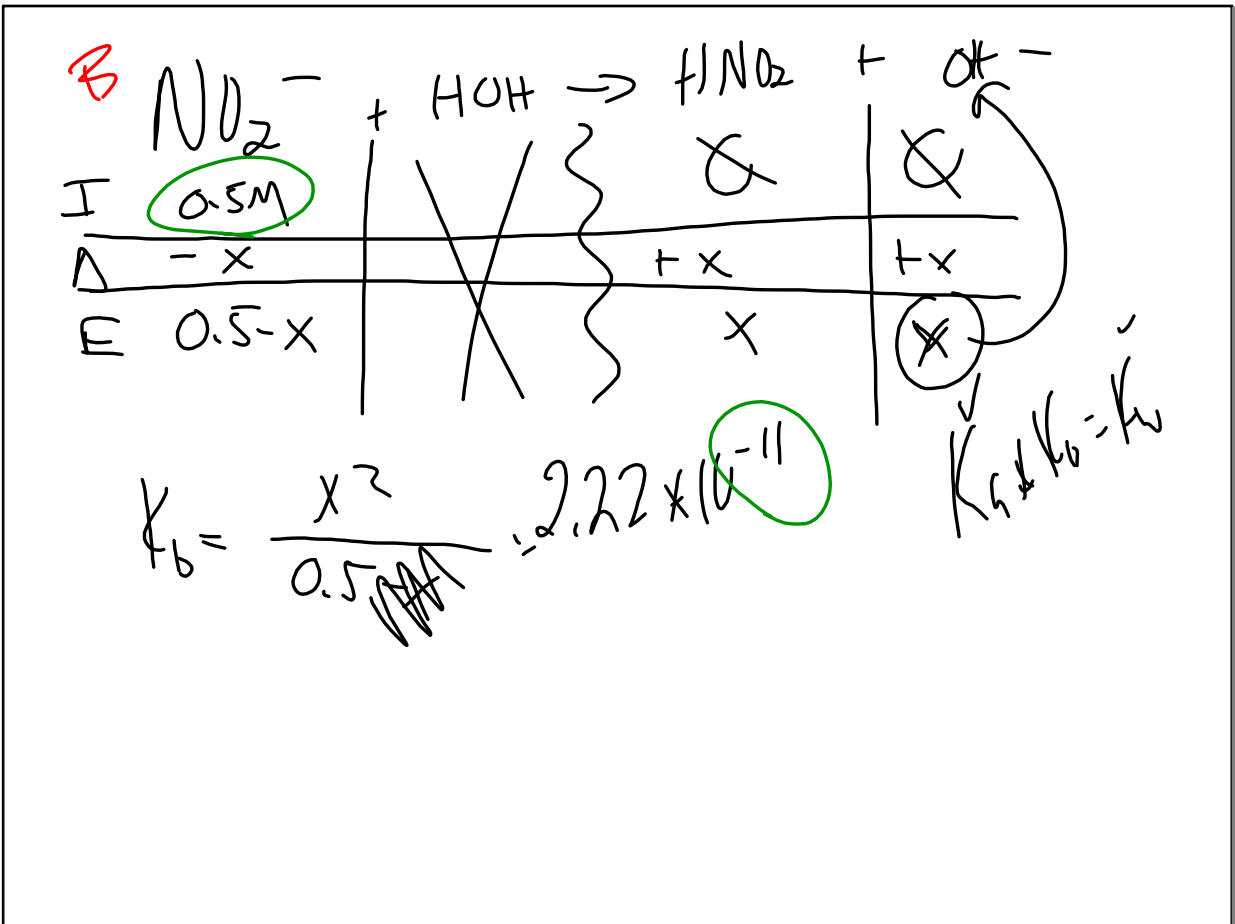
Mar 11-8:01 AM



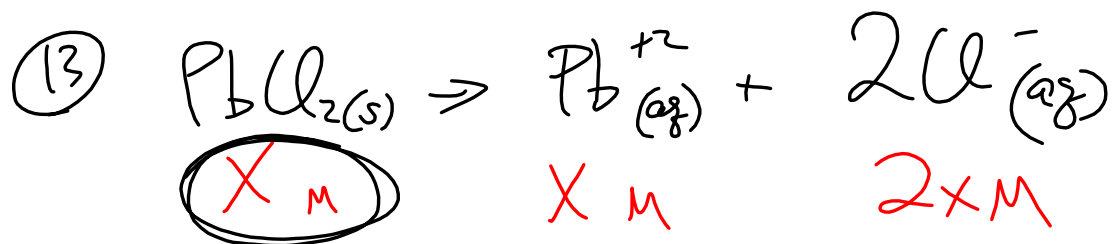
Mar 11-8:12 AM



Mar 11-8:17 AM



Mar 11-8:20 AM



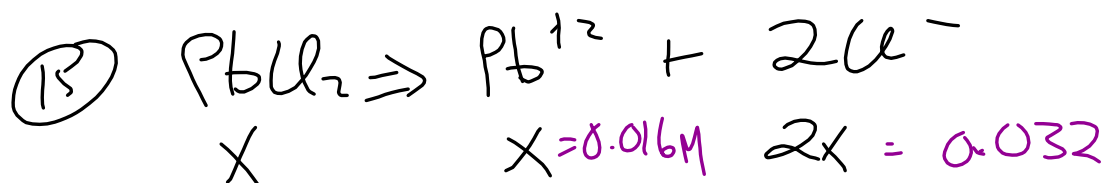
$$K_{sp} = [\text{Pb}^{+2}][\text{Cl}^{-}]^2 = 1.6 \times 10^{-6}$$

*

$$(X) (\cancel{2X} \cdot 0.5)^2 = 1.6 \times 10^{-6}$$

$$X = 6.4 \times 10^{-6}$$

Mar 11-8:27 AM



$$[\text{Pb}^{+2}][\text{Cl}^{-}]^2 = K_{sp}$$

$$\underbrace{(X) (2X)^2}_{\text{circled}} = K_{sp}$$

$$4X^3 = K_{sp}$$

$$4(0.016)^3 = K_{sp}$$

Mar 11-8:33 AM

(19)

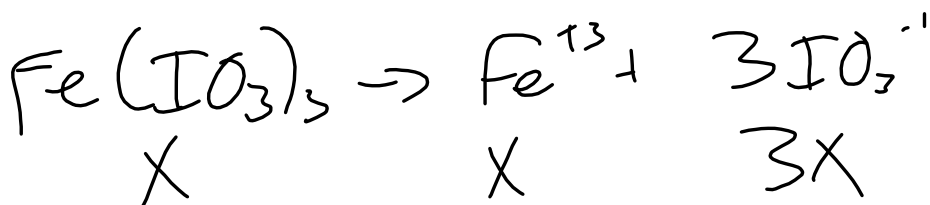
$$\text{mole } A = \text{mole } B$$

$$nM^m = nM^m$$

$$(2) M(2) = 1(0.375)(62.5)$$

Mar 11-8:42 AM

(23)



$$(\text{Fe}^{+3})(\text{IO}_3^{-})^3 = K_{sp}$$

$$(1 \times 10^{-4})(1 \times 10^{-5})^3 =$$

$$Q < K$$

$$Q_{sp} = 1 \times 10^{-19}$$

Mar 11-8:45 AM