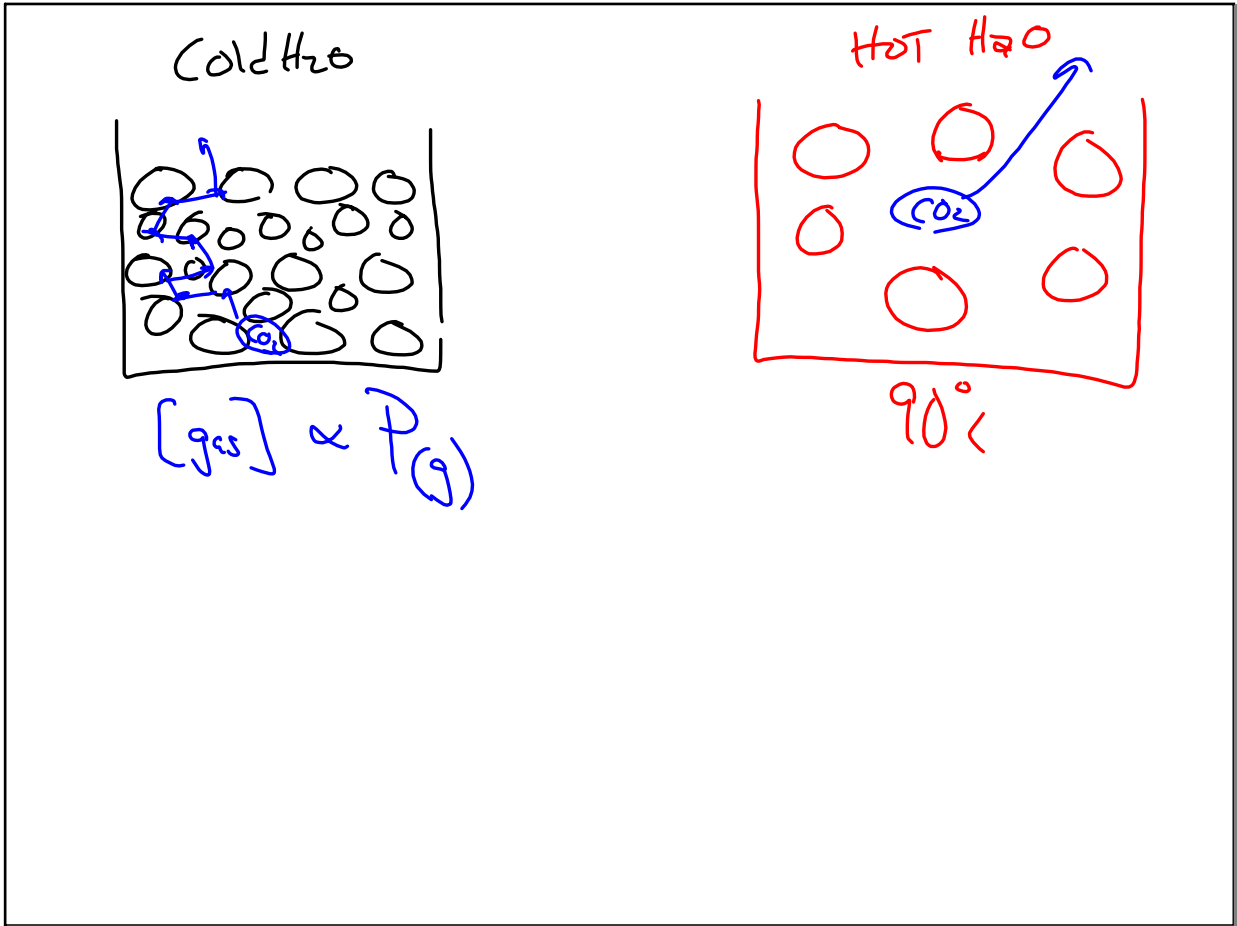


Jan 6-7:52 AM

- ↑ Solubility and RATE
- ① Stir
 - ② (cube → crystal → Powder)
 - ↑ Surface area solute
 - ③ Heat
 - ④ Like bonding solute - Solvent
 - Polar → Polar
 - Non-Polar → NP
 - ⑤ [] to start
 - concentration []
- Dissolve Faster
- ↑ KE

Jan 6-8:33 AM



Jan 6-9:00 AM

4.35g C₆H₁₂O₆ (180g/mole) in 250ml H₂O @ 25°C

density = $\frac{1g}{1ml}$

Find M + m

Moles solute / 2 solution → Molar Ratio

mole L ity

4.35g	1mole		
	180g	0.25L	

=

4.35g	1mole		
	180g	0.25L	0.097 _m

250ml H₂O

4.35g +

254.35ml = 0.095M

Jan 6-9:04 AM

Colligative Properties

" Rich get Richer, Poor get Poorer. "

When you add a Solute to a solvent }
 Hot gets hotter, Cold gets colder.

(↑ BP) (↓ FP)

Jan 6-9:16 AM

$\Delta T = (K \times m) \cdot i$

Change in BP/FP Constant K_b / K_f Molality $m = \frac{\text{moles solute}}{\text{Kg solvent}}$ Van't Hoff factor i (# of ions in solution)

$\text{NaCl} = \text{Na}^+ + \text{Cl}^-$
 $\text{CaCl}_2 = \text{Ca}^{2+} + 2\text{Cl}^-$ (2)
 $\text{CH}_4 \rightarrow$ (1) molecule
 $\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow$ (1) molecule

Jan 6-9:25 AM

13 / 43. SO a+b

Jan 6-9:31 AM